



# Mark Scheme (Final)

Summer 2023

Pearson Edexcel International Advanced  
Subsidiary Level In Chemistry (WCH12)

Paper 01

Unit 2: Energetics, Group Chemistry,  
Halogenoalkanes and Alcohol

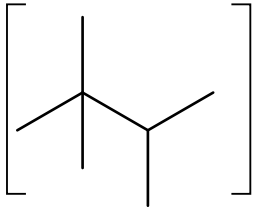
### Section A (Multiple Choice)

Question number	Answer	Mark
1	<p><b>The only correct answer is C</b> (<math>2\text{Cl(g)} \rightarrow \text{Cl}_2\text{(g)}</math>)</p> <p><i>A is incorrect because the diagram represents an exothermic reaction and atomisation is always endothermic</i></p> <p><i>B is incorrect because the diagram represents an exothermic reaction and ionisation is always endothermic</i></p> <p><i>D is incorrect because the diagram represents an exothermic reaction and dissolving <math>\text{NH}_4\text{NO}_3</math> is endothermic</i></p>	(1)

Question number	Answer	Mark
2(a)	<p><b>The only correct answer is C</b> (14.7 %))</p> <p><i>A is incorrect because <math>\pm 7.37</math> is an uncertainty based on halving the difference between the experimental and data book values and taking this as a percentage of the data book value</i></p> <p><i>B is incorrect because <math>\pm 8.65</math> is an uncertainty based on halving the difference between the experimental and data book values and taking this as a percentage of the experimental value</i></p> <p><i>D is incorrect because 17.3 compares the difference in values to the experimental rather than the data book value</i></p>	(1)

Question number	Answer	Mark
2(b)	<p><b>The only correct answer is B</b> (lowers the error in the final value obtained)</p> <p><i>A is incorrect because increasing the specific heat capacity increases the magnitude of the final value which will then be closer to the data book value</i></p> <p><i>C is incorrect because the difference is 8.6% which is significant</i></p> <p><i>D is incorrect because 8.6% is large compared with the measurement uncertainties</i></p>	(1)

Question number	Answer	Mark
3	<p><b>The only correct answer is A</b>(+491 kJ mol<sup>-1</sup>)</p> <p><i>B is incorrect because the <math>\Delta_r H^\ominus</math> value for the formation of carbon monoxide has not been tripled</i></p> <p><i>C is incorrect because -491kJ mol<sup>-1</sup> is the enthalpy change for the reverse reaction</i></p> <p><i>D is incorrect because -713kJ mol<sup>-1</sup> is a calculation for the reverse reaction in which the <math>\Delta_r H^\ominus</math> value for the formation of carbon monoxide has not been tripled</i></p>	(1)

Question number	Answer	Mark
4	<p>The only correct answer is D</p>  <p>A is incorrect because the alkane with the most carbon atoms will always have the highest boiling temperature</p> <p>B is incorrect because the alkane with the most carbon atoms will always have the highest boiling temperature</p> <p>C is incorrect because the alkane with the most carbon atoms will always have the highest boiling temperature</p>	(1)

Question number	Answer	Mark
5	<p>The only correct answer is C (cyclohexane molecules have a larger area of contact)</p> <p>A is incorrect because the cyclohexane molecules have fewer electrons than hexane molecules</p> <p>B is incorrect because these molecules have negligible permanent dipole-permanent dipole forces</p> <p>D is incorrect because neither molecule forms hydrogen bonds</p>	(1)

Question number	Answer	Mark
6	<p><b>The only correct answer is B</b> (propanal molecules have strong permanent dipole-permanent dipole forces)</p> <p><i>A is incorrect because butane molecules have more electrons than propanal molecules</i></p> <p><i>C is incorrect because the areas of contact between the molecules will be similar and any difference is too small to account for such a large difference in boiling temperature</i></p> <p><i>D is incorrect because neither molecule forms hydrogen bonds in the pure liquid</i></p>	(1)

Question number	Answer	Mark
7	<p><b>The only correct answer is A</b></p> <div style="text-align: center;"> </div> <p><i>B is incorrect because the O-----H—O bond angle should be 180°</i></p> <p><i>C is incorrect because the hydrogen atoms bonded to carbon atoms cannot form hydrogen bonds</i></p> <p><i>D is incorrect because the hydrogen atoms bonded to carbon atoms cannot form hydrogen bonds</i></p>	(1)

Question number	Answer	Mark
8	<p><b>The only correct answer is B</b> (chromium in <math>\text{K}_2\text{Cr}_2\text{O}_7</math> and <math>\text{K}_2\text{CrO}_4</math>)</p> <p><i>A is incorrect because chlorine is in oxidation state +7 in <math>\text{Cl}_2\text{O}_7</math> and +5 in <math>\text{Ca}(\text{ClO}_3)_2</math></i></p> <p><i>C is incorrect because manganese is in oxidation state +6 in <math>\text{K}_2\text{MnO}_4</math> and +7 in <math>\text{Mn}_2\text{O}_7</math></i></p> <p><i>D is incorrect because oxygen is in oxidation state <math>-2</math> in <math>\text{K}_2\text{O}</math> and <math>-1</math> in <math>\text{K}_2\text{O}_2</math></i></p>	(1)

Question number	Answer	Mark
9	<p><b>The only correct answer is D</b> (<math>\text{BaFeO}_4</math>)</p> <p><i>A is incorrect because <math>\text{Ba}_2\text{FeO}_3</math> is consistent with the oxidation number of Ba being +1 and of Fe being +4</i></p> <p><i>B is incorrect because <math>\text{BaFeO}_3</math> is consistent with the oxidation number of Fe being +4</i></p> <p><i>C is incorrect because <math>\text{Ba}_2\text{FeO}_4</math> is consistent with the oxidation number of Ba being +1</i></p>	(1)

Question number	Answer	Mark
10	<p><b>The only correct answer is C</b> (the sodium ion has a larger ionic radius than the lithium ion)</p> <p><i>A is incorrect because the greater reactivity of sodium does not lead to the formation of a peroxide</i></p> <p><i>B is incorrect because the lower first ionisation energy of sodium does not affect the oxide formed</i></p> <p><i>D is incorrect because the lithium ion has a higher charge density than the sodium ion</i></p>	(1)

Question number	Answer	Mark
11	<p><b>The only correct answer is D</b> (<math>0.200 \text{ mol dm}^{-3}</math>)</p> <p><i>A is incorrect because 0.025 is obtained by omitting to scale from <math>0.25 \text{ dm}^3</math> to <math>1 \text{ dm}^3</math> and halving the amount of NaOH formed</i></p> <p><i>B is incorrect because 0.050 is obtained by omitting to scale from <math>0.25 \text{ dm}^3</math> to <math>1 \text{ dm}^3</math></i></p> <p><i>C is incorrect because 0.100 is obtained by halving the amount of NaOH formed</i></p>	(1)

Question number	Answer	Mark
12	<p><b>The only correct answer is B</b> (2.97 g)</p> <p><i>A is incorrect because 1.48 g does not take into account that 2 mol <math>\text{Mg}(\text{NO}_3)_2</math> decompose</i></p> <p><i>C is incorrect because 7.42 g does not take into account that 2 mol <math>\text{Mg}(\text{NO}_3)_2</math> decompose or that 5 mol of gas form</i></p> <p><i>D is incorrect because 14.8 g does not take into account that 5 mol of gas form</i></p>	(1)

Question number	Answer	Mark
13	<p><b>The only correct answer is A</b> (electrons being excited to higher energy levels and emitting yellow light on returning to the ground state)</p> <p><i>B is incorrect because blue-violet is the complementary colour to yellow</i></p> <p><i>C is incorrect because this describes the formation of a blue-violet colour by the reflection of a beam of white light</i></p> <p><i>D is incorrect because this describes the formation of a yellow colour by the reflection of a beam of white light</i></p>	(1)

Question number	Answer	Mark
14	<p><b>The only correct answer is A (copper(II))</b></p> <p><i>B is incorrect because the presence of lead(II) slows the reaction down</i></p> <p><i>C is incorrect because sodium nitrate has no significant effect on the rate</i></p> <p><i>D is incorrect because zinc sulfate has no significant effect on the rate</i></p>	(1)

Question number	Answer	Mark
15	<p><b>The only correct answer is A(1.3 cm<sup>3</sup> min<sup>-1</sup>)</b></p> <p><i>B is incorrect because 2.5 cm<sup>3</sup> min<sup>-1</sup> is final total volume (20 cm<sup>3</sup>) divided by the total reaction time (8 min)</i></p> <p><i>C is incorrect because 4.4 cm<sup>3</sup> min<sup>-1</sup> is the average rate in reaching point X</i></p> <p><i>D is incorrect because 25 cm<sup>3</sup> min<sup>-1</sup> is the initial rate of formation of oxygen (taking a tangent at t = 0 to t = 1)</i></p>	(1)

Question number	Answer	Mark
16	<p><b>The only correct answer is C (R only)</b></p> <p><i>A is incorrect because P is a secondary halogenoalkane</i></p> <p><i>B is incorrect because Q is a primary halogenoalkane</i></p> <p><i>D is incorrect because Q is a primary halogenoalkane</i></p>	(1)



Question number	Answer	Mark
17	<p><b>The only correct answer is D</b> (water)</p> <p><i>A is incorrect because ethanol is a co-solvent</i></p> <p><i>B is incorrect because nitrate ions are not effective nucleophiles</i></p> <p><i>C is incorrect because the silver ion is positively charged and cannot be a nucleophile</i></p>	(1)

Question number	Answer	Mark
18(a)	<p><b>The only correct answer is C</b> (phosphoric(V) acid)</p> <p><i>A is incorrect because 50% sulfuric acid is not concentrated enough to effect dehydration</i></p> <p><i>B is incorrect because ethanolic potassium hydroxide cannot effect elimination of alcohols, only of halogenoalkanes</i></p> <p><i>D is incorrect because red phosphorus is used with iodine to bring about a substitution reaction</i></p>	(1)

Question number	Answer	Mark
18(b)	<p><b>The only correct answer is A</b>(elimination)</p> <p><i>B is incorrect because water is a product not a reactant</i></p> <p><i>C is incorrect because both hydrogen and oxygen are lost</i></p> <p><i>D is incorrect because no atom or group is replaced</i></p>	(1)

**TOTAL FOR SECTION A = 20 MARKS**

## Section B

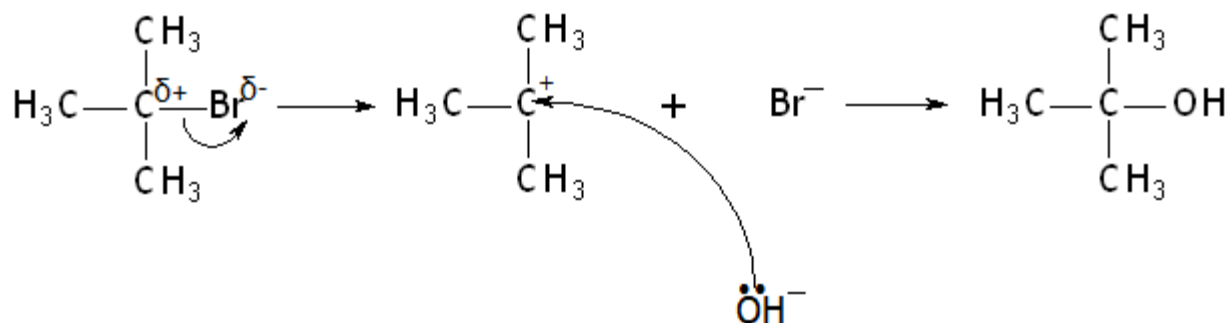
Question number	Answer	Additional guidance	Mark
19(a)	An answer that makes reference to the following <ul style="list-style-type: none"> <li>nucleophilic substitution</li> </ul>	Allow substitution nucleophilic	(1)

Question number	Answer	Additional guidance	Mark
19(b)	An answer that makes reference to the following <ul style="list-style-type: none"> <li>IUPAC name</li> <li>structure</li> </ul>	<p>(1) 2,2-dimethylpropanenitrile</p> <p>Ignore punctuation errors e.g. spaces. omission of commas or hyphen or 'e' in propane</p> <p>Allow 2,2-dimethylpropanonitrile / 2,2-dimethylpropanitrile</p> <p>(1) Accept displayed, skeletal or structural formula or any combination of these. e.g.</p> $  \begin{array}{c}  \text{CH}_3 \\    \\  \text{H}_3\text{C}-\text{C}-\text{C}\equiv\text{N} \\    \\  \text{CH}_3  \end{array}  $ <p>Accept CN for C≡N</p> <p>Ignore connectivity errors on alkyl group e.g. CH<sub>3</sub>—C</p> <p>Do not award C=N / C—N / NC</p> <p>Standalone marks. No TE</p>	(2)

Question number	Answer	Additional guidance	Mark
19(c)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>ammonia / <math>\text{NH}_3</math> (1)</li> <li>(heat) under pressure <b>and</b> in alcoholic solution (1)</li> </ul>	<p>Standalone marks</p> <p>Accept concentrated ammonia Allow concentrated aqueous ammonia Allow ammonia solution Ignore gas Do not award just aqueous ammonia Do not award dilute ammonia (solution) Allow heat <b>and</b> in a sealed tube / container (in alcohol) Accept alcohol or ethanolic or ethanol for alcoholic Do not award heat under reflux</p>	(2)

Question number	Answer	Additional guidance	Mark
19(d)(i)	<p>An answer that shows the following</p> <ul style="list-style-type: none"> <li>dipole on C—Br</li> <li>curly arrow from C—Br bond to Br atom or just beyond</li> <li>lone pair on oxygen of OH<sup>-</sup></li> <li>curly arrow from oxygen to positively charged carbon atom</li> </ul>	<p>All four points correct scores (2)</p> <p>Any two or three points correct scores (1)</p> <p>Do not award second curly arrow from negative charge</p> <p>Penalise use of half-arrows once only</p> <p>Ignore an extra OH<sup>-</sup> attacking the 2-bromo-2-methylpropane molecule.</p> <p>Ignore lone pair on this extra species</p> <p>Ignore extra curly arrows and lone pairs</p>	(2)

#### Example of mechanism

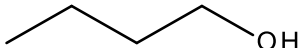
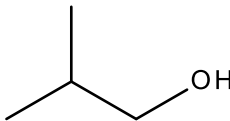
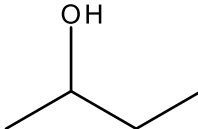


Question number	Answer	Additional guidance	Mark
19(d)(ii)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>heterolytic (bond) fission occurs (1)</li> <li>the (pair of) electrons in the C—Br bond are transferred to the bromine atom (1)</li> </ul>	<p>Standalone marks</p> <p>Allow heterolytic (bond) breaking</p> <p>Allow heterolysis</p> <p>Do not award heterolytic fission of any other bond.</p> <p>Ignore heterolytic reaction</p> <p>Allow ‘the <b>bonding</b> electron <b>pair</b> is transferred to the bromine (atom)’</p> <p>Allow ‘C—Br bond is broken <b>and</b> both electrons are transferred to the bromine (atom)’</p> <p>There must be an indication of <b>two</b> electrons (pair / both)</p>	(2)

Question number	Answer	Additional guidance	Mark
19(d)(iii)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>(shape is) trigonal planar (1)</li> <li>three bond pairs (and no lone pairs) <b>and</b> arranged to minimise repulsion (1)</li> </ul>	<p>Allow triangular planar</p> <p>Ignore just ‘trigonal’</p> <p>Ignore bond angles</p> <p>Allow regions of electrons</p> <p>Do not award lone pairs / atoms for bond pairs</p> <p>Allow for maximum separation</p> <p>No TE on incorrect numbers of valence electron pairs</p>	(2)

Question number	Answer	Additional guidance	Mark
19(d)(iv)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>two structures of isomeric alcohols <math>C_4H_{10}O</math> (1)</li> <li>third structure of isomeric alcohol <math>C_4H_{10}O</math> (1)</li> </ul>	<p>All three correct scores (2)</p> <p>Any two correct scores (1)</p> <p>Allow displayed, structural or skeletal formulae or any combination of these</p> <p>Allow <math>C_2H_5</math> for <math>CH_3CH_2</math></p> <p>M2 depends on M1 being scored</p> <p>Penalise connectivity errors only with horizontal OH and once only</p> <p>Penalise omission of H atoms in displayed structures once only</p> <p>Using <math>C_3H_7OH</math> scores zero</p>	(2)

#### Examples of structures

		
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Question number	Answer	Additional guidance	Mark
19(e)(i)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>all species correct (1)</li> <li>equation balanced (1)</li> </ul>	$\text{Ag}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{Ag}^+ + 2\text{OH}^-$ <p>Non-ionic equation scores (1)</p> $\text{Ag}_2\text{O} + \text{H}_2\text{O} \rightarrow 2\text{AgOH}$ <p>Allow reversible arrow</p> <p>Ignore state symbols even if incorrect</p>	(2)

Question number	Answer	Additional guidance	Mark
19(e)(ii)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>(because silver(I) oxide is insoluble) it can easily be separated from the reaction mixture (when reaction is complete)</li> </ul>	<p>Allow (because silver(I) oxide is insoluble) it can be reused</p> <p>Allow NaOH / KOH are caustic / corrosive (due to high concentration of hydroxide ions)</p> <p>Allow Ag<sub>2</sub>O is less corrosive (than NaOH)</p> <p>Ignore reference to reactivity of NaOH and KOH</p> <p>Ignore reference to reactivity of Na and K</p> <p>Ignore reference to other silver compounds</p>	(1)

(Total for Question 19 = 16 marks)



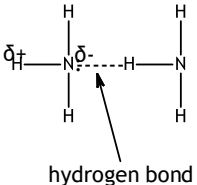
Question number	Answer	Additional guidance	Mark
20(a)(i)	<ul style="list-style-type: none"> <li>identification and number of bonds broken <b>and</b> the values needed</li> <li>evaluation of energy required</li> <li>identification of bonds formed <b>and</b> the values needed <b>and</b> evaluation of energy produced</li> <li>evaluation of enthalpy change of combustion</li> </ul>	<p>Here and throughout the paper</p> <p>Do not penalise correct premature rounding</p> <p>Penalise incorrect rounding only in their final answer</p> <p>Example of calculation</p> <p>Bonds broken:</p> $7 \times \text{C—C} + 18 \times \text{C—H} + 12.5 \times \text{O=O}$ <p>(1) <math>E</math> (bond breaking) = <math>7 \times 347 + 18 \times 413 + 12.5 \times 498</math></p> <p>(1) <math>= (+)16088 \text{ (kJ mol}^{-1}\text{)}</math></p> <p>TE only if at least 2 bonds used</p> <p>Bonds formed:</p> $E \text{ (bond forming)} = 16 \times \text{C=O} + 18 \times \text{O—H}$ $= 16 \times 805 + 18 \times 464$ <p>(1) <math>= (-)21232 \text{ (kJ mol}^{-1}\text{)}</math></p> <p>(1) <math>\Delta_c H = 16088 - 21232 = -5144 \text{ (kJ mol}^{-1}\text{)}</math></p> <p>TE at each stage (even if final value is positive)</p> <p>Ignore SF except 1 SF</p> <p>Units are not required but if given must be correct for the <b>final</b> value</p> <p>Correct answer with some working scores (4)</p>	(4)

Question number	Answer	Additional guidance	Mark
20(a)(ii)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>the bond enthalpies are averaged over a (large) number of compounds (1)</li> <li>bond enthalpies always refer to substances in the gas phase <b>and</b> octane and water are liquids when the value of <math>\Delta_c H^\ominus</math> is obtained (1)</li> </ul>	<p>Allow just bond enthalpies are average values</p> <p>Allow just octane is a liquid or water is a liquid</p> <p>Allow calculations using mean bond enthalpies do not include changes of state</p> <p>Ignore non-standard conditions</p> <p>Do not award explanations for experimental error such as heat loss, or incomplete combustion</p>	(2)

Question number	Answer	Additional guidance	Mark
20(a)(iii)	<ul style="list-style-type: none"> <li>calculation of molar mass of octane</li> <li>conversion of <math>\text{kJ mol}^{-1}</math> to <math>\text{MJ kg}^{-1}</math> <b>and</b> calculation of percentage efficiency</li> </ul>	<p>Example of calculation</p> <p>(1) <math>M = 8 \times 12 + 18 = 114 \text{ g mol}^{-1}</math></p> <p>Allow mol octane = <math>1000 \div (8 \times 12 + 18 \times 1) = 8.772</math></p> <p>Ignore just <math>(8 \times 12 + 18 \times 1)</math></p> <p><math>1000 \times 5470 \div 114 = 47982 \text{ kJ kg}^{-1} = 47.982 \text{ MJ kg}^{-1}</math></p> <p>(1) Efficiency = <math>100 \times 11 \div 47.982 = 22.925 / 22.9 / 23\%</math></p> <p>Allow conversion of <math>\text{MJ kg}^{-1}</math> to <math>\text{kJ mol}^{-1}</math></p> <p><math>= 114 \times 11 \times 1000 \div 1000 = 1254 \text{ (kJ mol}^{-1}\text{)}</math></p> <p>and</p> <p>Efficiency = <math>100 \times 1254 \div 5470 = 22.925 / 22.9 / 23\%</math></p> <p>Ignore SF except 1 SF</p> <p>TE unless % efficiency &gt; 100</p> <p>Correct answer with no working scores (2)</p> <p>Allow calculation using <math>\Delta_c H^\ominus</math> from mean bond enthalpy data (<math>-5144 \text{ kJ mol}^{-1}</math>):</p> <p>Efficiency = 24.378%</p> <p>Allow calculation using stated incorrect <math>\Delta_c H^\ominus</math> from mean bond enthalpy data unless % efficiency &gt; 100</p>	(2)

Question number	Answer	Additional guidance	Mark
20(a)(iv)	<p>An answer that makes reference to <b>two</b> of the following</p> <ul style="list-style-type: none"> <li>heat loss to the surroundings (1)</li> <li>energy is used to bring the engine to operating temperature (1)</li> <li>incomplete combustion (of the fuel) (1)</li> </ul>	<p>Accept specific examples such as</p> <p>Heat loss due to friction in the engine</p> <p>Heat loss via hot exhaust</p> <p>Allow just ‘converted to heat’</p> <p>Allow energy loss to the surroundings</p> <p>Ignore just ‘friction’</p> <p>Allow energy is used to warm up the engine</p> <p>Allow energy is used to start the engine</p> <p>Allow energy is used for aircon / electronic devices</p> <p>Allow combustion is not smooth</p> <p>Ignore inefficient combustion</p> <p>Ignore references to standard conditions</p> <p>Ignore fuel evaporates</p> <p>Ignore petrol not 100% octane</p> <p>Ignore the idea that some other force is moving the car e.g. car is going downhill</p>	(2)

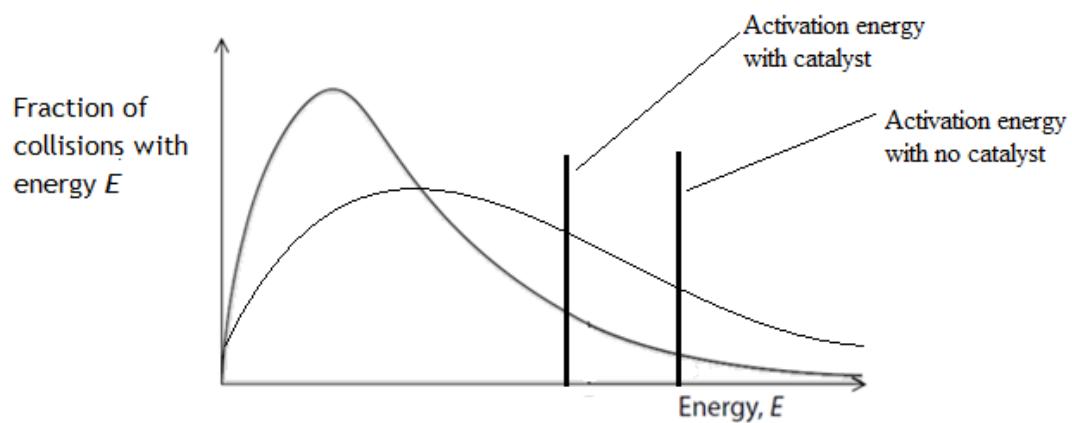
Question number	Answer	Additional guidance	Mark																				
*20(b)	<p>This question assesses the student’s ability to show a coherent and logically structured answer with linkages and fully sustained reasoning.</p> <p>Marks are awarded for indicative content and for how the answer is structured and shows lines of reasoning.</p> <p>The following table shows how the marks should be awarded for indicative content.</p> <table><tr><th>Number of indicative marking points seen in answer</th><th>Number of marks awarded for indicative marking points</th></tr><tr><td>6</td><td>4</td></tr><tr><td>5-4</td><td>3</td></tr><tr><td>3-2</td><td>2</td></tr><tr><td>1</td><td>1</td></tr><tr><td>0</td><td>0</td></tr></table> <p>The following table shows how the marks should be awarded for structure and lines of reasoning</p> <table><tr><th></th><th>Number of marks awarded for structure of answer and sustained lines of reasoning</th></tr><tr><td>Answer shows a coherent logical structure with linkages and fully sustained lines of reasoning demonstrated throughout</td><td>2</td></tr><tr><td>Answer is partially structured with some linkages and lines of reasoning</td><td>1</td></tr><tr><td>Answer has no linkages between points and is unstructured</td><td>0</td></tr></table>	Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points	6	4	5-4	3	3-2	2	1	1	0	0		Number of marks awarded for structure of answer and sustained lines of reasoning	Answer shows a coherent logical structure with linkages and fully sustained lines of reasoning demonstrated throughout	2	Answer is partially structured with some linkages and lines of reasoning	1	Answer has no linkages between points and is unstructured	0	<p>Guidance on how the mark scheme should be applied.</p> <p>The mark for indicative content should be added to the mark for lines of reasoning. For example, a response with five indicative marking points that is partially structured with some linkages and lines of reasoning scores 4 marks (3 marks for indicative content and 1 mark for partial structure and some linkages and lines of reasoning).</p> <p>If there were no linkages between the points, then the same indicative marking points would yield an overall score of 3 marks (3 marks for indicative content and no marks for linkages).</p> <p>In general it would be expected that 5 or 6 indicative points would get <b>2</b> reasoning marks 3 or 4 indicative points would get <b>1</b> reasoning mark 0, 1 or 2 indicative points would get <b>0</b> reasoning marks.</p> <p>If there is any incorrect chemistry, deduct mark(s) from the reasoning. If no reasoning mark(s) awarded do not deduct mark(s).</p> <p><b>Comment:</b> Look for the indicative marking points first, then consider the mark for the structure of the answer and sustained line of reasoning</p>	6
Number of indicative marking points seen in answer	Number of marks awarded for indicative marking points																						
6	4																						
5-4	3																						
3-2	2																						
1	1																						
0	0																						
	Number of marks awarded for structure of answer and sustained lines of reasoning																						
Answer shows a coherent logical structure with linkages and fully sustained lines of reasoning demonstrated throughout	2																						
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Answer has no linkages between points and is unstructured	0																						

Question	Answer	Additional guidance	Mark
*20(b) cont	<p>Indicative points</p> <p>Similarities</p> <ul style="list-style-type: none"> <li><b>IP1</b> both hydrogen and ammonia form London / dispersion forces</li> <li><b>IP2</b> a temporary dipole forms in a molecule and induces a dipole in an adjacent molecule</li> <li><b>IP3</b> the attraction (between the temporary dipoles) is small(er) in hydrogen because the H<sub>2</sub> electron cloud is not easily polarised</li> </ul> <p>Differences</p> <ul style="list-style-type: none"> <li><b>IP4</b> ammonia forms hydrogen bonds (because nitrogen is very electronegative)</li> </ul>  <ul style="list-style-type: none"> <li><b>IP5</b> Accept hydrogen bond forms between the nitrogen lone pair and the (δ+) hydrogen (of a different molecule)</li> <li><b>IP6</b> ammonia liquefies more easily than hydrogen because hydrogen bonds are stronger than London forces</li> </ul>	<p>Allow van der Waals forces</p> <p>Ignore permanent dipole-dipole forces in ammonia</p> <p>Do not award H<sub>2</sub> forms hydrogen bonds / permanent dipole-dipole forces</p> <p>Accept fluctuating electron clouds result in differences in electron density within the molecule</p> <p>Allow instantaneous dipole-dipole attractions between molecules</p> <p>Allow because H<sub>2</sub> has only two / few electrons</p> <p>Allow attraction is greater in ammonia because it has more electrons</p> <p>Do not award ammonia forms hydrogen bonds with water</p> <p>IP4 and IP5 may be scored by a diagram showing dipole and lone pair and with H bond labelled</p> <p>Allow N---H—N bond angle not equal to 180</p> <p>Allow hydrogen bond forms between the δ- nitrogen and the δ+ hydrogen (of a different molecule)</p> <p>Allow permanent dipole-dipole forces for H bonds here</p> <p>Ignore just 'H-bonding is the strongest IMF'</p> <p>Ignore reference to boiling temperatures</p> <p>Do not award hydrogen liquefies to form water</p> <p>Do not award energy is required to liquefy a gas</p>	6

(Total for Question 20 = 16 marks)

Question number	Answer	Additional guidance	Mark
21(a)(i)	<p>A diagram showing the following</p> <ul style="list-style-type: none"> <li>new line starting at the origin</li> </ul> <p><b>and</b></p> <p>lower maximum</p> <p><b>and</b></p> <p>peak moved to the right</p> <p><b>and</b></p> <p>asymptotic line above the original</p>	<p>Do not award if any part of the curve to the right of the peak is below the 300°C line</p> <p>Do not award if asymptote finishes higher than 25% the height of the new peak</p>	<p>(1)</p> <p><b>clip with (a)(ii)</b></p>

### Example of Maxwell-Boltzmann distribution



Question number	Answer	Additional guidance	Mark
21(a)(ii)	<p>An answer that makes reference to the following</p> <p>EITHER</p> <ul style="list-style-type: none"> <li>Increasing the temperature increases the (average kinetic) energy of collisions <b>and</b> so more collisions have energy greater than the activation energy (1)</li> <li>adding a catalyst lowers the activation energy <b>and</b> so more collisions have energy greater than the activation energy (1)</li> </ul> <p>OR</p> <ul style="list-style-type: none"> <li>(rate increases when) more collisions have an energy greater than (or equal to) the activation energy (1)</li> <li>the area under the curve beyond <math>E_a</math> is increased when the temperature is increased <b>and</b> when <math>E_a</math> becomes <math>E_{cat}</math> (1)</li> </ul>	<p>Allow particles / molecules / atoms for collisions throughout</p> <p>Allow just ‘increasing the temperature increases the number of collisions with energy greater than <math>E_a</math>’</p> <p>Standalone mark</p> <p>Allow particles / molecules / atoms for collisions</p> <p>In no other mark is scored increasing the temperature increases the (average kinetic) energy of collisions <b>and</b> adding a catalyst lowers the activation energy scores (1)</p> <p>M2 may be scored by clear labelling of the relevant areas under the curves</p>	<p>(2)</p> <p><b>clip with (a)(i)</b></p>



Question number	Answer	Additional guidance	Mark
21(b)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>• Advantage: hydrogen is a useful by-product (1)</li> <li>• Disadvantage: higher temperature means higher energy costs (1)</li> </ul>	<p>Allow higher atom economy  Ignore higher temperature gives higher rate  Ignore oxygen not needed</p> <p>Allow higher temperature is more expensive  Allow the energy required is more expensive  Allow higher temperature needs more expensive equipment  Allow more energy is required  Ignore just 'more expensive'  Ignore reference to yield  Do not award the hazards of hydrogen</p>	(2)

Question number	Answer	Additional guidance	Mark
21(c)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>(methanol) O—H stretch wavenumber range 3750-3200 (<math>\text{cm}^{-1}</math>)(1)</li> <li>(methanal) C=O stretch wavenumber range 1740-1720 (<math>\text{cm}^{-1}</math>)</li> </ul> <p>or</p> <p>(methanal) C—H stretch wavenumber range 2900-2820 (<math>\text{cm}^{-1}</math>)</p> <p>or</p> <p>(methanal) C—H stretch wavenumber range 2775-2700 (<math>\text{cm}^{-1}</math>) (1)</p>	<p>Allow OH or CO</p> <p>Ignore ROH / CH<sub>3</sub>OH / HCHO</p> <p>Both wavenumber ranges without bonds scores (1)</p> <p>Specific wavenumbers within the ranges or wavenumber ranges within the Data Booklet ranges scores max (1)</p> <p>If more than one wavenumber is given for methanal all must be correct</p> <p>Ignore alkyl C—H stretch wavenumber ranges</p>	(2)

(Total for Question 21 = 7 marks)

**TOTAL FOR SECTION B = 39 MARKS**

### Section C

Question number	Answer	Additional guidance	Mark
22(a)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>(safety) glasses / goggles / spectacles</li> </ul>	<p>Allow gloves</p> <p>Ignore laboratory coat / apron</p> <p>Ignore mask</p> <p>Ignore open window</p> <p>Ignore all explanations for the selected safety precaution</p> <p>Do not award use in a cupboard / fume cupboard</p>	(1)

Question number	Answer	Additional guidance	Mark
22(b)(i)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>ionic equation for acid-carbonate reaction</li> </ul>	<p>Example of equation</p> $2\text{H}^+ + \text{CaCO}_3 \rightarrow \text{Ca}^{2+} + \text{H}_2\text{O} + \text{CO}_2$ <p>Allow <math>2\text{H}^+ + \text{CO}_3^{2-} \rightarrow \text{H}_2\text{O} + \text{CO}_2</math></p> <p>Allow multiples</p> <p>Ignore state symbols even if incorrect</p> <p>Do not award inclusion of spectator ions</p>	(1)

Question number	Answer	Additional guidance	Mark
22(b)(ii)	<ul style="list-style-type: none"> <li>• calculation of mass of HCl in 50 cm<sup>3</sup></li> <li>• calculation of <math>M_r</math> of HCl</li> </ul> <p><b>and</b></p> <ul style="list-style-type: none"> <li>• calculation of moles of HCl in 50 cm<sup>3</sup></li> <li>• calculation of moles of CaCO<sub>3</sub></li> <li>• calculation of <math>M_r</math> of CaCO<sub>3</sub></li> </ul> <p><b>and</b></p> <ul style="list-style-type: none"> <li>• calculation of mass of CaCO<sub>3</sub></li> </ul> <div style="border: 1px solid black; padding: 5px; margin-top: 10px;"> <p>If moles of HCl in 750 cm<sup>3</sup> are calculated first  <math>= 85 \div 36.5 = 2.3288</math> (mol). This scores M2</p> </div>	<p>Example of calculation</p> <p>mass of HCl = <math>85.0 \times 50 \div 750 = 5.6667</math> (g)  HCl <math>M_r = 36.5</math></p> <p>mol HCl = <math>5.6667 \div 36.5 = 0.15525</math> (mol)</p> <p>mol CaCO<sub>3</sub> = <math>0.15525 \div 2</math>  <math>= 0.077626 / 7.7626 \times 10^{-2}</math> (mol)  CaCO<sub>3</sub> <math>M_r = 100.1</math></p> <p><math>0.077626 \times 100.1 = 7.7703 / 7.77 / 7.8</math> (g)  TE at each stage  Allow CaCO<sub>3</sub> <math>Mr = 100</math> giving 7.7626 (g)  Correct answer with some working scores (4)  Ignore SF except 1 SF  If all the HCl is used, mass CaCO<sub>3</sub> = 116.55 g  scores (3)  If the reacting ratio is used the wrong way round the  mass CaCO<sub>3</sub> = 31.08 g</p>	(4)

Question number	Answer	Additional guidance	Mark
22(c)(i)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>disproportionation (reaction) (1)</li> <li>chlorine oxidised from zero (in <math>\text{Cl}_2</math>) to +1 in <math>\text{HClO} / \text{ClO}^-</math> (1)</li> <li>chlorine reduced from zero (in <math>\text{Cl}_2</math>) to -1 in <math>\text{HCl} / \text{Cl}^-</math> (1)</li> </ul>	<p>Example of equation</p> <p>standalone mark</p> <p>Ignore 'redox'</p> <p>Ignore just 'chlorine both oxidised and reduced'</p> <p>For M2 and M3 allow oxidation numbers of the species shown in the equation</p> <p>If M2 and M3 not scored just 'chlorine oxidised from zero (in <math>\text{Cl}_2</math>) to +1 and reduced from zero (in <math>\text{Cl}_2</math>) to -1' scores (1)</p> <p>If M2 and M3 not scored the three oxidation numbers of chlorine scores (1). This may be shown on the equation</p> <p>Do not award M2 / M3 if changes of oxidation numbers of H and O given</p>	(3)

Question number	Answer	Additional guidance	Mark
22(c)(ii)	<ul style="list-style-type: none"> <li>• calculation of moles of oxygen (1)</li> <li>• deduction of moles of NaOCl(1)</li> <li>• calculation of <math>M_r(\text{NaOCl})</math> &amp; mass of NaOCl in <math>5 \text{ cm}^3</math>(1)</li> <li>• calculation of mass of NaOCl in <math>1 \text{ dm}^3</math> <b>and</b> answer given to 2 SF</li> </ul> <p>(1)</p>	<p>Example of calculation</p> $\text{mol O}_2 = 113 \div 24000$ $= 4.7083 \times 10^{-3} / 0.0047083 \text{ (mol)}$ $\text{mol NaOCl} = \text{mol O}_2 = 4.7083 \times 10^{-3} / 0.0047083 \text{ (mol)}$ $M_r(\text{NaOCl}) = 74.5$ $\text{mass of NaOCl in } 5 \text{ cm}^3 = 74.5 \times 4.7083 \times 10^{-3}$ $= 0.35077 \text{ (g)}$ $0.35077 \times 200 = 70.154 = 70 \text{ (g dm}^{-3}\text{)}$ <p>TE at each stage</p> <p>Correct answer with some working scores (4)</p> <p>The conversion from moles to mass omitted and the answer given to 2 SF as <math>0.94 \text{ (mol dm}^{-3}\text{)}</math> scores (3)</p>	(4)

Question number	Answer	Additional guidance	Mark
22(c)(iii)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>propane-1,2,3-triol forms <b>three</b> hydrogen bonds per molecule (1)</li> <li>so the molecules are strongly attracted to each other (and to surfaces ) (1)</li> </ul>	<p>Standalone marks</p> <p>Allow propane-1,2,3-triol has three OH groups <b>and</b> forms hydrogen bonds</p> <p>Allow 'propane-1,2,3-triol forms three hydrogen bonds'</p> <p>Do not award propane-1,2,3-triol has / forms OH<sup>-</sup> ions</p> <p>Allow so the intermolecular forces are (very) strong</p> <p>Allow so molecules do not move past each other (easily)</p> <p>Allow the intermolecular forces need a lot of energy to overcome</p> <p>Ignore just 'hydrogen bonding is the strongest intermolecular force'</p>	(2)

Question number	Answer	Additional guidance	Mark
22(d)	<p>An answer that makes reference to the following</p> <ul style="list-style-type: none"> <li>toxic chlorine would be produced (1)</li> <li>(because) the chlorine disinfectant equilibrium would move to the left (because the concentration of H<sup>+</sup> would increase (1)</li> </ul>	<p>Ignore just 'forms chlorine'</p> <p>Allow <math>\text{NaClO} + 2\text{HCl} \rightarrow \text{NaCl} + \text{H}_2\text{O} + \text{Cl}_2</math></p> <p>Allow <math>\text{HClO} + \text{HCl} \rightarrow \text{H}_2\text{O} + \text{Cl}_2</math></p> <p>Do not award <math>\text{NaClO} + \text{HCl} \rightarrow \text{NaOH} + \text{Cl}_2</math></p>	(2)

Question number	Answer	Additional guidance	Mark																									
22(e)	<div><div><ul style="list-style-type: none"><li>calculation of mol of each element</li><li>calculation of ratio</li><li><b>and</b></li><li>calculation of simplest whole number ratio</li><li>deduction of moles of water of crystallisation</li><li>deduction of formula</li></ul></div><div><div>(1)</div><div>(1)</div><div>(1)</div><div>(1)</div></div></div>	<div>Example of calculation</div> <table><tr><td></td><td>H</td><td>B</td><td>O</td><td>Na</td></tr><tr><td>%</td><td>5.2</td><td>11.3</td><td>71.4</td><td>12.1</td></tr><tr><td>mol</td><td>5.2/1 =5.2</td><td>11.3/10.8 =1.05</td><td>71.4/16 = 4.46</td><td>12.1/23 = 0.526</td></tr><tr><td>÷ 0.526</td><td>9.89</td><td>2.00</td><td>8.48</td><td>1</td></tr><tr><td>ratio</td><td>20</td><td>4</td><td>17</td><td>2</td></tr></table> <div>TE for only 3 elements</div> <div>Hydrogen is only in water so 10 H<sub>2</sub>O</div> <div>Na<sub>2</sub>B<sub>4</sub>O<sub>7</sub>.10H<sub>2</sub>O</div> <div>TE at each stage for formulae with H<sub>2</sub>O</div> <div>Correct answer with some working scores (4)</div> <div>Ignore SF except 1 SF</div> <div>Incorrect rounding leads to formulae such as NaB<sub>2</sub>H<sub>10</sub>O<sub>8</sub> which scores (1)</div> <div>Processing this to NaB<sub>2</sub>O<sub>3</sub>.5H<sub>2</sub>O scores (3)</div>		H	B	O	Na	%	5.2	11.3	71.4	12.1	mol	5.2/1 =5.2	11.3/10.8 =1.05	71.4/16 = 4.46	12.1/23 = 0.526	÷ 0.526	9.89	2.00	8.48	1	ratio	20	4	17	2	(4)
	H	B	O	Na																								
%	5.2	11.3	71.4	12.1																								
mol	5.2/1 =5.2	11.3/10.8 =1.05	71.4/16 = 4.46	12.1/23 = 0.526																								
÷ 0.526	9.89	2.00	8.48	1																								
ratio	20	4	17	2																								

(Total for Question 22 = 21 marks)

**TOTAL FOR SECTION C = 21 MARKS**

**TOTAL FOR PAPER = 80 MARKS**