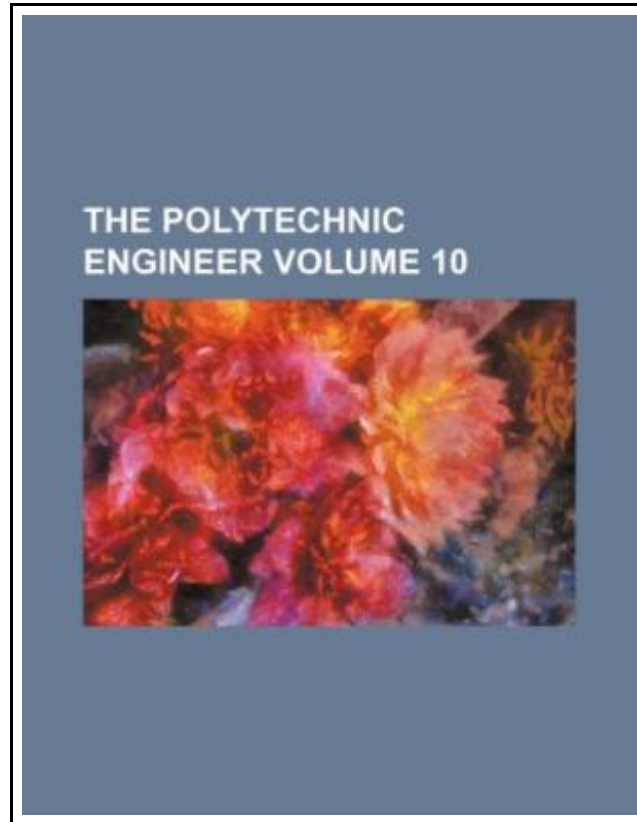


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Not Avail, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1910 Excerpt: .acetate whose chemical action is essentially the same as caustic soda and dried sodium acetate heated together, thus $\text{CH}_3\text{CO} \cdot \text{ONa} + \text{NaOH} = \text{CH}_4 + \text{Na}_2\text{CO}_3$. The thoroughly dried gas was passed over the previously dried metallic oxides placed in the bend of a U-shaped Marchand tube immersed in an oil bath. The bulb of the mercury thermometer was held in position against the bend of the tube. The following table shows the temperature at which dry ammonia gas and dry methane began to take oxygen from the metallic oxides and reduce them to the metallic state: III.--Table of Reductions with Ammonia and with Methane. Oxide. Formula. Temperature with Temperature with W 13. CrI4. Mercurous oxide Hg, O 67C. 220C. Mercuric oxide HgO 157 200-210 Cuprous oxide Cu2O 208 230 Oxide. Formula. Temperature with Temperature with Cupric oxide CuO 225 280 Lead suboxide Pb2O 202 202 Lead oxide PbO 299 210 Lead dioxide PbO2 198 45 Red lead Pb3O4 above 300 158 Zinc oxide ZnO 233 152-159 By examination of the above table, it will be observed that in each case the lowest oxide of the metal releases oxygen more easily or at a lower temperature than the next higher oxide; with the exception of mercurous oxide when reduced by methane; that in the case of mercury and copper oxides, the ammonia gas has a reducing power at a lower temperature than methane. The oxides of lead and zinc, however, give up oxygen more easily...



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