

# The formation of Mg-orthocarbonate through the reaction $\text{MgCO}_3 + \text{MgO} = \text{Mg}_2\text{CO}_4$ at Earth's lower mantle $P$ – $T$ conditions

Pavel N. Gavryushkin <sup>\*1,2</sup>, Dinara N. Sagatova <sup>1,2</sup>, Nursultan Sagatov <sup>1</sup>, and Konstantin D. Litasov <sup>3</sup>

<sup>1</sup>*Sobolev Institute of Geology and Mineralogy, Siberian Branch of Russian Academy of Sciences, prosp. acad. Koptiyuga 3, 630090 Novosibirsk, Russia*

<sup>2</sup>*Novosibirsk State University, Pirogova 2, Novosibirsk 630090, Russia*

<sup>3</sup>*Vereshchagin Institute for High Pressure Physics RAS, 108840, Troitsk, Moscow, Russian Federation*

## Abstract

Orthocarbonates of alkaline-earth metals are the newly discovered class of compounds stabilized at high pressures. Mg-orthocarbonates are the potential carbon host phases, transferring oxidized carbon in the Earth's lower mantle up to the core-mantle boundary. Here, we demonstrate the possibility for the formation of  $\text{Mg}_2\text{CO}_4$  in the lower mantle at pressures above 50 GPa, by *ab initio* calculations.  $\text{Mg}_2\text{CO}_4$  is formed by the reaction  $\text{MgCO}_3 + \text{MgO} = \text{Mg}_2\text{CO}_4$ , proceeding only at high-temperatures. At 50 GPa the reaction starts at 2200 K. The temperature decreases with pressure and drops down to 1085 K at the pressure of the Earth's core-mantle boundary, near 140 GPa. Two stable structures,  $\text{Mg}_2\text{CO}_4$ - $Pnma$  and  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$ , were revealed.  $\text{Mg}_2\text{CO}_4$ - $Pnma$  is isostructural to forsterite ( $\text{Mg}_2\text{SiO}_4$ ), while  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  is isostructural to larnite ( $\beta$ - $\text{Ca}_2\text{SiO}_4$ ). Transition pressure from  $\text{Mg}_2\text{CO}_4$ - $Pnma$  to  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  is around 80 GPa. Both phases are dynamically stable on decompression up to the ambient pressure and temperature. This assumes the possibility for their finding in natural samples of high-pressure rocks.  $\text{Mg}_2\text{CO}_4$ - $Pnma$  has a melting temperature more than 16% higher than the melting temperature of magnesite ( $\text{MgCO}_3$ ). At 23.7 GPa and 35.5 GPa,  $\text{Mg}_2\text{CO}_4$ - $Pnma$  melts at 2742 K and 2819 K, respectively. Acoustic wave velocities  $V_p$  and  $V_s$  of  $\text{Mg}_2\text{CO}_4$ - $Pnma$  are very similar to that of

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\*Electronic address: [gavryushkin@igm.nsc.ru](mailto:gavryushkin@igm.nsc.ru), [p.gavryushkin@g.nsu.ru](mailto:p.gavryushkin@g.nsu.ru); Corresponding author

magnesite, while universal anisotropy of  $\text{Mg}_2\text{CO}_4\text{-}Pnma$  is stronger than that of magnesite, as well as coefficient  $A^U$  is larger for orthocarboante .

## Introduction

During the recent decade, the crystal structure prediction technique became an integral part of high-pressure research. Numerous experimental synthesis were guided by this technique, for instance, the synthesis of the high-pressure phases of alkaline carbonates  $\text{Li}_2\text{CO}_3$ ,  $\text{Na}_2\text{CO}_3$ , and  $\text{K}_2\text{CO}_3$  [1–3] and alkaline-earth carbonates  $\text{CaCO}_3$ ,  $\text{MgCO}_3$ , and  $\text{CaMg}(\text{CO}_3)_2$  [4–10].

Orthocarbonates are another example for experimental confirmation of theoretically predicted structures. The stability of the alkaline-earth orthocarbonates has not been considered till the last year. The performed crystal structure prediction in the  $\text{MgO} - \text{CO}_2$  and  $\text{CaO} - \text{CO}_2$  systems have revealed  $sp^3$ -bonded structures with the intermediate stoichiometries,  $\text{Ca}_3\text{CO}_5$  and  $\text{CaC}_2\text{O}_5$ , stabilized at high pressures [11]. Thermodynamically stable structures of orthocarbonates with the  $\text{Ca}_2\text{CO}_4$  or  $\text{Mg}_2\text{CO}_4$  stoichiometry have not been revealed in these calculations. Assuming the stochastic nature of the used methods of crystal structure prediction and possibility that some favorable structures can be missed in the calculation, we have performed a thorough search using both evolutionary algorithms (USPEX) and *ab initio* random methods (AIRSS) within the stoichiometry of orthocarbonates  $\text{M}_2\text{CO}_4$ , where M is alkaline-earth metal, Mg, Ca, Sr, or Ba. As the result of this investigation, the structure  $\text{Ca}_2\text{CO}_4\text{-}Pnma$  was obtained [12]. This structure lies above the energetic convex hull in the  $\text{CaO} - \text{CO}_2$  system, i.e. it is stable relative to the decomposition on the known phases of this system, in particular to the mixture of  $\text{CaO} + \text{CaCO}_3$ . Further theoretical search has also shown that the similar structures of Sr and Ba orthocarbonates,  $\text{Sr}_2\text{CO}_4\text{-}Pnma$  and  $\text{Ba}_2\text{CO}_4\text{-}Pnma$ , are stable at relatively low pressures [13]. Recent experimental synthesis combined with single-crystal X-ray diffraction method have unambiguously confirmed the stability of the predicted  $\text{Ca}_2\text{CO}_4\text{-}Pnma$  and  $\text{Sr}_2\text{CO}_4\text{-}Pnma$  structures [?, ?].

Undoubtedly, the most interesting orthocarbonate for the Earth sciences is the orthocarbonate of Mg. This compound can be readily formed within the slab subducted in the Earth's lower mantle from the locally abundant  $\text{MgCO}_3$  and  $\text{MgO}$  by the simple reaction  $\text{MgCO}_3 + \text{MgO} = \text{Mg}_2\text{CO}_4$ . The formation of  $\text{Mg}_2\text{CO}_4$  with spinel structure have been speculatively suggested by Fyfe back in 1970 [14]. Then formation of  $\text{Mg}_2\text{CO}_4$  have been hypothesized by Irwing AJ and Wyllie PJ [15], and T. Katsura [16]. However, as no experiments [15], no calculations [?, 11] have revealed any intermediate phases in the system  $\text{MgO-MgCO}_3$ , the possibility for the formation of Mg-orthocarboantes at high pressures is unclear.

In the present study, using crystal structure prediction techniques, we revealed new structures of Mg-orthocarboante,  $\text{Mg}_2\text{CO}_4$ , stable against decomposition on  $\text{MgCO}_3$  and  $\text{MgO}$ , having chances to be formed at the Earth's lower mantle  $P$ – $T$  conditions.

## Methods

To increase the chances for finding the energetically favorable structures we have used both USPEX and AIRSS methods, each of which has apparent advantages [17]. *Ab initio* structure prediction was complemented by the prediction technique based on the known structures of Mg- and Zn- orthosilicates with isolated  $[\text{SiO}_4]$  tetrahedrons, as well as sulfates with isolated  $[\text{SO}_4]$  tetrahedrons.  $\text{Mg}_2\text{CO}_4$  structures were produced from these structures by the corresponding replacement of the cations and consequent local optimizations. The following crystal structures have been used (numbers in parentheses indicate the number of formula units (f.u.) in the unit cell): ringwoodite ( $\text{Mg}_2\text{SiO}_4$ - $Fd\bar{3}m$  (8)), porierite ( $\text{Mg}_2\text{SiO}_4$ - $Pmma$  (4)), wadsleyite ( $\text{Mg}_2\text{SiO}_4$ - $Imma$  (8)), forsterite ( $\text{Mg}_2\text{SiO}_4$ - $Pnma$  (4)) [18]  $\text{Zn}_2\text{SiO}_4$ - $Pbca$  (8),  $\text{Zn}_2\text{SiO}_4$ - $I\bar{4}2d$  (4) [19], ,  $\text{Na}_2\text{SO}_4$ - $Fddd$  (8) [20],  $\text{Li}_2\text{SO}_4$ - $P2_1/c$  (4) [21], and  $\text{Ca}_2\text{CO}_4$ - $Pnma$  (4) [12].

The calculations with the USPEX method [22–25] have been performed for 1-4 f.u. of  $\text{Mg}_2\text{CO}_4$  at 25, 50, and 100 GPa. The seeding technique, implemented in version 10.2 of the USPEX, has been employed in all calculations. The aforementioned structures based on silicates, sulfates, and orthocarbonate have been used as the seeds. Totally, around 3000 structures have been calculated at each pressure. Crystal structure prediction with AIRSS 0.9.1 [26, 27] has been performed only at 50 GPa for 2, 3, and 4 f.u. and a total of nearly 4000 structures have been generated in this calculation.

The energetic optimizations of the predicted structures have been performed within density functional theory (DFT), implemented in the VASP package [28, 29].

To take into account the temperature effect and calculate  $P$ - $T$  phase diagram, we used the lattice dynamics method within the quasi-harmonic approximation (QHA). For this task, the lattice vibration frequencies were calculated with the PHONOPY code [30]. By our experience with  $P$ - $T$  phase diagrams of carbonates, this technique reliably reproduces phase boundaries in the wide temperature range [2, 12, 31], up to 80-90 % of the melting temperature, if the process of dynamical disordering does not take place.

Melting temperatures of the predicted structures have been determined, using the so-called *Z-method*, based on the molecular dynamic (MD) simulations [32].

To determine the wave velocities  $V_p$  and  $V_s$  and assess anisotropy of Mg-orthocarbonates, static elastic stiffness tensor ( $C_{ij}$ ) was calculated from the stress ( $\sigma$ )- strain ( $\epsilon$ ) relation  $\sigma_i = C_{ij}\epsilon_j$ . Based on these  $C_{ij}$  data, we calculated averages of bulk (B) and shear (G) moduli using the Voigt-Ruess-Hill scheme [33, 34]. Then we have determined compression ( $A_B$ ), shear ( $A_G$ ), and universal anisotropy ( $A^U$ ) indexes.

The details of DFT, crystal structure predictions, thermodynamic and elastic property calculations, and MD simulations are given in *Supporting information*.

FindSym program [35] and instruments of Phonopy package have been used for the symmetry determination. VESTA and ToposPro [36, 37] programs have

been used for the visualization of the crystal structures and figures preparation. The topology of the structures was analysed with ToposPro and Robocrystallographer programs [37, 38].

## Results and discussion

### Predicted structures

In the crystal structure prediction calculations at 25 and 50 GPa, USPEX has revealed the  $\text{Mg}_2\text{CO}_4$ - $Pnma$  structure as the most energetically favorable. This structure has lower enthalpy than  $\text{Mg}_2\text{CO}_4$ - $Cm$ , predicted by AIRSS, and structures, obtained based on the crystal structures of silicates and sulfates (Figure 1 and ??).

At 100 GPa, USPEX has found another favorable structure with monoclinic symmetry,  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$ . The transition from  $\text{Mg}_2\text{CO}_4$ - $Pnma$  to  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  occurs at 52 GPa, and at higher pressures  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  is the most energetically favorable among considered structures (Figure 1 and ??).

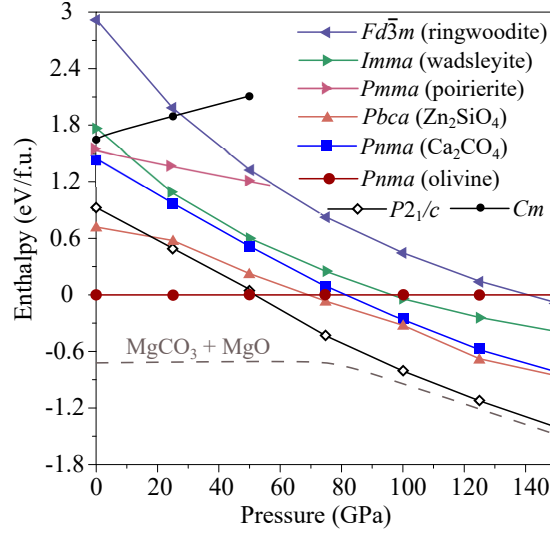


Figure 1: Enthalpy-pressure dependencies of  $\text{Mg}_2\text{CO}_4$  structures. Isostructural compounds are indicated in brackets.

Revealed  $\text{Mg}_2\text{CO}_4$ - $Pnma$  is isostructural to the mineral forsterite (Figure 2a,b). Atoms in both positions Mg1 and Mg2 are 6-coordinated, with the octahedral coordination polyhedron (Figure 2c).  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  is isostructural to  $\beta$ - $\text{Ca}_2\text{SiO}_4$  (larnite) and can be considered as the monoclinic analogue of  $\text{Ca}_2\text{CO}_4$ - $Pnma$ . In crystal structure of  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$ , there are two non-equivalent Mg sites. There is some ambiguity in determination of coordination numbers of Mg, due to the smooth variation of the bonds lengths. In the first

site Mg(1) is bonded to six oxygen atoms, arranged in highly deformed trigonal prism (Figure 2d). Lengths of the six Mg–O bonds within coordination polyhedron vary in the range 1.832–2.073 Å, while the seventh oxygen atoms is distant on 2.382 Å. In the second site, Mg(2) is bonded to eight oxygen atoms, forming relatively regular square antiprism. Lengths of these eight Mg–O bonds vary in the range of 1.887–2.126 Å, while the ninth Mg–O bond is sufficiently longer, being equal to 2.688 Å. In  $\text{Mg}_2\text{CO}_4$ - $Pnma$  structure, oxygen atoms in the first coordination sphere is distant on 1.815–1.888 Å for Mg(1) site and on 1.781–1.95 Å — for (Mg2) site. All the presented bond lengths correspond to the pressure of 100 GPa. In both  $Pnma$  and  $P2_1/c$  crystal structures, coordination polyhedrons around Mg atoms share both vertices and edges with  $[\text{CO}_4]$  tetrahedrons (Figure 2c,d). Transition from  $\text{Mg}_2\text{CO}_4$ - $Pnma$  to  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  is accompanied by the sufficient decrease of the volume equal to 5.7 % (Figure ??).

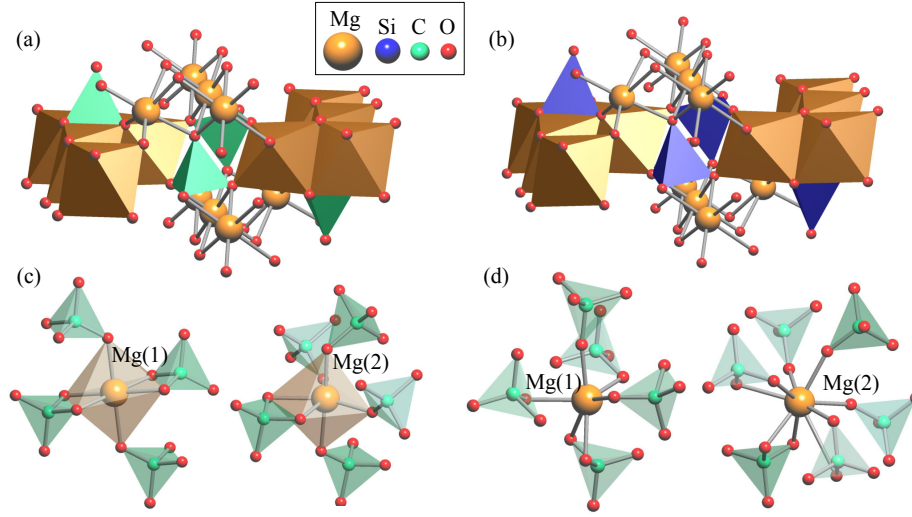


Figure 2: Comparison of crystal structures of  $\text{Mg}_2\text{CO}_4$ - $Pnma$  (a) and  $\text{Mg}_2\text{SiO}_4$ - $Pnma$  (b) in projection on the (010) plane, coordination environment around Mg(1) and Mg(2) sites in crystals structures  $\text{Mg}_2\text{CO}_4$ - $Pnma$  (c) and  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  (d).

Structural data for the  $\text{Mg}_2\text{CO}_4$ - $Pnma$  and  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  are given in the Table 1. Both structures are dynamically stable in the investigated pressure range 0-140 GPa, and there are no imaginary frequencies in their phonon dispersion curves (Figure 3, ??). The dynamical stability of the found  $\text{Mg}_2\text{CO}_4$  structures at ambient pressure assume the possibility for their extraction from the chamber of the diamond anvil cell (DAC) after the high pressure synthesis and the further investigation with the technique of transmission electron microscopy. In the *Supporting information*, we have also shown phonon dispersion

curves of  $\text{MgCO}_3$  and  $\text{MgO}$  (Figure ??) used for the construction of  $P$ - $T$  phase diagram.

Table 1: Structural data of predicted  $\text{Mg}_2\text{CO}_4$  phases at 0 K.

$P$ (GPa)	Space group	Lattice parameters ( $\text{\AA}$ , deg)			Atom	Coordinates		
						x	y	z
50	$Pnma$ (#62)	$a = 8.926$ $\alpha = 90.0$	$b = 5.565$ $\beta = 90.0$	$c = 4.221$ $\gamma = 90.0$	Mg1	0.000	0.000	0.500
					Mg2	0.721	0.250	0.531
					C1	-0.097	0.250	0.087
					O1	-0.091	0.250	0.770
					O2	0.548	0.250	0.284
					O3	0.169	0.552	0.786
100	$P2_1/c$ (#14)	$a = 4.408$ $\alpha = 90.0$	$b = 5.383$ $\beta = 117.65$	$c = 8.345$ $\gamma = 90.0$	Mg1	0.702	0.360	0.425
					Mg2	-0.022	0.000	0.693
					C1	0.355	0.282	0.082
					O1	0.146	0.334	0.638
					O2	0.681	0.245	0.197
					O3	0.272	0.168	-0.080
					O4	0.295	0.520	0.064

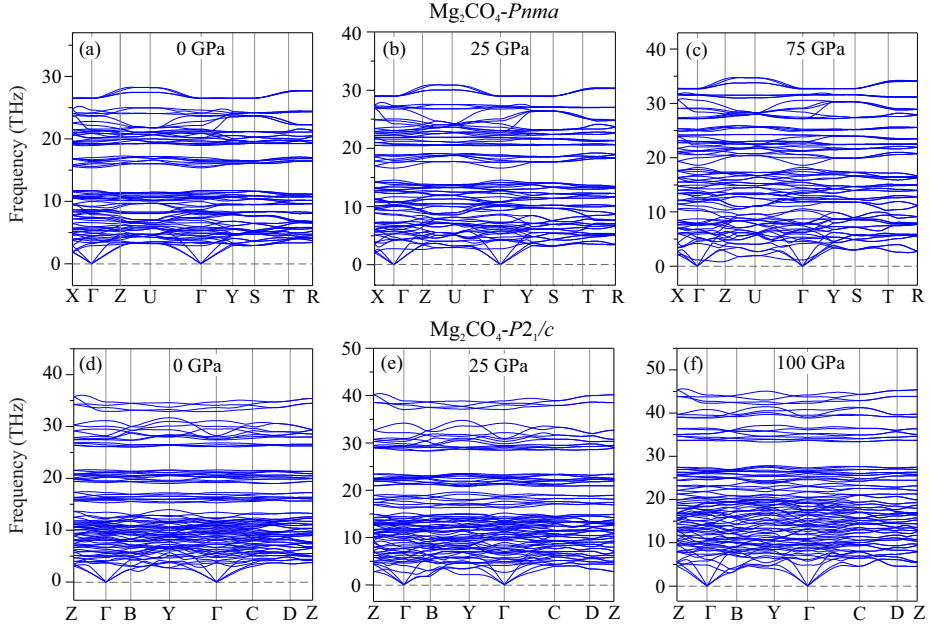


Figure 3: Phonon dispersion curves of  $\text{Mg}_2\text{CO}_4$ - $Pnma$  (a-c) and  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  (d-f) at various pressures and 0 K.

There is no analogy of the found transition from  $\text{Mg}_2\text{CO}_4$ - $Pnma$  to  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  in the  $\text{Mg}_2\text{SiO}_4$  system. However, there is one in  $\text{Ca}_2\text{SiO}_4$  system, where transformation of  $\text{Mg}_2\text{CO}_4$ - $Pnma$  into  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  corresponds to the transformation of  $\gamma$ - $\text{Ca}_2\text{SiO}_4$  into  $\beta$ - $\text{Ca}_2\text{SiO}_4$ , which is also realized on compression [39]. Heating of  $\beta$ - $\text{Ca}_2\text{SiO}_4$  produces the phase  $\alpha'_H$ - $\text{Ca}_2\text{SiO}_4$  and similar transition can be suggested for  $\text{Mg}_2\text{CO}_4$ . In this case, heating of  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  will produce the new hypothetical phase  $\text{Mg}_2\text{CO}_4$ - $Pnma$ -II similar to  $\alpha'_H$ - $\text{Ca}_2\text{SiO}_4$ . The performed calculations of phonon dispersion curves of  $\text{Mg}_2\text{CO}_4$ - $Pnma$ -II, constructed based on the structure of  $\alpha'_H$ - $\text{Ca}_2\text{SiO}_4$  by corresponding atomic replacements, have shown the dynamical instability of this phase (Figure ??). However, the stabilization of the structure by the factors, which are not considered within QHA can not be excluded.

### P–T phase diagram

The performed enthalpy calculations of the  $\text{Mg}_2\text{CO}_4$ ,  $\text{MgO}$ , and  $\text{MgCO}_3$  structures have shown, that both  $\text{Mg}_2\text{CO}_4$ - $Pnma$  and  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  lie above the energetic convex hull, i.e. they are unstable and decompose to the mixture ( $\text{MgO}+\text{MgCO}_3$ ) at 0 K (Figure ??).

However, performed calculations of the Gibbs free energies in the temperature range of 0–3000 K have shown that above some temperature  $\text{Mg}_2\text{CO}_4$  became more energetically favorable than the ( $\text{MgO}+\text{MgCO}_3$ ) mixture (Figure ??). At 20 GPa, this temperature is 2420 K, which is nearly equal to the melting temperature of magnesite ( $\text{MgCO}_3$ - $R\bar{3}c$ ) [40] (Figure 4). With increasing the pressure the temperature of transition decreases, and at 140 GPa  $\text{Mg}_2\text{CO}_4$ - $P2_1/c$  became more favourable than the mixture of ( $\text{MgO}+\text{MgCO}_3$ ) at temperatures higher than 1085 K. Thus, in the pressure range of 50–140 GPa,  $\text{Mg}_2\text{CO}_4$  can be synthesised at temperatures 250–1100 K higher than the corresponding temperatures of  $\text{MgCO}_3$  melting (Figure 4). It suggests the possibility for the formation of Mg-orthocarbonate in most part of the Earth's lower mantle, at pressures higher than 50 GPa.

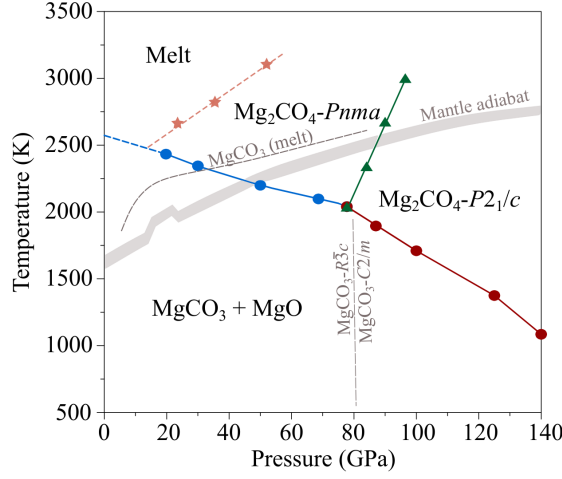


Figure 4:  $P$ - $T$  phase diagram of  $\text{Mg}_2\text{CO}_4$ . The asterisks represent calculated melting temperatures; the grey dash-dotted line — calculated phase transition boundary of  $\text{MgCO}_3$ ; the grey dashed line — melting curve of  $\text{MgCO}_3$  as reported by Solopova et al. [40]; grey solid line — mantle adiabat according to Katsura [41].

### Raman spectra

As Raman technique are now actively used for the identification of  $sp^3$  bonded carbonates and orthocarbonates in high-pressure DAC experiments [42, 43], we have calculated Raman spectra for the  $\text{Mg}_2\text{CO}_4$ - $Pnma$  phase.

The unit cells of  $\text{Mg}_2\text{CO}_4$ - $Pnma$  contain 28 atoms, i.e. there are 84 phonon modes. According to a group theoretical analysis, 36 Raman active modes are expected for  $\text{Mg}_2\text{CO}_4$ - $Pnma$ :  $\Gamma = 11A_g + 7B_{1g} + 11B_{2g} + 7B_{3g}$ . The most intense mode ( $B_{2g}$ ) corresponds to the bending and stretching vibrations in the  $[\text{CO}_4]$  tetrahedral groups and appears at  $1025 \text{ cm}^{-1}$ . The second ( $A_g$ ) and third ( $B_{1g}$ ) most intense modes appear at  $1089$  and  $1106 \text{ cm}^{-1}$ , correspondingly (Figure 5). The calculated Raman spectra at 60 GPa is shown in Figure 6.

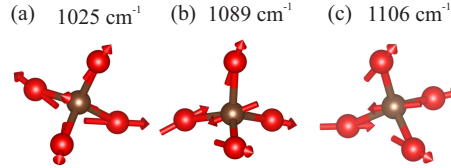


Figure 5: Displacement patterns in Raman modes of  $\text{Mg}_2\text{CO}_4$ - $Pnma$  at 60 GPa. Arrows indicate the displacement of the atoms during the specific vibration.



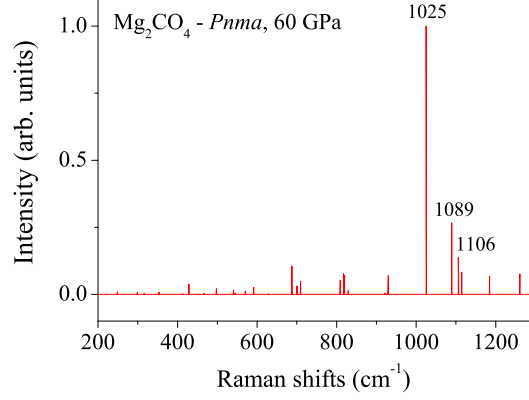


Figure 6: DFT-calculated Raman spectra of  $\text{Mg}_2\text{CO}_4$ -*Pnma* at 60 GPa.

#### Analysis of electronic denisty distribution

In both found structures, *Pnma* and *P2<sub>1</sub>/c*, four C–O bonds within [CO<sub>4</sub>] tetrahedron are of nearly the same length. At 100 GPa, these bond distances vary in the range 1.32–1.37 Å in both structures. The similar values of bond lengths assume the covalent nature for all four bonds. The performed analysis of electron density distribution confirms this assumption. The isosurface of the electron density difference clearly shows the accumulation of charge halfway along each of the four C–O bond (Figure 7).

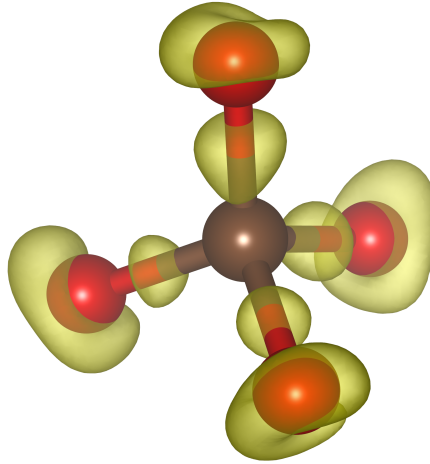


Figure 7: Isosurface of the electron density difference in  $\text{Mg}_2\text{CO}_4$ -*Pnma* structure. The isosurface shows those regions in which the electron density is larger than that obtained by overlapping electron densities of non-interacting atoms.

The isostructural character of  $\text{Mg}_2\text{CO}_4\text{-}Pnma$  and  $\text{Mg}_2\text{SiO}_4\text{-}Pnma$  gives a rare possibility for performing the comparison of electron density distribution in the structures of orthocarbonate and orthosilicate.

At 0 GPa, the Si–O bond is longer than C–O bond on 15-19 %. According to the classical rule of isomorphism, suggested by Goldschmidt, the isomorphism in the wide range of compositions require the difference of atomic radii on no more than 15 %. The obtained value is on the border of this limit and from the first glance the possibility of the (Si,C) isomorphism at high temperatures can not be excluded. However, it is not the case, as isomorphism requires similar nature of the bonds around atoms substituting each other. The obtained Bader charge of  $\text{C}^{4+}$  is predictably lower than that of  $\text{Si}^{4+}$ , 1.914 against 3.109. This indicates on the more covalent character of C–O bond in comparison with the Si–O bond. Bader charges of the other atoms are summarised in Table 2.

Table 2: Comparison of Bader charges on atoms in  $\text{Mg}_2\text{SiO}_4\text{-}Pnma$  and  $\text{Mg}_2\text{CO}_4\text{-}Pnma$  structures at 0 GPa (in unit of e).

	Mg1	Mg2	C/Si	O1	O2	O3
$\text{Mg}_2\text{SiO}_4$	1.731	1.745	3.109	-1.645	-1.657	-1.641
$\text{Mg}_2\text{CO}_4$	1.724	1.745	1.914	-1.412	-1.340	-1.315

### Melting temperature and seismic properties

To compare the melting temperature of Mg-carbonate and Mg-orthocarbonates, we have estimated melting temperatures of Mg-orthocarbonate. The obtained values of the  $\text{Mg}_2\text{CO}_4\text{-}Pnma$  melting temperatures at ... GPa, 23.7 GPa, and at 35.5 GPa equal ... K to 2742 K and 2819 K, respectively (Figure ??). These values are 16-18 % higher than the experimentally measured melting temperatures of magnesite and comparable with solidi of silicate rocks under lower mantle P-T conditions [44]. As can be seen from Figure 4 the difference in melting temperatures promptly increases with pressure, reaching 500 K at nearly 40 GPa.

To assess the effect of orthocarbonate formation on the seismic properties of carbonate, we have also calculated the elastic stiffness tensor and compressional/shear sound velocities ( $V_p/V_s$ ) for Mg-carbonate and Mg-orthocarbonate. We have not aimed the investigation of trends for the changes of elastic properties on compression or their description during  $Pnma \rightarrow P2_1/c$  transition, but only the rough comparison of carbonate and orthocarbonate properties. By this reason, calculations have been performed only at 50 GPa and 0 K for  $\text{Mg}_2\text{CO}_4\text{-}Pnma$  and magnesite structures.

The obtained elastic properties of  $\text{Mg}_2\text{CO}_4\text{-}Pnma$  and magnesite are summarized in Table ?? and Table ?. Obtained values of  $C_{ij}$ ,  $B$ ,  $G$ ,  $V_p$ , and  $V_s$  for magnesite are in good agreement with the previous results of theoretical calculations [45]. According to our results, the density and both bulk and shear moduli of  $\text{Mg}_2\text{CO}_4\text{-}Pnma$  are higher than that of magnesite by 4-5 %. Orthocarbonate

is similar to carbonate in sense of seismic velocities  $V_p$  and  $V_s$ , but orthocarbonate is more anisotropic, owning higher universal anisotropy ( $A^U$ ). The value of  $A^U$  for  $\text{Mg}_2\text{CO}_4\text{-}Pnma$  is 0.4772, while that for magnesite is 0.2838.

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The computations were performed using resources provided by the Novosibirsk State University Supercomputer Center.

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