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The Freezing-Point Lowering, Conductivity, and Viscosity of Solutions of Certain Electrolytes in Water Volume 180; Methyl Alcohol, Ethyl Alcohol, Acetone, and Glycerol, and in Mixtures of These Solvents with One Another

By Harry Clary Jones

Rarebookclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1913 Excerpt: . gives the conductivities of sodium bromide in methyl alcohol at both 0 and 25. At 0 we have complete dissociation probably at 6,400 liters. At 25 complete dissociation is not attained until a dilution of 12,800 liters is reached. Table 74 gives the conductivities of sodium bromide in ethyl alcohol at 0 and 25. Complete dissociation is reached at both temperatures at a dilution of 12,800 liters. At the volume 102,400 there is a marked decrease in conductivity, probably due to the great solvation at this extremely high dilution. The same fact may be noticed in several other tables where ethyl alcohol is used as a solvent. Table 74 gives the conductivity of lithium bromide in methyl alcohol. At 0 complete dissociation probably is reached at 6,400 liters. At 25 there is no maximum in conductivity, but the rate of increase is much smaller above...



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