
Thermodynamics and Statistical Mechanics

Contents

1	Motivation	1
1.1	Random Walks	1
2	Preliminaries	7
2.1	Condensed Matter Phenomenology	8
2.2	The Classical World	8
2.3	Then there were atoms	8
2.4	The Future	8
2.5	The Basics	8
2.6	Mathematics	8
3	The Four Commandments	9
4	Thermodynamics in Action	11
4.1	Some Preliminaries: Pfaffians and Integrability . . .	11
4.2	Equations of State	12
4.3	Some Empirical Equations of State	14
5	Phase Transitions	17
5.1	First Order Phase Transitions	17
6	Super Conductivity	21
6.1	Ginzburg-Landau Theory	21
7	Statistical Mechanics	23
7.1	Two State Systems	23
7.2	Equations of State	25
8	Bose-Einstein	27
8.1	Black Body Radiation	27
	Bibliography	31

Chapter 1

Motivation

We are going to start this whole thing by laying down the ideals and assumptions made for studying system at equilibrium.

When studying these systems, the quantities of interest will be their probabilities or the probability distribution of certain quantities. From probability distributions we will then be able to predict measurable quantities.

We will use our knowledge of the physical laws to try and predict the probability of a particle to be in state X. From here on we will be able to use a vast assortment of mathematical equations to calculate things such as the energy of the system, its heat capacity, magnetization, etc.

We will start our trip by looking at the "harmonic oscillator" of statistical mechanics: the random walk.

As mentioned above, we want to see how we can go from "we understand the behaviour of this one 'microscopic' thing" to "this is the probability distribution for a lot of 'microscopic' things to do something" and finally to "using the probability distribution we are able to make this prediction about the macroscopic state of our system".

This process is very important to understand as it is the only sensible way to study complex systems. And as such, let us begin!

1.1 Random Walks

Ensembles

As we begin looking into probabilities it is very useful to introduce the idea of an ensemble.

Let's say you want to calculate the probability of a coin landing in heads. What would you do? You would probably throw a coin a couple dozen times and count the number of times it landed heads and compare that to the total number of throws.

A similar setup would be to configure multiple dozens of universes, all the same as yours, in which you throw the same coin and then count the number of universes in which the coin landed in heads and compare that to the total number of universes we looked at. In this latter example, we setup an ensemble to measure the probability of our coin landing in heads.

This formulation, as trivial as it sounds, can be a lot more powerful when trying to make sense of our results. To make a quick example, have you checked the weather for today? What is the probability of rain?

We will repeatedly use this idea of ensembles to setup our thought experiments and to formulate our formulas. So whenever you see us mention that the probability of something is blah, think to yourself that the way we are calculating said probability is by configuring many many replicas of the thing we are studying and seeing how each one evolves and then using our results to get the probability.

A Coin Toss

We now begin by looking at the simplest system possible. Imagine you have a particle in a 1D plane. Our particle can only move left or right.

Based on our knowledge of the physical laws governing this particle, we will try and generalize by imagining an ensemble of said particles. This will allow us to figure out probability distributions, e.g., how likely is it that my particle ended X distance away from the start.

One thing you will see in complex systems is that even very simple building blocks will give rise to interesting phenomena. So let's get going with our study.

Again, our particle can only move left or move right. The distance it can move in any of those directions is constant (its steps are always of the same length). We want to figure out first, what is the probability of landing a distance x after N steps and then see if we can generalize this expression to obtain a probability distribution that we can use to calculate the probability of the particle being at any x distance from the origin after any given value of N steps.

Counting and Probabilities

So first step, just to get a feeling for the problem, at any given time our particle can either take a step to the left or a step to the right. We could decide by tossing a coin and moving to the right if it lands heads and move to the left otherwise.

With this procedure, what would be the probability of ending up 2 steps to the right of the origin after taking 8 steps?

If we ended up 2 steps to the right after taking 8 steps, then it means we took 2 steps more to the right than to the left. This could have happened if we took a total of 5 steps to the right and 3 to the left. And this is our first problem: what is the probability of this happening?

Well, let's go to the basics. We want to calculate

$$P(5 \text{ steps to the right out of } 8 \text{ steps taken})$$

From the basic real of probability, the above quantity is the ratio of the number of ways in which it is possible to take 5 steps to the right when taking 8 steps and the total number of possible outcomes after 8 steps - the outcome being the distance from the origin after N steps.

$$P\left(\frac{5}{8} \text{ steps}\right) = \frac{\# \text{ of ways we can take 5 steps}}{\# \text{ of possible outcomes}} \quad (1.1)$$

Again, here the possible outcomes could have been our particle ending 8 steps to the left, 8 to the right, 7 to the left, 7 to the right, etc.

Assumption The above formulation of our problem makes the assumption that all possible outcomes are equally likely. We will build upon this assumption and explicitly call it out if we ever deviate from it.

At this point we have a counting problem. We need to calculate the number of possible outcomes after 8 steps and we need to calculate how many different ways we could have taken 5 steps to the right in order to end 2 steps away from the origin.

In order to calculate the possible number of outcomes let's think about it. Each time that our particle takes a step it can go to the left or to the right, there are 2 possibilities at each step.

In mathematics you will see the tool for solving this listed under the name **multiplication rule**

The multiplication rule If something happens (i.e., take a step to the right or to the left, flip a coin) and this something has n_1 possible outcomes, and each of these resulting outcomes has n_2 possible outcomes, and so on, such that at the k th iteration of our experiment we have n_k possible outcomes then the total number of possible outcomes is

$$n_1 \times n_2 \times \dots \times n_k$$

In our case, each time we take a step we have 2 possible outcomes (we take a step to the left or to the right). Using the multiplication rule it means that we have

$$2 \times 2 \times 2 \times 2 \times 2 \times 2 \times 2 \times 2 = 2^8 = 256$$

possible outcomes. This is the value of the denominator in equation (1.1).

Now we need to figure out the different number of ways we can take 5 steps to the right when we take a total of 8 steps. The order in which the steps are taken doesn't matter, it only matters the total number of steps.

Let's think about it for a bit. We are taking N steps. One step we take can be taken at any of the N times we take a step. Another step could have been taken in the remaining $N - 1$ steps (we have now taken 2 steps to the right). Another step could have been taken in the remaining $N - 2$ steps (we have now taken 3 steps to the right), etc.

Using the multiplication rule, we see that we have

$$N \times (N - 1) \times (N - 2) \times (N - 3) \times (N - 4) \quad (1.2)$$

$$= 8 \times 7 \times 6 \times 5 \times 4 = 6720 \quad (1.3)$$

Is the total number of ways in which we could have taken our 5 steps to the right, if we take order into account. In our case, it doesn't matter what order we took our steps it only matters that we took them. To account for this fact, we need to divide the above number (the number of permutations of 5 steps to the right out of 8 total steps) by the number of ways in which we could have taken 5 steps.

The number of ways in which 5 steps could have been taken can also be obtained from the multiplication rule by following a similar reasoning. If we forget about order, 1 step could have been taken at 5 different times. Step 2 could have been taken at any of the remaining 4 times. Step 3 could have been taken at any of the remaining 3 times, and so on. Which means there are

$$5 \times 4 \times 3 \times 2 \times 1 = 120 \quad (1.4)$$

different ways in which the 5 steps could be ordered.

Taking the ratio of (1.3) and (1.4), which is 56, as the numerator for (1.1)

$$P\left(\frac{5}{8}\text{steps}\right) = \frac{\# \text{ of ways we can take 5 steps}}{\# \text{ of possible outcomes}} \quad (1.5)$$

$$= \frac{56}{256} = 0.21875 \quad (1.6)$$

However, the answer is not as important as what we can learn from it. Let's go back into our calculation.

We discovered that the number of possible outcomes was

$$2^N \tag{1.7}$$

If we go back through the reasoning for (1.2) we can see that for the number of different ways in which we can take k steps to the right in N trials is

$$N \times (N - 1) \times \dots \times (N - (k - 1)) \tag{1.8}$$

when order matters, is the total number of permutations. And the factor by which the total number of permutations overcounts is, (1.4),

$$k! \tag{1.9}$$

Putting together (1.8) and (1.9) we can then calculate the number of combinations (different ways in which we can take k steps in a given direction out of N trials)

$$\frac{N \cdot (N - 1) \cdot \dots \cdot (N - (k - 1))}{k!} \tag{1.10}$$

The numerator of (1.10) is almost $N!$. Let's do some manipulation to see if we can further simplify our formula for calculating the number of combinations.

$$\frac{N!}{(N - k)!} = N \cdot (N - 1) \cdot \dots \cdot (N - (k - 1)) \tag{1.11}$$

We added the denominator of $(N - k)!$ as this essentially cancels out all the terms of the right-hand side after $(N - (k - 1)) = (N - k + 1)$. Using (1.11) in (1.10) we arrive at the number of combinations (the different number of ways we can take k steps in a given direction if we take N steps) to be

$$\frac{N!}{(N - k)!k!} = \binom{N}{k} \tag{1.12}$$

Which is the expression for binomial coefficients.

Now that we have a general expression for calculating the probability of ending up at any given distance from the origin we can go ahead and move with the rest of our plan of study which is to study the probability distribution that arises from all of this and then figure out how we can use all these equations to calculate observable quantities - we already know how to calculate the probability of ending up some distance away from the origin so let's explore how much more information we can obtain from our system given what little we know about it (it moves left or right at

constant length steps).

TODO: add that one thing about how to calculate different moments of a distribution.[1]

Chapter 2

Preliminaries

Limits of the continuum model (Include discussion on Loschmidt's number).

Property (thermodynamic variable/coordinate) observables characteristics of the system.

State variable describe equilibrium states. Also, these are properties whose differential is exact (which begs the question of what an inexact differential is). The state of a system is defined as a condition uniquely specified by a set of thermodynamic coordinates. Like in any problem one is then interested in what is the number of required of thermodynamic coordinates to uniquely specify my system. For example, if we were to design an experiment, how many properties would one need to specify in order to properly describe in a non-redundant manner the experimental set up, so that the experiment could be replicated. It turns out that these can be seen from the resulting form of the first law of thermodynamics being used. Lets take for example a closed system with one component. The resulting form of the first law is simply:

$$dU = TdS - PdV. \tag{2.1}$$

Here one sees that we are looking at the internal energy (a state whose entropy and volume are constrained) and has two pairs of conjugate variables (think of classical mechanics where the generalized distance and momenta are needed in order to completely solve the problem). Each one of these pairs is made up of an Intensive property multiplying and (exact) differential of the conjugate extensive property. For such a system one would need to specify two thermodynamic coordinates temperature and pressure, entropy and volume, etc.

One last definition: an equation of state is a functional relationship among the state variables of a system. Taking the former example, the equation of state would have the following form:

$$f(P, V, T) = 0 \quad (2.2)$$

Using this relationship one then reduces the number of independent variables from three to two. (Include discussion from Holme's textbook on foundations of applied mathematics).

2.1 Condensed Matter Phenomenology

2.2 The Classical World

Carnot's interest on the efficiency of steam engines.

Mayer postulated the equivalence of heat and work.

Clausius and entropy.

Nernst and Planck with the third law and limitations of physical systems.

2.3 Then there were atoms

Maxwell's kinetic theory.

Boltzmann, Gibbs on statistical mechanics.

2.4 The Future

Non-equilibrium

Non-linear effects.

2.5 The Basics

Building models of study going from the simplest to the real. Isolated, closed and open systems.

Intensive versus extensive properties and euler's theorem.

Classical thermodynamics is a continuum theory. Invoking quantum effects and the limitations of our theory. Loschmit's number.

2.6 Mathematics

The state of a system is uniquely defined by a set of properties.

How many constraints do we need to completely specify our system?

Equilibrium versus non-equilibrium states.

State variables are properties that describe equilibrium states.

Equation of state is a functional relationship between state variables.

The differential of a state variable is exact. Mathematics and abstraction of thermodynamics coordinates for descriptive purposes.

Chapter 3

The Four Commandments

Here we are to go through the laws of thermodynamics.

Chapter 4

Thermodynamics in Action

4.1 Some Preliminaries: Pfaffians and Integrability

Legendre Transformation

$$Z = Z(x, y) \\ dZ = Xdx + Ydy,$$

where x , X , and y , Y are canonical conjugate pairs.

If we wish to replace only one of the variables, say y , by its canonically conjugate variable Y , we must consider the function

$$N(x, Y) = Z - yY$$

A curve in a plane can be equally well represented by pairs of coordinates (point geometry), or as the envelope of a family of tangent line (line geometry).

Building Thermodynamic Relations

$$dE = TdS - PdV + \mu dN$$

SEE chapter 1.3 to 1.5 from pathria on studying the thermodynamics of an ideal gas.

Maxwell Relations

Internal Energy: $U(S, V)$

$$\left(\frac{\partial T}{\partial V}\right)_S = -\left(\frac{\partial P}{\partial S}\right)_V$$

This result can be derived as follows:

$$\frac{\partial^2 U}{\partial V \partial S} = \left(\frac{\partial T}{\partial V}\right)_S = \frac{\partial^2 U}{\partial S \partial V} = -\left(\frac{\partial P}{\partial S}\right)_V$$

Enthalpy: $H(S, P)$

$$\left(\frac{\partial T}{\partial P}\right)_S = \left(\frac{\partial V}{\partial S}\right)_P$$

Helmholtz: $F(T, V)$

$$\left(\frac{\partial S}{\partial V}\right)_T = \left(\frac{\partial P}{\partial T}\right)_V$$

Gibbs: $G(T, P)$

$$\left(\frac{\partial S}{\partial P}\right)_T = -\left(\frac{\partial V}{\partial T}\right)_P$$

Their use

Suppose it is found experimentally that for a solid

$$\left(\frac{\partial V}{\partial T}\right)_p = a + bp + cp^2$$

at temperature T for p_A to p_B . How much will the entropy increase when the solid is compressed from point A to point B at a constant temperature. Recall, $(\partial S/\partial p)_T = -(\partial V/\partial T)_p$, the desired entropy is then given by

$$\begin{aligned} \int dS &= \int_{p_A}^{p_B} \left(\frac{\partial S}{\partial p}\right)_T dp = - \int \left(\frac{\partial V}{\partial T}\right)_p dp \\ &= - \left[a(p_B - p_A) + \frac{1}{2}b(p_B^2 - p_A^2) + \frac{1}{3}c(p_B^3 - p_A^3) \right] \end{aligned}$$

4.2 Equations of State

From Point Particles to Hard Spheres

In a classical gas of hard spheres (of diameter D), the spatial distribution of the particles is no longer uncorrelated. Roughly speaking, the presence of n particles in the system leaves only a volume $(V - nv_0)$ available for the $(n + 1)$ th particle; clearly, v_0 would be proportional to D^3 . Assuming that $Nv_0 \ll V$, determine the dependence of (N, V, E) on V (compare to equation $(\Omega(N, V, E) \propto V^N)$).

$$\Omega \propto V(V - v_0)(V - 2v_0) \dots (V - (N - 1)v_0),$$

so

$$\log \Omega \propto \log V + \log(V - v_0) + \log(V - 2v_0) + \dots + \log(V - (N - 1)v_0)$$

$$= \log V + \sum_{i=1}^{N-1} \log(V - iv_0) = N \log V + \sum_{i=1}^{N-1} \log\left(1 - \frac{iv_0}{V}\right)$$

Because $v_0 \ll V$, and using:

$$\log(1+x) = x - \frac{1}{2}x^2 + \frac{1}{3}x^3 - \dots \quad (4.1)$$

We get

$$\log \Omega \propto N \log V + \sum_{i=1}^{N-1} \left(-\frac{iv_0}{V}\right)$$

Here we use the series

$$\sum_{k=1}^N k = \frac{N(N+1)}{2} \quad (4.2)$$

and so

$$\log \Omega \propto N \log V - \frac{v_0 N^2}{2V}$$

Using the formula

$$S = k \log \Omega \quad (4.3)$$

One can easily arrive at,

$$\frac{P}{T} = \left(\frac{\partial S}{\partial v}\right)_{N,E} = k \left(\frac{\partial \log \Omega}{\partial V}\right)_{N,E} = k \left(\frac{N}{V} + \frac{v_0 N^2}{2V^2}\right)$$

Rearranging terms,

$$PV \left(1 + \frac{v_0 N}{2V^2}\right)^{-1} = NkT$$

Noting that the term in parenthesis can be expanded by using a binomial expansion

$$(a+x)^\gamma = a^\gamma + \gamma x a^{\gamma-1} + \frac{1}{2} \gamma(\gamma-1) x^2 a^{\gamma-2} + \frac{1}{3!} \gamma(\gamma-1)(\gamma-2) x^3 a^{\gamma-3} + \dots \quad (4.4)$$

So

$$\frac{1}{1 + \frac{v_0 N}{2V^2}} = 1 - \frac{v_0 N}{2V^2} + \dots$$

and thus

$$P(v-b) = NkT$$

where $b = \frac{1}{2} N v_0$. Comparing b with the the radius of a particle, we can see that this term is proportional to four times the actual volume occupied by a particle. taking into account that hard spheres exclude a volume relative to twice their actual radius

$$b = \frac{N}{2} \frac{4\pi}{3} \sigma^3 = 4N \frac{4\pi}{3} \left(\frac{1}{2} \sigma\right)^3$$

4.3 Some Empirical Equations of State

Van der Waals Gas

Dietereci's Gas

Analysing the Dependencies of the Equation of State

Suppose that we have an equation of state such that

$$P = f(V)T$$

Show that the internal energy is independent of volume. In order for us to show this we use the equation

$$dS = \frac{1}{T} (dU + pdV)$$

Thus

$$\left(\frac{\partial S}{\partial V} \right)_T = \frac{1}{T} \left(\frac{\partial U}{\partial V} \right)_T + \frac{p}{T}$$

and

$$\left(\frac{\partial S}{\partial T} \right)_V = \frac{1}{T} \left(\frac{\partial U}{\partial T} \right)_V$$

We now require for dS to be an exact differential, $\partial^2 S / \partial V \partial T = \partial^2 S / \partial T \partial V$, which gives us

$$\frac{\partial}{\partial T} \left(\frac{1}{T} \frac{\partial U}{\partial V} \right) + \frac{\partial}{\partial T} \left(\frac{p}{T} \right) = \frac{\partial}{\partial V} \left(\frac{1}{T} \frac{\partial U}{\partial T} \right)$$

Rearranging terms

$$\left(\frac{\partial U}{\partial V} \right)_T = T^2 \frac{\partial}{\partial T} \left(\frac{p}{T} \right)_V = T \left(\frac{\partial p}{\partial T} \right)_V - p = 0.$$

This shows that the Internal energy is independent of the volume.

Another Equation of State

$$p = \frac{1}{3}u(T)$$

Where $u(T)$ is the internal energy density.

Using the result from the previous section, we have

$$\left(\frac{\partial U}{\partial V} \right)_T = T \left(\frac{\partial p}{\partial T} \right)_V - p$$

and using the fact that $U = u(T)V$, we have

$$u(T) = \frac{1}{3}T \frac{du(T)}{dT} - \frac{1}{3}u(T)$$

Which is equivalent to

$$T \frac{du(T)}{dT} - 4u(T) = 0$$

After integrating this gives,

$$u(T) \propto T^4$$

This is the case for the equation of state describing the thermal radiation due to a boson gas.

The Thermodynamics of a Rubber Band

Suppose that we know the equation of state for a rubber band is given by,

$$U = \theta S^2 L / n^2$$

where θ is a constant, L is the length of the rubber band. Determine the chemical potential $\mu(T, L/n)$ and show that the equation of state satisfies the analogue of the Gibbs-Duhem equation.

$$f = \left(\frac{\partial E}{\partial L} \right)_{S,n}$$

$$\mu = \left(\frac{\partial E}{\partial n} \right)_{S,L}$$

$$T = \left(\frac{\partial E}{\partial S} \right)_{L,N}$$

The Gibbs-Duhem equation is $0 = SdT + Ldf + nd\mu$.

$$dT = \left(\frac{\partial T}{\partial S} \right)_{L,n} dS + \left(\frac{\partial T}{\partial L} \right)_{S,n} dL + \left(\frac{\partial T}{\partial n} \right)_{S,L} dn$$

$$df = \left(\frac{\partial f}{\partial S} \right)_{L,n} dS + \left(\frac{\partial f}{\partial n} \right)_{S,L} dn$$

$$d\mu = \left(\frac{\partial \mu}{\partial S} \right)_{L,T} dS + \left(\frac{\partial \mu}{\partial L} \right)_{S,T,n} dL + \left(\frac{\partial \mu}{\partial n} \right)_{S,L} dn$$

Maxwell Relations for Rubber Bands

$$dE = TdS - pdV + fdL$$

On Stability of Rubber Bands

It is known from experiments that a rubber band will heat up when stretched adiabatically. But what about if we took the same rubber band and we cooled it down keeping the tension constant. would the rubber band expand or contract?

The fact that the rubber band heats up when stretched adiabatically means that

$$\left(\frac{\partial T}{\partial L}\right)_{S,n} > 0$$

or

$$\left(\frac{\partial T}{\partial f}\right)_{S,n} > 0$$

and we are interested in knowing about

$$\left(\frac{\partial L}{\partial T}\right)_{f,n}.$$

Chapter 5

Phase Transitions

5.1 First Order Phase Transitions

7.1

The latent heat of fusion of ice at a pressure of 1 atm and 0 °C is $3.348 \times 10^5 \text{ J kg}^{-1}$. The density of ice under these conditions is 917 kg m^{-3} and the density of water is 999.8 kg m^{-3} . If one kilomole of ice is melted,

(a) During a phase change the temperature and pressure remain constant. The volume is the only thing that changes. Hence the work done is $P\Delta v$.

$$v_1 = \frac{M}{\rho_1} = \frac{18 \text{ kg}}{917 \text{ kg m}^{-3}}; \quad v_2 = \frac{M}{\rho_2} = \frac{18 \text{ kg}}{999.8 \text{ kg m}^{-3}} \quad (5.1)$$

so that,

$$W = P\Delta v = P(v_2 - v_1) = -164 \text{ J}. \quad (5.2)$$

(b) The change in internal energy can be calculated by the first law, $Q = \Delta U - W$ which can be written as,

$$\begin{aligned} \Delta U &= Q - W = m_2 \ell_{12} - W \\ &= 18 \text{ kg} \times 3.344 \times 10^5 \text{ J kg}^{-1} + 164.2 \text{ J} = 6.02 \times 10^6 \text{ J}. \end{aligned} \quad (5.3)$$

(c) The change in entropy,

$$\Delta S = \frac{Q}{T} = \frac{m_2 \ell_{12}}{T} = 2.21 \times 10^4 \text{ J kg}^{-1}. \quad (5.4)$$

Universality

The equation of state is plotted in the P-V plane. The critical point is the point where (on the $T = T_c$ curve):

$$\left(\frac{\partial p}{\partial V}\right)_T = 0$$

$$\left(\frac{\partial^2 p}{\partial V^2}\right)_T = 0$$

Van Der Waals

$$p = \frac{nrT}{V - bn} = \frac{an^2}{V^2}$$

Taking the first and second derivatives with respect to volume on the critical temperature surface, we obtain the following equations:

$$RT/(V - bn)^2 = 2an/V^3$$

and

$$RT/(V - bn)^3 = 3an/V^4$$

From these, one can obtain $V_c = 3bn$ and $RT_c = 8a/27b$. Plugging these results on the equation of state give $p_c = a/27b^2$.

Substituting $p = \mathcal{P}p_c$, $v = \mathcal{V}v_c$, and $T = \mathcal{T}T_c$, along with their corresponding values.

$$\left(\mathcal{P} + \frac{3}{\mathcal{V}^2}\right)(3\mathcal{V} - 1) = 8\mathcal{T}$$

In this form the material constant a and b do not appear explicitly. Hence all van der Waals type gases may be considered equivalent when the values \mathcal{P} , \mathcal{V} , and \mathcal{T} are the same. This is the law of corresponding states.

Imagine you were an experimentalist. If you wanted to describe the behaviour of multiple gases this would be a very useful representation!

In plotting the \mathcal{T} curves on the new \mathcal{P} - \mathcal{V} plane notice the following:

For the regions of negative pressure, $\mathcal{V} < \frac{1}{3}$ and $\mathcal{P} < -3/\mathcal{V}^2$, need not be considered.

For $\mathcal{V} \gg \frac{1}{3}$ and $\mathcal{V} \gg 9/8\mathcal{T}$, then the equation of state becomes $\mathcal{P}\mathcal{V} \approx \frac{8}{3}\mathcal{T}$, the gas behaves like an ideal gas for large \mathcal{V} . Although, when \mathcal{V} approaches $\frac{1}{3}$, \mathcal{P} becomes infinite.

For a given value of \mathcal{V} there is only one value of \mathcal{P} , but for each value of \mathcal{P} there are up to three values of \mathcal{V} .

Dieterici

7.2 Coexistence and Phenomenology

So an experimentalist at RPI measured the equation of state for a substance near the liquid-solid phase transition. She found that over a limited range of temperatures and densities, the liquid phase could be characterized by the following formula

$$A_{(l)}/V = \frac{1}{2}a(T)\rho_{(l)}^2 \quad (5.5)$$

Here A is the Helmholtz free energy, $\rho = n/V$ is the molar density, and

$$a(T) = \alpha/T$$

Where α is a constant. Similarly, for the solid phase she found that,

$$A_{(s)}/V = \frac{1}{3}b(T)\rho_{(s)}^3 \quad (5.6)$$

Here

$$b(t) = \beta/T$$

with β being another constant.

At any given temperature, the pressure of the liquid can be adjusted to a particular pressure, P_s , at which point the liquid freezes.

Again let's look at the Clausius-Clapeyron equation and determine the slope dP/dT at the liquid-solid coexistence line.

First of all let's determine the molar densities as a function of temperature. We know that at the coexistence line:

$$\mu_{(s)} = \mu_{(l)}, P_{(s)} = P_{(l)},$$

and

$$T = T_{(s)} = T_{(l)}$$

Taking the first constraint, we have

$$\left(\frac{\partial A_{(s)}}{\partial n}\right)_{V,T} = \left(\frac{\partial A_{(l)}}{\partial n}\right)_{V,T} \rightarrow \frac{\alpha}{T} \left(\frac{n}{V}\right)_{(l)} = \frac{\beta}{T} \left(\frac{n}{V}\right)_{(s)}^2$$

Taking the second constraint,

$$-\left(\frac{\partial A_{(l)}}{\partial V}\right)_{n,T} = -\left(\frac{\partial A_{(s)}}{\partial V}\right)_{n,T} \rightarrow \frac{2}{3} \frac{\beta}{T} \left(\frac{n}{V}\right)_{(s)}^3 = \frac{1}{2} \frac{\alpha}{T} \left(\frac{n}{V}\right)_{(l)}^2$$

Solving for $\rho_{(l)}$ we find that $\rho_{(l)} = \frac{\beta}{\alpha} \rho_{(s)}^2$. so

$$\rho_{(s)} = \frac{4}{3} \frac{\alpha}{\beta} \quad (5.7)$$

and

$$\rho_{(l)} = \frac{16}{9} \frac{\alpha}{\beta} \quad (5.8)$$

Now that we have an expression for the molar densities in terms of the experimental parameters we will be able to calculate more meaningful quantities. For example,

$$\Delta S_{l \rightarrow s} = \frac{s_{(s)} - s_{(l)}}{n} = \frac{1}{T^2} \left(\frac{\beta}{3} \rho_{(s)}^2 - \frac{\alpha}{2} \rho_{(l)} \right) = -\frac{8}{27} \frac{\alpha^2}{\beta T^2}$$

Where we used the expressions for the molar densities previously obtained and the fact that,

$$s_i = - \left(\frac{\partial A_i}{\partial T} \right)_{n,V}$$

Similarly, rearranging the definition for the molar density, $V_i = n/\rho_i$,

$$\Delta v_{l \rightarrow s} = \frac{3}{16} \frac{\beta}{\alpha}.$$

Notice that until now we have been setting up the stage to obtain a Clausius-Clapeyron equation for our system. In order to proceed we now need to determine P_s .

$$P_s = P_{(l)} = P_{(s)} = \frac{1}{2} \frac{\alpha}{T} \rho_{(l)}^2 = \frac{128}{81} \frac{\alpha^3}{\beta^2 T}$$

From this we are able to calculate dP_s/dT , and in turn are able to confirm the fact that

$$\frac{dP_s}{dT} = \frac{\Delta S_{l \rightarrow s}}{\Delta v}.$$

Chapter 6

Super Conductivity

6.1 Ginzburg-Landau Theory

3.1

According to the Ginzburg-Landau theory the conditional canonical potential associated with a magnetization pattern $M(r)$ is

$$\phi(M) = \phi_0 - \int_V d^3r (atM^2 + bM^4 + c|\nabla m|^2). \quad (6.1)$$

For temperatures not too close to the critical temperature, the observed magnetization pattern would be one that maximizes the value of $\phi(M)$. In a previous section we showed that the maximizing function, with any given boundary conditions, is a solution of a differential equation of the form (here written in terms of magnetization)

$$-c\nabla^2 M + atM + 2bM^3 = 0 \quad (6.2)$$

Here we will consider the case in which the temperature is below the critical temperature, that is, $t < 0$. In this case the absolute minimum of $\phi(M)$ is given by the trivial solution of equation(6.2) in which $M(r) = M_0$ is a constant. The value of M_0 is given by equation(6.2) without the gradient term

$$-a|t|M_0 + 2bM_0^3 = 0 \quad (6.3)$$

or

$$M_0 = \pm \sqrt{\frac{a|t|}{2b}} \quad (6.4)$$

It is easy to see that the other trivial solution to equation(6.3), namely $M_0 = 0$, gives a local minimum, rather than a maximum.

A more interesting solution of equation(6.2) is the one that describes a two phase state. To describe such a state, one can

take the function $M(x)$ that depends only on \mathbf{x} and has the form describing a domain wall.

First of all how about we set out to determine an exact solution to equation(6.2). Assume that $M(-x) = -M(x)$ and that $M(x) \rightarrow M_0$ as $x \rightarrow \infty$.

If M as solely a dependence on x , then for $t < 0$, equation(6.2) becomes

$$-c \frac{\partial^2 M}{\partial x^2} - a|t|M + 2bM^3 = 0$$

Chapter 7

Statistical Mechanics

7.1 Two State Systems

An ensemble is a collection of all the possible configurations of the system. In this chapter we will focus on two level systems. Thus we look at the question of what does the ensemble for multiple coins look like? For a system of spins? etc.

This problem is tractable in the fact that we can enumerate the possible outcomes.

Assume we have four coins. These coins are "classical" coins, watch one can be distinguished from the other three by putting a label on it.

In this case we have five different macro-states. After being thrown all four coins could have landed heads, this is one macro-state, all but one could have landed heads, this is another macro-state, only two coins landed on heads, only one coin landed heads or no coin landed heads.

Each one of these macro-states is composed of multiple microstates. Microstates do not represent observable quantities such as the macrostates. These instead represent the underlying detail. For example, one can measure the average magnetization of a material, the given magnetization would correspond to the macroscopic state of the system. The underlying atomic detail of the system, the orientations of each of the spins.

The number of possible microstates for any given macrostate can be calculated using the binomial coefficient $\binom{N}{n}$.

SEE APPENDIX (include discussion on pascal's triangle, play with cards, etc.)

So for example, if we want to calculate the number of microstates which have on average 400 spins pointing up out of 1000 atoms that would be equal to

$$\binom{1000}{400} = \frac{1000!}{400!600!}$$

From a more fundamental point of view, here the binomial coefficient corresponds to the degeneracy of the system, a sort of symmetry. How many ways can I arrange my systems so that I still have the same experimental result?

Notice that the degeneracy is called by different symbols in different places, here are some common notations:

$$\binom{N}{n} \quad g_i \quad \omega_i \quad \Omega(\epsilon)$$

For a spin up spin down non-interacting system the degeneracy is two.

If we wanted to generalize the previous result to a multinomial, for example a cube. Then for each face we could get n_1, n_2, n_3, n_4, n_5 , and n_6 , although because the system only needs five degrees of freedom to be completely determined we obtain

$$\binom{N}{n_1 n_2 n_3 n_4 n_5} = \frac{N!}{n_1! n_2! n_3! n_4! n_5! (N - n_1 - n_2 - n_3 - n_4 - n_5)!}$$

Going back to the spin system with four atoms. Recall the results we already obtained. Let's take the spin up to have an energy $+\epsilon$ and the spin down atom to have an energy of zero. We are free to choose our ground state as we desired for most of the time we are only interested in measuring energy differences.

Thus, we would have the following macrostates:

One state with $E = 4\epsilon$.

Four states with $E = 3\epsilon$.

Six states with $E = 2\epsilon$.

Four states with $E = \epsilon$.

One state with $E = 0$.

$$U = \sum_i n_i \epsilon_i$$

For a system at equilibrium the entropy will be maximized such that

$$\begin{aligned} \frac{S(N, E)}{k_b} &= \log N! - \log N_1! - \log(N - N_1)! \\ &= N \log N - N_1 \log N_1 - (N - N_1) \log(N - N_1) \end{aligned}$$

Here we used Stirling's approximation along with the result from the previous section for the number of macrostates.

Notice that we can define $N_1 = E/\epsilon$. So the entropy can be expressed as

$$\frac{S(N, E)}{k_b} = N \log N - \frac{NE}{N\epsilon} \log \frac{NE}{N\epsilon} - \left(N - \frac{NE}{N\epsilon}\right) \log \left(N - \frac{NE}{N\epsilon}\right)$$

In order to simplify the result lets write the previous expression in terms of the relative occupation $x = E/N\epsilon$. Hence,

$$\begin{aligned} \frac{S(N, E)}{k_b} &= N \left(\log N - \frac{E}{N\epsilon} \log \left(\frac{NE}{N\epsilon} \right) - \left(1 - \frac{E}{N\epsilon} \right) \log \left(N \left(1 - \frac{E}{N\epsilon} \right) \right) \right) \\ &= N (\log N - x \log (Nx) - (1 - x) \log (N(1 - x))) \end{aligned}$$

So,

$$\frac{S(N, E)}{k_b} = -N (x \log x + (1 - x) \log (1 - x))$$

7.2 Equations of State

Ultra-relativistic Gas

Dispersion relation

$$\epsilon = pc. \tag{7.1}$$

$$Q_N = \frac{1}{N!} Q_1^N \tag{7.2}$$

$$Q_1 = \frac{1}{h^3} \int d^3p d^3q e^{-\beta pc} = \frac{V}{h^3} \int 4\pi p^2 dp e^{-\beta pc}$$

Using the fact that, (see Appendix A)

$$\int dx x^n e^{-x} = n!$$

$$Q_1 = \frac{8\pi V}{(\beta hc)^3}$$

and so

$$Q_N = \frac{1}{N!} \left(8\pi V \left(\frac{kT}{hc} \right)^3 \right)^N \tag{7.3}$$

How does this compare to what we know from thermodynamics? For starters, given the fact that

$$A = -k_b t \log Q_N \tag{7.4}$$

$$P = - \left(\frac{\partial A}{\partial V} \right)_{N,T}$$

Also,

$$U = kT^2 \left(\frac{\partial \log Q}{\partial T} \right)_{N,V}$$

$$S = k \log Q + kT \left(\frac{\partial \log Q}{\partial T} \right)_{N,V}$$

$$c_v = \frac{\partial U}{\partial T}$$

and

$$c_P = \frac{\partial (U + PV)}{\partial T}$$

Chapter 8

Bose-Einstein

8.1 Black Body Radiation

It is a fact of everyday life that hot objects lose energy by the process of radiation. In this case we can regard radiation as a gas of photon. Because photons do not interact with each other but only mediate interactions where fermions are couple to the electromagnetic field, it is possible to approximate their density of states by taking into consideration a number of wave functions in a box.

Solving the one dimensional box potential one arrives at the result that the wave number k has the following restrictions,

$$k = \frac{2\pi}{\lambda} = \frac{n\pi}{L}, n = 1, 2, 3, \dots \quad (8.1)$$

Which is equivalent to,

$$n = \frac{2L}{\lambda}. \quad (8.2)$$

Expanding this result to three dimensions,

$$n^2 = (n_x^2 + n_y^2 + n_z^2). \quad (8.3)$$

The quantum numbers n_i must be positive thus we restrict the possible solutions to the top right quadrant of a Cartesian plane, the zone where all $n_i > 0$, that is an eighth of the three dimensional plane.

Putting all together we are able to calculate the number density,

$$g(\lambda)d\lambda = \gamma \frac{1}{8} 4\pi n^2 dn = \gamma \left(\frac{\pi}{2}\right) \left(\frac{2L}{\lambda}\right)^2 \left(\frac{2L}{\lambda^2} d\lambda\right) \quad (8.4)$$

Here we used the fact that

$$d|n| = \frac{2L}{\lambda^2} d|\lambda| \quad (8.5)$$

In order to calculate the energy density we need a couple facts. First of all, the degeneracy constant γ is equal to two due to the fact that photons have only two (transverse) modes of polarization. Furthermore we need to recall that photons have integer spins and thus follow Bose-Einstein statistics.

Hence,

$$u(\lambda)d\lambda = E_\gamma f(\lambda)g(\lambda)d\lambda \quad (8.6)$$

Here E_γ is the energy per photon and $f(\lambda)$ is the bose-einstein distribution of states.

Thus,

$$\begin{aligned} u(\lambda)d\lambda &= \left(\frac{hc}{\lambda}\right) \left(\frac{1}{e^{hc/\lambda k_b T} - 1}\right) \left[\pi \left(\frac{2^3 L^3}{\lambda^4}\right) d\lambda\right] \\ &= 8\pi hcV \frac{d\lambda}{\lambda^5 (e^{hc/\lambda k_b T} - 1)} \end{aligned}$$

The key to these sort of problems is that once we have an integral in terms of desired observables we define a dimensionless order parameter that will simplify the problem.

Looking at the argument in the distribution we define $x = hc/\lambda k_b T$. Making this substitution,

$$\frac{U}{V} = \frac{8\pi}{h^3 c^3} (k_b T)^4 \int_0^\infty \frac{x^3 dx}{e^x - 1} = \frac{8\pi^5 k_b^4}{15 h^3 c^3} T^4 = a T^4 \quad (8.7)$$

One can see that indeed the total radiation energy is proportional to the fourth power of the absolute temperature.

Taking a second look at the energy density,

$$u(\lambda)d\lambda = 8\pi hcV \frac{d\lambda}{\lambda^5 (e^{hc/\lambda k_b T} - 1)} \quad (8.8)$$

It is important to notice that for long wavelengths, that is for, $hc/\lambda \ll k_b T$ we can approximate eq(8.8) with a two term expansion of the exponential term, the first term becomes negligible to the second and we arrive obtain the Rayleigh-Jeans formula

$$u(\lambda)d\lambda = V \frac{8\pi k_b T}{\lambda^4} d\lambda \quad (8.9)$$

This formula exhibits an ultraviolet catastrophe, for as the wavelength approaches zero the energy density approaches infinity.

On the other hand, in the limit of short wavelengths,

$$e^{hc/\lambda k_b T} \gg 1.$$

So,

$$u(\lambda)d\lambda = V \frac{8\pi hc}{\lambda^5} e^{-hc/\lambda k_b T} d\lambda \quad (8.10)$$

This last expression is the so called Wien's law.

More Realistic Models

The flaw in the derivation of the Rayleigh-Jeans law is the classical assumption that the radiation modes vary continuously, which implies there is an average energy per mode. Planck recognized this flaw and instead used the fact that energy is quantized in such a manner that the eigenvalues of energy have are described as follows,

$$E = nh\nu. \quad (8.11)$$

Here n is the quantum number describing how far a given quantum of energy is from the ground state.

Using the formalism from statistical mechanics we are then able to calculate the average energy for a black body,

$$\bar{E} = \frac{\sum_{n_s}^{\infty} E_{n_s} e^{\beta E_{n_s}}}{\sum_{n_s}^{\infty} e^{\beta E_{n_s}}} = \frac{\sum h\nu_i e^{\beta h\nu_i}}{\sum e^{\beta h\nu_i}} \quad (8.12)$$

Here $\beta (= 1/k_b T)$ is the inverse temperature and the sums on eq(8.12) are over all quantum states.

Furthermore, one can see that in the numerator of the previous expression the energy occurs twice. Simplifying this we obtain,

$$\bar{E} = \frac{\frac{-\partial}{\partial(\beta)} \sum e^{\beta h\nu_i}}{\sum e^{\beta h\nu_i}} = -\frac{\partial}{\partial(\beta)} \log \left(\sum e^{\beta h\nu_i} \right).$$

It is important to recognize that the term inside the logarithm is a geometric series and so,

$$\sum_i^{\infty} e^{\beta h\nu_i} = \frac{1}{1 - e^{-\beta h\nu}}. \quad (8.13)$$

Using this fact, we arrive at the final result,

$$\bar{E} = \frac{\partial}{\partial(\beta)} \log (1 - e^{-\beta h\nu}) = \frac{h\nu e^{-\beta h\nu}}{1 - e^{-\beta h\nu}} = h\nu \frac{1}{e^{\beta h\nu} - 1}$$

citing things

Bibliography

- [1] F. Reif, *Fundamentals of Statistical and Thermal Physics*.