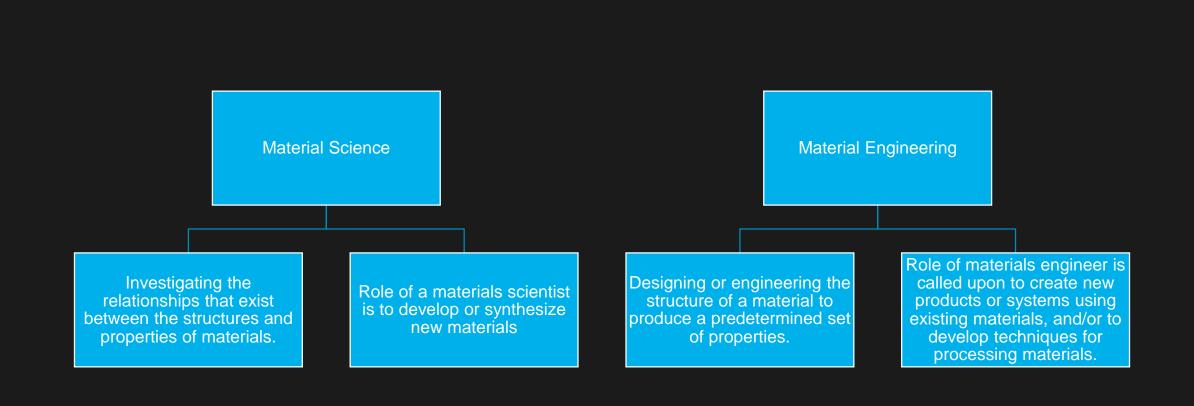


Material Sciences Review

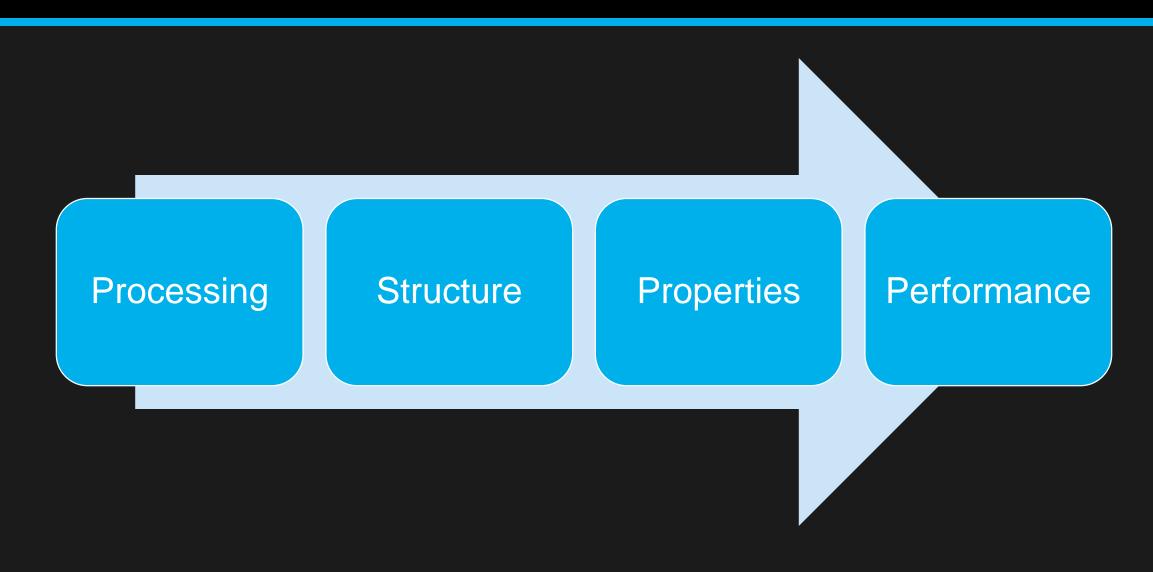
References/ Credit

- Material Science and Engineering an Introduction by William Callister & David Rethwisch.
- Alan Asper Guzik , NeurlPS 2019.

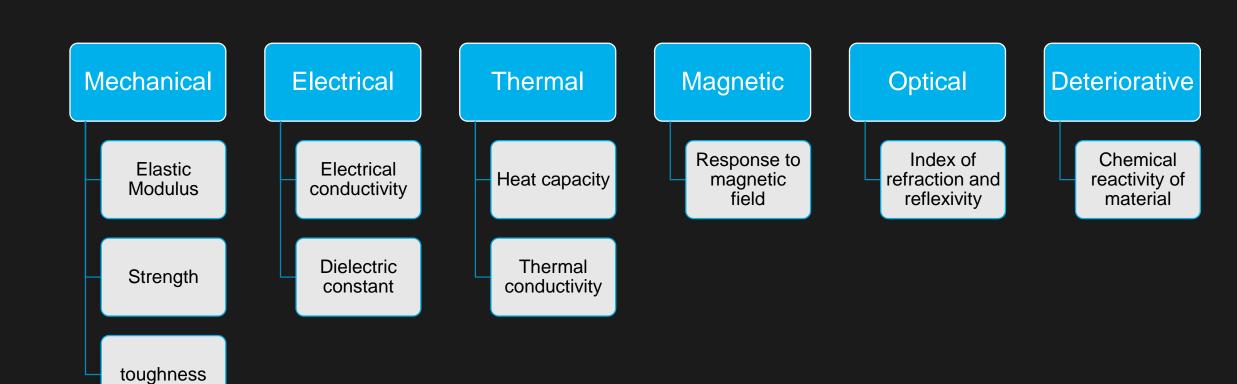
Material Science and Engineering



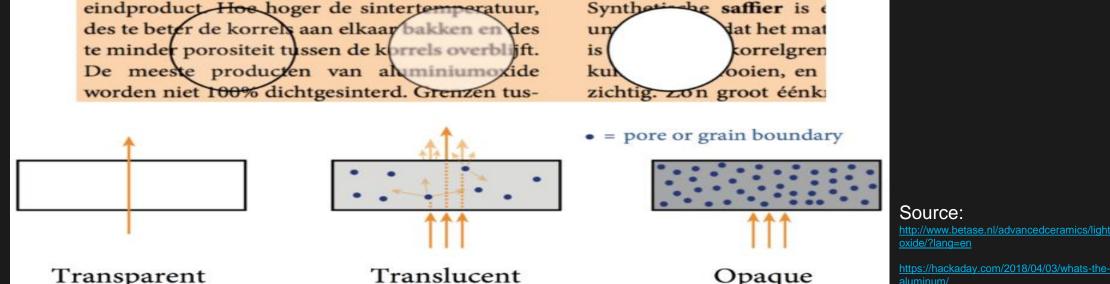
Relationship between components of material science



Properties of materials



Example of inter relation of components of material science



http://www.betase.nl/advancedceramics/light-transmission-of-aluminium-

https://hackaday.com/2018/04/03/whats-the-deal-with-transparent-

The ceramic material aluminium oxide, also known as alumina, is completely opaque, translucent or transparent - depending on the structure, and therefore on the manufacturing process

3 Criteria for Selection of Material

Selection based on in-service conditions and trade off properties.

Deterioration during service conditions.

Economics

Classification of materials



Metallic elements with non metallic elements in relatively small amount.

Stiff, strong, ductile, good conductor of heat and electricity.

Fe, Co & Ni have magnetic properties.



Between Metallic and Non Metallic (such as Oxides , Nitrites & Carbides)

Traditional ceramics – clay, porcelain and glass.

Stiff, strong, hard, insulative to heat and electricity & resistant to high temperature



Plastic and rubber materials. Many are organic compounds based on carbon, hydrogen etc.

Large molecular structure.



Wood and bone are naturally occurring composites.

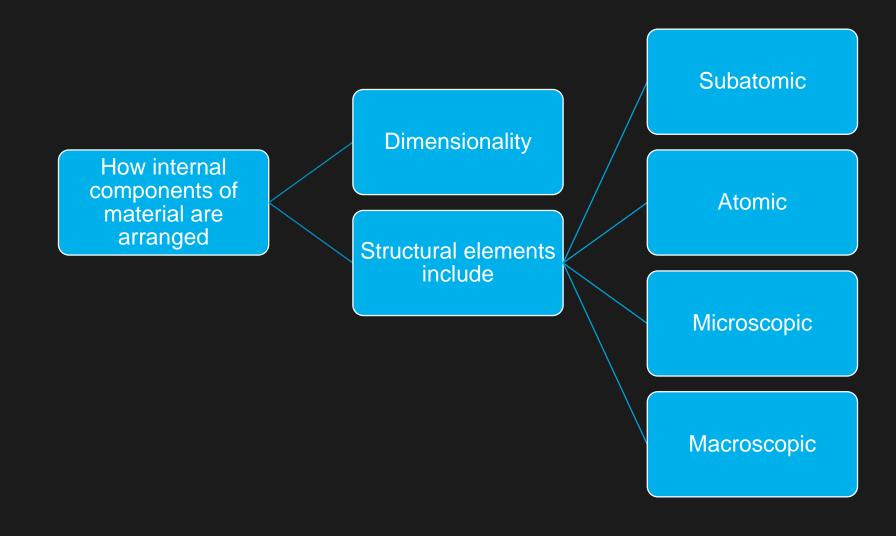
Fiberglass and Carbon fiber reinforced polymer



Semiconductor

Bio Material Smart Material Nano Material

Our Objective: Find relationship between Structure Properties



Our Objective: Find relationship between Atomic Structure Properties

4 Quantum Numbers Participate in bonding **Ground State: All Quantum Mechanics** between atoms to form electrons occupy Model atomic and molecular lowest possible energy aggregates Physical and chemical Valence Electrons: properties of solids is Occupy outermost based on valence shell electrons States within valence electron shell are completely filled Stable Electron Configuration Inert or Noble Gases

Periodic Table

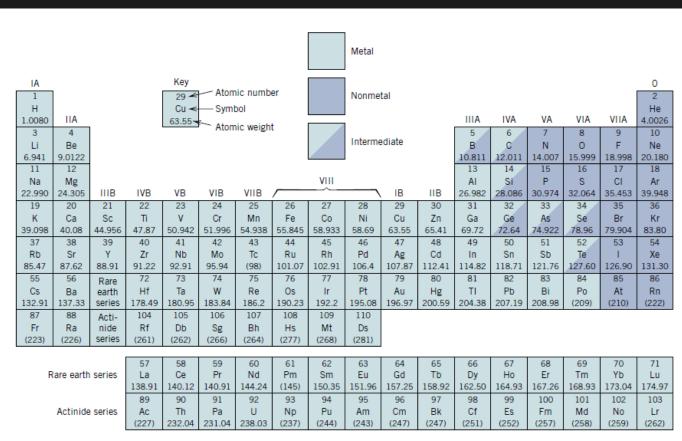


Figure 2.6 The periodic table of the elements. The numbers in parentheses are the atomic weights of the most stable or common isotopes.

- Most Elements come under Metal / Electropositive elements.
 - Capable of giving up few valence electrons to become positively charged.
- Elements on right side of periodic table are electronegative elements.

Variables

Controlled variables

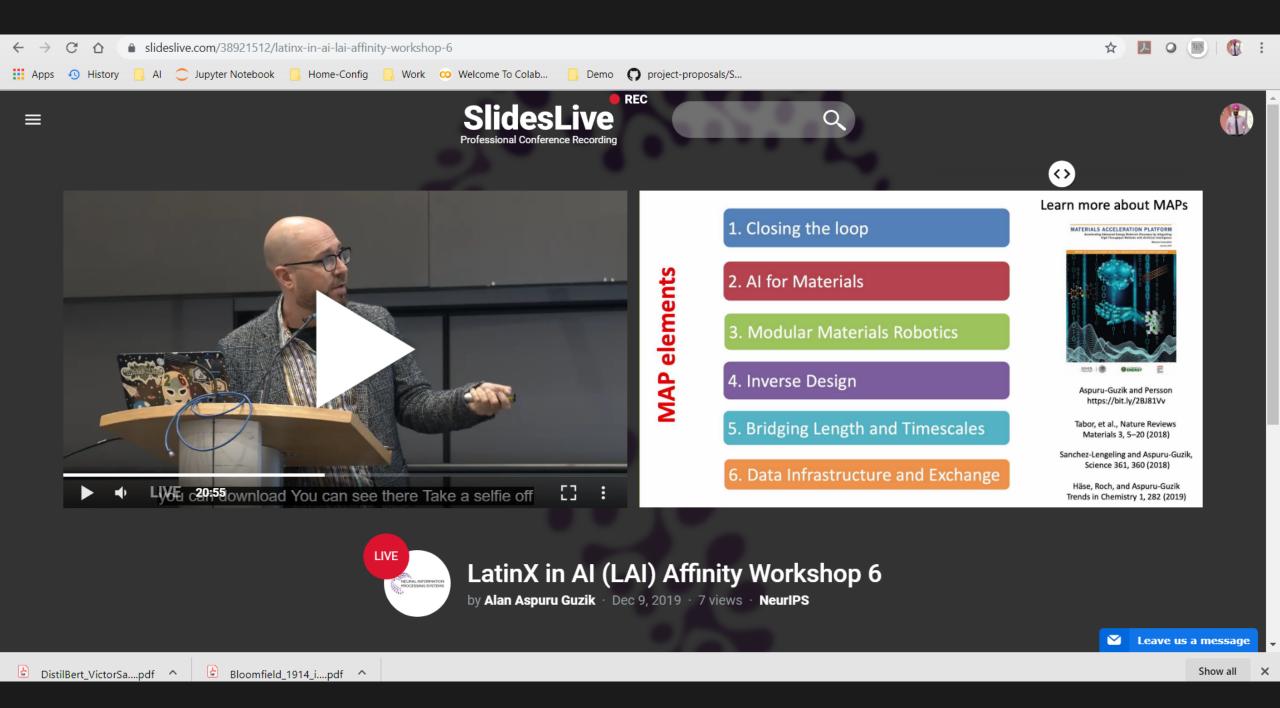
 These are kept the same throughout your experiments

Independent variable

 The one variable you purposely change and test

Dependent variable

- The measure of change observed because of independent variable
- Decide how you will measure the change

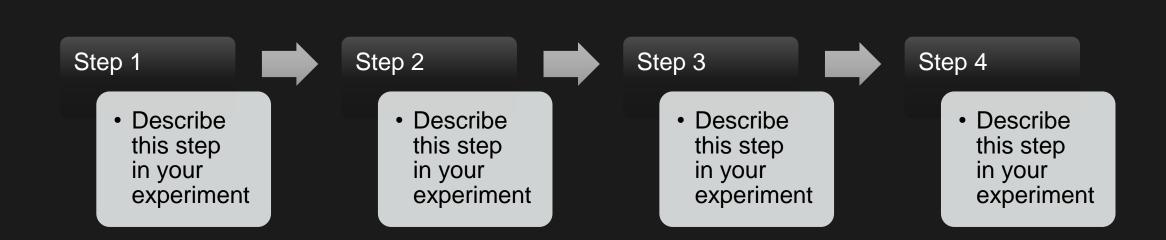


Add your answer / solution here

Materials

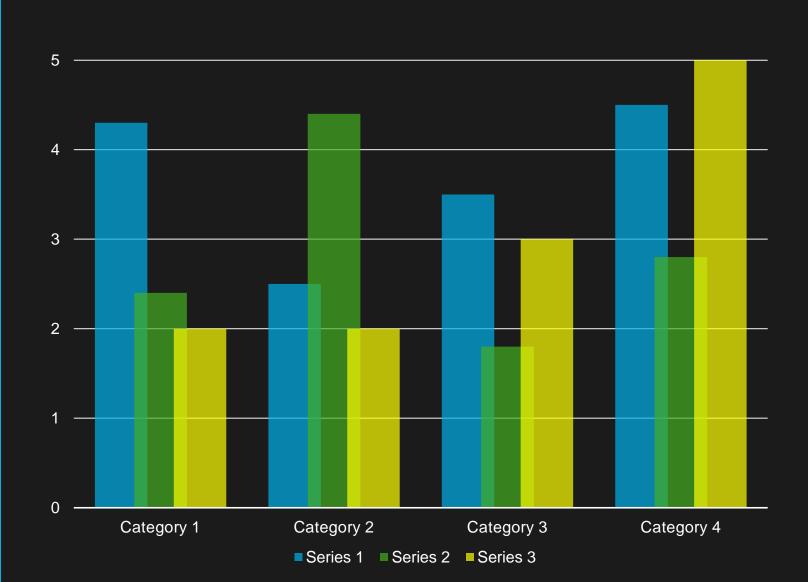
Materials (detailed list)	Quantity (be specific)
Item	Amount

Procedure



Data/Observations

Include Observation based on the data from your experiments



Conclusion

- Brief summary of what you discovered based on results
- Indicate and explain whether or not the data supports your hypothesis

Works Cited

Include print and electronic sources in alphabetical order