

Second Law of Thermodynamics

Heat engine

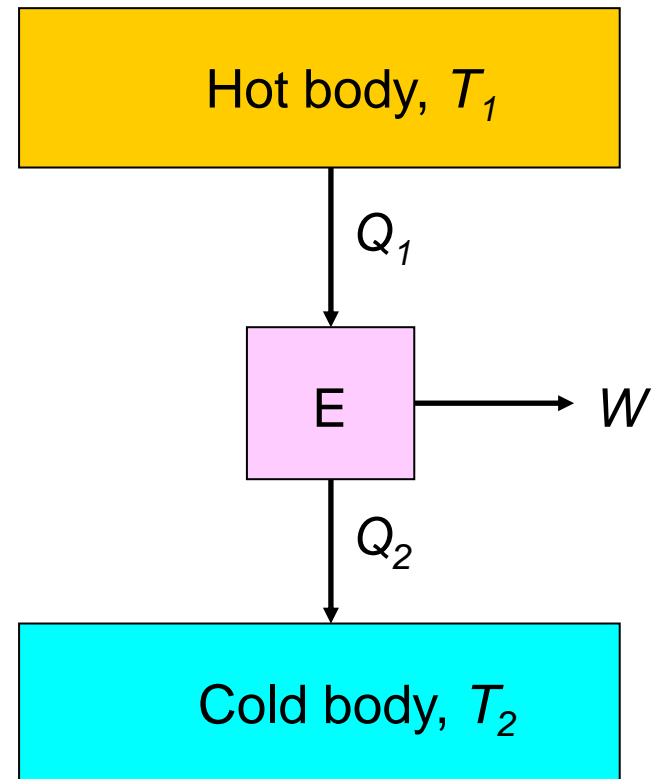
A heat engine converts heat into work in a **cyclic process** in which the working substance is unchanged.

The efficiency, $\zeta = W/Q_1$

In a cyclic process $\Delta U = 0$

First law gives $W = Q_1 - Q_2$

Hence, $\zeta = 1 - Q_2/Q_1$



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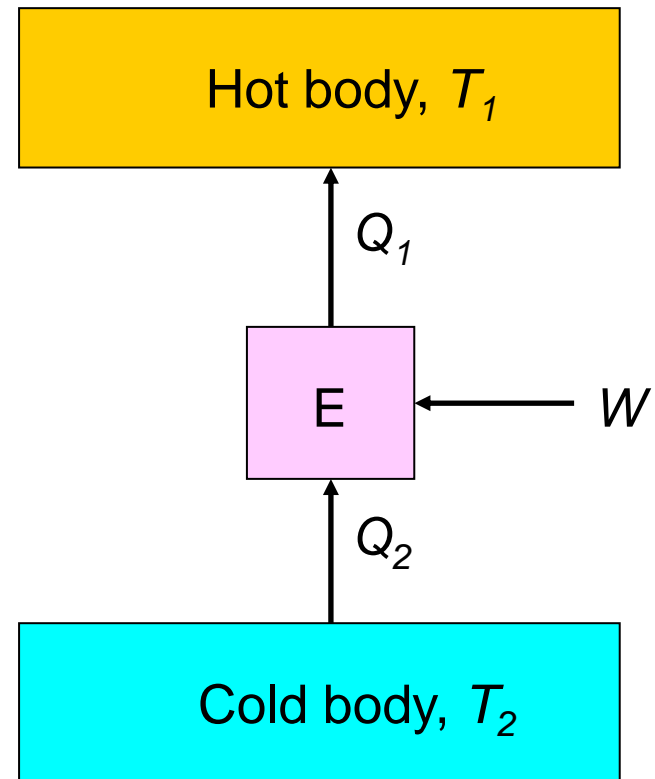
Refrigerator & Heat Pump

A refrigerator extracts heat from a body at lower temperature than surroundings.

$$\zeta^R = \frac{Q_2}{W}$$

A heat pump delivers heat to a body at a higher temperature than surroundings

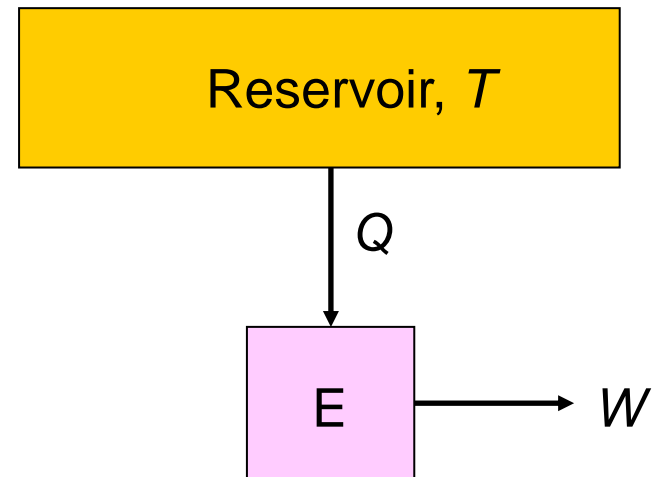
$$\zeta^{HP} = \frac{Q_1}{W}$$



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Kelvin-Planck Statement

It is impossible to construct a device operating in a cycle, which will produce no effect other than the extraction of heat from a single body at a uniform temperature and the performance of an equivalent amount of work.

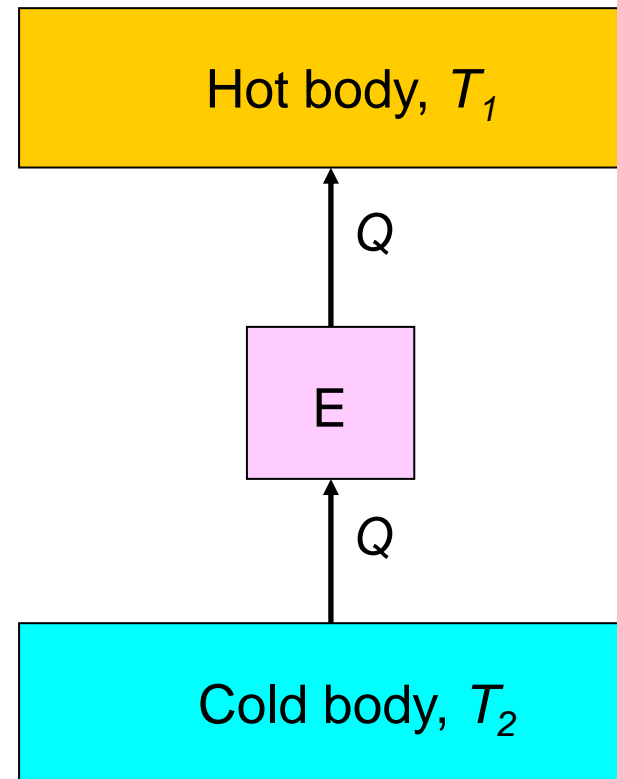


IMPOSSIBLE!

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Clausius Statement

It is impossible to construct a device operating in a cycle, which will produce no effect other than the transfer of heat from a colder to a hotter body.



IMPOSSIBLE!

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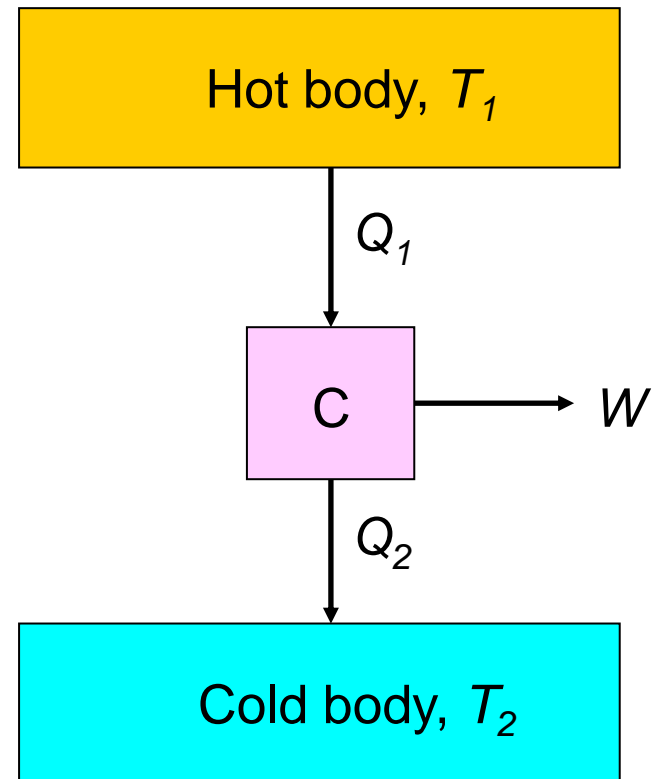
Carnot engine

A Carnot engine, C , is a **reversible** engine operating between only **two temperatures**.

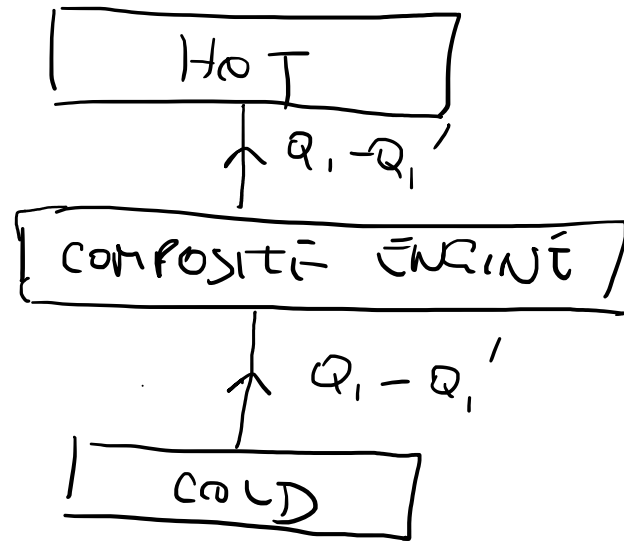
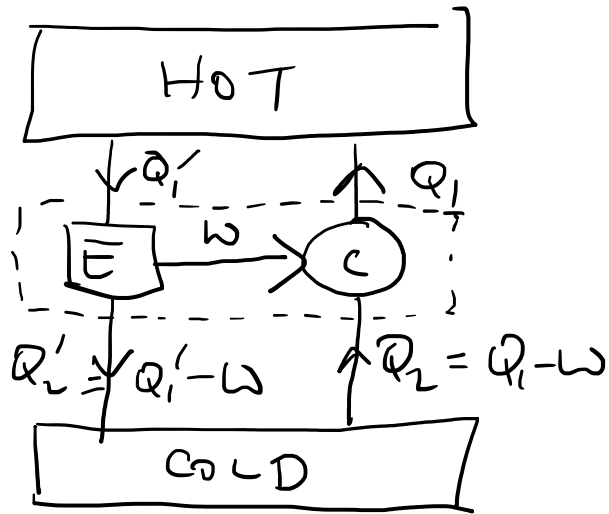
Carnot's Theorem

No engine operating between two reservoirs can be more efficient than a Carnot engine operating between those same two temperatures.

Equivalent to 2nd law



Proof by contradiction. Suppose $\eta^{HE} > \eta^C$



Efficiency of engine $\eta^{HE} = \frac{W}{Q_1'}$; $\eta^C = \frac{W}{Q_1}$

If $\eta^{HE} > \eta^C$, since W same

$$Q_1 > Q_1'$$

VIOLATES CLAUSIUS STATEMENT OF 2ND LAW

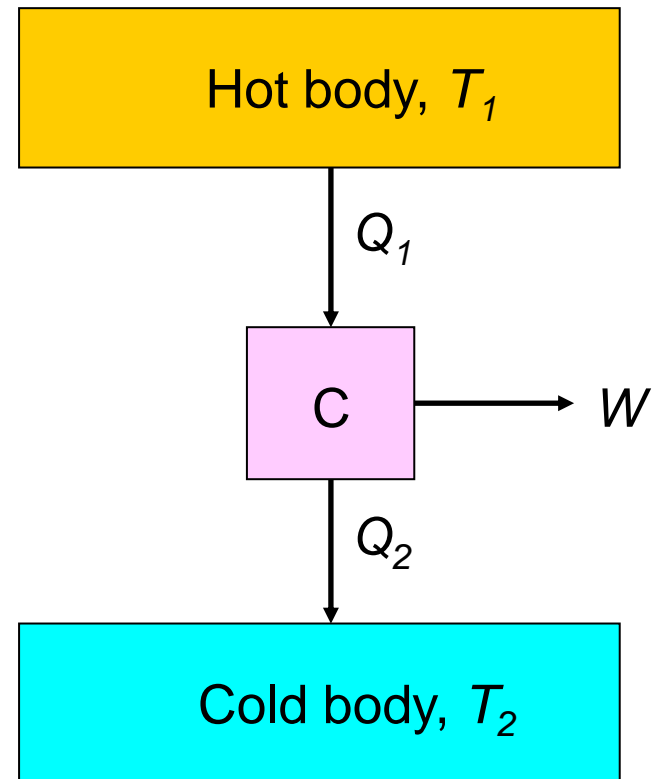
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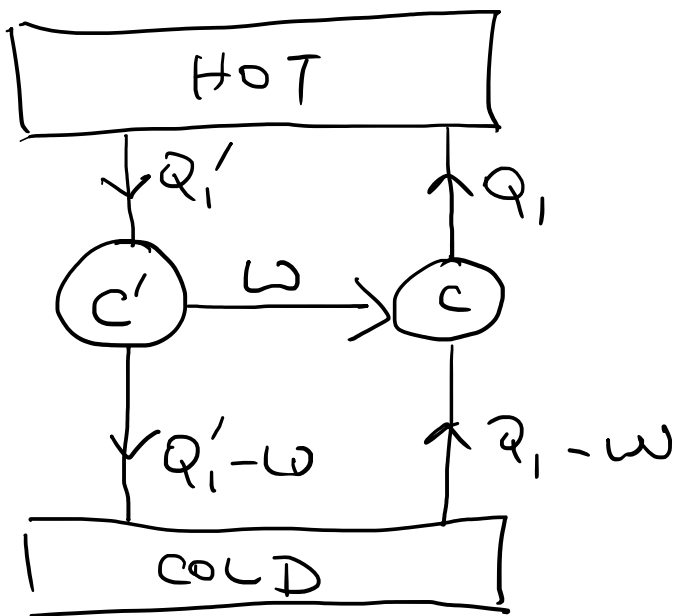
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Corollary to Carnot's Theorem

All Carnot engines operating between the same two reservoirs have the same efficiency.





o o
o o

$$\eta^{c'} \leq \eta^c$$

Run in opposite direction

$$\eta^{c'} \geq \eta^c$$

$$\eta^{c'} = \eta^c$$