

LANGUAGE OF CHEMISTRY

Q.1 Name the following :

1. The substances that reacts with each other.
2. The insoluble solid substance formed that settles to the bottom of the solution.
3. A substance that alters the rate of a chemical reaction without itself undergoing any chemical change.
4. The reactions that absorb energy are called as

Q.2 Write a word equation for the following reactions.

- a) Zinc metal reacts with dilute hydrochloric acid to form zinc chloride and liberate hydrogen gas
- b) Iron reacts with hydrochloric acid to give iron chloride and hydrogen gas.
- c) Mercuric oxide on strong heating decomposes to form mercury and oxygen.
- d) Calcium carbonate on reaction with nitric acid given calcium nitrate, water and carbon dioxide gas.

Q.3 Write your observations for the following chemical reactions and name the products formed.

- a) When zinc granules reacts with dilute hydrochloric acid.
- b) When calcium carbonate reacts with nitric acid
- c) When iron reacts with hydrochloric acid
- d) When ammonia gas reacts with hydrogen chloride

Q. 4 Answer the following questions :

- 1) What are characteristics of a chemical reaction?
- 2) Define endothermic reaction and give one example.
- 3) Explain the change of colour occurring in a chemical reaction with the help of an example.
- 4) State the four conditions necessary for chemical reactions to take place.

Q.5 Give reasons :

1. A characteristic change in colour is seen when a piece of iron is added to copper sulphate solution.
2. Electrolytic decomposition of water is a reaction involving change of state.

Q.6 Fill in the blanks

1. The substances which undergo chemical change are called _____
2. The substances formed as a result of a chemical reaction are called _____
3. During a chemical reaction transfer of _____ takes place.
4. The basic conditions necessary for a chemical reaction is _____.

Q 7. Complete the following word equations:

1. Magnesium + _____ \rightarrow Magnesium Oxide + heat
2. Carbon dioxide + water \rightarrow _____ + Oxygen
3. Mercuric Oxide (heating) \rightarrow _____ + _____
4. Copper carbonate (heat) \rightarrow + _____ + _____