

Demonstration of potentiometric pH determination using the filament of an incandescent lightbulb as indicator electrode

János Klucsik, Venkat Rajat Rao Yalamarti, and András Kiss*

Department of General and Physical Chemistry, University of Pécs, Ifjúság útja 6, 7622 Pécs, Hungary

E-mail: akiss@gamma.ttk.pte.hu

Abstract

One of the earliest concepts of chemistry taught in school is pH. It is usually introduced during the discussion of acids and bases. Students are taught to detect the acidity of substances using pH indicators, usually in the form of indicator paper. While this method is convenient and easy for students to relate to, apart from not being very accurate, it does not give any deeper insight into the meaning of pH. To truly comprehend pH, students need to learn to approach the concept from an electroanalytical point of view. The aim of this paper is to avail teachers of an additional interactive and demonstrative experiment in order to facilitate this transition.

We demonstrate the fabrication of a simple and economical metal/metal-oxide type pH indicator electrode using simple household items. The sensing element is made from the tungsten alloy filament of an incandescent lightbulb, that is sandwiched between two pieces of Plexiglass using superglue.

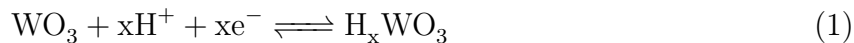
Keywords

pH, tungsten electrode, light-bulb, potentiometry

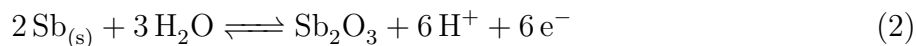
Introduction

The glass electrode is the most well-known ion-selective electrode, because the negative logarithm of hydrogen-ion activity that is pH, has special importance. It is found in virtually every laboratory, and students are likely to encounter it even during their highschool education. However, besides the glass electrode, highschool students and even undergraduates usually don't learn about other hydrogen-ion selective electrodes. The exclusivity of the glass electrode is most likely due to its relatively low cost and excellent characteristics. Other ion-selective electrodes usually cost a lot more and don't function as well. For educational purposes however, they are excellent tools. There are several low cost fabrication techniques described in the literature, which allow students even in a highschool chemistry laboratory to fabricate their own electrodes, such as ionophore based indicator electrodes^{1,2} or simple reference electrodes.³ There is even a procedure to fabricate low cost glass electrodes from microscope slide cover sheets.⁴ These electrodes can be manufactured and calibrated by students in a few laboratory sessions and can provide a great learning experience for them. They allow a view on pH and ion-selective electrodes from another angle, and students are more likely to understand the function of a device when they fabricate it themselves.

The electrode demonstrated in this paper is an ion-selective electrode that is made using the filament of an incandescent light bulb. It is well known that the filament of incandescent lightbulbs is made of an alloy called „*tungsten bronze*” that is mostly tungsten⁵⁻⁷ with a small amount of potassium to improve ductility. It is a metal/metal-oxide type electrode that is sensitive to pH. Its function is based on the intercalation and deintercalation of hydrogen ions into the oxide layer followed by a change in the oxidation state of tungsten oxide:⁸



The metal/metal-oxide type pH electrodes are usually used in research if the application of glass electrodes is challenging or impossible. These include measuring pH at high temperature or in hydrogen-fluoride solution, or application in the food industry. Besides the tungsten electrode, one of the most often used type of these is the Ir/IrO₂ electrode.⁹ The oldest is certainly the Sb/Sb₂O₃ electrode, its initial characterization dating back to 1923.¹⁰ It is pH sensitive because hydrogen ions participate in the equilibrium between antimony and its oxide on the surface:



In this paper, the fabrication and calibration of the above mentioned tungsten filament electrode is described. The components for the electrode can be obtained from a local hardware store, and the construction is very simple, requiring only an incandescent lightbulb, a small sheet of PlexiGlas and superglue. Electrodes prepared by students will be shown as well as the results of the calibration and selectivity study performed by 1st and 2nd year university BSc students of chemistry. Altogether 8 students participated in the demonstration that involved a 2 hour seminar about ion-selective electrodes and pH and two 3 hour laboratory sessions fabricating and characterizing the electrodes. A typical notebook prepared by a 2nd year chemistry BSc student about the demonstration is shown as example in the supportin material.

Theory

pH is one of the most important concepts in chemistry, originally defined by Sørensen in 1909¹¹ as the negative logarithm of the concentration of hydrogen ion:

$$\text{pH} = -\log_{10} c_{\text{H}^+} \quad (3)$$

At first glance, it looks like a deceptively simple concept. Complications start to arise when we replace concentration with activity, taking into account that hydrogen ions have charge, therefore their behaviour correlates more closely with activity than concentration:

$$\text{pH} = -\log_{10} a_{\text{H}^+} = -\log_{10} c_{\text{H}^+} \gamma_{\text{H}^+} \quad (4)$$

where a_{H^+} is the activity and γ_{H^+} is the activity coefficient of hydrogen ions. This definition however raises another problem. It includes the activity coefficient of a single ion, which cannot be separated from the activity coefficient of counterions. The IUPAC recommended solution¹² is the *Bates–Guggenheim* convention.¹³ In order to provide a basis for the determination of single ion activity coefficients, the activity coefficient of chloride ion is defined using the extended Debye–Hückel equation:

$$-\log_{10} \gamma = \frac{Az^2\sqrt{I}}{1 + Ba\sqrt{I}} \quad (5)$$

where A and B are constants, a is ionic radius and I is ionic strength calculated with the following equation:

$$I = \frac{1}{2} \sum_{i=1}^n c_i z_i^2 \quad (6)$$

where c_i and z_i is the concentration and charge of the i th ion, and n is the total number of ionic species in the solution.

In practice, pH is most commonly measured with a glass electrode. It has been known for more than a hundred years that the potential difference measured across a glass membrane depends on the ratio of the hydrogen ion activity in the two solutions that the membrane separates.^{14,15} The glass electrode is usually combined with a reference electrode for conve-

nience. Such an electrode pair is called the „*combined glass electrode*”, and this the most common type the students might encounter. The glass electrode is the best electrochemical sensor and one of the best sensors ever made, with a linear response of over more than 13 orders of magnitude and excellent selectivity. However, it’s not perfect. One of the imperfections is that although to a negligible extent, it responds to cations other than hydrogen ion. This behaviour is described by the Nikolsky–equation that takes the effect of interfering ions into account with the so-called „*selectivity coefficient*”:¹⁶

$$E = E^\theta + \frac{RT}{z_i F} \ln \left[a_i + \sum_j \left(k_{ij} a_j^{z_i/z_j} \right) \right] \quad (7)$$

Materials and Methods

Electrode fabrication. To obtain the sensing element of the electrode, a 100 W incandescent lightbulb (Tungsum brand, purchased at local hardware store) was carefully wrapped in cloth and broken with a mallet. The contact wire that still had the filament attached was cut near the stem (Fig. 1a). The electrode body was prepared from two, 4 mm thick, ≈ 0.5 cm \times ≈ 2 cm, identically shaped Plexiglass pieces (Fig. 1b). In one of them, a ≈ 1 mm groove was cut longitudinally, in which the filament with stem was layed (Fig. 1c). Then, superglue was applied onto the Plexiglas surface that contained the groove with the filament and stem. The other Plexiglass piece was pressed onto the superglued surface, and was held firmly until the superglue cured (Fig. 1d). The electrode was sanded and polished from all sides except where the electrical lead, the stem was protruding (Fig. 1e).

Calibration. The fabricated tungsten filament electrodes were calibrated by students in a series of Britton–Robinson buffers with the following pH values: 2.19, 3.04, 4.10, 5.12, 6.09, 7.19, 8.08, 8.89, 9.87, 10.88, 11.81. The pH was measured by the instructor beforehand with a WTW SenTix combined pH electrode connected to a WTW inoLab pH 740p high input impedance pH meter (Xylem Analytics Germany Sales GmbH & Co. KG, Weilheim

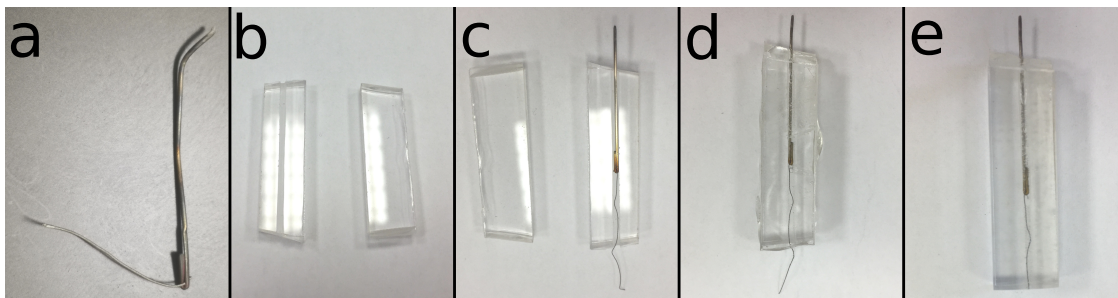


Figure 1: The fabrication process of the tungsten filament electrode.

Germany), that was calibrated with three high accuracy pH buffers (4.00, 7.00, 10.00 at 20 °C, Scharlab,S.L. Barcelona, Spain). The glass and the tungsten filament electrodes were calibrated by measuring their potential against the internal reference electrode of the glass electrode in the buffer series. The potential values were recorded when the instrument indicated a steady reading. This ensured that they were taken at the same point of the response curve. The students then plotted the potential values as a function of the corresponding pH, and performed a linear regression using a commonly available open source spreadsheet program (LibreOffice).

Selectivity study. The selectivity coefficient was determined using the „*separate solution method II (SSMII)*”.¹⁷ Besides the pH buffer series, several other series of solutions (KCl, NaCl, CaCl₂ and MgCl₂) were prepared by the students, each containing an interfering ion. Since the primary ion for the glass and the tungsten filament electrodes is already present in water at an activity that would influence the selectivity study, a buffer had to be used to lower the hydrogen ion activity. Most of the common pH buffers contain alkaline metals, and they could not be used to carry the selectivity study, because alkaline and alkaline earth metals are the most common interferences. For this reason, a pH = 8 TRIS buffer was used as a solvent to create the dilution series using KCl, NaCl, CaCl₂ and MgCl₂. In each cases, the activities ranged from 10⁻⁶ M to 10⁻¹ M with tenfold increments in activity. The electrodes were immersed in each of the solutions and the resulting potential differences were measured and recorded with the already mentioned pH meter.

To determine the selectivity coefficients, first the activities had to be calculated from the concentrations. This involved the calculation of activity coefficients which in turn required the calculation of ionic strength for each solution. This is an important step, and usually requires guidance from the laboratory instructor. To calculate the ionic strength, eq. 6 was used. Then the extended Debye-Hückel equation 5 was applied to calculate the activity coefficients. To obtain the activities, concentrations have to be multiplied with the corresponding activity coefficients.

To calculate the selectivity coefficient, a common potential was picked and the corresponding activity of the primary and interfering ions were calculated. The selectivity coefficient is then given by the ratio (SSMII):

$$k_{i,j} = \frac{a_i}{a_j^{z_i/z_j}} \quad (8)$$

where a_i and a_j are the activities of the primary and interfering ion at the same potential, and z_i and z_j is the charge of the primary and interfering ions.

Hazards

When salvaging the tungsten filament from the lightbulb, care must be taken. The lightbulb must be wrapped several times in a cloth before it is hit with a mallet. Care must be taken when the filament is picked up from the broken pieces of glass. When the Plexiglass body is prepared, a saw is used to cut the groove, then superglue is used to enclose the filament in the groove. During these steps, close supervision of the students is necessary to avoid any harm. Lastly, HCl and NaOH is used during the preparation of the pH buffer series. If these get in contact with skin or the eye, flush with plenty of water for several minutes.

Results and Discussion

To make sure that the measured pH values are accurate, the glass electrode was first calibrated with high accuracy pH buffers. The resulting slope was 57.19 mV / pH at 21 °C. Then, the pH values of the created Britton–Robinson buffer series were measured. Once these were known with high accuracy, the same buffer series was used again to calibrate the tungsten filament electrode. The resulting calibration plot can be seen in Fig. 2B. The slope of the calibration equation is 40.05 mV / pH, which is significantly lower than that of the theoretically expected value from the Nernst-equation, or the one that was measured for the glass electrode in this study. The lower value might be explained by the alloying elements in the tungsten filament. It is well known, that the filaments of incandescent lightbulbs are not made of pure tungsten, but a so-called „tungsten bronze” alloy. The alloy consists of tungsten and potassium or sodium, which improves ductility that helps during the manufacturing process.^{5–7} Even high purity metal/metal-oxide electrodes don’t reach the nernstian slope, which is well documented in the literature Nonetheless, a slope of 40.05 mV / pH is enough to measure pH reliably. The electrode response was slightly non-linear, but it can be said that it works well within the tested pH range of 2–12.

Because of these alloying elements, the tungsten filament electrode was expected to be less selective towards hydrogen ions than the glass electrode. The selectivity study was performed in dilution series of KCl, NaCl, CaCl₂ and MgCl₂ from 10⁻¹ M to 10⁻⁶ M with tenfold increases in concentration. A buffer was used to lower the hydrogen ion activity, and because most of the alkaline buffers contain an ion that interfere with hydrogen ion-selective electrodes, the TRIS buffer was selected at pH = 8. The TRIS buffer had a concentration of 1.25 M. This concentration was more than enough to increase the ionic strength to a value at which the activity coefficients for the members of the buffers series can be regarded as constant, and concentration can be used in the selectivity calculations instead of activity. The measured potential values for the glass and tungsten filament electrodes for the interfering ions can be seen in Fig. 2A and B, respectively. As expected, no interference can be observed

for any of the tested ions when the glass electrode was used. However, when measured with the tungsten filament electrode, some of the tested ions interfered with the pH measurement at higher than 10^{-4} M concentration. Below this concentration the potential corresponding to pH = 8 buffer was observed (≈ -315 mV). Above 10^{-4} M, most of the ions were interfering to a small extent. Interestingly, K^+ caused the potential to drop at higher concentration. This result was replicated several times. The selectivity coefficients were calculated with the separate solution method from activities corresponding to the same potential. The selectivity coefficient values are summarized in Table 1.

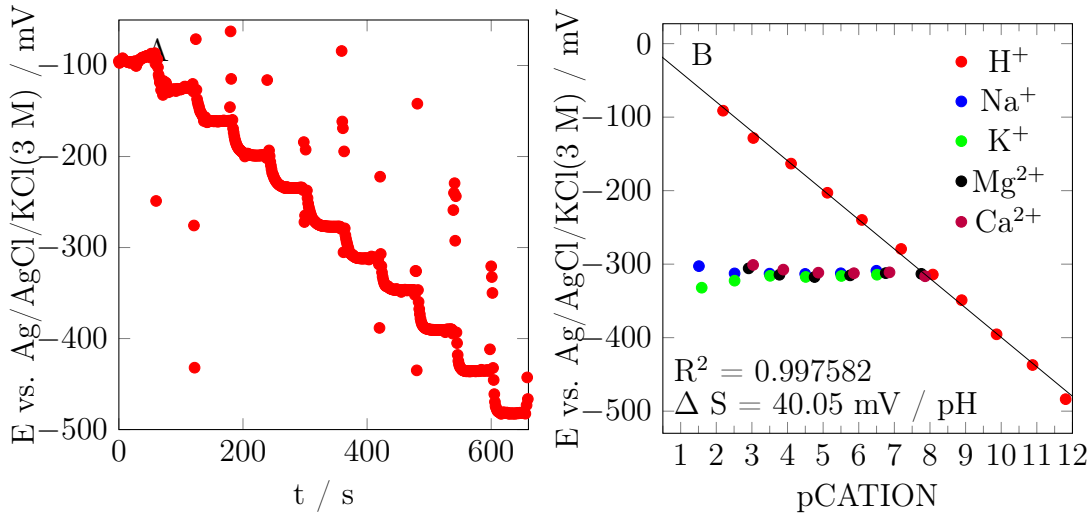


Figure 2: Calibration and selectivity study of the glass electrode (A), and the tungsten filament electrode (B).

Table 1: The selectivity coefficients of the tungsten filament electrode for the different cations.

Interfering ion	Selectivity coefficient
Na^+	$8.57 \cdot 10^{-7}$
K^+	$1.91 \cdot 10^{-7}$
Ca^{2+}	$9.48 \cdot 10^{-7}$
Mg^{2+}	$6.17 \cdot 10^{-7}$

Conclusion

The construction and usage of an economical pH sensitive electrode was presented that is highly selective towards hydrogen ions. It has been shown that the electrode performs very well, with a 40.05 mV / pH slope in the pH range 2–12. There was no significant interference from K^+ , Na^+ , Mg^{2+} and Ca^{2+} ions. The potentiometric cell utilizing the electrode can be used as an engaging demonstration of the potentiometric determination of pH. The participating students showed high motivation during the electrode fabrication process. They were excited to use electrodes that they've prepared themselves to measure pH. As a result of the interactive demonstration and evaluation, they showed a deeper understanding of pH and its potentiometric determination.

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References

- (1) Goldcamp, M. J.; Conklin, A.; Nelson, K.; Marchetti, J.; Brashear, R.; Epure, E. Inexpensive and disposable pH electrodes. *Journal of chemical education* **2010**, *87*, 1262–1264.
- (2) Marafie, H. M.; Shoukry, A. F.; Alshatti, L. A. Plastic Membrane Sensor from a Disposed Combined Glass Electrode. A Project for Graduate and Fourth-Year Undergraduate Students of Analytical Chemistry. *Journal of chemical education* **2007**, *84*, 793.

- (3) Riyazuddin, P. Low-Cost Reference Electrode for Potentiometric Titrimetry. *Journal of chemical education* **1994**, *71*, 167.
- (4) Yong, F.; Zhu, Q.; Zhang, G.; Tao, G.; Qin, S. Simple and Economical Procedure To Assemble pH Glass Membrane Electrodes Used in Chemical Education. *Journal of Chemical Education* **2019**,
- (5) Cisternas, R.; Kahlert, H.; Wulff, H.; Scholz, F. The Electrode Responses of a Tungsten Bronze Electrode differ in Potentiometry and Voltammetry and give Access to the Individual Contributions of Electron and Proton Transfer. *Electrochemistry Communications* **2015**, *56*, 34–37.
- (6) Wechter, M.; Shanks, H.; Carter, G.; Ebert, G.; Guglielmino, R.; Voigt, A. Use of Metal Tungsten Bronze Electrodes in Chemical Analysis. *Analytical chemistry* **1972**, *44*, 850–853.
- (7) Schade, P. 100 Years of Doped Tungsten Wire. *International Journal of Refractory Metals and Hard Materials* **2010**, *28*, 648–660.
- (8) Fenster, C.; Smith, A. J.; Abts, A.; Milenkovic, S.; Hassel, A. W. Single tungsten nanowires as pH sensitive electrodes. *Electrochemistry Communications* **2008**, *10*, 1125–1128.
- (9) Beyenal, H.; Davis, C. C.; Lewandowski, Z. An improved Severinghaus-type carbon dioxide microelectrode for use in biofilms. *Sensors and Actuators B: Chemical* **2004**, *97*, 202–210.
- (10) Uhl, A.; Kestranek, W. The electrometric titration of acids and bases with the antimony indicator electrodes. *Monatshefte für Chemie – Chemical Monthly* **1923**, *44*, 29–34.
- (11) Sørensen, S. P. L. Über die Messung und die Bedeutung der Wasserstoffionenkonzentration bei enzymatischen Prozessen. *Biochemische Zeitschrift* **1909**, *21*, 131–304.

- (12) Buck, R.; Rondinini, S.; Covington, A.; Baucke, F.; Brett, C. M.; Camoes, M.; Milton, M.; Mussini, T.; Naumann, R.; Pratt, K., et al. Measurement of pH. Definition, standards, and procedures (IUPAC Recommendations 2002). *Pure and applied chemistry* **2002**, *74*, 2169–2200.
- (13) Bates, R.; Guggenheim, E. Report on the standardization of pH and related terminology. *Pure and Applied Chemistry* **1960**, *1*, 163–168.
- (14) Haber, F.; Klemensiewicz, Z. Über elektrische Phasengrenzkräfte. *Zeitschrift für Physikalische Chemie* **1909**, *67*, 385–431.
- (15) Haber, F.; Klemensiewicz, Z. Concerning electrical phase boundary forces. *Zeitschrift für Physikalische Chemie, Stöchiometrie und Verwandtschaftslehre* **1909**, *67*, 385–431.
- (16) Nicolsky, B. Theory of the glass electrode. I. *Acta Physicochimica USSR* **1937**, *7*, 597–610.
- (17) Buck, R. P.; Lindner, E. Recommendations for nomenclature of ionselective electrodes (IUPAC Recommendations 1994). *Pure and Applied Chemistry* **1994**, *66*, 2527–2536.