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Chemistry in High Schools

By Eugene Paul Schoch

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Excerpt: .and one other product (salt); while there is no action with the metals below hydrogen in the table. (b) Oxidizing acids (nitric acid, concentrated sulphuric acid) act on almost all metals except the lowest in the table (gold, platinum); but hydrogen does not appear as one of the products, and the action is altogether more complex than any action with non-oxidizing acids. The above general facts (a and b) should be learned and retained by the student. He should realize that by this means he learns a great number of important facts easily. In connection with (a) the following quantitative data is given: $\text{Fe} + 2\text{HCl} = \text{FeCl}_2 + 2\text{H}$ $\text{Zn} + 2\text{HCl} = \text{ZnCl}_2 + 2\text{H}$ $\text{Al} + 3\text{HCl} = \text{AlCl}_3 + 3\text{H}$ $\text{Fe} + \text{H}_2\text{SO}_4 = \text{FeSO}_4 + 2\text{H}$ $\text{Zn} + \text{H}_2\text{SO}_4 = \text{ZnSO}_4 + 2\text{H}$ $2\text{Al} + 3\text{H}_2\text{SO}_4 = \text{Al}_2(\text{SO}_4)_3 + 6\text{H}$. These equations should be committed to memory and drilled on by means of numerical problems. 2. Preparation and collection of the gas to show some of its common properties. Perform as given in...



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