

The effect of electric field on potentiometric Scanning Electrochemical Microscopic images

András Kiss^{1,*}, Dániel Filotás¹, Géza Nagy¹

^a*Department of General and Physical Chemistry, Faculty of Sciences, University of Pécs, 7624 Pécs, Ifjúság útja 6, Hungary*

Abstract

Scanning Electrochemical Microscopy (SECM) is an invaluable tool in corrosion science. It allows the selective imaging of a particular ionic species being released at the anodic sites, using ion-selective microelectrodes (ISMEs) as scanning probes. An often studied phenomenon is galvanic corrosion, which involves two metals in electrical contact, immersed in the same electrolyte. The measured potential of the ISME is thought to depend only on the activity of primary ion. However, an electric field is also formed as a result of the potential difference between the surfaces of the galvanic pair, which has a direct influence on the potential of the microelectrode; the measured potential is the sum of these two. The potential difference caused by the electric field can be substantially large, exceeding that of the potential difference associated with the activity of the primary ion. In this paper, we present experimental evidence of this, and investigate the extent to which it influences the final image.

Keywords: scanning electrochemical microscopy, potentiometry, galvanic corrosion, electric field

1. Introduction

In the past decade, potentiometric SECM - or in other words used by the experts of this field Scanning Ion Selective Electrode Technique (SIET) - has become very popular among corrosion scientists[1, 2, 3, 4, 5, 6]. The most broad spread application is the visualization of galvanic corrosion[7, 8, 9, 10]. Galvanic corrosion occurs when two dissimilar metals are connected both electrically and immersed in the same electrolyte. The electric coupling results the preferential and accelerated dissolution of the anode, while reduces corrosion rate of the cathode. The spatial separation of the anodic and the cathodic sites makes the complex corrosion processes easily interpretable and due to the increased corrosion rate conveniently short exposure times are sufficient to obtain convincing concentration distributions in the solution adjacent to the corroding sample.

Despite these beneficial circumstances, quantitative evaluation of galvanic corrosion using potentiometric SECM often fails due to – up to now – unrevealed reasons. Izquierdo et. al. reported discrepant results comparing vertical approaching curves towards the cathode of the Mg-Fe galvanic couple obtained by amperometric O_2 detection and potentiometric pH measurements [11]. Local alkalization could be detected even at 2 mm tip-substrate distance, whereas oxygen concentration reached the bulk level at ca. 900 μm height. The phenomenon was explained by the contribution of the electric field to the potentiometric signal. In another works, Mg^{2+} above Mg alloy disc galvanically coupled to iron detected with Mg ISME highly exceeded the upper limit of detection of the probe [12, 13, 14]. On the other hand, pMg values fallen behind lower limit of detection of Mg ISMEs scanning above cathodically polarized magnesium strips [15].

These phenomenon could be explained by a contribution of the electric field to the measured potential. The potential difference between the surfaces of the anode and the cathode causes an electric field to be

*Corresponding author

Email addresses: akiss@gamma.ttk.pte.hu (András Kiss), filotasdaniel@gmail.com (Dániel Filotás), g-nagy@gamma.ttk.pte.hu (Géza Nagy)

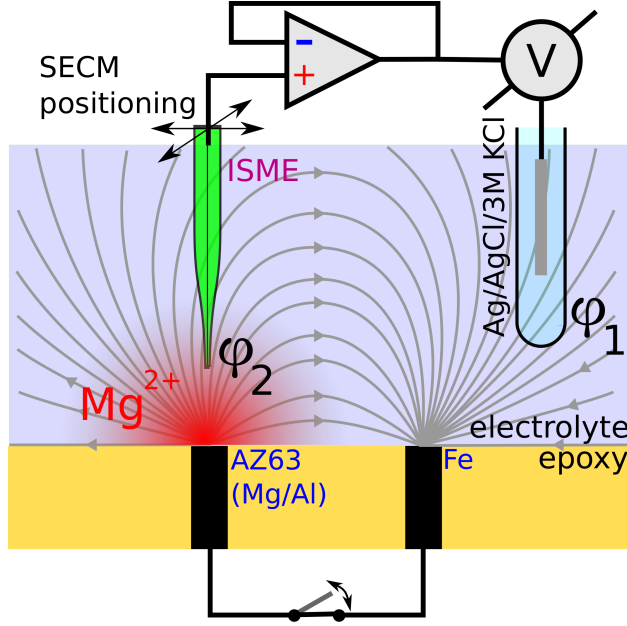


Figure 1: Caption.

formed. The potential difference between the points where the electrodes are located is added to the potential difference associated with the primary ion activity at the tip of the measuring electrode:

$$\Delta E = E_M - E_R + (\phi_M - \phi_R) \quad (1)$$

where ΔE is the measured potential difference, E_R is the potential of the reference electrode, ϕ_M and ϕ_R are the potentials in the electric field at the measuring and reference electrodes, respectively. E_M is the potential of the measuring electrode:

$$E_M = S \times \lg[Mg^{2+}] + E_M^o \quad (2)$$

where S is the slope of the calibration curve of the potentiometric cell with respect to the primary ion, and E_M^o is the standard potential.

The effect of the electric field on the measured potential difference has been investigated in this paper. The galvanic corrosion of the AZ63 Mg-Al alloy and iron was used as a model system.

2. Material and methods

The preparation of solid contact Mg^{2+} selective microelectrodes was adopted from a previous work [14]. Micropipettes were pulled from borosilicate capillaries (outer diameter $\varnothing = 1.5$ mm, inner dia. $\varnothing = 1.0$ mm, obtained from Hilgenberg GmbH, Malsfeld, Germany) with a Sutter Instruments P-30 type vertical capillary puller (Novato, CA, USA). The micropipette were soaked in 1:1 $H_2SO_4:H_2O_2$ solution and washed with double deionized water. The capillaries were silanized by 1 hour exposition the saturated vapour of dichloro-dimethyl-silane in closed Petri dishes at 120 °C. A poly-ethylen-dioxy-thiophene (PEDOT) coated carbon fiber of 33 μm diameter (obtained as a generous gift from Specialty Materials, Lowell, MA, USA) served as the solid contact of the ISME. The PEDOT was electrochemically polymerized to the carbon fiber in 0.1 M EDOT-containing $BMIM - PF_6$ ionic liquid solution. 10 consecutive cyclic voltammetry cycles were taken in $-0.9 \leq E \leq 1.3$ V range. The doping step was performed in 0.1 M KCl aqueous solution by applying 15 consecutive potential cycles in the $-0.9 \leq E \leq 0.8$ V range. The membrane

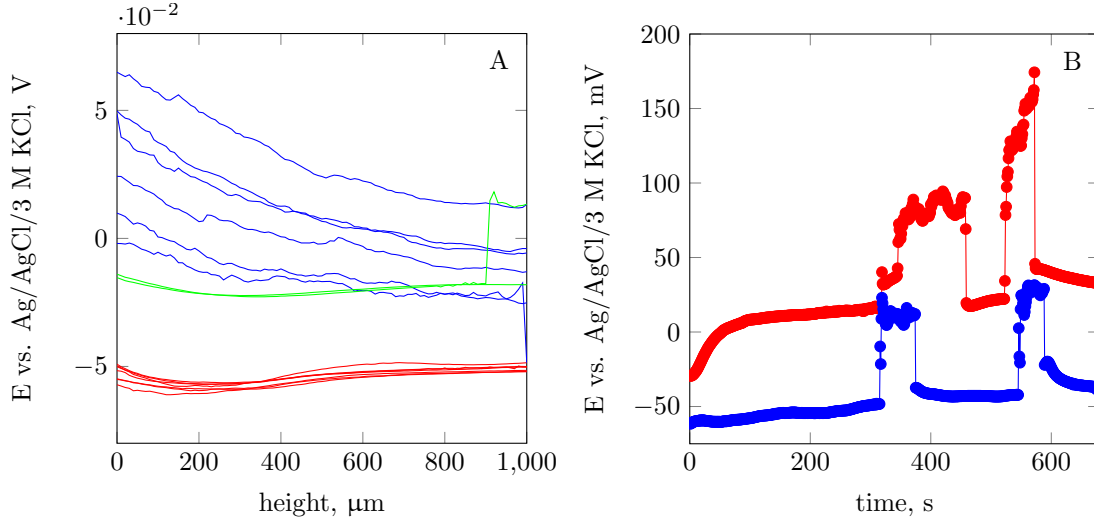


Figure 2: Transient response of the antimony microelectrode to analyte activity step. The measuring and reference electrodes were dipped into buffer solutions with $\text{pH} = 4$ before the measurements started, and $\text{pH} = 6$ at $t = 0$ s, respectively. Eq. ?? was fitted (red line) on the measurement (gray marks) from the pH step to the end of the curve when potential reaches equilibrium in the $\text{pH} = 6$ buffer.

components were purchased from Fluka (Buchs, Switzerland). The cocktail contains N,N'' -Octamethylene-bis(N' -heptyl- N' -methyl-methylmalonamide inophore, 2-nitrophenyl-octyl ether emollient, PVC, potassium-[tetrakis-4-chlorophenyl]-borate and tetrahydrofuran. Eventually, the micropipette was frontfilled with the cocktail and the PEDOT coated carbon fiber was inserted in the lumen of the capillary. The Mg ISMEs were calibrated by measuring their potential against $\text{Ag}/\text{AgCl}/\text{KCl}$ (3M) reference electrode in tenfold diluted MgCl_2 solutions between 10^{-7} and 10^{-1} M concentrations. The activities were calculated using the Debye-Hückel theory. Nernstian relationship was found in between 10^{-1} and 10^{-5} M, the equation of the linear portion of the calibration curve is $-29.5\text{mV/decade} + 98.3\text{mV}$ ($R^2 = 0.9997$). The lower limit of detection was $\text{pMg} = 5.3$.

The (Mg/Al)/Fe galvanic couple target was prepared from AZ63 Mg/Al alloy and high purity Fe wires with identical diameter of 0.76 mm. The wires were mounted in an epoxy resin sleeve (Struers, Ballerup, Denmark), exposing only the disk shaped surfaces and allowing to make electric contact at the rear of the mould. Frontal surface of the mould was first polished with SiC paper down to 4000 grit, then with 1.0, 0.3 μm alumina powder.

SECM experiments were carried out employing homemade instrument operated with custom software. The sample was placed at the bottom of the electrochemical cell, whereas the local potential values of the Mg ISMEs were measured with against a Ag/AgCl (3M KCl) reference electrode. All the measurements were performed using a high input impedance eDAQ pH ISE isoPod USB (eDAQ Pty Ltd, Australia).

3. Results and discussion

First, consecutive approaching curves were recorded above the corroding AZ63 sample, while the galvanic connection with the iron sample was...

The moment the galvanic connection was established, there was an immediate rise of about 140 mV in the measured potential of the microelectrode [fig], which cannot possibly be attributed to the increase of Mg^{2+} activity that far from the source. Also, a 140 mV rise would mean an increase of about 3.5 orders of magnitude in Mg^{2+} activity in less than a second. Even if one argues it's possible 100 μm from the source, it cannot be the case 1000 μm from it. The only plausible explanation is that sudden change is due to the electric field formed between the two metals.

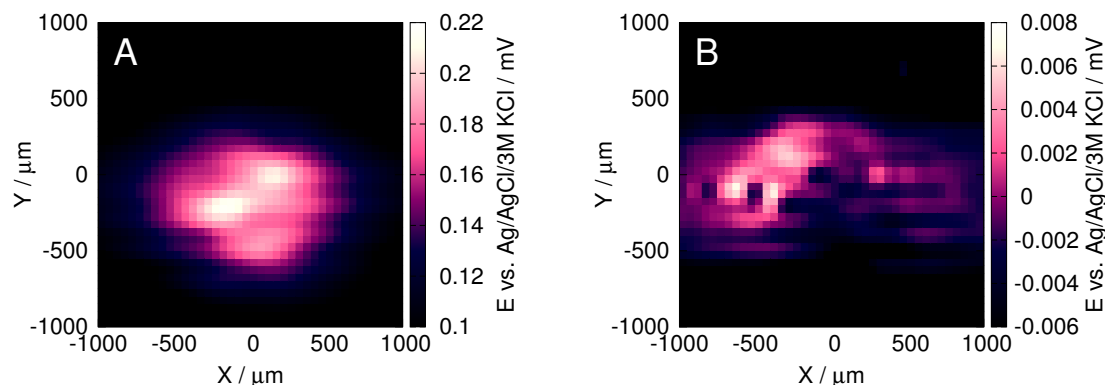


Figure 3: Caption.

4. Conclusions

Acknowledgements

This research was supported by the European Union and the State of Hungary, co-financed by the European Social Fund in the framework of TÁMOP-4.2.4.A/ 2-11/1-2012-0001 'National Excellence Program' and TÁMOP-4.2.2.A-11/1/KONV-2012-0065.

References

- [1] S.V. Lamaka, M.G. Taryba, M.L. Zheludkevich, M.G.S. Ferreira, Novel Solid-Contact Ion-Selective Microelectrodes for Localized Potentiometric Measurements *Electroanalysis* 21 (2009) 2447-2453.
- [2] J. Izquierdo, L. Nagy, Á.s Varga, I. Bitter, G. Nagy, R. M. Souto Scanning electrochemical microscopy for the investigation of corrosion processes: Measurement of Zn^{2+} spatial distribution with ion selective microelectrodes *Electrochimica Acta* 59 (2012) 398-403.
- [3] E.L. Silva, A.C. Bastos, M.A. Neto, R.F. Silva, M.G.S. Ferreira, M.L. Zheludkevich, F.J. Oliveira Novel diamond micro-electrode for pH sensing *Electrochemistry Communications* 40 (2014) 31-34.
- [4] A. Alvarez-Pampliega, S.V. Lamaka, M.G. Taryba, M. Madani, J. De Strycker, E. Tourwé, M.G.S. Ferreira, H. Terryn Cut-edge corrosion study on painted aluminum rich metallic coated steel by scanning vibrating electrode and micro-potentiometric techniques *Electrochimica Acta* 61 (2012) 107-117.
- [5] E. A. Zdrachek, A. G. Karotkaya, V. A. Nazarov, K. A. Andronchyk, L. S. Stanishevskii, V. V. Egorov, M. G. Taryba, D. Snihirova, M. Kopylovich, S. V. Lamaka H^{+} -selective microelectrodes with optimized measuring range for corrosion studies *Sensors and Actuators B* 207 (2015) 967-975.
- [6] H. Shi, Z. Tian, T. Hu, F. Liu, E. H. Han, M. Taryba, S.V. Lamaka Simulating corrosion of Al₂CuMg phase by measuring ionic currents, chloride concentration and pH *Corrosion Science* 88 (2014) 178-186.
- [7] J. Izquierdo, L. Nagy, J. J. Santana, G. Nagy, R. M. Souto A novel microelectrochemical strategy for the study of corrosion inhibitors employing the scanning vibrating electrode technique and dual potentiometric/amperometric operation in scanning electrochemical microscopy: Application to the study of the cathodic inhibition by benzotriazole of the galvanic corrosion of copper coupled to iron *Electrochimica Acta* 58 (2011) 707-716.
- [8] V. A. Nazarov, M. G. Taryba, E. A. Zdrachek, K. A. Andronchyk, V. V. Egorov, S. V. Lamaka Sodium- and chloride-selective microelectrodes optimized for corrosion studies *Journal of Electroanalytical Chemistry* 706 (2013) 13-24.
- [9] E. Tada, S. Satoh, H. Kaneko, The spatial distribution of Zn^{2+} during galvanic corrosion of a Zn/steel couple *Electrochimica Acta* 49 (2004) 2279-2285.
- [10] A.G. Marques M. Taryba A.S. Panˆao S. Lamaka A.M. Simoes Application of scanning electrode techniques for the evaluation of iron-zinc corrosion in nearly neutral chloride solutions *Corrosion Science* 104 (2016) 123-131.
- [11] J. Izquierdo, L. Nagy, I. Bitter, Ricardo M. Souto, G. Nagy Potentiometric scanning electrochemical microscopy for the local characterization of the electrochemical behaviour of magnesium-based materials *Electrochimica Acta* 87 (2013) 283-293.
- [12] R. M. Souto, A. Kiss, J. Izquierd, L. Nagy, I. Bitter, G. Nagy Spatially-resolved imaging of concentration distributions on corroding magnesium-based materials exposed to aqueous environments by SECM *Electrochemistry Communications* 26 (2013) 25-28.

- [13] R.M. Souto, J Izquierdo, J.J. Santana, A Kiss, L Nagy, G Nagy Progress in Scanning Electrochemical Microscopy by Coupling Potentiometric and Amperometric Measurement Modes In: Gabor Nagy, Gyula Pinczes, Gabor Pinter, Istvan Poci, Jozsef Prokisch, Gaspar Banfalvi A Méndez-Vilas (szerk.) Current microscopy contributions to advances in science and technology. Spain: Formatex Research Center, 2012. pp. 1407-1415.
- [14] J. Izquierdo, A. Kiss, J. J. Santana, L. Nagy, I. Bitter, H. S. Isaacs, G. Nagy, R. M. Souto Development of Mg^{2+} Ion-Selective Microelectrodes for Potentiometric Scanning Electrochemical Microscopy Monitoring of Galvanic Corrosion Processes Journal of the Electrochemical Society 160 (2013) 451-459.
- [15] J. Izquierdo, B. M. Fernández-Pérez, D. Filotás, Z. Őri, A. Kiss, R.T. Martín-Gómez, L. Nagy, G. Nagy, R. M. Souto Imaging of Concentration Distributions and Hydrogen Evolution on Corroding Magnesium Exposed to Aqueous Environments Using Scanning Electrochemical Microscopy Electroanalysis 28 (2016) 1–14.