

TABEL PERIODIK HTML CSS DAN JS

1. HTML

```
<!DOCTYPE html>
<html lang="id">
  <head>
    <meta charset="UTF-8">
    <link rel="stylesheet" href="html.css">
  </head>
</html><div id="status" style="display:none;position:fixed;background:rgba(0, 0, 0,
0.75);color:white;top:0;left:0;width:100%;height:auto;z-index:5;text-align:center;"></div>
<table>
  <caption id="title">
    TABEL PERIODIK UNSUR-UNSUR KIMIA</caption>
  <thead>
    <tr class="gol">
      <th class="indiv">Golongan
        <br/>Periodik</th>
      <th class="gol">
        <div class="content">1A
          <br/>Alkali</div>
      </th>
      <th class="gol">
        <div class="content">2A
          <br/>Alkali Tanah</div>
      </th>
      <th class="gol">
        <div class="content">3B</div>
      </th>
      <th class="gol">
        <div class="content">4B</div>
      </th>
      <th class="gol">
        <div class="content">5B</div>
      </th>
      <th class="gol">
        <div class="content">6B</div>
      </th>
      <th class="gol">
        <div class="content">7B</div>
      </th>
      <th class="gol" colspan="3">8B</th>
      <th class="gol">
        <div class="content">1B</div>
      </th>
      <th class="gol">
        <div class="content">2B</div>
      </th>
      <th class="gol">
        <div class="content">3A</div>
      </th>
```

```

<th class="gol">
  <div class="content">4A</div>
</th>
<th class="gol">
  <div class="content">5A</div>
</th>
<th class="gol">
  <div class="content">6A</div>
</th>
<th class="gol">
  <div class="content">7A
    <br/>Halogen</div>
</th>
<th class="gol">
  <div class="content">8A
    <br/>Gas Mulia</div>
</th>
</tr>
</thead>
<tbody>
<tr>
  <th class="per">
    <div class="content-per">1</div>
  </th>
  <td class="g">
    <div class="content-tp">
      <div class="na">1</div>
      <div class="ma">1.008</div>
      <br/>
      <div class="lu">H</div>
      <div class="nu">Hidrogen</div>
    </div>
  </td>
  <td colspan="16"></td>
  <td class="gm">
    <div class="content-tp">
      <div class="na">2</div>
      <div class="ma">4.003</div>
      <br/>
      <div class="lu">He</div>
      <div class="nu">Helium</div>
    </div>
  </td>
</tr>
<tr>
  <th class="per">
    <div class="content-per">2</div>
  </th>
  <td class="l">
    <div class="content-tp">
      <div class="na">3</div>

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        <div class="ma">6.939</div>
        <br/>
        <div class="lu">Li</div>
        <div class="nu">Litium</div>
    </div>
</td>
<td class="l">
    <div class="content-tp">
        <div class="na">4</div>
        <div class="ma">9.012</div>
        <br/>
        <div class="lu">Be</div>
        <div class="nu">Berilium</div>
    </div>
</td>
<td colspan="10"></td>
<td class="m">
    <div class="content-tp">
        <div class="na">5</div>
        <div class="ma">10.811</div>
        <br/>
        <div class="lu">B</div>
        <div class="nu">Boron</div>
    </div>
</td>
<td class="l">
    <div class="content-tp">
        <div class="na">6</div>
        <div class="ma">12.011</div>
        <br/>
        <div class="lu">C</div>
        <div class="nu">Karbon</div>
    </div>
</td>
<td class="g">
    <div class="content-tp">
        <div class="na">7</div>
        <div class="ma">14.007</div>
        <br/>
        <div class="lu">N</div>
        <div class="nu">Nitrogen</div>
    </div>
</td>
<td class="g">
    <div class="content-tp">
        <div class="na">8</div>
        <div class="ma">15.999</div>
        <br/>
        <div class="lu">O</div>
        <div class="nu">Oksigen</div>
    </div>

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</td>
<td class="h">
  <div class="content-tp">
    <div class="na">9</div>
    <div class="ma">18.998</div>
    <br/>
    <div class="lu">F</div>
    <div class="nu">Fluor</div>
  </div>
</td>
<td class="gm">
  <div class="content-tp">
    <div class="na">10</div>
    <div class="ma">20.18</div>
    <br/>
    <div class="lu">Ne</div>
    <div class="nu">Neon</div>
  </div>
</td>
</tr>
<tr>
  <th class="per">
    <div class="content-per">3</div>
  </th>
  <td class="l">
    <div class="content-tp">
      <div class="na">11</div>
      <div class="ma">22.989</div>
      <br/>
      <div class="lu">Na</div>
      <div class="nu">Natrium</div>
    </div>
  </td>
  <td class="l">
    <div class="content-tp">
      <div class="na">12</div>
      <div class="ma">24.305</div>
      <br/>
      <div class="lu">Mg</div>
      <div class="nu">Magnesium</div>
    </div>
  </td>
  <td colspan="10"></td>
  <td class="l">
    <div class="content-tp">
      <div class="na">13</div>
      <div class="ma">26.981</div>
      <br/>
      <div class="lu">Al</div>
      <div class="nu">Aluminium</div>
    </div>
  </td>

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</td>
<td class="m">
  <div class="content-tp">
    <div class="na">14</div>
    <div class="ma">28.086</div>
    <br/>
    <div class="lu">Si</div>
    <div class="nu">Silikon</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">15</div>
    <div class="ma">30.974</div>
    <br/>
    <div class="lu">P</div>
    <div class="nu">Fosfor</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">16</div>
    <div class="ma">32.066</div>
    <br/>
    <div class="lu">S</div>
    <div class="nu">Belerang</div>
  </div>
</td>
<td class="h">
  <div class="content-tp">
    <div class="na">17</div>
    <div class="ma">35.453</div>
    <br/>
    <div class="lu">Cl</div>
    <div class="nu">Klor</div>
  </div>
</td>
<td class="gm">
  <div class="content-tp">
    <div class="na">18</div>
    <div class="ma">39.948</div>
    <br/>
    <div class="lu">Ar</div>
    <div class="nu">Argon</div>
  </div>
</td>
</tr>
<tr>
  <th class="per">
    <div class="content-per">4</div>
  </th>

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```
<td class="1">
  <div class="content-tp">
    <div class="na">19</div>
    <div class="ma">39.098</div>
    <br/>
    <div class="lu">K</div>
    <div class="nu">Kalium</div>
  </div>
</td>
<td class="1">
  <div class="content-tp">
    <div class="na">20</div>
    <div class="ma">40.076</div>
    <br/>
    <div class="lu">Ca</div>
    <div class="nu">Kalsium</div>
  </div>
</td>
<td class="1">
  <div class="content-tp">
    <div class="na">21</div>
    <div class="ma">44.956</div>
    <br/>
    <div class="lu">Sc</div>
    <div class="nu">Skandium</div>
  </div>
</td>
<td class="1">
  <div class="content-tp">
    <div class="na">22</div>
    <div class="ma">47.88</div>
    <br/>
    <div class="lu">Ti</div>
    <div class="nu">Titanium</div>
  </div>
</td>
<td class="1">
  <div class="content-tp">
    <div class="na">23</div>
    <div class="ma">50.942</div>
    <br/>
    <div class="lu">V</div>
    <div class="nu">Vanadium</div>
  </div>
</td>
<td class="1">
  <div class="content-tp">
    <div class="na">24</div>
    <div class="ma">51.996</div>
    <br/>
    <div class="lu">Cr</div>
```

```
<div class="nu">Kromium</div>
</div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">25</div>
    <div class="ma">54.938</div>
    <br/>
    <div class="lu">Mn</div>
    <div class="nu">Mangan</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">26</div>
    <div class="ma">55.847</div>
    <br/>
    <div class="lu">Fe</div>
    <div class="nu">Besi</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">27</div>
    <div class="ma">58.933</div>
    <br/>
    <div class="lu">Co</div>
    <div class="nu">Kobalt</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">28</div>
    <div class="ma">58.69</div>
    <br/>
    <div class="lu">Ni</div>
    <div class="nu">Nikel</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">29</div>
    <div class="ma">63.546</div>
    <br/>
    <div class="lu">Cu</div>
    <div class="nu">Tembaga</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">30</div>
```

```
<div class="ma">65.39</div>
<br/>
<div class="lu">Zn</div>
<div class="nu">Seng</div>
</div>
</td>
<td class="c">
  <div class="content-tp">
    <div class="na">31</div>
    <div class="ma">69.723</div>
    <br/>
    <div class="lu">Ga</div>
    <div class="nu">Galium</div>
  </div>
</td>
<td class="m">
  <div class="content-tp">
    <div class="na">32</div>
    <div class="ma">72.61</div>
    <br/>
    <div class="lu">Ge</div>
    <div class="nu">Germanium</div>
  </div>
</td>
<td class="m">
  <div class="content-tp">
    <div class="na">33</div>
    <div class="ma">74.922</div>
    <br/>
    <div class="lu">As</div>
    <div class="nu">Arsen</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">34</div>
    <div class="ma">78.96</div>
    <br/>
    <div class="lu">Se</div>
    <div class="nu">Selenium</div>
  </div>
</td>
<td class="h">
  <div class="content-tp">
    <div class="na">35</div>
    <div class="ma">79.904</div>
    <br/>
    <div class="lu">Br</div>
    <div class="nu">Bromin</div>
  </div>
</td>
```



```
<td class="gm">
  <div class="content-tp">
    <div class="na">36</div>
    <div class="ma">83.8</div>
    <br/>
    <div class="lu">Kr</div>
    <div class="nu">Krypton</div>
  </div>
</td>
</tr>
<tr>
  <th class="per">5</th>
  <td class="l">
    <div class="content-tp">
      <div class="na">37</div>
      <div class="ma">85.4678</div>
      <br/>
      <div class="lu">Rb</div>
      <div class="nu">Rubidium</div>
    </div>
  </td>
  <td class="l">
    <div class="content-tp">
      <div class="na">38</div>
      <div class="ma">87.62</div>
      <br/>
      <div class="lu">Sr</div>
      <div class="nu">Strontium</div>
    </div>
  </td>
  <td class="l">
    <div class="content-tp">
      <div class="na">39</div>
      <div class="ma">88.906</div>
      <br/>
      <div class="lu">Y</div>
      <div class="nu">Yttrium</div>
    </div>
  </td>
  <td class="l">
    <div class="content-tp">
      <div class="na">40</div>
      <div class="ma">91.224</div>
      <br/>
      <div class="lu">Zr</div>
      <div class="nu">Zirkonium</div>
    </div>
  </td>
  <td class="l">
    <div class="content-tp">
      <div class="na">41</div>
```



```
<td class="l">
  <div class="content-tp">
    <div class="na">47</div>
    <div class="ma">107.868</div>
    <br/>
    <div class="lu">Ag</div>
    <div class="nu">Perak</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">48</div>
    <div class="ma">112.41</div>
    <br/>
    <div class="lu">Cd</div>
    <div class="nu">Kadmium</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">49</div>
    <div class="ma">114.82</div>
    <br/>
    <div class="lu">In</div>
    <div class="nu">Indium</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">50</div>
    <div class="ma">118.71</div>
    <br/>
    <div class="lu">Sn</div>
    <div class="nu">Timah</div>
  </div>
</td>
<td class="m">
  <div class="content-tp">
    <div class="na">51</div>
    <div class="ma">121.76</div>
    <br/>
    <div class="lu">Sb</div>
    <div class="nu">Antimon</div>
  </div>
</td>
<td class="m">
  <div class="content-tp">
    <div class="na">52</div>
    <div class="ma">127.6</div>
    <br/>
    <div class="lu">Te</div>
```

```

        <div class="nu">Telurium</div>
    </div>
</td>
<td class="h">
    <div class="content-tp">
        <div class="na">53</div>
        <div class="ma">126.904</div>
        <br/>
        <div class="lu">I</div>
        <div class="nu">Yodium</div>
    </div>
</td>
<td class="gm">
    <div class="content-tp">
        <div class="na">54</div>
        <div class="ma">131.29</div>
        <br/>
        <div class="lu">Xe</div>
        <div class="nu">Xenon</div>
    </div>
</td>
</tr>
<tr>
    <th class="per">
        <div class="content-per">6</div>
    </th>
    <td class="c">
        <div class="content-tp">
            <div class="na">55</div>
            <div class="ma">132.905</div>
            <br/>
            <div class="lu">Cs</div>
            <div class="nu">Sesium</div>
        </div>
    </td>
    <td class="l">
        <div class="content-tp">
            <div class="na">56</div>
            <div class="ma">137.327</div>
            <br/>
            <div class="lu">Ba</div>
            <div class="nu">Barium</div>
        </div>
    </td>
    <td class="lan">
        <div class="content-tp">
            <div class="na">57-71</div>
            <div class="ma">138-174</div>
            <br/>
            <div class="nu">Rangkaian Lantanida</div>
            <br/>

```

```
<div class="nu-desc">Nomor 57 sampai 71</div>
</div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">72</div>
    <div class="ma">178.49</div>
    <br/>
    <div class="lu">Hf</div>
    <div class="nu">Hafnium</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">73</div>
    <div class="ma">180.947</div>
    <br/>
    <div class="lu">Ta</div>
    <div class="nu">Tantalum</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">74</div>
    <div class="ma">137.327</div>
    <br/>
    <div class="lu">Ba</div>
    <div class="nu">Barium</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">75</div>
    <div class="ma">186.207</div>
    <br/>
    <div class="lu">Re</div>
    <div class="nu">Renyum</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">76</div>
    <div class="ma">190.2</div>
    <br/>
    <div class="lu">Os</div>
    <div class="nu">Osmium</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">77</div>
```

```
<div class="ma">192.22</div>
<br/>
<div class="lu">Ir</div>
<div class="nu">Iridium</div>
</div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">78</div>
    <div class="ma">195.08</div>
    <br/>
    <div class="lu">Pt</div>
    <div class="nu">Platina</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">79</div>
    <div class="ma">196.97</div>
    <br/>
    <div class="lu">Au</div>
    <div class="nu">Emas</div>
  </div>
</td>
<td class="c">
  <div class="content-tp">
    <div class="na">80</div>
    <div class="ma">200.59</div>
    <br/>
    <div class="lu">Hg</div>
    <div class="nu">Air Raksa</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">81</div>
    <div class="ma">204.37</div>
    <br/>
    <div class="lu">Tl</div>
    <div class="nu">Talium</div>
  </div>
</td>
<td class="l">
  <div class="content-tp">
    <div class="na">82</div>
    <div class="ma">207.2</div>
    <br/>
    <div class="lu">Pb</div>
    <div class="nu">Timbal</div>
  </div>
</td>
```

```

<td class="l">
  <div class="content-tp">
    <div class="na">83</div>
    <div class="ma">208.98</div>
    <br/>
    <div class="lu">Bi</div>
    <div class="nu">Bismut</div>
  </div>
</td>
<td class="m">
  <div class="content-tp" title="Elemen ini termasuk dalam kategori Radioaktif">
    <div class="na">84</div>
    <div class="ma">(209)</div>
    <br/>
    <div class="lu">Po</div>
    <div class="nu">Barium</div>
  </div>
</td>
<td class="h">
  <div class="content-tp" title="Elemen ini termasuk dalam kategori Radioaktif">
    <div class="na">85</div>
    <div class="ma">(210)</div>
    <br/>
    <div class="lu">At</div>
    <div class="nu">Astatin</div>
  </div>
</td>
<td class="gm">
  <div class="content-tp">
    <div class="na">86</div>
    <div class="ma">(222)</div>
    <br/>
    <div class="lu">Rn</div>
    <div class="nu">Radon</div>
  </div>
</td>
</tr>
<tr>
  <th class="per">
    <div class="content-per">7</div>
  </th>
  <td class="c">
    <div class="content-tp" title="Elemen ini termasuk dalam kategori Radioaktif">
      <div class="na">87</div>
      <div class="ma">(223.02)</div>
      <br/>
      <div class="lu">Fr</div>
      <div class="nu">Fransium</div>
    </div>
  </td>
  <td class="l">

```

```

<div class="content-tp" title="Elemen ini termasuk dalam kategori Radioaktif">
  <div class="na">88</div>
  <div class="ma">(226)</div>
  <br/>
  <div class="lu">Ra</div>
  <div class="nu">Radium</div>
</div>
</td>
<td class="akt">
  <div class="content-tp" title="Semua elemen Aktinida termasuk dalam kategori Radioaktif">
    <div class="na">89-103</div>
    <div class="ma">227-262</div>
    <br/>
    <div class="nu">Rangkaian Aktinida</div>
    <br/>
    <div class="nu-desc">Nomor 89 sampai 103</div>
  </div>
</td>
<td class="b" title="Elemen ini termasuk dalam kategori Radioaktif">
  <div class="content-tp">
    <div class="na">104</div>
    <div class="ma">[267]</div>
    <br/>
    <div class="lu">Rf</div>
    <div class="nu">Rutherfordium</div>
  </div>
</td>
<td class="b" title="Elemen ini termasuk dalam kategori Radioaktif">
  <div class="content-tp">
    <div class="na">105</div>
    <div class="ma">[268]</div>
    <br/>
    <div class="lu">Db</div>
    <div class="nu">Dubnium</div>
  </div>
</td>
<td class="b" title="Elemen ini termasuk dalam kategori Radioaktif">
  <div class="content-tp">
    <div class="na">106</div>
    <div class="ma">[269]</div>
    <br/>
    <div class="lu">Sg</div>
    <div class="nu">Seaborgium</div>
  </div>
</td>
<td class="b" title="Elemen ini termasuk dalam kategori Radioaktif">
  <div class="content-tp">
    <div class="na">107</div>
    <div class="ma">[270]</div>
    <br/>
    <div class="lu">Bh</div>

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        <div class="nu">Bohrium</div>
    </div>
</td>
<td class="b" title="Elemen ini termasuk dalam kategori Radioaktif">
    <div class="content-tp">
        <div class="na">108</div>
        <div class="ma">[269]</div>
        <br/>
        <div class="lu">Hs</div>
        <div class="nu">Hassium</div>
    </div>
</td>
<td class="undef" title="Elemen ini termasuk dalam kategori Radioaktif">
    <div class="content-tp">
        <div class="na">109</div>
        <div class="ma">[278]</div>
        <br/>
        <div class="lu">Mt</div>
        <div class="nu">Meitnerium</div>
    </div>
</td>
<td class="undef" title="Elemen ini termasuk dalam kategori Radioaktif">
    <div class="content-tp">
        <div class="na">110</div>
        <div class="ma">[281]</div>
        <br/>
        <div class="lu" title="Kode lain: Uun">Ds</div>
        <div class="nu" title="Nama lain: Ununnilium">Darmstadtium</div>
    </div>
</td>
<td class="undef">
    <div class="content-tp" title="Elemen ini termasuk dalam kategori Radioaktif">
        <div class="na">111</div>
        <div class="ma">[281]</div>
        <br/>
        <div class="lu" title="Kode lain: Uuu">Rg</div>
        <div class="nu" title="Nama lain: Unununium">Roentgenium</div>
    </div>
</td>
<td class="l">
    <div class="content-tp" title="Elemen ini termasuk dalam kategori Radioaktif">
        <div class="na">112</div>
        <div class="ma">[285]</div>
        <br/>
        <div class="lu" title="Kode lain: Uub">Cn</div>
        <div class="nu" title="Nama lain: Ununbium">Kopernesium</div>
    </div>
</td>
<td class="undef">
    <div class="content-tp">
        <div class="na">113</div>

```

```
<div class="ma">[284]</div>
<br/>
<div class="lu" title="Kode lain: -">Uut</div>
<div class="nu" title="Nama lain: -">Ununtrium</div>
</div>
</td>
<td class="undef">
  <div class="content-tp" title="Elemen ini termasuk dalam kategori Radioaktif">
    <div class="na">114</div>
    <div class="ma">[289]</div>
    <br/>
    <div class="lu" title="Kode lain: Uuq">Fl</div>
    <div class="nu" title="Nama lain: Ununquadium">Flerovium</div>
  </div>
</td>
<td class="undef">
  <div class="content-tp">
    <div class="na">115</div>
    <div class="ma">[288]</div>
    <br/>
    <div class="lu" title="Kode lain: -">Uup</div>
    <div class="nu" title="Nama lain: -">Ununpentium</div>
  </div>
</td>
<td class="undef">
  <div class="content-tp" title="Elemen ini termasuk dalam kategori Radioaktif">
    <div class="na">116</div>
    <div class="ma">[293]</div>
    <br/>
    <div class="lu" title="Kode lain: Uuh">Lv</div>
    <div class="nu" title="Nama lain: Ununheksium">Livermorium</div>
  </div>
</td>
<td class="undef">
  <div class="content-tp">
    <div class="na">117</div>
    <div class="ma">[294]</div>
    <br/>
    <div class="lu" title="Kode lain: -">Uus</div>
    <div class="nu" title="Nama lain: -">Ununseptium</div>
  </div>
</td>
<td class="undef">
  <div class="content-tp" title="Elemen ini termasuk dalam kategori Radioaktif">
    <div class="na">118</div>
    <div class="ma">[294]</div>
    <br/>
    <div class="lu" title="Kode lain: -">Uuo</div>
    <div class="nu" title="Nama lain: -">Ununoktium</div>
  </div>
</td>
```

```

</tr>
<tr>
  <th></th>
  <td colspan="18">
    <div class="content-tp"></div>
  </td>
</tr>
<tr>
  <th></th>
  <td></td>
  <td class="lan">
    <div class="content-tp">
      <div class="na">&nbsp;</div>
      <div class="ma">&nbsp;</div>
      <br/>
      <div class="nu rlra">Rangkaian Lantanida</div>
    </div>
  </td>
  <td class="lan">
    <div class="content-tp">
      <div class="na">57</div>
      <div class="ma">138.906</div>
      <br/>
      <div class="lu">La</div>
      <div class="nu">Lantanum</div>
    </div>
  </td>
  <td class="lan">
    <div class="content-tp">
      <div class="na">58</div>
      <div class="ma">140.115</div>
      <br/>
      <div class="lu">Ce</div>
      <div class="nu">Serium</div>
    </div>
  </td>
  <td class="lan">
    <div class="content-tp">
      <div class="na">59</div>
      <div class="ma">140.908</div>
      <br/>
      <div class="lu">Pr</div>
      <div class="nu">Praseodimium</div>
    </div>
  </td>
  <td class="lan">
    <div class="content-tp">
      <div class="na">60</div>
      <div class="ma">144.24</div>
      <br/>
      <div class="lu">Nd</div>
    </div>
  </td>

```

```

        <div class="nu">Neodimium</div>
    </div>
</td>
<td class="lan-b">
    <div class="content-tp">
        <div class="na">61</div>
        <div class="ma">(145)</div>
        <br/>
        <div class="lu">Pm</div>
        <div class="nu">Prometium</div>
    </div>
</td>
<td class="lan">
    <div class="content-tp">
        <div class="na">62</div>
        <div class="ma">150.36</div>
        <br/>
        <div class="lu">Sm</div>
        <div class="nu">Samarium</div>
    </div>
</td>
<td class="lan">
    <div class="content-tp">
        <div class="na">63</div>
        <div class="ma">151.96</div>
        <br/>
        <div class="lu">Eu</div>
        <div class="nu">Europium</div>
    </div>
</td>
<td class="lan">
    <div class="content-tp">
        <div class="na">64</div>
        <div class="ma">157.25</div>
        <br/>
        <div class="lu">Gd</div>
        <div class="nu">Gadolinium</div>
    </div>
</td>
<td class="lan">
    <div class="content-tp">
        <div class="na">65</div>
        <div class="ma">158.923</div>
        <br/>
        <div class="lu">Tb</div>
        <div class="nu">Terbium</div>
    </div>
</td>
<td class="lan">
    <div class="content-tp">
        <div class="na">66</div>

```

```
<div class="ma">162.5</div>
<br/>
<div class="lu">Dy</div>
<div class="nu">Disprosium</div>
</div>
</td>
<td class="lan">
  <div class="content-tp">
    <div class="na">67</div>
    <div class="ma">164.93</div>
    <br/>
    <div class="lu">Ho</div>
    <div class="nu">Holmium</div>
  </div>
</td>
<td class="lan">
  <div class="content-tp">
    <div class="na">68</div>
    <div class="ma">167.26</div>
    <br/>
    <div class="lu">Er</div>
    <div class="nu">Erbium</div>
  </div>
</td>
<td class="lan">
  <div class="content-tp">
    <div class="na">69</div>
    <div class="ma">168.93</div>
    <br/>
    <div class="lu">Tm</div>
    <div class="nu">Tulium</div>
  </div>
</td>
<td class="lan">
  <div class="content-tp">
    <div class="na">70</div>
    <div class="ma">173.04</div>
    <br/>
    <div class="lu">Yb</div>
    <div class="nu">Iterbium</div>
  </div>
</td>
<td class="lan">
  <div class="content-tp">
    <div class="na">71</div>
    <div class="ma">174.967</div>
    <br/>
    <div class="lu">Lu</div>
    <div class="nu">Lutelium</div>
  </div>
</td>
```

```

</tr>
<tr>
  <th></th>
  <td></td>
  <td class="akt">
    <div class="content-tp">
      <div class="na">&nbsp;</div>
      <div class="ma">&nbsp;</div>
      <br/>
      <div class="nu rlr">Rangkaian Aktinida</div>
    </div>
  </td>
  <td class="akt">
    <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
      <div class="na">89</div>
      <div class="ma">(227)</div>
      <br/>
      <div class="lu">Ac</div>
      <div class="nu">Aktinium</div>
    </div>
  </td>
  <td class="akt">
    <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
      <div class="na">90</div>
      <div class="ma">232.038</div>
      <br/>
      <div class="lu">Th</div>
      <div class="nu">Torium</div>
    </div>
  </td>
  <td class="akt">
    <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
      <div class="na">91</div>
      <div class="ma">231.035</div>
      <br/>
      <div class="lu">Pa</div>
      <div class="nu">Protaktinium</div>
    </div>
  </td>
  <td class="akt">
    <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
      <div class="na">92</div>
      <div class="ma">238.029</div>
      <br/>
      <div class="lu">U</div>
      <div class="nu">Uranium</div>
    </div>
  </td>
  <td class="akt-b">
    <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
      <div class="na">93</div>

```

```
<div class="ma">(237)</div>
<br/>
<div class="lu">Np</div>
<div class="nu">Neptunium</div>
</div>
</td>
<td class="akt-b">
  <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
    <div class="na">94</div>
    <div class="ma">(244)</div>
    <br/>
    <div class="lu">Pu</div>
    <div class="nu">Plutonium</div>
  </div>
</td>
<td class="akt-b">
  <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
    <div class="na">95</div>
    <div class="ma">(243)</div>
    <br/>
    <div class="lu">Am</div>
    <div class="nu">Amerisium</div>
  </div>
</td>
<td class="akt-b">
  <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
    <div class="na">96</div>
    <div class="ma">(247)</div>
    <br/>
    <div class="lu">Cm</div>
    <div class="nu">Kurium</div>
  </div>
</td>
<td class="akt-b">
  <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
    <div class="na">97</div>
    <div class="ma">(247)</div>
    <br/>
    <div class="lu">Bk</div>
    <div class="nu">Berkelium</div>
  </div>
</td>
<td class="akt-b">
  <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
    <div class="na">98</div>
    <div class="ma">(251)</div>
    <br/>
    <div class="lu">Cf</div>
    <div class="nu">Kalifornium</div>
  </div>
</td>
```

```

<td class="akt-b">
  <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
    <div class="na">99</div>
    <div class="ma">(252)</div>
    <br/>
    <div class="lu">Es</div>
    <div class="nu">Einsteinium</div>
  </div>
</td>
<td class="akt-b">
  <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
    <div class="na">100</div>
    <div class="ma">(257)</div>
    <br/>
    <div class="lu">Fm</div>
    <div class="nu">Fermium</div>
  </div>
</td>
<td class="akt-b">
  <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
    <div class="na">101</div>
    <div class="ma">(258)</div>
    <br/>
    <div class="lu">Md</div>
    <div class="nu">Mendelevium</div>
  </div>
</td>
<td class="akt-b">
  <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
    <div class="na">102</div>
    <div class="ma">(259)</div>
    <br/>
    <div class="lu">No</div>
    <div class="nu">Nobelium</div>
  </div>
</td>
<td class="akt-b">
  <div class="content-tp" title="Elemen ini termasuk ke dalam kategori Radioaktif">
    <div class="na">103</div>
    <div class="ma">(262)</div>
    <br/>
    <div class="lu">Lr</div>
    <div class="nu">Lawrencium</div>
  </div>
</td>
</tr>
</tbody>
</table>

```


2. CSS

```
body {
  font-size: 11pt;
  line-height: initial;
}

table {
  width: 96%;
  background: #607D8B;
  font-family: roboto;
  border-style: none;
  border-collapse: separate;
  border-spacing: 2px;
}

th {
  width: 400px;
  max-width: 400px;
  font-family: roboto;
  text-align: center;
}

abbr {
  text-decoration: underline;
}
/* Layouting */

#title {
  background-color: #455A64;
  color: white;
  text-align: center;
}

.indiv {
  font-size: 0.9em;
  font-weight: bold;
}

.content {
  width: 90px;
  max-width: 90px;
}

.content-per {
  height: 90px;
  max-height: 90px;
}

.content-tp {
  width: 90px;
  max-width: 90px;
  height: 90px;
  max-height: 90px;
}

.gol {
  font-size: 0.9em;
  font-weight: bold;
}

.per {
  font-size: 4em;
  font-weight: bold;
}
/* For periodic Elements Inside */

.na {
  font-size: 0.7em;
  float: left;
}

.ma {
  font-size: 0.7em;
  float: right;
}

.lu {
  font-size: 2.75em;
  font-weight: bold;
}
```

```

    text-align: right;
    width: 97%;
    margin-right: 5px;
}

.sp {
    font-size: 0.9em;
    float: left;
}

.nu {
    font-size: 10.1pt;
    text-align: center;
}

.rlra {
    font-size: 12pt;
    margin-top: 5px;
}
td
{
    box-shadow:0 1.5px 4px rgba(0, 0, 0, 0.24), 0 1.5px 6px rgba(0, 0, 0, 0.12)
}
td[colspan], td:empty
{
    box-shadow:none;
}
.nu-desc {
    font-size: 7pt;
    text-align: center;
}

.kotak {
    width: 85px;
    height: 85px;
}
/* CSS for Element Category Starts */

.l {
    background: #FFE33B;
}

.c {
    background: #2196F3;
}

.g {
    background: #4CAF50;
}

.b {
    background: #F44336;
}

.h {
    background: #9E9E9E;
}

.lan {
    background: #E91E63;
}

.lan-b {
    background: #F44336;
}

.akt {
    background: #9C27B0;
    color: #fff;
}

.akt-b {
    background: mediumpurple;
}

.m {
    background: brown;
}

.gm {

```

```
    background: LightskyBlue;
}

.undef {
    background: silver;
}
```

3. JS

```
function showstatus(text, color) {

    $('#status').stop().slideUp().css('color', color).html(text).slideDown().delay(2000).slideUp()
}

window.onbeforeunload = function () {

    if (navigator.onLine) {} else {

        showstatus('<marquee>No network connection. Please do not refresh until this message is
disappear</marquee>', 'red')

    }

}

window.oncontextmenu = function () {

    showstatus('Right click is disabled', '#F44336');

    return false

}

shortcut.add('CTRL+U', function () {

    showstatus('View source is disabled!', '#F44336')

}),

shortcut.add('CTRL+Shift+I', function () {

    showstatus('Inspect Element is disabled!', '#F44336')

}),

shortcut.add('CTRL+Shift+J', function () {

    showstatus('JS Console is disabled!', '#F44336')

}),

shortcut.add('CTRL+Shift+C', function () {

    showstatus('Inspect Element is disabled!', '#F44336')

}),

shortcut.add('F12', function () {

    showstatus('JS Console is disabled!', '#F44336')

}),

shortcut.add('Meta+Alt+U', function () {

    showstatus('View source is disabled!', '#F44336')

}),

shortcut.add('Meta+Alt+I', function () {

    showstatus('Inspect Element is disabled!', '#F44336')
```

```

)),

shortcut.add('Meta+Alt+J', function () {

    showstatus('JS Console is disabled!', '#F44336')

}),

shortcut.add('Meta+Shift+C', function () {

    showstatus('Inspect Element is disabled!', '#F44336')

}),

shortcut.add('Meta+P', function () {

    showstatus('Preparing to print...', 'white');

    setTimeout(function () {

        window.print()

    }, 3000)

}),

shortcut.add('Ctrl+P', function () {

    showstatus('Preparing to print...', 'white');

    setTimeout(function () {

        window.print()

    }, 3000)

});

```

Hasil

| Golongan Periodik | 1A Alkali | 2A Alkali Tanah | 3B | 4B | 5B | 6B | 7B | 8B | 9B | 10B | 11B | 12B | 3A | 4A | 5A | 6A | 7A Halogen | 8A Gas Mulia |
|-------------------|----------------|--------------------|---------------------|-----------------|------------------|------------------|-----------------|------------------|--------------------|-------------------|-------------------|------------------|-----------------|--------------------|-------------------|--------------------|-------------------|-----------------|
| 1 | H Hydrogen | | | | | | | | | | | | | | | | | He Helium |
| 2 | Li Lithium | Be Beryllium | | | | | | | | | | | B Boron | C Karbon | N Nitrogen | O Oksigen | F Fluor | Ne Neon |
| 3 | Na Natrium | Mg Magnesium | | | | | | | | | | | Al Aluminium | Si Silikon | P Fosfor | S Belerang | Cl Klor | Ar Argon |
| 4 | K Kalium | Ca Kalsium | Sc Skandium | Ti Titanium | V Vanadium | Cr Kromium | Mn Mangan | Fe Besi | Co Kobalt | Ni Nikel | Cu Tembaga | Zn Seng | Ga Gallium | Ge Germanium | As Arsen | Se Selenium | Br Bromin | Kr Krypton |
| 5 | Rb Rubidium | Sr Stronsium | Y Yttrium | Zr Zirkonium | Nb Niobium | Mo Molibdenum | Tc Teknesium | Ru Rutenium | Rd Rodium | Rb Rubidium | Ag Perak | Cd Kadmium | In Indium | Sn Timah | Sb Antimon | Te Telurium | I Yodium | Xe Xenon |
| 6 | Cs Sesium | Ba Barium | Ra Radium | Hf Hafnium | Ta Tantalum | Ba Barium | Re Rhenium | Os Osmium | Ir Iridium | Pt Platina | Au Emas | Hg Air Raksa | Tl Thallium | Pb Timbal | Bi Bismut | Po Polonium | At Astatin | Rn Radon |
| 7 | Fr Francium | Ra Radium | Rf Rutherfordium | Db Dubnium | Sg Seaborgium | Bh Bohrium | Hs Hassium | Mt Meitnerium | Ds Darmstadtium | Rg Roentgenium | Cn Kopernisium | Uut Ununtrium | Fl Flerovium | Uup Ununseptium | Lv Livermorium | Uus Ununseptium | Uuo Ununoctium | |

| | | | | | | | | | | | | | | | |
|---------------------|----------------|--------------|--------------------|-----------------|-----------------|-----------------|-----------------|------------------|-----------------|-------------------|-------------------|---------------|-------------------|-----------------|------------------|
| Rangkaian Lantanida | La Lantanum | Ce Seriun | Pr Praseodimium | Nd Neodimium | Pm Prometium | Sm Samarium | Eu Europium | Gd Gadolinium | Tb Terbium | Dy Dysprosium | Ho Holmium | Er Erbium | Tm Thulium | Yb Ytterbium | Lu Lutetium |
| Rangkaian Aktinida | Ac Aktinium | Th Torium | Pa Protaktinium | U Uranium | Np Neptunium | Pu Plutonium | Am Americium | Cm Kurium | Bk Berkelium | Cf Kalifornium | Es Einsteinium | Fm Fermium | Md Mendelevium | No Nobelium | Lr Lawrensium |

Penjelasan

Golongan **utama A(1A)** terbagi menjadi 4

1. Alkali (H,Li,Na,Rb,Cs,Fr)
2. Alkali Tanah (Be,Mg,Ca,Sr,Ba,Ra)
3. Boron (B,Al,Ga,In,Tl)

4. Gas Mulia (He,Ne,Ar,Kr,Xe,Rn)
5. Karbon(C,Si,Ge,Sn,Pb)
6. Nitrogen(N,P,As,Sb,Bi)
7. Hologen(F,Cl,Br,At,Ts)
8. Kalkogen(O,S,Se,Te,Po)

Golongan B

Logam Transisi mencangkup

Sc, Y, Ti, Zr, Hf, V, Nb, Ta, Cr, Mo, W, Mn, Tc, Re, Fe, Ru, Os, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Au, Zn, Cd, Hg

Untuuk golongan Utama atau 1A di beri garis coretan hitam

[illegible]

Berarti sisanya yang tidak ada coretan adalah golongan B yaitu golongan 3 sampai 12.

| | | | | | | | | | | | | | | |
|---|--|--|---|--|--|--|--|---|--|---|--|--|---|---|
| <div>Region Lanthanide</div> <div><div>La</div><div>Lanthanum</div></div> <div>71</div> | <div>Region Lanthanide</div> <div><div>Ce</div><div>Cerium</div></div> <div>72</div> | <div>Region Lanthanide</div> <div><div>Pr</div><div>Praseodymium</div></div> <div>73</div> | <div>Region Lanthanide</div> <div><div>Nd</div><div>Neodymium</div></div> <div>74</div> | <div>Region Lanthanide</div> <div><div>Pm</div><div>Promethium</div></div> <div>75</div> | <div>Region Lanthanide</div> <div><div>Sm</div><div>Samarium</div></div> <div>76</div> | <div>Region Lanthanide</div> <div><div>Eu</div><div>Europium</div></div> <div>77</div> | <div>Region Lanthanide</div> <div><div>Gd</div><div>Gadolinium</div></div> <div>78</div> | <div>Region Lanthanide</div> <div><div>Tb</div><div>Terbium</div></div> <div>79</div> | <div>Region Lanthanide</div> <div><div>Dy</div><div>Dysprosium</div></div> <div>80</div> | <div>Region Lanthanide</div> <div><div>Ho</div><div>Holmium</div></div> <div>81</div> | <div>Region Lanthanide</div> <div><div>Er</div><div>Erbium</div></div> <div>82</div> | <div>Region Lanthanide</div> <div><div>Tm</div><div>Thulium</div></div> <div>83</div> | <div>Region Lanthanide</div> <div><div>Yb</div><div>Ytterbium</div></div> <div>84</div> | <div>Region Lanthanide</div> <div><div>Lu</div><div>Lutetium</div></div> <div>87</div> |
| <div>Region Actinide</div> <div><div>Ac</div><div>Actinium</div></div> <div>89</div> | <div>Region Actinide</div> <div><div>Th</div><div>Thorium</div></div> <div>90</div> | <div>Region Actinide</div> <div><div>Pa</div><div>Protactinium</div></div> <div>91</div> | <div>Region Actinide</div> <div><div>U</div><div>Uranium</div></div> <div>92</div> | <div>Region Actinide</div> <div><div>Np</div><div>Neptunium</div></div> <div>93</div> | <div>Region Actinide</div> <div><div>Pu</div><div>Plutonium</div></div> <div>94</div> | <div>Region Actinide</div> <div><div>Am</div><div>Americium</div></div> <div>95</div> | <div>Region Actinide</div> <div><div>Cm</div><div>Curium</div></div> <div>96</div> | <div>Region Actinide</div> <div><div>Bk</div><div>Berkelium</div></div> <div>97</div> | <div>Region Actinide</div> <div><div>Cf</div><div>Californium</div></div> <div>98</div> | <div>Region Actinide</div> <div><div>Es</div><div>Einsteinium</div></div> <div>99</div> | <div>Region Actinide</div> <div><div>Fm</div><div>Fermium</div></div> <div>100</div> | <div>Region Actinide</div> <div><div>Md</div><div>Mendelevium</div></div> <div>101</div> | <div>Region Actinide</div> <div><div>No</div><div>Nobelium</div></div> <div>102</div> | <div>Region Actinide</div> <div><div>Lr</div><div>Lawrencium</div></div> <div>103</div> |

Lantanida dan **Aktinida** adalah dua kelompok unsur yang terletak di bagian bawah tabel periodik dan **tidak termasuk dalam golongan-golongan 1 hingga 18** yang biasa kita lihat di tabel periodik utama. Mereka berada dalam **blok f** dan sering ditempatkan terpisah dari tabel utama untuk menjaga kejelasan tabel periodik.

```
body {  
  font-size: 11pt;  
  line-height: initial;  
}
```

Mengatur ukuran font di seluruh halaman menjadi 11pt dan tidak menetapkan line-height secara eksplisit (pakai default).

```
table {  
  width: 96%;  
  background: #d2e7f1;  
  font-family: roboto;  
  border-style: none;  
  border-collapse: separate;  
  border-spacing: 2px;  
}
```

- Tabel selebar 96% dari kontainer.
- Warna latar belakang biru muda.
- Menggunakan font Roboto.
- Tidak ada garis tepi (border), dan sel-sel tabel tidak digabung (separate).

```
th {  
  width: 400px;  
  max-width: 400px;  
  text-align: center;  
}
```

Lebar maksimum untuk header adalah 400px dan teks di dalamnya disejajarkan ke tengah.

```
#title {  
  background-color: #455A64;  
  color: white;  
  text-align: center;  
}
```

- Digunakan untuk judul halaman.
- Latar belakang abu kebiruan dan teks berwarna putih.