

Determination of Pressure Inside of a Beverage Can

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Abstract—This experiment uses knowledge of the two-dimensional principal stress state to determine the internal pressure in an unopened soda can. Thin-walled pressure vessel theory is used in conjunction with theories on principal stresses to report the internal can pressure. This result is compared to another method based upon the ideal gas law. The hoop stress in the can is 78.11 ± 1.128 MPa and the longitudinal stress is 46.27 ± 0.677 MPa. The internal pressure reported by the mechanics of materials approach is 252 ± 43.4 kPa. The internal pressure reported by the ideal gas law method is 258 ± 57.2 kPa.

Index Terms—plane stress, principal stresses, strain gauge rosette, thin-walled pressure vessel

I. INTRODUCTION

CYLINDRICAL pressure vessels are an effective way to store gasses or liquids which must be kept at high pressures. As such, knowing the operating and current pressures in these vessels is imperative to ensuring user safety; catastrophic failures in these vessels can sometimes be fatal. The stresses in the walls of these vessels can be used to determine the overall pressure within the vessel. This study seeks to determine the pressure in a sealed soda can.

The can is a thin-walled pressure vessel. The diameter of the can is much greater than its wall thickness, leading to the assumption that there is no stress in the direction of the thickness of the can, so the stresses in the vessel wall can be analyzed as a two-dimensional stress state, or the plane stress condition.

To analyze this stress state, strain gauges can be affixed in a 0-45-90° rosette on the side of the can to record the strains on the walls. Using multiple strain gauges allows an experimenter to determine the complete strain state, including the shear strain. As the strain gauge is a transducer, the strains are recorded as a voltage difference, ΔV_g , and are calculated according to (1), where V_s is the excitation voltage for the circuit, G_f is the gauge factor of the strain gauge, given by the manufacturer, and A is the gain of the amplifier [1].

$$\epsilon = \frac{4\Delta V_g}{AV_s G_f} \quad (1)$$

These strain gauges work in conjunction with a quarter Wheatstone bridge circuit to record meaningful voltage

differences (Fig. 1). The strain gauge acts as a variable resistor. The change in voltage across the bridge as the strain gauge moves and varies its resistance is the voltage used to compute strain in (1). Each gauge in the strain gauge rosette has its own individual Wheatstone bridge circuit fed by an excitation voltage.

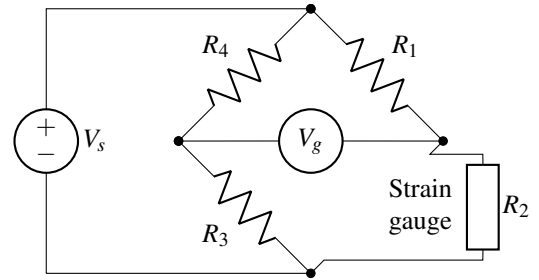


Fig. 1. A simple quarter Wheatstone bridge circuit. Ideally, all resistances are equal.

In plane strain, the shear strain γ_{xy} can be computed using the 45° strain gauge value according to (2), where ϵ_x and ϵ_y are the strains from the 0° and 90° strain gauges, respectively, and θ is 45° for this rosette [2].

$$\epsilon_\theta = \frac{\epsilon_x + \epsilon_y}{2} - \frac{\epsilon_x - \epsilon_y}{2} \cos(2\theta) + \gamma_{xy} \sin(2\theta) \quad (2)$$

While the experimenter should attempt to align the strain gauge rosette along the principal axes, along the can's central axis and along its circumference, the true principal strain should be found according to (3), where $\epsilon_{1,2}$ are the principal strains for the plane strain condition. This is based upon the concepts of Mohr's circle, a way to identify principal stresses and strains in a material geometrically.

$$\epsilon_{1,2} = \frac{\epsilon_x + \epsilon_y}{2} \pm \sqrt{\left(\frac{\epsilon_x - \epsilon_y}{2}\right)^2 + \left(\frac{\gamma_{xy}}{2}\right)^2} \quad (3)$$

Knowing the principal strains, the principal stresses in the vessel can be computed. These principal stresses in the wall arise from the internal pressure pushing on the can wall, so knowing the principal stresses in the wall allows the pressure within the can to be calculated. In 2D strain, Hooke's law relating stress to strain is given directionally as (4a) and (4b),

where σ_1 and σ_2 are the principal stress on the wall, E is the modulus of elasticity of the can material, which is aluminum 3004-H19 [3], ν is Poisson's ratio of the can material, and ϵ_1 and ϵ_2 are the principal strains, found by using (3) with the recorded strains [1].

$$\sigma_1 = \frac{E}{1-\nu^2} (\epsilon_1 + \nu\epsilon_2) \quad (4a)$$

$$\sigma_2 = \frac{E}{1-\nu^2} (\epsilon_2 + \nu\epsilon_1) \quad (4b)$$

For thin-walled pressure vessels, the 2D stress state is described by the longitudinal stress, along the vessel's central axis, and hoop stress, along its circumference, which can be computed with the principal strains above. The free-body diagram (Fig. 2) can be a helpful tool in relating these stresses to the can's internal pressure [1].

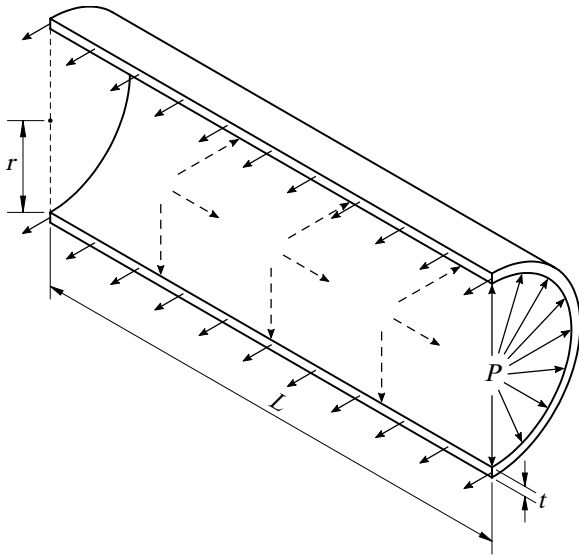


Fig. 2. Free-body diagram of the pressures acting on the pressure vessel internally.

The force balance for this diagram is (5),

$$(2rL)P = (2Lt)\sigma_H \quad (5)$$

as pressure P acts on the can inner surface area, $2rL$, and is resisted by the circumferential hoop stress σ_H acting on wall area $2Lt$. The expression (5) can be solved for the hoop stress σ_H using (6).

$$\sigma_H = \frac{Pr}{t} \quad (6)$$

This hoop stress is related to its orthogonal longitudinal stress σ_L with (7), known from thin-walled pressure vessel theory [1].

$$\sigma_L = \frac{1}{2}\sigma_H \quad (7)$$

These hoop and longitudinal stresses are assumed to be the principal stresses σ_1 and σ_2 . Thus, the relationships from (6) and (4) can be rearranged to yield (8), a functional relationship between principal strain and internal pressure [1].

$$\epsilon_1 = \frac{Pr}{2Et}(2-\nu) \quad (8)$$

Soda is carbonated by being bottled with carbon dioxide (CO_2) under high pressure [4]. Because the soda itself is negligibly compressible in comparison to the gaseous CO_2 , the pressure can be assumed to be caused only by the CO_2 . Therefore, the pressure in the vessel can be determined by measuring the CO_2 that escapes during depressurization. The ideal gas law, (9), can be used to estimate this pressure.

$$PV = nRT \quad (9)$$

where P is pressure, V is volume, n is the number of moles, R is the gas constant, and T is the temperature. Solving for pressure gives (10).

$$P = \frac{nRT}{V} \quad (10)$$

By weighing the can before and after depressurization, the amount of escaped CO_2 , n , can be calculated using (11), where W_{ci} is the initial weight of the can, W_{cf} is the weight of the can after depressurization, and M is the molar mass of CO_2 .

$$n = \frac{W_{ci} - W_{cf}}{M} \quad (11)$$

The volume occupied by the CO_2 can be calculated by measuring the weight of water required to fill that space using (12), where W_{wi} is the initial weight of the water container, W_{wf} is the final weight of the water container, and ρ is the density of water.

$$V = \frac{W_{wi} - W_{wf}}{\rho} \quad (12)$$

Assuming the can is left in a room for a significant amount of time, the temperature of that room can be measured and will be accurate for the can's contents [4].

II. PROCEDURE

A. Supplies Required

The supplies needed for this experiment include a data acquisition system (DAQ) and a soda can. A strain gauge rosette and cyanoacrylate glue is also required. A container with an arbitrary amount of water, packing tape, painter's tape, fine sandpaper, and methanol for cleaning the can surface are needed. A micrometer is needed to measure the diameter of the can and the thickness its material. A scale is needed to measure the weight of the can and the container of water. A thermometer is needed to measure the temperature of the room. Three amplifiers, nine known resistors with the same resistance as the strain gauge rosette, a breadboard, various wires, solder, a soldering iron, and an external power supply are needed to make the Wheatstone bridge circuits. A LabVIEW virtual instrument (VI) is used to record the data. Table I shows the details of the equipment used in this specific lab.

TABLE I
SPECIFIC EQUIPMENT USED IN THIS LAB

Equipment	Specification
DAQ	Out of the Box SADI
350 Ω strain gauge rosette	Omega SGD-3/350-RYT23
3 Amplifiers	Texas Instruments THS4522
9 Resistors	350 Ω

B. Instrumenting the Soda Can

The can must first be outfitted with the strain gauge rosette. First, the decorative wrapping on the beverage can is scrubbed off with fine sandpaper so that this wrapping does not affect the strain response of the can. Once the surface is properly cleaned using methanol, painter's tape is used to create an approximately 90° L-shaped guideline towards the outer left and bottom edges of the cleaned surface for strain gauge placement.

To place the strain gauge rosette, clear packing tape is placed over the top of the rosette, opposite the side that will contact the surface of the can. The tacky side of the taped strain rosette is then gently laid onto the can surface, such that the 0° and 90° strain gauges are aligned with the taped guidelines. When the placement of the strain rosette is properly aligned, part of the taped side is lifted, and glue is applied to the can. The strain gauge rosette is replaced over top of the glue and is left to cure.

C. Wheatstone Bridge

Each strain gauge works as a variable resistor in a Wheatstone bridge, where the voltage difference across the bridge is recorded for the strain (Fig. 1). Three Wheatstone bridge configurations are wired onto a simple breadboard. Each bridge also makes use of an amplifier, as the voltages passing through the circuit are expected to be very small. An excitation voltage V_s of 3.3 V is wired so as to pass across each bridge from the external power supply due to the DAQ being limited to 30 mA [5]. This excitation voltage V_s and the amplified bridge voltage V_{amp} for each of the three bridges is recorded with the DAQ.

D. Data Acquisition

The DAQ which records the voltage difference is fed into the LabVIEW VI. This VI reads the V_{amp} for each bridge as V_{amp_x} for the 0° gauge, V_{amp_θ} for the 45° gauge, and V_{amp_y} for the 90° gauge. These are observed with a ± 5.12 V gain window.

First, the VI is run on the undisturbed instrumented soda can and voltage data is collected for taring purposes. The V_{amp} values are then tared in order to eliminate bias error. The V_{amp} values for each bridge and the excitation voltage V_s are recorded by the VI before, during, and after the can is opened.

E. Ideal Gas Law Pressure Determination Method

A secondary method based upon the ideal gas law is also used. For this method, the temperature of the room is measured and the can is allowed to reach thermodynamic equilibrium with the room. The unopened, instrumented can is weighed on a scale, and its initial weight is recorded. After the can is opened, the instrumented can is weighed again and its final weight is recorded. Next, a bottle filled with an arbitrary amount of water is weighed on the scale. Water is poured from the bottle into the open can until the headspace between the soda and the top of the can is filled. The bottle is weighed again [4].

III. RESULTS

The beverage can that the group originally instrumented during this lab suffered a short in the circuit. Thus, all V_{amp} readings recorded the excitation voltage rather than the voltage differences in the strain rosette. As such, the data used for the instrumented can is supplied by Dr. Ridgeway, and collected by a different student group from the Fall 2021 semester of EML3301C. The type of beverage that this data represents was not reported. The ideal gas law method data was collected in lab by the authors.

A. Can Properties

The can properties measured in lab for use in the ideal gas law method are presented in Table II. All weights were measured by the same digital scale. The ambient temperature of the room was measured by a mercury thermometer and the can was left unrefrigerated in the room long enough to reach ambient temperature.

TABLE II
CAN PROPERTIES FOR IDEAL GAS LAW METHOD

Property	Value
Initial Can Weight	406.8 \pm 0.1 g
Final Can Weight	406.2 \pm 0.1 g
Initial Water Weight	152.0 \pm 0.1 g
Final Water Weight	141.1 \pm 0.1 g
Ambient Temperature	22.78 \pm 0.28°C

The can properties measured by the Fall 2021 student group and used in this report for the mechanics of materials method are presented in Table III. These values were measured four times with a micrometer and a 95% confidence interval was calculated for the uncertainties.

TABLE III
CAN PROPERTIES FOR MECHANICS OF MATERIALS METHOD

Property	Value (mm)
Can Thickness	0.091 \pm 0.016
Can Diameter	66.63 \pm 0.89

B. Recorded Voltages

By examining the data from Fall 2021, it is determined that four voltages were continuously recorded before, during, and after the can was opened. The data is assumed to have been taken at a frequency of 0.5 kHz. This assumption does not impact the conclusions drawn from the data, as only the steady state values before and after the can was opened are considered. The three amplified voltages from the strain rosette show the can being opened just after one second into data collection (Fig. 3). The steady state data before and after can opening are isolated and individually reported. The mean and standard deviation for the three amplified voltages before and after can opening and the excitation voltage, V_s , are presented in Table IV.

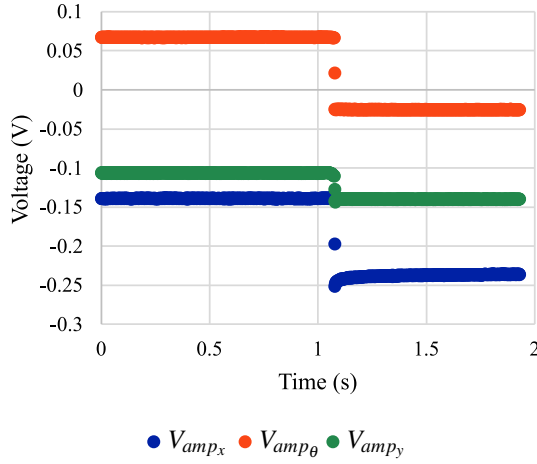


Fig. 3. Captured voltages versus time.

TABLE IV
RECORDED VOLTAGES

Voltage	Initial (V)	Final (V)
V_{amp_x}	-0.1389 ± 0.00056	-0.2367 ± 0.0011
V_{amp_θ}	0.0674 ± 0.00039	-0.0253 ± 0.00027
V_{amp_y}	-0.1064 ± 0.00052	-0.1398 ± 0.00055
V_s	3.349 ± 0.002	

IV. DISCUSSION

A. Limitations

The can instrumented by the group originally suffered a short, indicated by the original recorded data. All V_{amp} values are within tolerance of one another, and read nominally the same as the 3.3 V excitation voltage. This could have occurred for multiple reasons. Firstly, the DAQ used in this experiment may have been damaged, not allowing proper connections with the jumper wires from the breadboard. The jumper wires themselves were also rigid, and the solder used to join the jumper wires had to be bent at awkward angles to properly place the pins in their proper places in the circuit. Also, the wires that were in contact with the strain rosette and near the can were separated by painter's tape; perhaps this was insufficient in insulating the wires from the aluminum can or from each other.

B. Principal Stress State

The rosette strains, ϵ_x , ϵ_y , and ϵ_θ calculated from (1), rosette shear strain, γ_{xy} calculated from (2), principal strains, $\epsilon_{1,2}$ calculated from 3, and the principal stresses, $\sigma_{1,2}$ calculated from (4) are reported in Table V. The principal stresses found from the principal strains can be used to craft a 2D principal stress element (Fig. 4) and 2D Mohr's circle (Fig. 5).

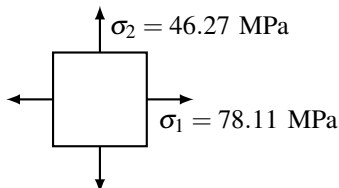


Fig. 4. 2D principal stress element for pressure vessel.

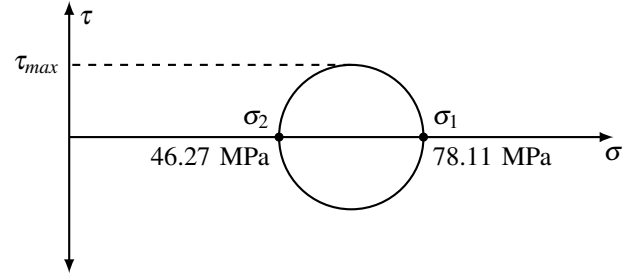


Fig. 5. Principal stresses plotted on Mohr's circle diagram.

These principal stresses are comparable to the longitudinal and hoop stresses in the pressure vessel. Theoretically, the ratio between hoop and longitudinal stress should be 2, so the relationship given by (7) should hold for this pressure vessel. However, this ratio is found to be 1.68 experimentally. This may be in part because the strain gauges are used to measure compressive rather than tensile strains when opening the can. The 90° strain gauge sees less change than the 0° strain gauge does, so the change in ϵ_x is more difficult to record. There may also be losses due to the adhesive used to affix the strain gauges to the can, as this may affect the strain measured in the aluminum.

TABLE V
CALCULATED STRESSES AND STRAINS

Property	Value
ϵ_x	0.000873 ± 0.000014 mm/mm
ϵ_θ	0.000828 ± 0.000010 mm/mm
ϵ_y	0.000299 ± 0.000005 mm/mm
γ_{xy}	0.000242 ± 0.000013 mm/mm
ϵ_1	0.000897 ± 0.000014 mm/mm
ϵ_2	0.000274 ± 0.000006 mm/mm
σ_1	78.11 ± 1.128 MPa
σ_2	46.27 ± 0.677 MPa

C. Pressure from Mechanics of Materials

Three different pressure estimates are calculated under different assumptions about the calculated principal stress state. The first pressure, P_{ϵ_1} , is calculated from the first principal strain, ϵ_1 , by solving (8) for pressure and substituting one half the diameter, d , for radius, r .

$$P_{\epsilon_1} = \frac{4Et\epsilon_1}{d(2-\nu)} \quad (13)$$

This pressure estimates that the associated principal stress state describes the hoop and longitudinal stresses of the pressure vessel. By extension, this pressure estimate also assumes that σ_1 is double σ_2 despite this not being true for the associated principal stress state.

The second pressure, P_{σ_1} , is calculated from the first principal stress by solving (6) for pressure and substituting one half the diameter, d , for radius, r .

$$P_{\sigma_1} = \frac{2t\sigma_1}{d} \quad (14)$$

This pressure estimate assumes that the first principal stress, σ_1 , is representative of the hoop stress σ_H .

The third pressure, P_{σ_2} , is calculated from the second principal stress by substituting (7) into (6), solving for pressure, and substituting one half the diameter, d , for radius, r .

$$P_{\sigma_2} = \frac{4t\sigma_2}{d} \quad (15)$$

This pressure estimate assumes that the second principal stress, σ_2 , is representative of the longitudinal stress σ_L .

The decision on which of these three pressure estimates to use is dependent on the context in which it is used. This report assumes that determining the pressure drop in a pressure vessel would be used in the design of a system to handle this pressure. For this assumption, the highest estimated pressure, P_{σ_2} , is reported as the pressure drop estimate using the mechanics of materials method.

TABLE VI
CALCULATED PRESSURES

Property	Value (kPa)
P_{e1}	205 ± 35.3
P_{σ_1}	213 ± 36.6
P_{σ_2}	252 ± 43.4
P_{IGL}	258 ± 57.2

D. Pressure from Ideal Gas Law

The pressure calculated from the ideal gas law method is $P_{IGL} = 258 \pm 57.15$ kPa. This secondary method was performed on a different beverage can than the mechanics of materials method can, so this was initially to be useful as an order of magnitude comparison rather than an exact validation. However, the pressure result of the second method falls within 10 kPa of the mechanics of materials reported result, allowing for some level of confidence in the pressure results.

E. Comparing Pressure Calculation Methods

The uncertainty in the mechanics of materials method pressure is ± 43.4 kPa, and the uncertainty in the ideal gas method is ± 57.2 kPa, 14 kPa more than the uncertainty in the mechanics of materials method. This may indicate that the mechanics of materials method is more effective than the ideal gas method in determining can pressure. However, these values are near enough to one another that either method could be used to determine can pressure with comparable reliability.

F. Comparison Across Course-wide Data Set

The mean pressure excluding outliers across all lab groups in the Spring 2022 section is 261.32 kPa, with a standard deviation of 86.04 kPa. The calculated pressure for the selected data set is within 10 kPa of this mean value, and well within the standard deviation. The standard deviation is rather large, almost a third of the mean value, so this does not necessarily communicate accuracy, but this does indicate that the computed pressures above are within the proper order of magnitude for the other beverage cans. The mean pressure value for Dr. Pepper is 256.1 kPa and the mean value for Pepsi is 260.8 kPa, but with standard deviations of 88.9 kPa and 57.7 kPa respectively, it is difficult to say with confidence the type of beverage used in the selected data set.

V. CONCLUSION

This study uses the principles of mechanics of materials to determine the pressure in an unopened soda can treated as a thin-walled pressure vessel. Thin-walled pressure vessel theory applied to principal stress theory allows an experimenter to calculate the internal pressure. This result is compared to a secondary method based upon the ideal gas law. The hoop stress in the can is 78.11 ± 1.128 MPa and the longitudinal stress is 46.27 ± 0.677 MPa. The internal pressure reported by the mechanics of materials approach is 252 ± 43.4 kPa. The internal pressure reported by the ideal gas law method is 258 ± 57.2 kPa.

APPENDIX A: UNCERTAINTY CALCULATION

The uncertainty of each variable is given by Table VII. The uncertainties in V_{amp_x} , V_{amp_θ} , V_{amp_y} , and V_s were calculated by finding the standard deviation of a large sample set of these measurements and approximating a 95% confidence interval. The uncertainties in d and t were determined from the measuring device used, which was a micrometer. The uncertainties in the weights of the can and the water were determined from the scale that was used. The uncertainties in E , ν , R , M , and ρ were estimated to be $\pm 1\%$ of the given value due to variability associated with these material properties.

TABLE VII
UNCERTAINTY IN EACH VARIABLE

Variable (Symbol)	Value	Uncertainty
Initial Amplified Voltage x ($V_{amp_{xi}}$)	-0.1389 V	± 0.00056 V
Initial Amplified Voltage θ ($V_{amp_{\theta i}}$)	0.0674 V	± 0.00039 V
Initial Amplified Voltage y ($V_{amp_{yi}}$)	-0.1064 V	± 0.00052 V
Final Amplified Voltage x ($V_{amp_{xf}}$)	-0.2367 V	± 0.0011 V
Final Amplified Voltage θ ($V_{amp_{\theta f}}$)	-0.0253 V	± 0.00027 V
Final Amplified Voltage y ($V_{amp_{yf}}$)	-0.1398 V	± 0.00055 V
Excitation Voltage (V_s)	3.349 V	± 0.002 V
Can Diameter (d)	66.63 mm	± 0.89 mm
Can Thickness (t)	0.091 mm	± 0.016 mm
Modulus of Elasticity (E) [6]	69 GPa	± 0.69 GPa
Poisson's Ratio (ν) [6]	0.35	± 0.0035
Gauge Factor (G_f)	2.09	± 0.01045
Gain (A)	64	± 0.64
Initial Can Weight (W_{ci})	406.8 g	± 0.1 g
Final Can Weight (W_{cf})	406.2 g	± 0.1 g
Initial Water Weight (W_{wi})	152.0 g	± 0.1 g
Final Water Weight (W_{wf})	141.4 g	± 0.1 g
Temperature (T)	22.78°C	$\pm 0.28^\circ\text{C}$
Gas Constant (R) [7]	8.31 J/(K·mol)	± 0.083 J/(K·mol)
CO ₂ Molar Mass (M) [8]	44 g/mol	± 0.44 g/mol
Density of Water (ρ) [9]	0.998 g/cm ³	± 0.00998 g/cm ³

Uncertainty in ΔV_{amp}

The change in V_{amp} for each strain gauge can be calculated using (16), where $V_{amp_{xi}}$, $V_{amp_{\theta i}}$, and $V_{amp_{yi}}$ are the initial voltages and $V_{amp_{xf}}$, $V_{amp_{\theta f}}$, and $V_{amp_{yf}}$ are the final voltages.

$$\Delta V_{amp_x} = V_{amp_{xf}} - V_{amp_{xi}} \quad (16a)$$

$$\Delta V_{amp_\theta} = V_{amp_{\theta f}} - V_{amp_{\theta i}} \quad (16b)$$

$$\Delta V_{amp_y} = V_{amp_{yf}} - V_{amp_{yi}} \quad (16c)$$

The uncertainty in each ΔV_{amp} can be calculated using the root sum square (RSS) method and (16), where U is the uncertainty.

$$U_{\Delta V_{amp}} = \left[U_{V_{amp_i}}^2 \left(\frac{\partial(\Delta V_{amp})}{\partial V_{amp_i}} \right)^2 + U_{V_{amp_f}}^2 \left(\frac{\partial(\Delta V_{amp})}{\partial V_{amp_f}} \right)^2 \right]^{\frac{1}{2}} \quad (17)$$

For V_{amp_x} , using the values given in Table VII give an uncertainty in ΔV_{amp_x} of ± 0.27 V.

Uncertainty in Strain

Strain is given by (1). The uncertainty in strain can be calculated by using (18).

$$U_{\epsilon} = \left[U_{\Delta V_{amp}}^2 \left(\frac{\partial \epsilon}{\partial(\Delta V_{amp})} \right)^2 + U_A^2 \left(\frac{\partial \epsilon}{\partial A} \right)^2 + U_{V_s}^2 \left(\frac{\partial \epsilon}{\partial V_s} \right)^2 + U_{G_f}^2 \left(\frac{\partial \epsilon}{\partial G_f} \right)^2 \right]^{\frac{1}{2}} \quad (18)$$

For ΔV_{amp_x} , using the values given in Table VII give an uncertainty in strain ϵ_x of ± 0.000014 mm/mm.

Uncertainty in Shear Strain

Shear strain for a 0-45-90° rosette is given by (19).

$$\gamma_{xy} = \epsilon_{\theta} - \frac{\epsilon_x + \epsilon_y}{2} \quad (19)$$

The uncertainty in shear strain can be calculated by using (20).

$$U_{\gamma_{xy}} = \left[U_{\epsilon_x}^2 \left(\frac{\partial \gamma_{xy}}{\partial \epsilon_x} \right)^2 + U_{\epsilon_{\theta}}^2 \left(\frac{\partial \gamma_{xy}}{\partial \epsilon_{\theta}} \right)^2 + U_{\epsilon_y}^2 \left(\frac{\partial \gamma_{xy}}{\partial \epsilon_y} \right)^2 \right]^{\frac{1}{2}} \quad (20)$$

Using Table V, the uncertainty in shear strain is ± 0.000013 mm/mm.

Uncertainty in Principle Strain

Principle strain is given by (3). The uncertainty in principle strain can be calculated by using (21).

$$U_{\epsilon_{1,2}} = \left[U_{\epsilon_x}^2 \left(\frac{\partial \epsilon_{1,2}}{\partial \epsilon_x} \right)^2 + U_{\epsilon_y}^2 \left(\frac{\partial \epsilon_{1,2}}{\partial \epsilon_y} \right)^2 + U_{\gamma_{xy}}^2 \left(\frac{\partial \epsilon_{1,2}}{\partial \gamma_{xy}} \right)^2 \right]^{\frac{1}{2}} \quad (21)$$

For ϵ_1 , using Table V gives an uncertainty in principle strain ϵ_1 of ± 0.000014 mm/mm.

Uncertainty in Principle Stress

Principle stress is given by (4). The uncertainty in principle stress can be calculated by using (22).

$$U_{\sigma_{1,2}} = \left[U_E^2 \left(\frac{\partial \sigma_{1,2}}{\partial E} \right)^2 + U_V^2 \left(\frac{\partial \sigma_{1,2}}{\partial V} \right)^2 + U_{\epsilon_1}^2 \left(\frac{\partial \sigma_{1,2}}{\partial \epsilon_1} \right)^2 + U_{\epsilon_2}^2 \left(\frac{\partial \sigma_{1,2}}{\partial \epsilon_2} \right)^2 \right]^{\frac{1}{2}} \quad (22)$$

For σ_1 , using Table V and the values given in Table VII gives an uncertainty in principle stress σ_1 of ± 0.677 MPa.

Uncertainty in Pressure from Stress

Pressure from stress is given by (14) and (15). The uncertainty in pressure can be calculated by using (23).

$$U_P = \left[U_{\sigma_{1,2}}^2 \left(\frac{\partial P}{\partial \sigma_{1,2}} \right)^2 + U_t^2 \left(\frac{\partial P}{\partial t} \right)^2 + U_d^2 \left(\frac{\partial P}{\partial d} \right)^2 \right]^{\frac{1}{2}} \quad (23)$$

For σ_1 , using Table III, Table V, and the values given in Table VII gives an uncertainty in pressure from stress P_{σ_1} of ± 36.6 kPa.

Uncertainty in Amount of Substance

The amount of substance n is given by (11). The uncertainty in the amount of substance can be calculated by using (24).

$$U_n = \left[U_{W_{ci}}^2 \left(\frac{\partial n}{\partial W_{ci}} \right)^2 + U_{W_{cf}}^2 \left(\frac{\partial n}{\partial W_{cf}} \right)^2 + U_M^2 \left(\frac{\partial n}{\partial M} \right)^2 \right]^{\frac{1}{2}} \quad (24)$$

Using Table II and the values given in Table VII gives an uncertainty in the amount of substance of ± 0.0032 mol.

Uncertainty in Volume

Volume is given by (12). The uncertainty in volume can be calculated by using (25).

$$U_V = \left[U_{W_{wi}}^2 \left(\frac{\partial V}{\partial W_{wi}} \right)^2 + U_{W_{wf}}^2 \left(\frac{\partial V}{\partial W_{wf}} \right)^2 + U_{\rho}^2 \left(\frac{\partial V}{\partial \rho} \right)^2 \right]^{\frac{1}{2}} \quad (25)$$

Using Table II and the values given in Table VII gives an uncertainty in volume of ± 0.178 cm³.

Uncertainty in Pressure from Ideal Gas Law

Pressure from the ideal gas law is given by (10). The uncertainty in pressure can be calculated by using (26).

$$U_{P_{GL}} = \left[U_n^2 \left(\frac{\partial P_{GL}}{\partial n} \right)^2 + U_R^2 \left(\frac{\partial P_{GL}}{\partial R} \right)^2 + U_T^2 \left(\frac{\partial P_{GL}}{\partial T} \right)^2 + U_V^2 \left(\frac{\partial P_{GL}}{\partial V} \right)^2 \right]^{\frac{1}{2}} \quad (26)$$

Using the values given in Table VII, and the uncertainties calculated in the previous sections gives an uncertainty in pressure from the ideal gas law of ± 57.2 kPa.

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