

The formation of Mg-orthocarbonate through the reaction $\text{MgCO}_3 + \text{MgO} = \text{Mg}_2\text{CO}_4$ at Earth's lower mantle P – T conditions

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Abstract

Orthocarbonates of alkaline earth metals are the newly discovered class of compounds stabilized at high pressures. Mg-orthocarbonates are the potential carbon host phases, transferring oxidized carbon in the Earth's lower mantle up to the core-mantle boundary. Here, we demonstrate the possibility for the formation of Mg_2CO_4 in the lower mantle at pressures above 50 GPa, by *ab initio* calculations. Mg_2CO_4 is formed by the reaction $\text{MgCO}_3 + \text{MgO} = \text{Mg}_2\text{CO}_4$, proceeding only at high-temperatures. At 50 GPa the reaction starts at 2200 K. The temperature decreases with pressure and drops down to 1085 K at the pressure of the Earth's core-mantle boundary, near 140 GPa. Two stable structures, Mg_2CO_4 -*Pnma* and Mg_2CO_4 -*P2₁/c*, were revealed using crystal structure prediction technique. Mg_2CO_4 -*Pnma* is isostructural to mineral forsterite (Mg_2SiO_4), while Mg_2CO_4 -*P2₁/c* is isostructural to mineral larnite (β - Ca_2SiO_4). Transition pressure from Mg_2CO_4 -*Pnma* to Mg_2CO_4 -*P2₁/c* is around 80 GPa. Both phases are dynamically stable on decompression down to the ambient pressure and can be preserved in the samples of natural high-pressure rocks or the products of experiments. Mg_2CO_4 -*Pnma* has a melting temperature more than 16% higher than the melting temperature of magnesite (MgCO_3). At 23.7 GPa, 35.5 GPa, and 52.2 GPa, Mg_2CO_4 -*Pnma* melts at 2661 K, 2819 K, and 3109 K, respectively.

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Acoustic wave velocities V_p and V_s of $\text{Mg}_2\text{CO}_4\text{-}Pnma$ are very similar to that of magnesite, while universal anisotropy of $\text{Mg}_2\text{CO}_4\text{-}Pnma$ is stronger than that of magnesite, as well as coefficient A^U is larger for orthocarbonate. Calculated Raman spectra of $\text{Mg}_2\text{CO}_4\text{-}Pnma$ will help its identification in high pressure experiments.

Introduction

During the recent decade, the crystal structure prediction technique became an integral part of high-pressure research. Numerous experimental synthesis were guided by this technique, for instance, the synthesis of the high-pressure phases of alkaline [1–3] and alkaline earth carbonates [4–10]. Orthocarbonates of alkaline earth metals are another example for experimental confirmation of theoretically predicted structures.

For a long time, Mg-orthocarbonate attracted attention of the Earth scientists as the possible phase of the deep Earth’s interiors. To the best of our knowledge, Fyfe was the first, who suggested the formation of Mg_2CO_4 phase with spinel structure through the reaction $\text{MgCO}_3 + \text{MgO} = \text{Mg}_2\text{CO}_4$ [11]. Then, the possibility of this reaction has been hypothesized by Irwing and Wyllie [12], and by Katsura [13]. However, recently performed *ab initio* crystal structure prediction calculations have not found any structures of Mg_2CO_4 stable at pressures of the Earth’s mantle [14]. The enthalpies of all the found structures were lower than the enthalpy of the mechanical mixture of $\text{MgCO}_3 + \text{MgO}$, i.e. they were thermodynamically unstable.

Our calculations on crystal structure prediction have revealed stable crystal structures of Ca-, Sr-, and -Ba, orthocarboates, Ca_2CO_4 [15], Sr_2CO_4 [16], and Ba_2CO_4 [16]. Subsequent experimental synthesis have confirmed the existence of Ca_2CO_4 and Sr_2CO_4 [17,18], and Mg_2CO_4 remains the sole representative of alkaline earth orthocarbonates for which stable structures are not yet known.

In the present work, we continue our investigation of orthocarbonates, find stable structures of Mg_2CO_4 , calculate their melting temperatures, seismic properties and P – T phase diagram, as well as Raman spectra for the identification in high-pressure experiments, which are currently underway.

Methods

To increase the chances for finding the energetically favorable structures we have used both USPEX and AIRSS methods, each of which has apparent advantages [19]. *Ab initio* structure prediction was complemented by the prediction technique based on the known structures of Mg- and Zn- orthosilicates with isolated $[\text{SiO}_4]$ tetrahedrons, as well as sulfates with isolated $[\text{SO}_4]$ tetrahedrons. Mg_2CO_4 structures were produced from these structures by the corresponding replacement of the cations and consequent local optimizations. The following crystal structures have been used (numbers in parentheses indicate

the number of formula units (f.u.) in the unit cell): ringwodite (Mg_2SiO_4 - $Fd\bar{3}m$ (8)), porierite (Mg_2SiO_4 - $Pmma$ (4)), wadsleyite (Mg_2SiO_4 - $Imma$ (8)), forsterite (Mg_2SiO_4 - $Pnma$ (4)) [20], Zn_2SiO_4 - $Pbca$ (8), Zn_2SiO_4 - $I\bar{4}2d$ (4) [21], Na_2SO_4 - $Fddd$ (8) [22], Li_2SO_4 - $P2_1/c$ (4) [23], and Ca_2CO_4 - $Pnma$ (4) [15].

The calculations with the USPEX method [24–27] have been performed for 1–4 f.u. of Mg_2CO_4 at 25, 50, and 100 GPa. The seeding technique, implemented in version 10.2 of the USPEX, has been employed in all calculations. The aforementioned structures based on silicates, sulfates, and orthocarbonate have been used as the seeds. Totally, around 3000 structures have been calculated at each pressure. Crystal structure prediction with AIRSS 0.9.1 [28, 29] has been performed only at 50 GPa for 2, 3, and 4 f.u. and a total of nearly 4000 structures have been generated in this calculation.

The energetic optimizations of the predicted structures have been performed within density functional theory (DFT), implemented in the VASP package [30, 31].

To take into account the temperature effect and calculate P - T phase diagram, we used the lattice dynamics method within the quasi-harmonic approximation (QHA). For this task, the lattice vibration frequencies were calculated with the PHONOPY code [32]. By our experience with P - T phase diagrams of carbonates, this technique reliably reproduces phase boundaries in the wide temperature range [2, 15, 33], up to 80–90 % of the melting temperature, if the process of dynamical disordering does not take place.

Melting temperatures of the predicted structures have been determined, using the so-called *Z-method*, based on the molecular dynamic (MD) simulations [34].

To determine the wave velocities V_p and V_s and assess anisotropy of Mg-orthocarbonates, static elastic stiffness tensor (C_{ij}) was calculated from the stress (σ)- strain (ϵ) relation $\sigma_i = C_{ij}\epsilon_j$. Based on these C_{ij} data, we calculated averages of bulk (B) and shear (G) moduli using the Voigt-Ruess-Hill scheme [35, 36]. Then we have determined compression (A_B), shear (A_G), and universal anisotropy (A^U) indexes.

The details of DFT, crystal structure predictions, thermodynamic and elastic property calculations, and MD simulations are given in *Supporting information*.

In order to obtain the Raman spectra, the polarizability tensors for each crystal mode were calculated using the *vasp_raman.py* code [37]. The atomic charges were analyzed by the grid-based bader analysis algorithm [38, 39], in which the grid is obtained by decomposition of the charge density by the static self-consistent calculation for the optimized structures. FindSym program [40] and instruments of Phonopy package have been used for the symmetry determination. VESTA and ToposPro [41, 42] programs have been used for the visualization of the crystal structures and figures preparation. The topology of the structures were analysed with ToposPro and Robocrystallographer programs [42, 43].

Results and discussion

Predicted structures

In the crystal structure prediction calculations at 25 and 50 GPa, USPEX has revealed the Mg_2CO_4 - $Pnma$ structure as the most energetically favorable. This structure has lower enthalpy than Mg_2CO_4 - Cm , predicted by AIRSS, and structures, obtained based on the crystal structures of silicates and sulfates (Figure 1 and S1).

At 100 GPa, USPEX has found another favorable structure with monoclinic symmetry, Mg_2CO_4 - $P2_1/c$. The transition from Mg_2CO_4 - $Pnma$ to Mg_2CO_4 - $P2_1/c$ occurs at 52 GPa, and at higher pressures Mg_2CO_4 - $P2_1/c$ is the most energetically favorable among considered structures (Figure 1 and Figure S1).

Other structures constructed based on the polymorphs of Mg_2SiO_4 , ringwoodite, wadsleyite and poirierite, do not have pressure ranges of stability. However, in the stability field of Mg_2CO_4 - $P2_1/c$, the series of phase transitions, forsterite \rightarrow poirierite \rightarrow wadsleyite \rightarrow ringwoodite, is realised on compression, similarly as it takes place in Mg_2SiO_4 ([44] and references therein).

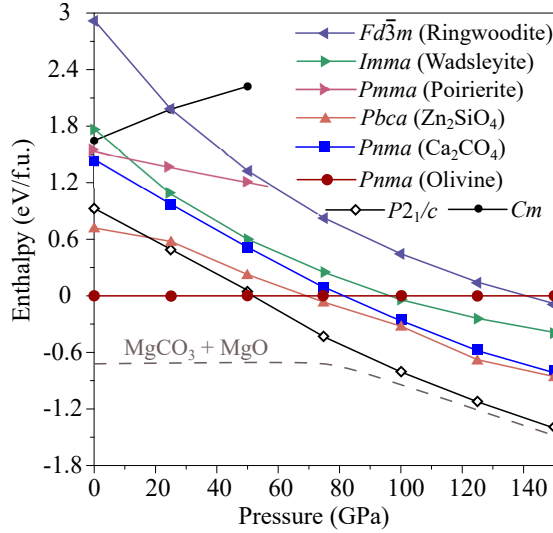


Figure 1: Enthalpy-pressure dependencies of Mg_2CO_4 structures. Isostructural compounds are indicated in brackets.

Revealed Mg_2CO_4 - $Pnma$ is isostructural to the mineral forsterite (Figure 2a,b). Atoms in both positions Mg(1) and Mg(2) are 6-coordinated, with the octahedral coordination polyhedron (Figure 2c). Mg_2CO_4 - $P2_1/c$ is isostructural to β - Ca_2SiO_4 (larnite) and can be considered as the monoclinic analogue of Ca_2CO_4 - $Pnma$. In crystal structure of Mg_2CO_4 - $P2_1/c$, there are two non-equivalent Mg sites. There is some ambiguity in determination of coordination numbers of Mg, due to the smooth variation of the bonds lengths. In the

first site Mg(1) is bonded to six oxygen atoms, arranged in highly deformed trigonal prism (Figure 2d). Lengths of the six Mg–O bonds within coordination polyhedron vary in the range 1.832–2.073 Å, while the seventh oxygen atoms is distant on 2.382 Å. In the second site, Mg(2) is bonded to eight oxygen atoms, forming relatively regular square antiprism. Lengths of these eight Mg–O bonds vary in the range of 1.887–2.126 Å, while the ninth Mg–O bond is sufficiently longer, being equal to 2.688 Å. For comparison, in $\text{Mg}_2\text{CO}_4\text{-}Pnma$ structure, oxygen atoms in the first coordination sphere is distant on 1.815–1.888 Å for Mg(1) site and on 1.781–1.95 Å — for (Mg2) site. All the presented bond lengths correspond to the pressure of 100 GPa. In both $Pnma$ and $P2_1/c$ crystal structures, coordination polyhedrons around Mg atoms share both vertices and edges with $[\text{CO}_4]$ tetrahedrons (Figure 2c,d). Transition from $\text{Mg}_2\text{CO}_4\text{-}Pnma$ to $\text{Mg}_2\text{CO}_4\text{-}P2_1/c$ is accompanied by the sufficient decrease of the volume equal to 5.7 % (Figure S2).

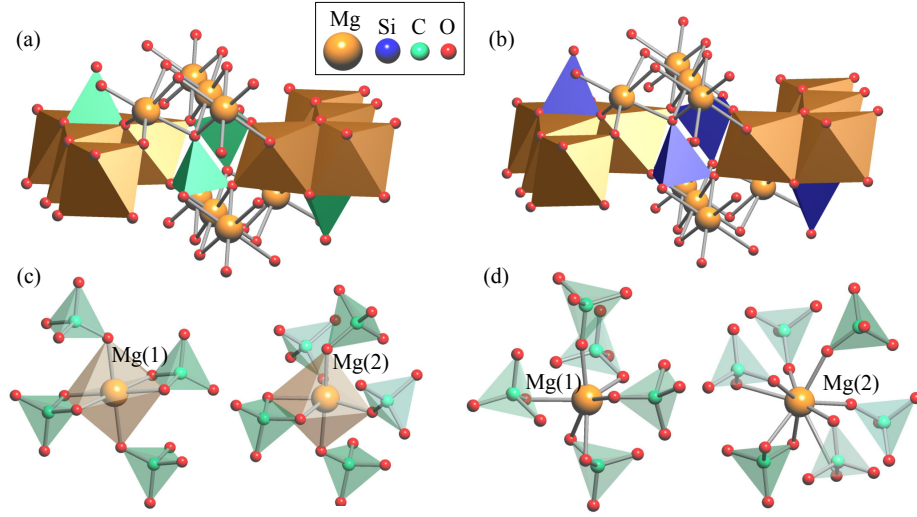


Figure 2: Comparison of crystal structures of $\text{Mg}_2\text{CO}_4\text{-}Pnma$ (a) and $\text{Mg}_2\text{SiO}_4\text{-}Pnma$ (b) along [010]; coordination environment around Mg(1) and Mg(2) sites in crystals structures of $\text{Mg}_2\text{CO}_4\text{-}Pnma$ (c) and $\text{Mg}_2\text{CO}_4\text{-}P2_1/c$ (d).

Structural data for the $\text{Mg}_2\text{CO}_4\text{-}Pnma$ and $\text{Mg}_2\text{CO}_4\text{-}P2_1/c$ are given in the Table 1. Both structures are dynamically stable in the investigated pressure range 0-140 GPa, and there are no imaginary frequencies in their phonon dispersion curves (Figure 3, S3). The dynamical stability of both structures at ambient pressure assume the possibility for their extraction from the chamber of the diamond anvil cell (DAC) after the high pressure synthesis and the further laboratory investigation. In the *Supporting information*, we have also shown phonon dispersion curves of MgCO_3 and MgO (Figure S4) used for the construction of P – T phase diagram.

Table 1: Mg_2CO_4 structures introduced in the current manuscript.

P (GPa)	Space group	Lattice parameters (\AA , deg)			Atom	Coordinates		
						x	y	z
50	$Pnma$ (#62)	$a = 8.926$ $\alpha = 90.0$	$b = 5.565$ $\beta = 90.0$	$c = 4.221$ $\gamma = 90.0$	Mg1	0.000	0.000	0.500
					Mg2	0.721	0.250	0.531
					C1	-0.097	0.250	0.087
					O1	-0.091	0.250	0.770
					O2	0.548	0.250	0.284
					O3	0.169	0.552	0.786
100	$P2_1/c$ (#14)	$a = 4.408$ $\alpha = 90.0$	$b = 5.383$ $\beta = 117.65$	$c = 8.345$ $\gamma = 90.0$	Mg1	0.702	0.360	0.425
					Mg2	-0.022	0.000	0.693
					C1	0.355	0.282	0.082
					O1	0.146	0.334	0.638
					O2	0.681	0.245	0.197
					O3	0.272	0.168	-0.080
					O4	0.295	0.520	0.064

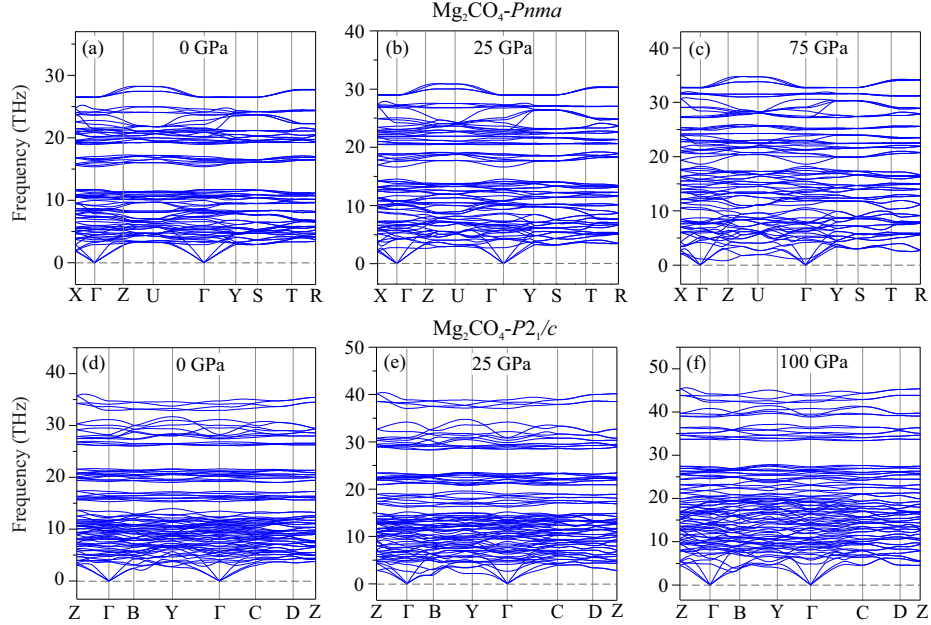


Figure 3: Phonon dispersion curves of Mg_2CO_4 - $Pnma$ (a-c) and Mg_2CO_4 - $P2_1/c$ (d-f) at several pressures.

There is no analogy of the found transition from Mg_2CO_4 - $Pnma$ to Mg_2CO_4 - $P2_1/c$ in the Mg_2SiO_4 system. However, there is one in Ca_2SiO_4 . As γ - Ca_2SiO_4 is isostructural to Mg_2CO_4 - $Pnma$ and β - Ca_2SiO_4 — to Mg_2CO_4 - $P2_1/c$, the

transformation of γ -Ca₂SiO₄ into β -Ca₂SiO₄ realized on compression [45], is the analogue of Mg₂CO₄-*Pnma* \rightarrow Mg₂CO₄-*P2₁/c* transformation. Based on the similarity of high pressure phase transitions of Mg₂CO₄ and Ca₂SiO₄, the analogue of the transition β -Ca₂SiO₄ \rightarrow α'_H -Ca₂SiO₄ can be suggested. In this case, on heating, Mg₂CO₄-*P2₁/c* will transform into the hypothetical phase Mg₂CO₄-*Pnma*-II, analogue of α'_H -Ca₂SiO₄. The performed calculations of phonon dispersion curves of Mg₂CO₄-*Pnma*-II, constructed based on the structure of α'_H -Ca₂SiO₄ by corresponding atomic replacements, have shown the dynamical instability of this phase (Figure S6). However, the stabilization of the structure by the factors, which are not considered in QHA can not be excluded.

***P-T* phase diagram**

The performed enthalpy calculations of the Mg₂CO₄, MgO, and MgCO₃ structures have shown, that both Mg₂CO₄-*Pnma* and Mg₂CO₄-*P2₁/c* lie above the energetic convex hull, i.e. they are unstable and decompose to the mixture (MgO+MgCO₃) at 0 K (Figure S7).

However, performed calculations of the Gibbs free energies in the temperature range of 0–3000 K have shown that above some temperature Mg₂CO₄ became more energetically favorable than the (MgO+MgCO₃) mixture (Figure S8). At 20 GPa, this temperature is 2420 K, which is nearly equal to the melting temperature of magnesite (MgCO₃-*R $\bar{3}$ c*) [46] (Figure 4). With increasing the pressure the temperature of transition decreases, and at 140 GPa Mg₂CO₄-*P2₁/c* became more favourable than the mixture of (MgO+MgCO₃) at temperatures higher than 1085 K. Thus, in the pressure range of 50–140 GPa, Mg₂CO₄ can be synthesised at temperatures 250–1100 K higher than the corresponding temperatures of MgCO₃ melting (Figure 4). It suggests the possibility for the formation of Mg-orthocarbonate in most part of the Earth's lower mantle, at pressures higher than 50 GPa.

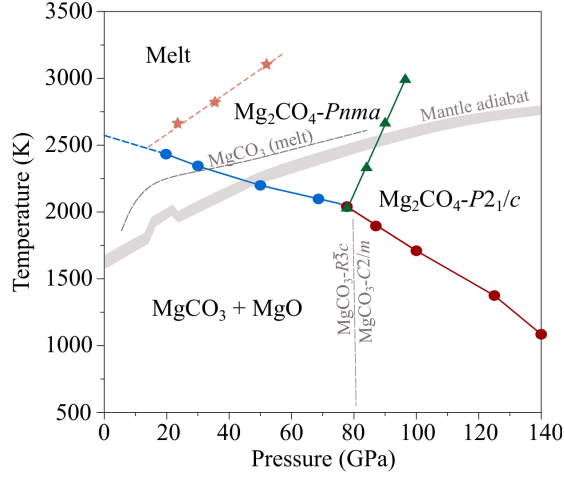


Figure 4: P - T phase diagram of Mg_2CO_4 . The asterisks represent calculated melting temperatures; the grey dash-dotted line — calculated phase transition boundary of MgCO_3 ; the grey dashed line — melting curve of MgCO_3 as reported by Solopova et al. [46]; grey solid line — mantle adiabat according to Katsura [47].

Raman spectra

As Raman technique are now actively used for the identification of sp^3 bonded carbonates and orthocarbonates in high-pressure DAC experiments [48, 49], we have calculated Raman spectra for the Mg_2CO_4 - $Pnma$ phase.

The unit cells of Mg_2CO_4 - $Pnma$ contain 28 atoms, i.e. there are 84 phonon modes. According to a group theoretical analysis, 36 Raman active modes are expected for Mg_2CO_4 - $Pnma$: $\Gamma = 11A_g + 7B_{1g} + 11B_{2g} + 7B_{3g}$. The most intense mode (B_{2g}) corresponds to the bending and stretching vibrations in the $[\text{CO}_4]$ tetrahedral groups and appears at 1025 cm^{-1} . The second (A_g) and third (B_{1g}) most intense modes appear at 1089 and 1106 cm^{-1} , correspondingly (Figure 5). The calculated Raman spectra at 60 GPa is shown in Figure 6.

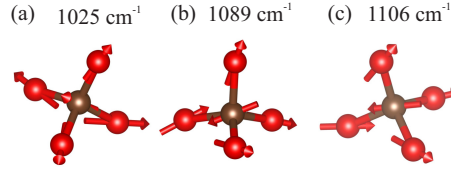


Figure 5: Displacement patterns in Raman modes of Mg_2CO_4 - $Pnma$ at 60 GPa. Arrows indicate the displacement of the atoms during the specific vibration.

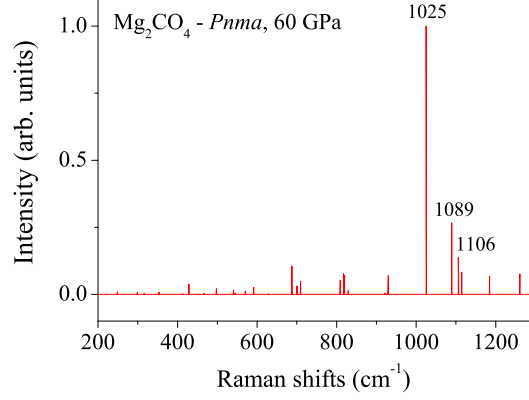


Figure 6: DFT-calculated Raman spectra of Mg_2CO_4 -*Pnma* at 60 GPa.

Analysis of electronic denisty distribution

In both found structures, *Pnma* and *P2₁/c*, four C–O bonds within $[\text{CO}_4]$ tetrahedron are of nearly the same length. At 100 GPa, these bond distances vary in the range 1.32–1.37 Å in both structures. The similar values of bond lengths assume the covalent nature of all four bonds. The performed analysis of electron density distribution confirms this assumption. The isosurface of the electron density difference clearly shows the accumulation of charge halfway along each of the four C–O bond (Figure 7).

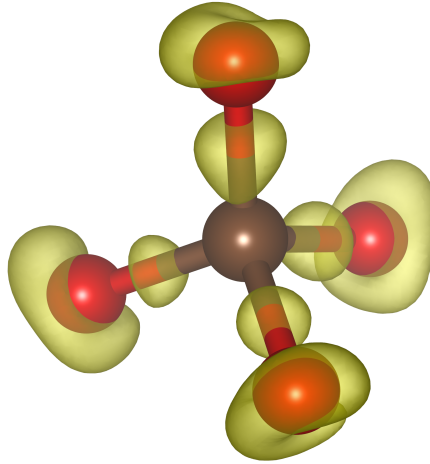


Figure 7: Isosurface of the electron density difference in Mg_2CO_4 -*Pnma* structure. The isosurface shows those regions in which the electron density is larger than that obtained by overlapping electron densities of non-interacting atoms.

The isostructural character of $\text{Mg}_2\text{CO}_4\text{-}Pnma$ and $\text{Mg}_2\text{SiO}_4\text{-}Pnma$ gives a rare possibility for performing the comparison of electron density distribution in the structures of carbonate and silicate.

At 0 GPa, the Si–O bond is longer than C–O bond on 15-19 %. The obtained Bader charge of C^{4+} is predictably lower than that of Si^{4+} , 1.914 and 3.109, for C^{4+} and Si^{4+} , respectively. This indicates on the more covalent character of C–O bond in comparison with the Si–O bond. This difference in the nature of the bonds together with difference of atomic sizes will obstruct (Si,C) isomorphism, despite the isostructural character of orthosilicates and orthocarbonates of Mg, Ca [15], and Sr [17]. Bader charges of the other atoms are summarised in Table 2.

Table 2: Comparison of Bader charges on atoms in $\text{Mg}_2\text{SiO}_4\text{-}Pnma$ and $\text{Mg}_2\text{CO}_4\text{-}Pnma$ structures at 0 GPa (in unit of e).

	Mg1	Mg2	C/Si	O1	O2	O3
Mg_2SiO_4	1.731	1.745	3.109	-1.645	-1.657	-1.641
Mg_2CO_4	1.724	1.745	1.914	-1.412	-1.340	-1.315

Melting temperature and seismic properties

To compare the melting temperatures of carbonate and orthocarbonate of Mg, we have estimated the melting temperatures of Mg-orthocarbonate. The obtained values of the $\text{Mg}_2\text{CO}_4\text{-}Pnma$ melting temperatures at 23.7 GPa, 35.5 GPa and 52.2 GPa equal to 2661 K, 2819 K, and 3109 K, respectively (Figure S9). These values are 14-22 % higher than the experimentally measured melting temperatures of magnesite and comparable with solidi of silicate rocks under lower mantle P - T conditions [50]. As can be seen from Figure 4 the difference in melting temperatures promptly increases with pressure, reaching 500 K at nearly 40 GPa.

To assess the effect of orthocarbonate formation on the seismic properties of carbonate, we have also calculated the elastic stiffness tensor and compressional/shear sound velocities (V_p/V_s) for Mg-carbonate and Mg-orthocarbonate. We have not aimed the investigation of trends for the changes of elastic properties on compression nor their description during $Pnma \rightarrow P2_1/c$ transition, but only the rough comparison of carbonate and orthocarbonate properties. By this reason, calculations have been performed only at 50 GPa and 0 K for $\text{Mg}_2\text{CO}_4\text{-}Pnma$ and magnesite structures.

The obtained elastic properties of $\text{Mg}_2\text{CO}_4\text{-}Pnma$ and magnesite are summarized in Table S1 and Table S2. Obtained values of C_{ij} , B , G , V_p , and V_s for magnesite are in good agreement with the previous results of theoretical calculations [51]. According to our results, the density and both bulk and shear moduli of $\text{Mg}_2\text{CO}_4\text{-}Pnma$ are higher than that of magnesite by 4-5 %. Orthocarbonate is similar to carbonate in sense of seismic velocities V_p and V_s , but orthocarbonate is more anisotropic, owning higher universal anisotropy (A^U). The value of

A^U for $\text{Mg}_2\text{CO}_4\text{-}Pnma$ is 0.4772, while that for magnesite is 0.2838.

Acknowledgments

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