

Computational Thermodynamics of the CoNiGa High Temperature Shape Memory Alloy system

Arpita Chari^a, Ebubekir Dogan^a, Anjana Talapatra^a, Avinash.R.Chivukula^a, Andres Garay^c, I. Karaman^{a,b}, R. Arróyave^{a,b,1}

^a Department of Mechanical Engineering, Texas A&M University, USA, 77843

^b Department of Materials Science and Engineering, Texas A&M University, USA, 77843

^c Centro de Investigación de Materiales Avanzados (CIMAV), Monterrey, Nuevo León, México,

Abstract

The CoNiGa ternary system has recently emerged as a promising alternative to currently used high temperature Shape Memory Alloys (SMA), with possible applications in the aerospace and automotive industries. In this work, we use thermodynamic modeling based on the CALPHAD approach to investigate the thermodynamic properties, phase stability and shape memory properties of these alloys. The thermodynamic model parameters have been obtained by fitting experimental and ab initio data. The stability of the β phase at high temperatures was enforced accurately by remodeling the CoGa system. Also, the phase stabilities of Ni_5Ga_3 (δ) and $\text{Ni}_{13}\text{Ga}_9$ (ϵ) were determined as they are known to cause detrimental effects in the ternary at low temperatures.

Key words: CALPHAD, CoNiGa, SMA, Thermodynamics, Computational Modeling, first-principles calculations

1. Introduction

Recent alternatives to the conventional Heusler Ni_2MnGa alloys have been the CoNiAl and CoNiGa systems [1]. These alloy systems have shown to exhibit shape memory and pseudoelastic properties at low temperatures. Also, CoNiGa alloys are commercially important as SMAs operating at temperatures above 250 °C. Limited experimental information is available for this system for a wide range of composition, and temperature regions [1–3]. Thus far, a reliable and consistent thermodynamic model of the CoNiGa system has not been calculated using a computational approach, and it would be useful to have a complete description of phase stability over a wide range of compositions and temperatures.

It is known that the ductility of the brittle CoNiGa- β alloys can be improved by the addition of the γ (fcc) phase [1]. Hence, a study of the stability of the β phase (which undergoes the martensitic transformation in the ternary and has a wide stability range in the NiGa and CoGa binaries), in conjunction with the existence of the two-phase $\beta + \gamma$ equilibrium in the Co-rich and Ni-rich regions

¹Corresponding author. E-mail: rarrayave@tamu.edu

in the binaries, is imperative in this investigation in order to design a ductile ferromagnetic CoNiGa SMA.

The CoGa binary model developed by Su and Tedenac [4] accurately describes the low temperature thermodynamics and it is seen that the β phase becomes stable at temperatures higher than 1000 °C. Although this artifact does not pose a difficulty in the binary, the extrapolation to the ternary CoNiGa system becomes difficult. Our previous work [5], dealt with this inconsistency and the CoGa system was remodeled with the β phase becoming stable at only lower temperatures. In this work, the binary phase descriptions of the CoGa [5], NiGa [6] and CoNi [7] systems have been used for the modeling of the ternary system. The thermodynamic parameters obtained in the binary become useful when trying to calculate the properties of the ternary system. However, a simple extrapolation of these models cannot be performed, owing to the effect of higher order ternary interaction parameters in the system. The ternary system was hence modeled using data available from experiments and information from the lower order binary systems.

The objective of this work is to develop a thermodynamic model of the CoNiGa SMA with accurate descriptions of the Gibbs energies of its phases. A self-consistent thermodynamic model of the CoNiGa Shape Memory Alloy has been obtained by combining the CALPHAD approach with *ab-initio* methods. The binary interaction parameters were obtained from previous models [5–7] while the ternary interaction parameters have been derived using experimental information in the optimization module ‘Parrot’ of the ThermoCalc package. The ternary model was then used to calculate isothermal sections at various temperatures, activities, partial enthalpies and vertical sections. Experiments were conducted to corroborate the results obtained from the calculated thermodynamic model.

The organization of this paper is as follows. In Section 2 we present a concise review of work carried out on the ternary CoNiGa system and describe the associated phase equilibria and stabilities. Section 3 we outline the various binary models used to build the ternary system and identify the assumptions made during the modeling process. We discuss the sample details, experimental conditions and procedures used to conduct the relevant experiments in Section 4. In Section 5 we describe the optimization process followed to obtain an accurate assessment by fitting the experimental and *ab-initio* data. Section 7 is devoted to some concluding remarks.

2. Literature Review

There have been several reports on Ni_2MnGa alloys being used as ferromagnetic shape memory materials. The main disadvantage found in these Heusler type alloys is that they are brittle in the polycrystalline state. This prevents them from being used as Ferromagnetic Shape Memory Alloys (FSMAs). Studies by Ishida et al. [8] showed that the ductility of the brittle CoNiGa β alloys can be improved by the introduction of γ phase (disordered fcc A1) of Ni solid solution on the grain boundaries of the β phase. This was also reported by Liu et al. and Oikawa et al [1, 9]. A two-stage annealing process was found to be an effective method in order to develop such a microstructure, among other methods [1, 2, 9] for this alloy. This type of microstructural control was also developed in several other Shape Memory Alloys such as CoNiAl[10, 11], NiFeAl[12, 13] and NiFeGa alloys[14–16]. Studies by Prusik et al. [17] showed that the precipitation of γ particles on the grain boundaries had a positive effect to increase ductility as compared to when the particles were dispersed inside the grains of the matrix. Another advantage of the CoNiGa Shape Memory Alloys is that their Curie temperature (T_c) is higher, thus having higher magnetization when compared to Ni_2MnGa alloys.

The equilibria among the phases in the CoNiGa system are similar to that in the CoNiAl system because Al and Ga belong to the same group *Rmnun3b* in the periodic table. Also, the β phase has a wide homogeneity range and an equilibrium exists with the γ phase in the Co-rich and Ni-rich compositions [1, 3, 18]. Mikula et al. [19], showed through their XRD studies that the β phase has a B2 structure at 900 °C. This was also reported by Booth et al.[20] at 550 °C and 830 °C. They concluded that Ni atoms occupy Co sites in the β phase and the Co atoms that were displaced in the Co sublattice move to the Ga sublattice.

The β phase and γ phases occupy a wide concentration range of the phase diagram and it is known that the β phase is the one that undergoes the martensitic transformation to the L1₀ phase [14, 16]. It is thus important to correctly ascertain the composition region where this martensitic phase is stable. Ducher et al. [14, 21, 22] performed experiments using the diffusion triple technique to determine the phase equilibria in the ternary at isothermal sections of 700 °C and 1000 °C. Their study focused on the Ni-rich composition regions where the martensitic transformations are observed. Equilibrium concentrations (at. %) for the two temperatures were accurately determined, which are used in the present work. At temperatures of 700 °C and below, Ducher et al. [23] and Ipser et al.[6] reported the presence of the δ -Ni₅Ga₃ and ϵ -Ni₁₃Ga₉ phases, in their models of the NiGa binary system. It was found that these phases penetrate into the ternary, continuing in the CoGa direction parallel to CoNi. Also, these phases are known to cause detrimental effects in the ternary. Vertical sections of 20 at. % Ni with varying compositions of Co and Ga was calculated by Ducher et al. [14]. This section forms a basis for the microstructural control in the CoNiGa SMA.

The crystal structure of the ternary β phase is an ordered cubic, B2-type structure. Different chemical models were suggested [24–28] to describe the β phase with point defects and vacancy concentration using a triple defect structure in the binary CoGa and NiGa subsystems. The models suggest that this phase is a substitutional solid solution for higher compositions of Co and deviates from this substitutional behavior at higher compositions of Ga, in which large concentrations of vacancies are responsible for the observed deviations from non-stoichiometry. This model was used by Su and Tedenac [4] in their binary CoGa system, where the cobalt and gallium atoms combined on the 1st sublattice and cobalt and vacancies combined on the 2nd sublattice. The two-sublattice model for the β phase is given by (Co,Ga)_{0.5}(Co,Va)_{0.5}. We also incorporated the same defect structure for the β phase while remodeling the CoGa system[5] Incidentally, Yuan and Ipser[6] in their model for NiGa, have also described the β phase using the same description for its defect structure, which we have used to model the ternary system. Gröbner [29] did not use the triple defect model.

However, limited information exists in literature on the thermodynamic properties and vacancy concentration of the ternary intermetallic β phase which exhibits triple defects. Mikula et al.[19] through their emf studies were able to predict these properties for a wide range of concentrations. They calculated the activity curves and partial enthalpies of the ternary alloy for varying at.% Ga. They were also able to predict the thermodynamic properties of the β phase (in the temperature range of 800-1000 °C) that showed triple-defects in the NiGa and CoGa binary systems. These calculations were also extended to the ternary. Three sections with constant ratios of x_{Co}/x_{Ni} taken as 3:1, 1:1 and 1:3 with varying Ga content of 40,45,50,55 and 60 at.% were studied by Mikula et al.[19]. The emfs obtained through their experiments were used in this work to calculate the activity and partial molar enthalpy. The data for the partial molar enthalpy and activities were then used as data points to check the consistency of the model by comparing the experimental data and the values predicted in the present work.

The order-disorder transformation phenomenon between the γ and γ' phases has not been

included in the present work as it is known to hinder the shape memory performance characteristics [3, 15, 30] and Magnetic Field-Induced Strain (MFIS) [3, 31]. This will be incorporated in future work.

2.1. Two-phase tie-lines

Tie lines are those that connect the composition of phases with equal chemical potentials μ_i . Compositions of the β , γ and γ' phases were taken from the experimental investigations of Oikawa et al. and Ducher et al. [1, 2, 14], in order to optimize the model parameters and calculate the equilibria among the phases.

The experimental data for these tie lines was obtained from the work of Ducher et al. [14] who used the diffusion triple (DT) method to determine phase equilibria over a wide composition range at 700 °C and 1000 °C. Oikawa et al. [1] measured the composition of the β and γ phases in equilibrium for a temperature range of 1000-1300 °C using energy dispersion X-ray spectroscopy (EDX). The experimental tie-lines indicate an equality in chemical potential of the element that is common to the two binary subsystems [32].

3. Thermodynamic model

The phase description of the binary systems that make up the ternary have been taken from the work of Ipser et al. [6] for the NiGa binary, Chari et al. [5] for the CoGa system and the thermodynamic description of the CoNi system has been taken from the SGTE solution database.

The thermodynamic phase description of the CoGa system by Chari et al. [5] is a remodeled system based on the previous work by Su and Tedenac [4]. The re-assessment was necessary to account for the stability conditions of the β phase at higher temperatures. The remodeled phase diagram is shown in Fig.1. The β phase in the binary systems is non-magnetic and has been modeled as such. The β phase in the ternary system is known to be ferromagnetic. However the Curie temperature across most of the compositional range is below 150 °C [2], hence the magnetic behavior of the β phase may be safely neglected since the assessment focuses on elevated temperatures. Ipser et al. [6] described the order-disorder transformation between the non stoichiometric Ni_3Ga (γ') and fcc(γ) phases, while modeling the NiGa system. The NiGa phase diagram as modeled by Ipser is shown in Fig.2. In order to consistently model the ternary, this phenomenon has not been taken into consideration. The Ni_5Ga_3 (δ) and Ni_3Ga_2 (ϵ) phases were modeled as stoichiometric phases entering into the ternary. They were modeled as non-stoichiometric phases as given by the experimental description of these phases by Ducher et al. [14] at 700 °C and this could be incorporated in future work. The phase diagram of the CoNibinary system as described in the SGTE solution database is shown in Fig.3.

4. Experimental procedure

In order to validate thermodynamic models for the CoNiGa ternary system, different CoNiGa alloy compositions (in at.%) were prepared using vacuum arc-melting of 99.9% purity Co, 99.95% Ni and 99.999% purity Ga. The compositions were chosen from different regions of the ternary system. From our experience on arc-melting of CoNiGa, usually Ga loss is around 3% of all Ga. Therefore, some extra Ga was added (\approx 3% more) to have a composition very close to target composition. WDS analyses confirmed that the measured compositions were very close to nominal compositions. Water cooled Cu crucible was used. No reaction with Ga was expected. Most importantly, Ga was

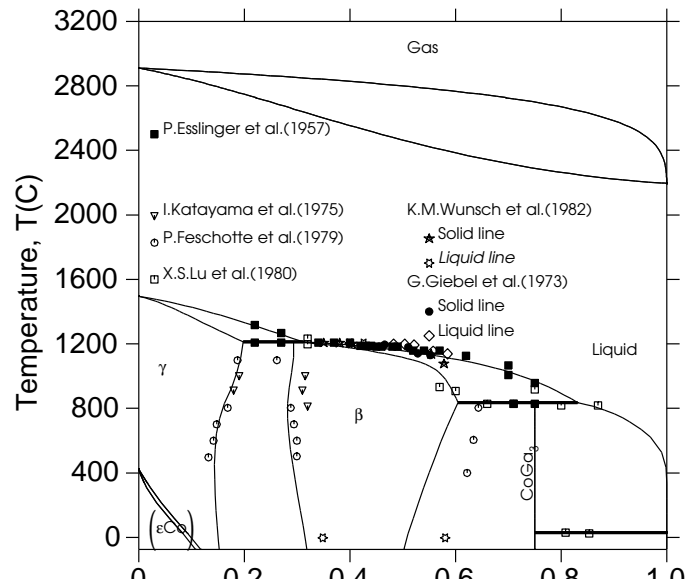


Figure 1: (

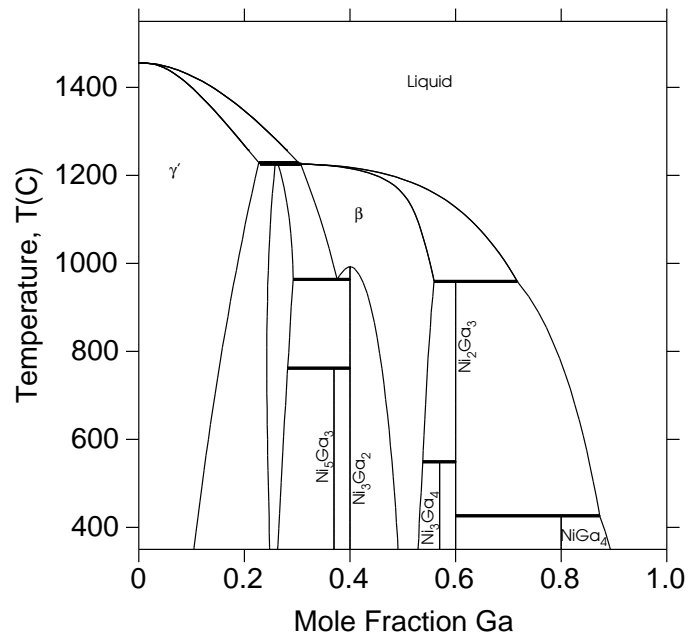


Figure 2: Binary NiGa phase diagram as calculated by Ipser et al [6]

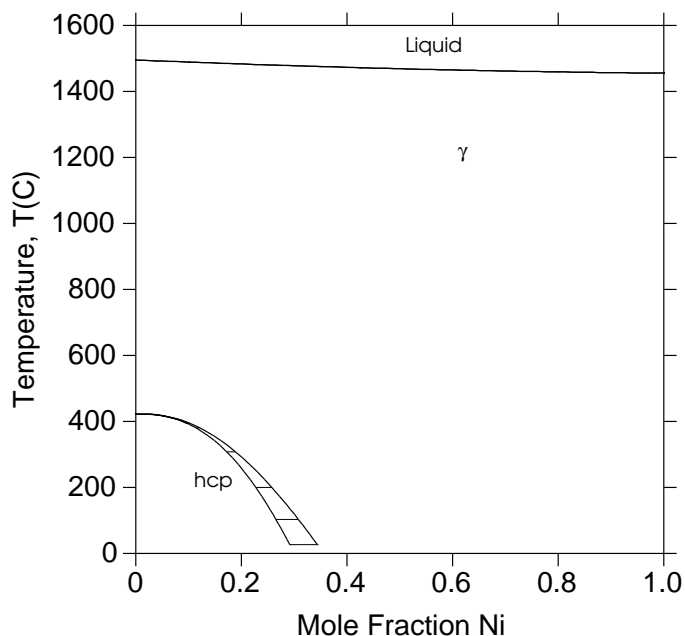


Figure 3: Binary CoNi phase diagram sourced from the SGTE solution database of ThermoCalc

added on top of other elements to minimize any reaction with crucible. In addition, no indication of a reaction on the crucible surface after melting was seen. These alloys were prepared in the form of arc melted buttons. The expected phases according to the thermodynamic model and experimentally observed phases at certain temperatures have been listed in Table 1.

The microstructural evolution and chemical analysis were examined using a digital Keyence VH-Z100 optical microscopy (OM) and a Cameca SX50 scanning electron microscopy (SEM) equipped with four wavelength dispersive X-Ray spectrometers (WDS). The samples examined using OM were etched in 50% hydrochloric acid, 33% ethanol, 8.5% copper sulfate and 8.5% water solution. Crystal structure of different phases were determined using a Bruker-AXS D8 X-ray diffractometer (XRD) with $\text{CuK}\alpha$ (0.15406 nm) radiation.

5. Optimization procedure

Experimental data including activities, volume fractions, tie-lines, triple points and partial molar enthalpies [1–3, 14, 19] were included in the optimization process, which was performed using the software ThermoCalc. The software uses a technique of global minimization of Gibbs energy to assess the thermodynamic model parameters such as the Gibbs energy functions and interaction parameters. The binary system parameters were taken from previously developed models of the NiGa and CoGa systems [5, 6] and these values remained fixed throughout the optimization process. Initial weights were given to the experimental data based on the accuracy required for the calculations. The weights were increased as the optimization process continued, depending on whether the results evaluated were within the uncertainty limits of the measurements. This was to ensure that self-consistent results were obtained.

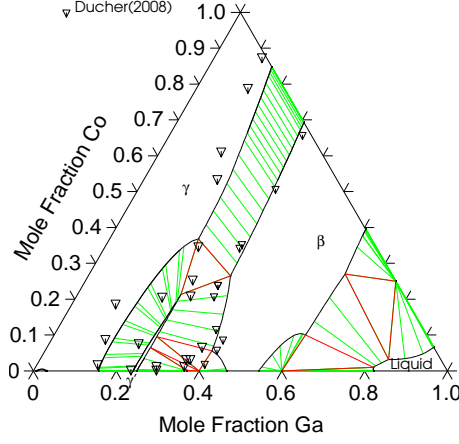


Figure 4: Isothermal section schematic phase diagram of the CoNiGa ternary system at 700 °C

All the model parameters were optimized together, starting with a small number of optimization iterations and increasing this number as the optimization progressed. Negative driving force constraints for some phases were enforced in regions of the ternary to prevent their stability in those regions.

The initial values of the ternary interaction parameters were determined by a trial and error method, until a reasonable fit was obtained with respect to the experimental data. The experimental data for vertical sections and volume fractions were given low weights throughout the optimization process, with higher weights given to the triple points and tie lines. As the optimization process progressed, the weights for the triple points were increased as described earlier.

In order for the binary system parameters to impact the ternary, it was necessary to include ternary interaction parameters as well as the enthalpies of formation for the Gibbs energy functions of the end members of the intermetallic phases of the Ni-Ga system. The enthalpies were calculated using first-principles within the GGA approximation with projector augmented-wave (PAW) pseudo-potentials, as implemented in the Vienna ab initio simulation package (VASP) [33, 34]. In the case of GGA, the PW91 corrections have been used [35]. Brillouin zone integrations were performed using a Monkhorst–Pack mesh [36]. Full relaxations were performed using the Methfessel–Paxton order 1 smearing method [37] and a final self-consistent static calculation with the tetrahedron smearing method with Blöchl corrections [38] was performed. The energy cutoff and k-point mesh used ensured convergence in the total energies calculated within less than 1 meV. The DFT calculations for the end members were included in the optimization process. The enthalpies of formation of the relevant phases are listed in Table 3. Experimental values obtained from literature for the same are also listed in the table.

6. Results and discussion

6.1. Isothermal sections at 700°C, 1000°C and 1200°C

6.1.1. Microstructure and two-phase equilibrium

The ternary phase diagram was optimized for a wide range of temperatures and the isothermal sections are shown in Figs. 4,5,6 as compared with experimental data. Equilibrium concentrations

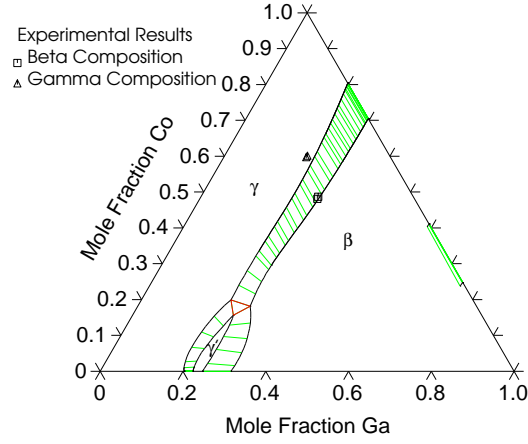


Figure 5: Isothermal section schematic phase diagram of the CoNiGa ternary system at 1000 °C

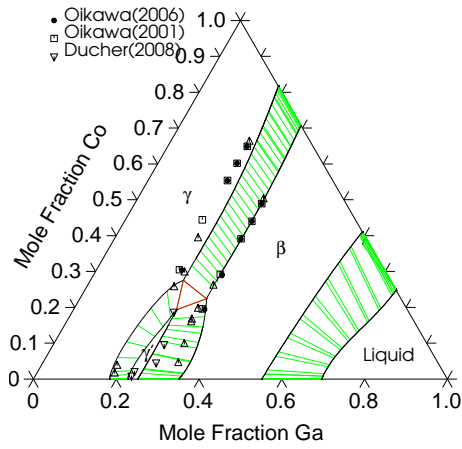


Figure 6: Isothermal section schematic phase diagram of the CoNiGa ternary system at 1200 °C

of the γ , β and γ' phases at 700 °C, 1000 °C and 1200 °C have been listed in Table 2. The optimized binary and ternary interaction parameters, including the Gibbs energy functions have been calculated using the model and are listed in Table 4. It can be observed from the three isothermal sections of the ternary that there are two distinct phases present: γ and β . This two-phase region extends into the ternary over a wide range of composition, from the CoGa side towards the Ni-rich side. The γ phase is present in the Co and Ni rich regions and the β phase is present in the central region of the ternary. The γ' phase can be observed at compositions up to 30 at. % Ni. Additional phases δ and ϵ are also observed at 700 °C (Fig. 4). These intermetallic phases are observed below 700 °C between 30-40 at.% Ga as also reported by Ducher et.al[14]. The single phase regions are parallel to the CoNi region, with no solubility change in Ga. The $\beta+\gamma$ two-phase region agrees well with the results reported by Ducher [14]. At 1000 °C (Fig 5), the concentrations at the two-phase $\beta+\gamma$ phase boundaries compare well with the experimental results by Oikawa et. al [1, 2]. The equilibrium data obtained between the β,γ and γ' phases (three-phase triple point region) is in agreement with the results by Ducher [14]. At 1200 °C, it can be seen that the β and γ phase compositions compare well with our experimental work (Fig 6). It is observed that the two-phase $\beta+\gamma$ region widens with decreasing temperature as also reported by Oikawa [1, 2].

6.1.2. Experimental results

Fig.7 displays an optical micrograph of the $\text{Co}_{0.30}\text{Ni}_{0.45}\text{Ga}_{0.25}$ sample after homogenization at 1077 °C for 24 hrs followed by water quenching indicating two phase microstructure ($\beta+\gamma$). Figures 8-9 show X-ray diffraction pattern of $\text{Co}_{0.20}\text{Ni}_{0.65}\text{Ga}_{0.15}$, $\text{Co}_{0.30}\text{Ni}_{0.45}\text{Ga}_{0.25}$, $\text{Co}_{0.80}\text{Ni}_{0.15}\text{Ga}_{0.05}$ and $\text{Co}_{0.05}\text{Ni}_{0.62}\text{Ga}_{0.33}$ alloys, respectively. Overall, there is a good agreement between expected and

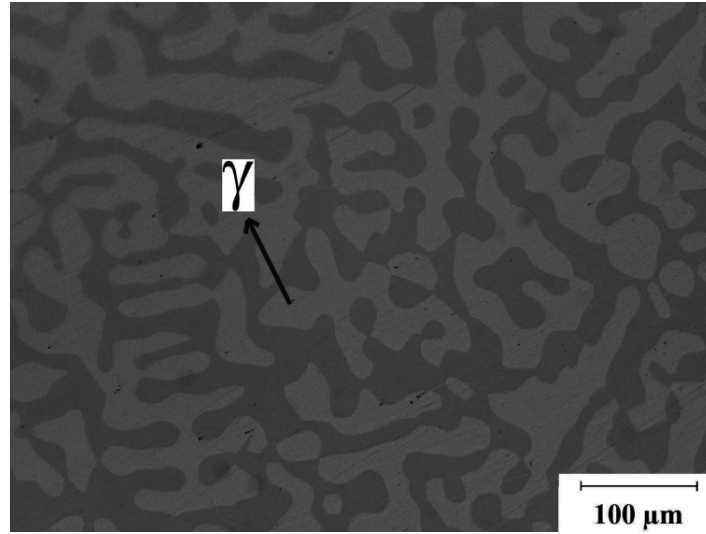


Figure 7: Optical micrograph of $\text{Co}_{0.30}\text{Ni}_{0.45}\text{Ga}_{0.25}$ sample after homogenization at 1077 °C for 24 hrs followed by water quenching indicating two phase microstructure ($\beta+\gamma$)

observed phases at the temperatures considered (Table 1). According to the thermodynamics model the stable phase at 1200 °C for $\text{Co}_{0.05}\text{Ni}_{0.62}\text{Ga}_{0.33}$ and at 1127 °C for $\text{Co}_{0.20}\text{Ni}_{0.65}\text{Ga}_{0.15}$ are β and γ phase respectively, which agree well with experimental results. For $\text{Co}_{0.20}\text{Ni}_{0.65}\text{Ga}_{0.15}$, the

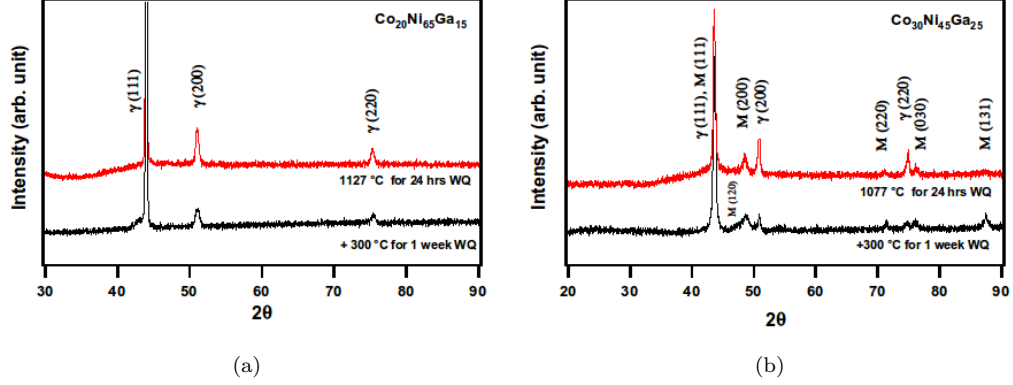


Figure 8: X-ray diffraction pattern of the (a) $\text{Co}_{0.2}\text{Ni}_{0.65}\text{Ga}_{0.15}$ and (b) $\text{Co}_{0.3}\text{Ni}_{0.45}\text{Ga}_{0.25}$ samples indicating the structures of the constitutive phases after different heat treatments. M: L10 Martensite, γ : A1 structure (disordered fcc)

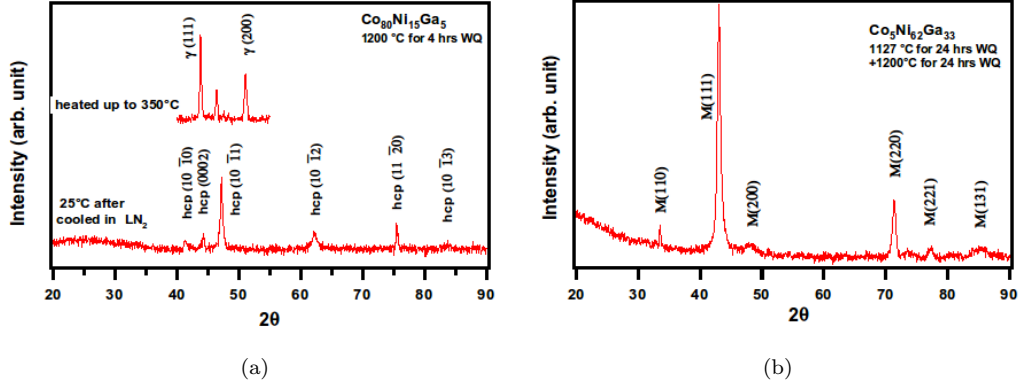


Figure 9: X-ray diffraction pattern of the (a) $\text{Co}_{0.8}\text{Ni}_{0.15}\text{Ga}_{0.05}$ and (b) $\text{Co}_{0.05}\text{Ni}_{0.62}\text{Ga}_{0.33}$ samples indicating the structures of the constitutive phases after different heat treatments. M: L10 Martensite, hcp: A3 structure

thermodynamic model indicates two phase microstructure γ and (γ') phases at 300 °C . Ageing at 300 °C for 1 week does not change the microstructure and stable structure is only the γ phase. For $\text{Co}_{0.30}\text{Ni}_{0.45}\text{Ga}_{0.25}$, the X-Ray diffraction pattern (Fig.8) reveals two phase microstructure of β and γ phases, which has a good agreement with the thermodynamics model for 1077 °C. However, aging at 300 °C for one week does not change the microstructure and the constitutive phases Fig.8. There is an inconsistency with the thermodynamics model which indicates Ni_5Ga_3 formation from β structure and γ phase when the sample is aged at 300 °C. The discrepancy between the model and the experiment for 300 °C could be attributed to the possibility that an aging time of one week is not sufficient to form such precipitates. Otherwise the thermodynamics model should be revised for this temperature. The expected phase at 1200 °C for $\text{Co}_{0.80}\text{Ni}_{0.15}\text{Ga}_{0.05}$ alloy is single γ phase structure however hcp phase has been observed experimentally as shown in Fig.9. This is due to martensitic transformation of γ phase to hcp phase. Fig. 9a shows XRD pattern of $\text{Co}_{0.80}\text{Ni}_{0.15}\text{Ga}_{0.05}$ after heating up to 350 °C, which indicated that hcp structure transforms back to fcc γ structure upon heating up to 350 °C.

6.2. Vertical sections, activities, partial enthalpy, phase fractions and sublattice site fraction of vacancy plots for different alloy compositions

Fig 10 shows a vertical section at 20 at. % Ni. Studies [1, 14] have shown that this section is the most important section for microstructure control in the CoNiGa ternary. The calculated phase boundaries have been compared with previously developed work. From Fig 10 it can be seen that the solubility of Ga in the β phase decreases with a decrease in temperature.

The chemical activity of Ga (Ref.State:Liquid) in the ternary β phase was determined for varying compositions of Ga at 900 °C and can be seen in Figs. 11, 12 and 13. It can be observed that the values of activities agree well with the results by Mikula et al.[19] for 45, 50 and 55 at.% Ga. Deviations are seen at 40 and 60 at.% Ga. The partial enthalpy of the ternary β phase was determined for varying compositions of Ga at 900 °C and can be seen in Figs. 14,15 and 16. It can be observed that the partial enthalpies are closer to Mikula's results [19] for 40, 45 and 50 at.% Ga, than at 55 and 60 at.% Ga. Mikula et al. [19] through their emf studies of the β phase, noticed considerable deviations of partial enthalpy and activity of Ga at 55 and 60 at. % when compared to theoretical calculations. A conclusion was hence made that that the phase boundary of the β phase in the ternary, could be close to or below 55 at. % Ga. In the present work, the β phase boundaries were found between 38 and 55 at. % Ga, hence agreeing with the experimental work by Mikula.

In order to understand the triple defect structure of the ternary β phase, the sublattice site fraction of vacancies in the second sublattice of the β phase with varying composition and temperature was calculated and is as shown in Figs. 17 and 18 respectively. Also, the phase fractions of all the phases at the compositions $\text{Co}_5\text{Ga}_{33}\text{Ni}_{62}$, $\text{Co}_{20}\text{Ga}_{15}\text{Ni}_{65}$, $\text{Co}_{30}\text{Ga}_{25}\text{Ni}_{45}$, $\text{Co}_{80}\text{Ga}_5\text{Ni}_{15}$ and 1200 °C have been calculated (based on the experimental compositions from Dr. Karaman's group) as shown in Figs. 19, 20, 21 and 22 respectively.

7. Conclusion

The phase equilibria among the γ (fcc), β , γ' (Ni_3Ga), δ (Ni_5Ga_3) and ϵ ($\text{Ni}_{13}\text{Ga}_9$) were determined at various temperatures using a computational (CALPHAD) approach coupled with experimental results in the CoNiGa High Temperature Shape Memory Alloy. The β phase was found to be observed in the central region of the ternary, parallel to the two-phase $\gamma+\beta$ region over a

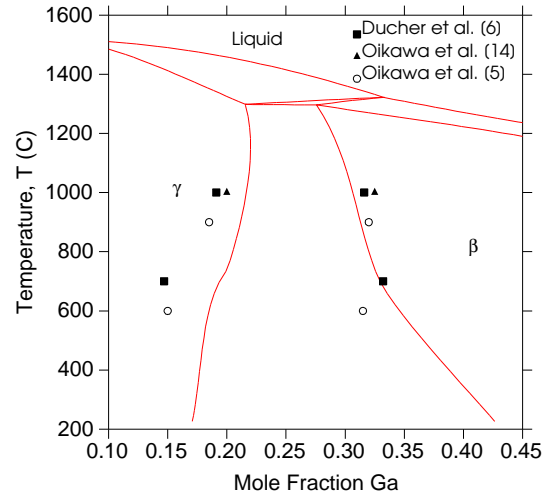


Figure 10: Vertical section of the CoNiGa system at 20 at. % Ni

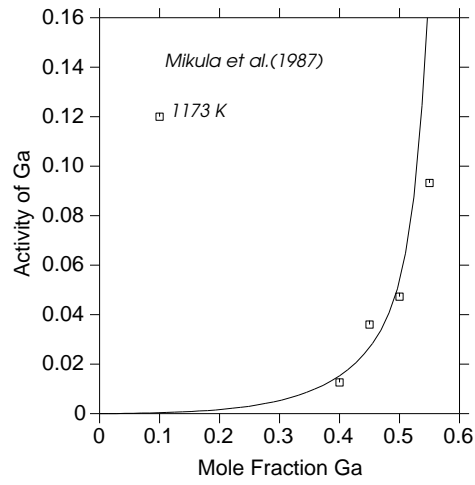


Figure 11: Activity of Ga in the β phase at $x_{Co}/x_{Ni}=3:1$ and 900 °C

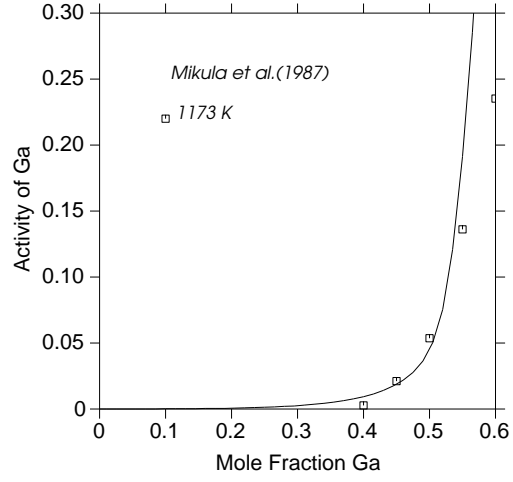


Figure 12: Activity of Ga in the β phase at $x_{Co}/x_{Ni}=1:3$ and 900°C

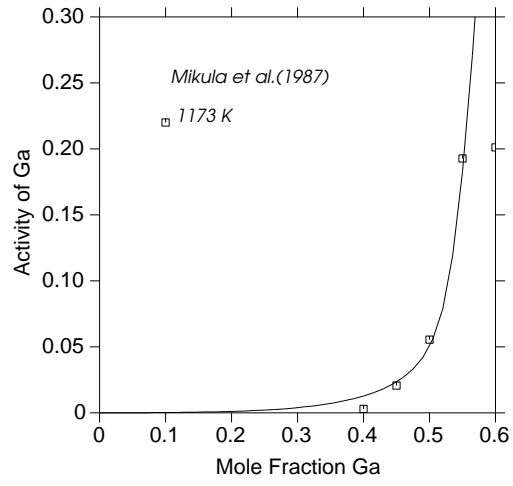


Figure 13: Activity of Ga in the β phase at $x_{Co}/x_{Ni}=1:1$ and 900°C

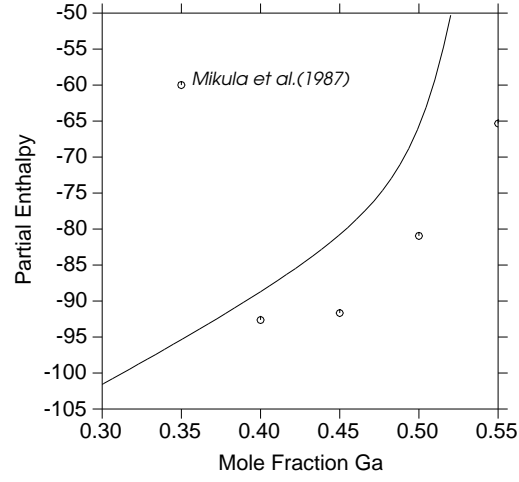


Figure 14: Partial Enthalpy of β phase at $x_{Co}/x_{Ni}=3:1$ and 900°C

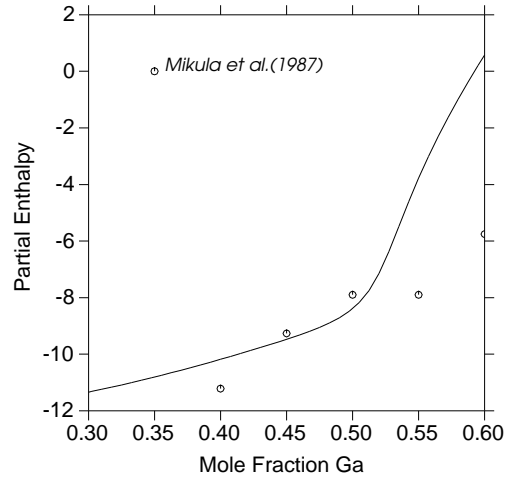


Figure 15: Partial Enthalpy of β phase at $x_{Co}/x_{Ni}=1:3$ and 900°C

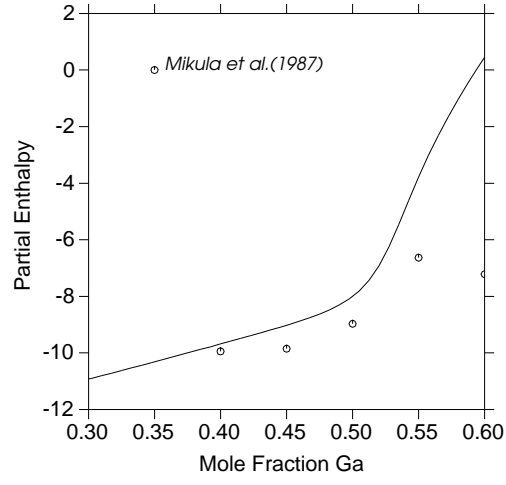


Figure 16: Partial Enthalpy of β phase at $x_{Co}/x_{Ni}=1:1$ and 900°C

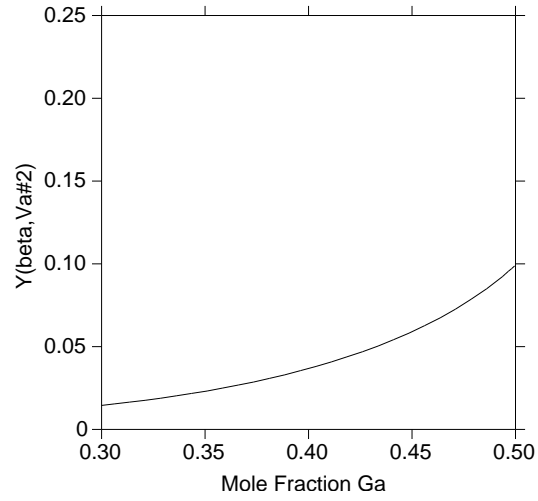


Figure 17: Sublattice site fraction of vacancies in the second sublattice of the ternary β phase as a function of composition at 900°C

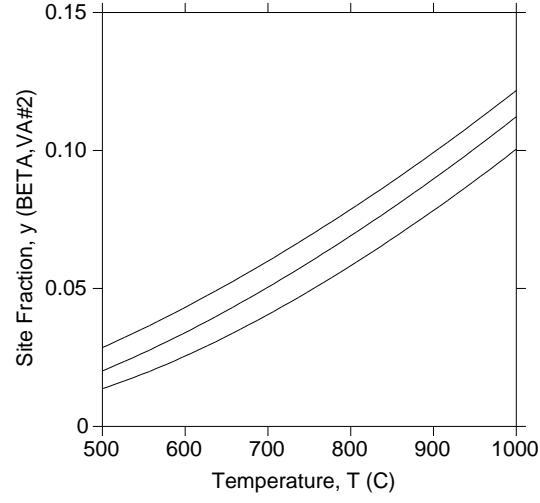


Figure 18: Sublattice site fraction of vacancies in the second sublattice of the ternary β phase as a function of temperature at 50, 52, 54 at. % Co and 50 at. % Ga

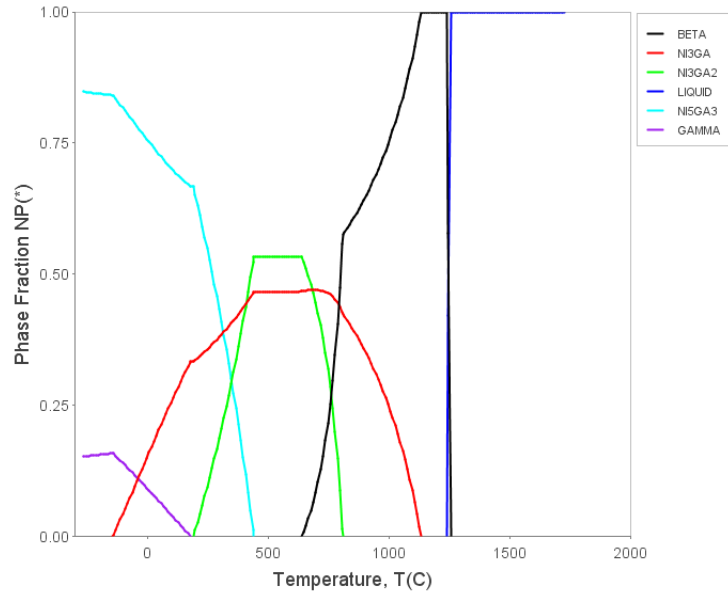


Figure 19: Phase fractions of all the phases in the ternary at $\text{Co}_{0.05}\text{Ga}_{0.33}\text{Ni}_{0.62}$ and 1200°C

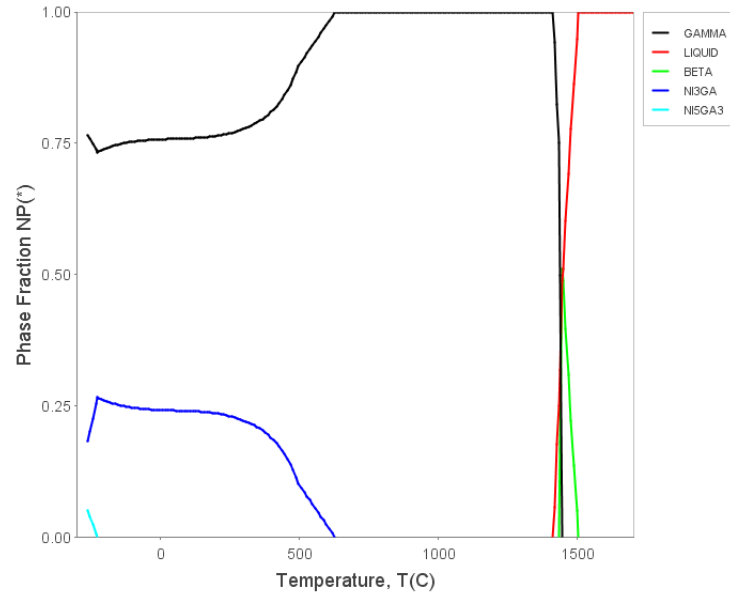


Figure 20: Phase fractions of all the phases in the ternary at $\text{Co}_{20}\text{Ga}_{15}\text{Ni}_{65}$ and 1200°C

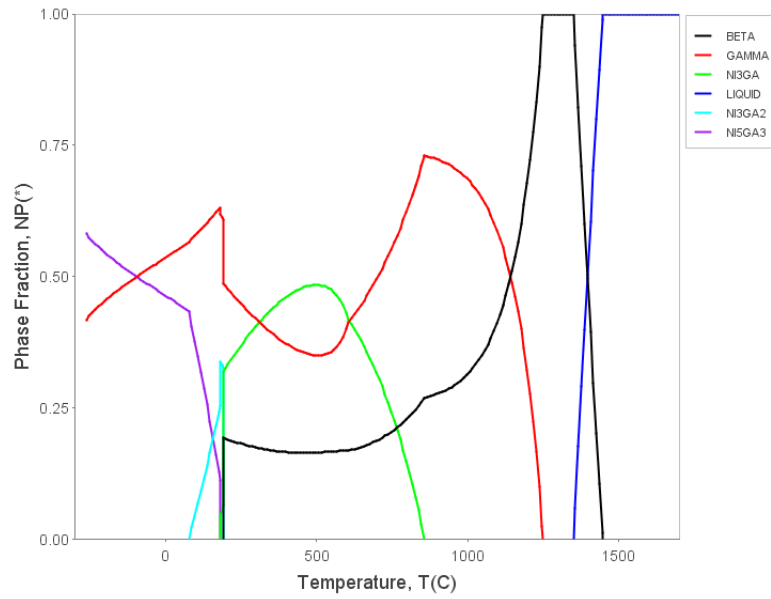


Figure 21: Phase fractions of all the phases in the ternary at $\text{Co}_{30}\text{Ga}_{25}\text{Ni}_{45}$ and 1200°C

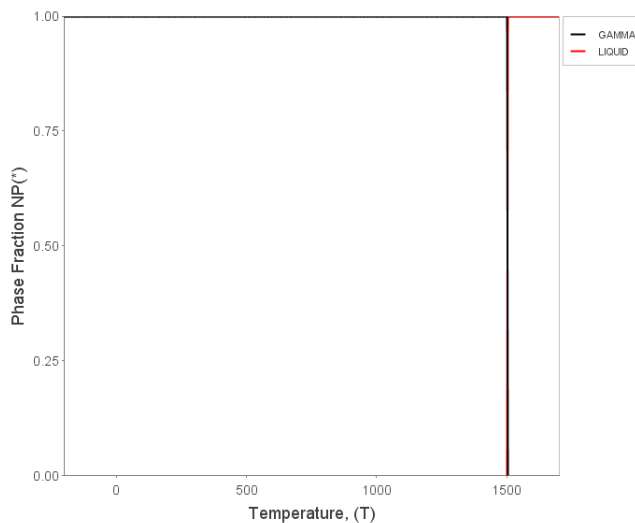


Figure 22: Phase fractions of all the phases in the ternary at $\text{Co}_{80}\text{Ga}_5\text{Ni}_{15}$ and 1200°C

wide range of compositions, extending from the CoGa side to the NiGa side as seen in the CoNiAl system and it decreases in width with an increase in temperature. The phases expected from the thermodynamic model compared well with the observed phases from the experimental work. Activities and partial enthalpies calculated from the model compare well with experimental results. This model can be improved in future work by including order-disorder phase transformations in the γ and γ' phases.

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Table 1: Heat treatments performed for six chosen alloys to corroborate phase stability in the calculated phase diagram

Composition	Temperature	Duration	Expected Phases	Observed Phases	γ phase vol.fraction(%)
Co _{0.05} Ni _{0.62} Ga _{0.33}	1127 °C	1 day	β	$\beta, \gamma', \text{hcp}$ (Ni _{1.235} Ga _{0.665})	0
	+1200 °C	1 day	β	β	0
	+300 °C	1 week	$\gamma', \gamma, \text{Ni}_5\text{Ga}_3$	β, hcp (Ni _{1.235} Ga _{0.665})	
Co _{0.15} Ni _{0.8} Ga _{0.05}	1200 °C	4 hours	γ	γ	100
Co _{0.2} Ni _{0.65} Ga _{0.15}	1127 °C	1 day	γ	γ	100
	300 °C	1 week	γ, γ'	γ	
Co _{0.3} Ni _{0.45} Ga _{0.25}	1077 °C	1 day	β, γ	β, γ	50
	300 °C	1 week	$\gamma, \text{Ni}_5\text{Ga}_3$	β, γ	
Co _{0.6} Ni _{0.1} Ga _{0.3}	1027 °C	1 day	β	β	0
Co _{0.8} Ni _{0.15} Ga _{0.05}	1200 °C	4 hours	γ	hcp (Co ₃ Ni) $\leftarrow \gamma$	

Table 2: Composition of β , γ and γ' phases at various temperatures in the CoNiGa system, compared with experimental data. Values of the experimental data are in brackets

T, °C	β Phase		γ Phase		γ' Phase		Ref.
	At. % Co	At. % Ga	At. % Co	At. % Ga	At. % Co	At. % Ga	
700 °C	33.10 (34.4)	33.54 (32.6)	49.70 (53.6)	21.42 (17.6)	-	-	[14]
	34.61 (35.4)	33.30 (32.7)	54.53 (61.3)	20.53 (14.7)	-	-	
	51.21 (50.8)	32.39 (33.1)	71.74 (79.1)	17.18 (12.3)	-	-	
	67.21 (65.8)	30.25 (32.1)	80.49 (87.6)	16.22 (11.4)	-	-	
	-	-	7.45 (9.00)	15.29 (12.9)	6.54 (2.56)	23.64 (25.7)	
	-	-	2.02 (2.00)	15.46 (14.5)	5.08 (2.12)	23.34 (27.4)	
	21.03 (20.8)	36.01 (33.2)	-	-	19.22 (20.8)	25.45 (20.7)	
1000 °C	25.56	30.44	35.82	23.01	-	-	[1] ¹ , [2] ²
	(19.6 ¹ , 19.49 ²)	(31.0 ¹ , 31.6 ²)	(30.5 ¹ , 30.4 ²)	(20.0 ¹ , 20.67 ²)	-	-	
	28.58	30.46	39.14	22.63	-	-	[1] ¹ , [2] ²
	(29.2 ¹ , 29.1 ²)	(30.5 ¹ , 31.06 ²)	(44.40 ¹ , 53.31 ²)	(18.61, 10.10 ²)	-	-	
	38.54	30.57	51.21	22.45	-	-	[1] ¹ , [2] ²
	(39.1 ¹ , 39.08 ²)	(30.6 ¹ , 30.58 ²)	(55.4 ¹ , 55.34 ²)	(19.2 ¹ , 19.21 ²)	-	-	
	43.37	30.70	56.34	21.92	-	-	[1] ¹ , [2] ²
	(44.0 ¹ , 44.01 ²)	(30.80 ¹ , 30.77 ²)	(60.2 ¹ , 60.17 ²)	(19.1 ¹ , 19.11 ²)	-	-	
	48.80	30.54	61.47	21.40	-	-	[1] ¹ , [2] ²
	(48.9 ¹ , 48.88 ²)	(30.7 ¹ , 30.72 ²)	(64.9 ¹ , 64.85 ²)	(19.2 ¹ , 19.17 ²)	-	-	
	-	-	4.13 (3.70 ³)	17.97	2.32 (1.00 ³)	21.93	[14] ³
	-	-	2.32 (1.50 ³)	(18.40 ³)		(23.20 ³)	
1200 °C	9.26 (9.70 ³)	34.76	-	-	0.51 (2.30 ³)	22.58	[14] ³
		(31.60 ³)	-	-	10.17 (9.80 ³)	25.65	
	4.73 (4.40 ³)	34.48	-	-	5.34 (4.70 ³)	25.51	[14] ³
		(32.80 ³)	-	-		(27.30 ³)	
1200 °C	49.1 (48.65)	28.35 (28.36)	58.46 (59.8)	22.14 (20.19)	-	-	Exp
	47.89 (48.13)	28.19 (28.4)	57.85 (59.36)	22.19 (20.08)	-	-	Exp

Table 3: Enthalpy of formation calculated with LDA, PBE and GGA approximations and comparison with experimental results

Phase	Structure	H_f (J/mol)			
		GGA	PBE	LDA	Other Work
Ni ₃ Ga	Al ₃ Zr	-26919	-27594	-32901	-27800(298 K)[29] ¹ -23191.5(300 K)[41] ² -33883(873 K)[42] ³ -33064.4(1223 K)[43] ³ -28138.7(298 K)[44] ²
	L12	-26726	-27594	32226	-
Ni ₅ Ga ₃	Pt ₅ Ga ₃	-34349	-35699	-41971	-35500(298 K)[29] ¹ -40080(298 K)[44] ²
Ni ₃ Ga ₄	Ni ₃ Ga ₄	-35603	-37725	-44576	-35700(298 K)[29] ¹
Ni ₂ Ga ₃	Ni ₂ Al ₃	-37243	-39173	-45541	-33600(298 K)[29] ¹ -45106.4(300 K)[41] ² -50341.1(873 K)[42] ³
	Bi ₂ Te ₃	-16209	-18428	-22191	-
NiGa ₄	Cu ₅ Zn ₈	10710	8587	10710	-24300(298 K)[29] ¹ -22127.7(300 K)[41] ²

Table 4: Optimized parameters for the phases in the CoNiGa system

Phase	Parameter	Description	Ref.
Gas	${}^0G_{\text{Co}}^{\text{Gas}}$	$F6985T + RT\ln(1 * 10^{-5}P)$	[5]
	${}^0G_{\text{Ga}}^{\text{Gas}}$	$F9633T + RT\ln(1 * 10^{-5}P)$	[5]
	${}^0G_{\text{Ga2}}^{\text{Gas}}$	$F9695T + RT\ln(1 * 10^{-5}P)$	[5]
Liquid	${}^0L_{\text{Co,Ga}}^{\text{Liq}}$	$-61807 + 7.985T$	[5]
	${}^1L_{\text{Co,Ga}}^{\text{Liq}}$	12605.5	[5]
	${}^0L_{\text{Co,Ga,Ni}}^{\text{Liq}}$	$66048.054 + 3.162T$	[This work]
	${}^0L_{\text{Co,Ni}}^{\text{Liq}}$	1331	[SGTE]
	${}^0L_{\text{Ga,Ni}}^{\text{Liq}}$	$-122488.59 + 35.72T$	[6]
	${}^1L_{\text{Ga,Ni}}^{\text{Liq}}$	$-29685 + 14T$	[6]
	${}^2L_{\text{Ga,Ni}}^{\text{Liq}}$	$-30751.9 + 22.1T$	[6]
γ (fcc)	${}^0L_{\text{Co,Ga}}^{\text{fcc}}$	$-125202.28 + 54.131T$	[5]
	${}^1L_{\text{Co,Ga}}^{\text{fcc}}$	$30657 - 25.625T$	[5]
	${}^0L_{\text{Co,Ga,Ni}}^{\text{fcc}}$	$-1.7622 - 38.595T$	[This work]
	${}^1L_{\text{Co,Ga,Ni}}^{\text{fcc}}$	$0.8857 - 1.176T$	[This work]

Continued on next page

¹Calculated

²Experimental-Calorimetry

³Experimental-emf

Table4–Continued

	${}^2L_{\text{Co,Ga,Ni}}^{\text{fcc}}$	$0.45 - 53.422T$	[This work]
	${}^0L_{\text{Co,Ni}}^{\text{fcc}}$	$-800 + 1.2629T$	[SGTE]
	${}^0L_{\text{Ga,Ni}}^{\text{fcc}}$	$-130526 + 40T$	[6]
Hcp	${}^0L_{\text{Co,Ga}}^{\text{hcp}}$	$-87051 + 22.438T$	[5]
	${}^0L_{\text{Co,Ni}}^{\text{hcp}}$	$-1620 - 0.385T$	[SGTE]
CoGa ₃	${}^0G_{\text{Co:Ga}}^{\text{CoGa}_3} - 0.25{}^0G_{\text{Co}}^{\text{hcp}} - 0.75{}^0G_{\text{Ga}}^{\text{ort}}$	$-30770 + 3.043T$	[5]
β	${}^0G_{\text{Co:Co}}^{\beta} = {}^0G_{\text{Co}}^{\text{bcc}}$	${}^0G_{\text{Co}}^{\text{hcp}} + 2938 - 0.7138T$	[4]
	${}^0G_{\text{Ga:Co}}^{\beta} - 0.5{}^0G_{\text{Ga}}^{\text{ort}} - 0.5{}^0G_{\text{Co}}^{\text{hcp}}$	$-42125 + 9.519T$	[5]
	${}^0G_{\text{Ni:Co}}^{\beta} - 0.5{}^0G_{\text{Ni}}^{\text{bcc}} - 0.5{}^0G_{\text{Co}}^{\text{bcc}}$	39.147	[This work]
	${}^0G_{\text{Co:Ni}}^{\beta} - 0.5{}^0G_{\text{Ni}}^{\text{bcc}} - 0.5{}^0G_{\text{Co}}^{\text{bcc}}$	$4853.355 - 8.378T$	[This work]
	${}^0G_{\text{Ga:Ni}}^{\beta} - 0.5{}^0G_{\text{Ga}}^{\text{bcc}} - 0.5{}^0G_{\text{Ni}}^{\text{bcc}}$	$-54030.75 + 16.5T$	[6]
	${}^0G_{\text{Ni:Ni}}^{\beta}$	${}^0G_{\text{Ni}}^{\text{bcc}}$	[6]
	${}^0G_{\text{Co:Va}}^{\beta}$	$0.5{}^0G_{\text{Co}}^{\text{hcp}} + 52313 - 16.5828T$	[5]
	${}^0G_{\text{Ga:Va}}^{\beta}$	$0.5{}^0G_{\text{Ga}}^{\text{ort}} + 7250 - 6.35T$	[5]
	${}^0G_{\text{Ni:Va}}^{\beta}$	${}^0G_{\text{Ni:Va}}^{\text{bcc}}$	[6]
	${}^0L_{\text{Co,Ga:Co}}^{\beta}$	$-11752 + 3.505T$	[5]
	${}^0L_{\text{Co,Ni:Co}}^{\beta}$	$68.236 - 0.5264T$	[This work]
	${}^0L_{\text{Co:Co,Va}}^{\beta}$	$-6847 + 0.6913T$	[5]
	${}^0L_{\text{Co:Co,Ni}}^{\beta}$	$-14.586 + 1.7T$	[This work]
	${}^0L_{\text{Ga,Ni:Co}}^{\beta}$	$-50.724 - 5.817T$	[This work]
	${}^0L_{\text{Ga:Co,Va}}^{\beta}$	$-24462 + 9.677T$	[5]
	${}^0L_{\text{Co,Ga:Ni}}^{\beta}$	$-12540.193 + 0.2745T$	[This work]
	${}^0L_{\text{Ga,Ni:Ni}}^{\beta}$	$-8724 - 2.38T$	[6]
	${}^0L_{\text{Ga:Ni,Va}}^{\beta}$	$-35016.42 + 20.31T$	[6]
	${}^0L_{\text{Ni:Ni,Va}}^{\beta}$	$-35016.42 + 20.31T$	[6]
	${}^0L_{\text{Co,Ga:Va}}^{\beta}$	$-7557 - 0.3907T$	[5]
	${}^0L_{\text{Ga,Ni:Va}}^{\beta}$	$-8724 - 2.38T$	[6]
Ni ₂ Ga ₃	${}^0G_{\text{Ni:Ga}}^{\text{Ni}_2\text{Ga}_3} - 0.4{}^0G_{\text{Ni}}^{\text{fcc}} - 0.6{}^0G_{\text{Ga}}^{\text{ort}}$	$-47426.09 + 8.94T$	[6]
γ' (Ni ₃ Ga)	${}^0G_{\text{Co:Co}}^{\text{Ni}_3\text{Ga}}$	${}^0G_{\text{Co}}^{\text{fcc}}$	[SGTE]
	${}^0G_{\text{Ga:Co}}^{\text{Ni}_3\text{Ga}} - 0.25{}^0G_{\text{Co}}^{\text{hcp}} - 0.75{}^0G_{\text{Ga}}^{\text{ort}}$	9116.964	[5]
	${}^0G_{\text{Co:Ga}}^{\text{Ni}_3\text{Ga}} - 0.75{}^0G_{\text{Co}}^{\text{hcp}} - 0.25{}^0G_{\text{Ga}}^{\text{ort}}$	-5758.22	[5]
	${}^0G_{\text{Ga:Ga}}^{\text{Ni}_3\text{Ga}}$	${}^0G_{\text{Ga}}^{\text{fcc}}$	[SGTE]
	${}^0G_{\text{Ni:Ga}}^{\text{Ni}_3\text{Ga}} - 0.75{}^0G_{\text{Ni}}^{\text{fcc}} - 0.25{}^0G_{\text{Ga}}^{\text{fcc}}$	$-27789 + 5.24T$	[6]
	${}^0G_{\text{Co:Ni}}^{\text{Ni}_3\text{Ga}} - 0.75{}^0G_{\text{Co}}^{\text{hcp}} - 0.25{}^0G_{\text{Ni}}^{\text{fcc}}$	314	[This work]
	${}^0G_{\text{Ga:Ni}}^{\text{Ni}_3\text{Ga}} - 0.75{}^0G_{\text{Ga}}^{\text{fcc}} - 0.25{}^0G_{\text{Ni}}^{\text{fcc}}$	3200	[6]
	${}^0G_{\text{Ni:Ni}}^{\text{Ni}_3\text{Ga}}$	${}^0G_{\text{Ni}}^{\text{fcc}}$	[SGTE]
	${}^0L_{\text{Co,Ni:Ga}}^{\text{Ni}_3\text{Ga}}$	-24445.144	[This work]
	${}^0L_{\text{Ga,Ni:Ga}}^{\text{Ni}_3\text{Ga}}$	$-25578 + 3T$	[This work]

Continued on next page

Table4–Continued			
	${}^0L_{\text{Ga},\text{Ni:Ni}}^{\text{Ni}_3\text{Ga}}$	$-25578 + 3T$	[This work]
	${}^0L_{\text{Ga},\text{Ni:Co}}^{\text{Ni}_3\text{Ga}}$	$-25578 + 3T$	[This work]
	${}^1L_{\text{Ga},\text{Ni:Ga}}^{\text{Ni}_3\text{Ga}}$	$14040 - 8T$	[This work]
	${}^1L_{\text{Ga},\text{Ni:Ni}}^{\text{Ni}_3\text{Ga}}$	$14040 - 8T$	[This work]
	${}^1L_{\text{Ga},\text{Ni:Co}}^{\text{Ni}_3\text{Ga}}$	$14040 - 8T$	[This work]
	${}^1L_{\text{Ga:Ga},\text{Ni}}^{\text{Ni}_3\text{Ga}}$	$7841.97 - 4.133T$	[This work]
	${}^1L_{\text{Ni:Ga},\text{Ni}}^{\text{Ni}_3\text{Ga}}$	$7841.97 - 4.133T$	[This work]
	${}^1L_{\text{Co:Ga},\text{Ni}}^{\text{Ni}_3\text{Ga}}$	$7841.97 - 4.133T$	[This work]
Ni_3Ga_2	${}^0G_{\text{Ni:Ga}}^{\text{Ni}_3\text{Ga}_2} - 0.4{}^0G_{\text{Ga}}^{\text{ort}} - 0.6{}^0G_{\text{Ni}}^{\text{fcc}}$	$-39753.31 + 5.55T$	[6]
Ni_3Ga_4	${}^0G_{\text{Ni:Ga}}^{\text{Ni}_3\text{Ga}_4} - 0.57{}^0G_{\text{Ga}}^{\text{ort}} - 0.43{}^0G_{\text{Ni}}^{\text{fcc}}$	$-47790.8 + 9.04T$	[6]
Ni_5Ga_3	${}^0G_{\text{Ni:Ga}}^{\text{Ni}_5\text{Ga}_3} - 0.37{}^0G_{\text{Ga}}^{\text{ort}} - 0.63{}^0G_{\text{Ni}}^{\text{fcc}}$	$-37658.61 + 5.34T$	[6]
NiGa_4	${}^0G_{\text{Ni:Ga}}^{\text{NiGa}_4} - 0.8{}^0G_{\text{Ga}}^{\text{ort}} - 0.2{}^0G_{\text{Ni}}^{\text{fcc}}$	$-24367.51 - 2.71T$	[6]

A. Appendix : Functions in the gas phase description

F6985:

$$\begin{aligned}
& 418307.916 - 34.9920304T - 20.82348T\ln T - .00804072T^2 + 1.94275e^{-6}T^3 + 69435.55T^{-1} \\
& \hspace{20em} (298.14 < T < 600.00) \\
& 417194.563 - 4.48128666T - 25.91852T\ln T - 3.2169625e^{-4}T^2 + 1.22796583e^{-8}T^3 \\
& + 69799.8T^{-1} \hspace{15em} (600.00 < T < 1600.00) \\
& 405652.322 + 60.9499874T - 34.4754T\ln T + .0022698485T^2 - 1.11742517e^{-7}T^3 \\
& + 2845478T^{-1} \hspace{15em} (1600.00 < T < 5300.00) \\
& 616085.763 - 445.433806T + 24.5681T\ln T - .00520688T^2 + 6.847335e^{-8}T^3 \\
& - 1.3989465e^8T^{-1} \hspace{15em} (5300.00 < T < 10000.00)
\end{aligned}$$

F9633:

$$\begin{aligned}
& 259072.278 + 88.0130706T - 38.71057T\ln T + .01053784T^2 - 9.86907833e^{-7}T^3 \\
& + 338489.2T^{-1} \hspace{15em} (298.14 < T < 600.00) \\
& 263812.519 + 33.4871435T - 30.75007T\ln T + .00537745T^2 - 5.46534e^{-7}T^3 \\
& - 150942.65T^{-1} \hspace{15em} (600.00 < T < 1400.00) \\
& 270292.501 - 28.1810494T - 21.9834T\ln T + 3.192416e^{-4}T^2 - 1.46299133e^{-8}T^3 \\
& - 992093T^{-1} \hspace{15em} (1400.00 < T < 6000.00) \\
& 340110.007 - 140.262257T - 9.704267T\ln T - 4.5138725e^{-4}T^2 - 1.13427367e^{-8}T^3 \\
& - 68387950T^{-1} \hspace{15em} (6000.00 < T < 10000.00)
\end{aligned}$$

F9695:

$$\begin{aligned}
& 422882.385 - 36.0787973T - 33.72863T\ln T - .009368525T^2 + 7.62775167e^{-7}T^3 \\
& - 19520.385T^{-1} \hspace{15em} (298.14 < T < 1100.00) \\
& 419324.178 + 8.33965897T - 40.33555T\ln T - .0041854135T^2 + 2.679565e^{-8}T^3 \\
& + 312119.6T^{-1} \hspace{15em} (1100.00 < T < 2500.00)
\end{aligned}$$

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