



Mark Scheme (Results)

October 2021

Pearson Edexcel International Advanced Level
In Chemistry (WCH16)
Paper 01: Practical Skills in Chemistry II

Question Number	Answer	Additional Guidance	Mark
1(a)(i)	<p>A description that makes reference to the following point:</p> <ul style="list-style-type: none"> blue precipitate (forms) 	<p>Accept light / pale blue Allow ppt / ppte / solid for precipitate</p> <p>Ignore formulae even if incorrect</p> <p>Do not award mixed colours e.g. blue / green Do not award dark blue / royal blue / navy blue</p>	(1)

Question Number	Answer	Additional Guidance	Mark
1(a)(ii)	<ul style="list-style-type: none"> (the initial blue solution goes) green (1) (this changes to a) yellow (solution) (1) 	<p>Allow any shades of colours e.g. bright Ignore formulae even if incorrect Penalise green precipitate or yellow precipitate once only</p> <p>Ignore mention of blue precipitate</p> <p>If no other mark is awarded, allow (1) for green-yellow / yellow-green (solution)</p>	(2)

Question Number	Answer	Additional Guidance	Mark
1(b)	<p>A description that makes reference to the following points:</p> <ul style="list-style-type: none"> addition of (dilute) hydrochloric acid / HCl(aq) and barium chloride (solution) / $\text{BaCl}_2\text{(aq)}$ white precipitate (forms) 	<p>Allow names or formulae of reagents but if both are given, both must be correct</p> <p>Allow HCl / acidified / H^+ / dilute nitric acid / $\text{HNO}_3\text{(aq)}$ for hydrochloric acid Allow barium nitrate solution / $\text{Ba(NO}_3)_2\text{(aq)}$ for barium chloride (solution) Ignore concentration of acid Do not award sulfuric acid</p> <p>Conditional on use of barium chloride or barium nitrate with or without any acid Allow ppt / ppte / solid for precipitate Ignore cloudy Ignore incorrect name / formula of precipitate Do not award just 'turns white'</p>	(2)

Question Number	Answer	Additional Guidance	Mark
1(c)(i)	<ul style="list-style-type: none"> calculation of E^\ominus_{cell} 	<p>Example of calculation: $0.77 - 0.34$ $= (+)0.43 \text{ (V)}$</p> <p>Correct answer with no working scores (1)</p> <p>Do not award -0.43 (V)</p>	(1)

Question Number	Answer	Additional Guidance	Mark
1(c)(ii)	<ul style="list-style-type: none"> low voltage supply and replace with (high resistance) voltmeter (1) (platinum) wire and replace with salt bridge (1) iron (electrode) and replace with platinum (1) 	<p>Penalise additional incorrect changes The mistakes can be in any order</p> <p>Allow potentiometer / Wheatstone bridge Do not award voltmeter</p> <p>Allow a description of a salt bridge containing potassium / sodium / ammonium with nitrate / chloride</p>	(3)

Question Number	Answer	Additional Guidance	Mark
1(d)(i)	<ul style="list-style-type: none"> balanced equation 	<p>Example of equation: $\text{Zn} + 4\text{HNO}_3 \rightarrow \text{Zn}(\text{NO}_3)_2 + 2\text{NO}_2 + 2\text{H}_2\text{O}$ Allow $\text{Zn} + 4\text{H}^+ + 2\text{NO}_3^- \rightarrow \text{Zn}^{2+} + 2\text{NO}_2 + 2\text{H}_2\text{O}$ </p> <p>Allow multiples</p> <p>Ignore state symbols, even if incorrect</p> <p>Do not award equation with copper</p>	(1)

Question Number	Answer	Additional Guidance	Mark
1(d)(ii)	<ul style="list-style-type: none"> (10 cm³) measuring cylinder 	Allow 25 cm ³ measuring cylinder but no bigger size specified Allow measurement on the side of a beaker Do not award burette / pipette / volumetric flask	(1)

Question Number	Answer	Additional Guidance	Mark
1(d)(iii)	<ul style="list-style-type: none"> (when the solution is) straw coloured / pale yellow 	Allow near / approaching / just before the end point Ignore at the end point / before the end point Do not award just yellow / pale brown	(1)

Question Number	Answer	Additional Guidance	Mark
1(d)(iv)	<ul style="list-style-type: none"> calculation of mol $\text{S}_2\text{O}_3^{2-}$ (1) calculation of mol of Cu^{2+} in 25.0 cm^3 (1) calculation of mol Cu^{2+} in 250 cm^3 (1) calculation of mass of Cu (1) calculation of percentage of copper in brass and answer to 2 / 3 SF (1) 	<p>Example of calculation: Mol $\text{S}_2\text{O}_3^{2-}$ used = $\frac{28.60 \times 0.100}{1000}$ = $0.00286 / 2.86 \times 10^{-3}$</p> <p>(Mol I_2 formed = 0.00143) Mol of Cu^{2+} in 25.0 cm^3 = $0.00286 / 2.86 \times 10^{-3}$ TE on mol $\text{S}_2\text{O}_3^{2-}$</p> <p>Mol Cu^{2+} in 250 cm^3 = 0.00286×10 = $0.0286 / 2.86 \times 10^{-2}$ TE on mol Cu^{2+} in 25.0 cm^3</p> <p>Mass of Cu = 0.0286×63.5 = 1.8161 (g) TE on mol Cu^{2+} in 250 cm^3</p> <p>Percentage of copper = $\frac{1.8161}{3.90} \times 100 = 46.567$ = 46.6 / 47 (%) TE on mass Cu unless percentage >100%</p> <p>Allow answer to 2 / 3 SF from earlier correct rounding 1.82 g gives 46.7 / 47(%) 1.8 g gives 46.2 / 46(%) Correct answer with no or some working scores (5)</p>	(5)

(Total for Question 1 = 17 marks)

Question Number	Answer	Additional Guidance	Mark
2(a)(i)	<ul style="list-style-type: none"> • Test 1: aldehyde or ketone / -CHO or C=O (1) • Test 2: aldehyde / -CHO (1) 	Allow structural / displayed / skeletal formulae Ignore COH / C=O for aldehyde in M1 and M2 Both needed for the mark Allow carbonyl (compound) Do not award methyl ketone / specific aldehydes and ketones for M1 only	(2)

Question Number	Answer	Additional Guidance	Mark
2(a)(ii)	<ul style="list-style-type: none"> • copper(I) oxide / Cu_2O 	If name and formula given, both must be correct Ignore copper oxide / Cu^+	(1)

Question Number	Answer	Additional Guidance	Mark
2(b)(i)	<ul style="list-style-type: none"> • C_2H_5^+ / CH_3CH_2^+ or CHO^+ 	Penalise additional incorrect formulae Accept brackets around the formulae Allow charge anywhere on the ion Allow symbols in any order e.g. H_5C_2^+ / COH^+ Do not award bond from formula e.g. $\text{-C}_2\text{H}_5^+$ Symbols and the charge are needed	(1)

Question Number	Answer	Additional Guidance	Mark
2(b)(ii)	<ul style="list-style-type: none"> m/z value = 58 and structure of propanal 	Example of structure: $\text{CH}_3\text{CH}_2\text{CHO}$ Allow any combination of structural and displayed formula / skeletal formula Ignore $\text{CH}_3\text{CH}_2\text{COH}$	(1)

Question Number	Answer	Additional Guidance	Mark
2(c)(i)	<ul style="list-style-type: none"> Test 3: blue (1) Test 4: amine (1) 	Allow blue-green / dark green / purple Ignore indigo / violet / mauve Allow amino Ignore classification of amine Do not award ammine / amide	(2)

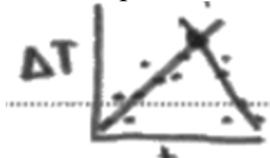
Question Number	Answer	Additional Guidance	Mark
2(c)(ii)	<ul style="list-style-type: none"> structure of B 	Example of structure: $\text{CH}_3\text{CH}_2\text{CH}_2\text{NH}_2$ / $\text{CH}_3\text{CH}_2\text{NHCH}_3$ / $(\text{CH}_3)_2\text{CHNH}_2$ / $(\text{CH}_3)_3\text{N}$ Allow any combination of structural and displayed formula / skeletal formula	(1)

(Total for Question 2 = 8 marks)

Question Number	Answer	Additional Guidance	Mark
3(a)	<ul style="list-style-type: none"> polystyrene / it is a better / good insulator or reduces / minimises heat loss (to the surroundings) or cup has a low heat capacity 	Allow description of insulation Allow glass is a poor insulator Ignore reference to polystyrene does not break Ignore prevents / no heat loss Do not award low specific heat capacity	(1)

Question Number	Answer	Additional Guidance	Mark
3(b)	<ul style="list-style-type: none"> calculation of heat produced (1) calculation of moles of LiCl (1) calculation of enthalpy change and sign and units (1) 	<p>Example of calculation: Heat produced = $25.0 \times 4.18 \times 12.5$ $= 1306.25 \text{ (J) / } 1.30625 \text{ (kJ)}$ Ignore sign</p> <p>Moles of LiCl = $\frac{2.12}{6.9 + 35.5} = 0.0500 / 5.00 \times 10^{-2} \text{ (mol)}$ Allow 0.05 / 0.04988 (from 7 for Li)</p> <p>Enthalpy change = $-\frac{1306.25}{0.0500} = -26125 \text{ J mol}^{-1}$ Or $-\frac{1.30625}{0.0500} = -26.125 \text{ kJ mol}^{-1}$ TE on heat produced and moles LiCl Ignore SF except 1 SF</p> <p>Allow answer from earlier correct rounding to at least 2 SF e.g. $-26.2 \text{ kJ mol}^{-1}$ from 1.31 kJ Correct answer with sign and units and no working scores (3)</p>	(3)

Question Number	Answer	Additional Guidance	Mark
3(c)	<ul style="list-style-type: none"> calculation of percentage uncertainty 	<p>Example of calculation:</p> $\frac{2 \times 0.25}{12.5} \times 100 = (\pm)4 \%$ <p>Correct answer with no working scores (1)</p>	(1)

Question Number	Answer	Additional Guidance	Mark
3(d)	<p>A description that makes reference to the following points:</p> <ul style="list-style-type: none"> • (start a stop watch / clock and) measure the temperature of the water every 30 s for 2½ minutes • add the lithium chloride at exactly 3 min • (stir and) record the temperature every 30 s for another 5 minutes • plot a graph of temperature against time • (join the two sets of points with 2 best fit straight lines and) extrapolate the lines to the time of mixing and determine the maximum temperature change / rise at that time 	<p>Allow any reasonable specified times / time intervals in M1, M2 and M3 Ignore any other changes to the apparatus</p> <p>(1) Allow idea of more than one reading to stabilise temperature Allow use of a lid / additional insulation</p> <p>(1) Allow start stop watch when LiCl is added</p> <p>(1) Stand alone mark for idea of record / measure temperature at regular time intervals</p> <p>(1) Do not award if time is on y axis Allow an annotated sketch graph for M4 and M5</p> <p>Do not award graph that shows extrapolated lines with temperature increasing then decreasing e.g.</p> 	(5)

(Total for Question 3 = 10 marks)

Question Number	Answer	Additional Guidance	Mark
4(a)	<ul style="list-style-type: none"> error: (conical) flask correction: change to pear-shaped flask error: thermometer should not be in the reaction mixture / liquid / flask correction: thermometer (bulb) should be level with entrance / opening to condenser error: apparatus should not be sealed / there would be a build-up of pressure correction: EITHER remove stopper from boiling tube / test tube OR use a bend with a vent / collection tube with side arm 	<p>If the error is omitted but the correction clearly indicates the error, then award the mark e.g removing the stopper implies what the error was Allow errors and corrections shown on diagram Penalise additional incorrect errors e.g. water wrong way in condenser once only</p> <p>(1) Allow change to round-bottomed flask</p> <p>(1) Allow move thermometer bulb until level with entrance to condenser Ignore just 'thermometer should be higher' / 'near to the top' unless shown where on diagram Do not award thermometer at neck of flask</p> <p>(1) Allow replace sealed test tube with beaker / measuring cylinder/ unstoppered container</p> <p>Ignore just 'change test tube to flask' unless mention of open / no bung</p>	(3)

Question Number	Answer	Additional Guidance		Mark
4(b)	A description that makes reference to the following points:	Examples of tests:		(2)
	• test (1)	PCl ₅ / phosphorus(V) chloride / phosphorus pentachloride	Observation with alcohol Steamy fumes Allow white / misty fumes Allow gas turns blue litmus red Do not award white smoke	
	• observation with alcohol (1)	Ethanoic acid / any carboxylic acid and sulfuric / hydrochloric acid (and heat)	Fruity smell	
		Allow Na / sodium	Effervescence / fizzing / bubbles	
		Allow name or formula for reagent but if both are given both must be correct		
	Observation conditional on correct or ‘near miss’ reagent e.g. acid missing in ester test			
	Ignore acidified potassium / sodium dichromate(VI) Ignore additional conditions e.g, heat Ignore names of gases in observation			

Question Number	Answer	Additional Guidance	Mark
4(c)(i)	<ul style="list-style-type: none"> balanced equation 	<p>Examples of equation:</p> $\text{C}_6\text{H}_5\text{COONa} + \text{HCl} \rightarrow \text{C}_6\text{H}_5\text{COOH} + \text{NaCl}$ <p>Or</p> $\text{C}_6\text{H}_5\text{COONa} + \text{H}^+ \rightarrow \text{C}_6\text{H}_5\text{COOH} + \text{Na}^+$ <p>Or</p> $\text{C}_6\text{H}_5\text{COO}^- + \text{H}^+ \rightarrow \text{C}_6\text{H}_5\text{COOH}$ <p>Or</p> $\text{C}_6\text{H}_5\text{COO}^- + \text{HCl} \rightarrow \text{C}_6\text{H}_5\text{COOH} + \text{Cl}^-$ <p>Allow multiples</p> <p>Allow displayed / skeletal formulae / combination of structural, displayed and skeletal formulae for organic reactant / product</p> <p>Ignore molecular formulae for organic reactant / product</p> <p>Ignore state symbols even if incorrect</p> <p>Ignore reversible arrow</p> <p>Do not award –O–Na in reactant</p>	(1)

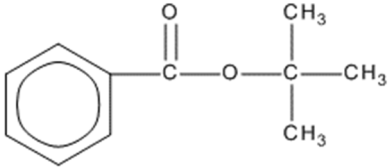
Question Number	Answer	Additional Guidance	Mark
4(c)(ii)	<ul style="list-style-type: none"> filter (under reduced pressure) 	<p>Penalise mention of hot filtration</p> <p>Allow any other type of filtration e.g. suction filtration</p> <p>Allow description of filtration using any type of funnel (except separating funnel) and filter paper</p> <p>Allow diagram of filtration</p> <p>Ignore decanting / rinsing / drying</p>	(1)

Question Number	Answer	Additional Guidance	Mark
4(d)	<ul style="list-style-type: none"> dissolve the benzoic acid / solid / crystals and in the minimum amount / volume and of boiling / hot water 	Allow mix / add / form a (saturated) solution for dissolve Allow solvent for water Allow small amount / volume Ignore missing amount / volume Do not award incorrect solvent e.g. ethanol	(1)

Question Number	Answer	Additional Guidance	Mark
4(e)	An answer that makes reference to the following points: <ul style="list-style-type: none"> (melting temperature / it) is lower (1) (it melts over) a range of temperatures / (the melting temperature / it) is not sharp (1) 	Penalise mention of boiling temperature / distillation once only	(2)

Question Number	Answer	Additional Guidance	Mark
4(f)(i)	<ul style="list-style-type: none"> (alkyl group is) C₄H₉ 	Allow structural / displayed formula for any C ₄ H ₉ group Ignore working Do not award C ₄ H ₉ ⁺	(1)

Question Number	Answer	Additional Guidance	Mark						
4(f)(ii)	<ul style="list-style-type: none">four alcohols with formula C₄H₉OH	<p>(2)</p> <p>Examples of alcohols:</p> <table><tr><td>$\begin{array}{ccccccc} & \text{H} & & \text{H} & & \text{H} & & \text{H} \\ & & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - & \text{C} - \text{OH} \\ & & & & & & & \\ & \text{H} & & \text{H} & & \text{H} & & \text{H} \end{array}$</td><td>$\begin{array}{ccccccc} & \text{H} & & \text{H} & & \text{H} & \\ & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - \text{OH} \\ & & & & & & \\ & \text{H} & & \text{CH}_3 & & \text{H} & \end{array}$</td></tr><tr><td></td><td></td><td>$\begin{array}{ccccccc} & \text{H} & & \text{H} & & \text{OH} & & \text{H} \\ & & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - & \text{C} - \text{H} \\ & & & & & & & \\ & \text{H} & & \text{H} & & \text{H} & & \text{H} \end{array}$</td><td>$\begin{array}{ccccccc} & \text{H} & & \text{OH} & & \text{H} & \\ & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - \text{H} \\ & & & & & & \\ & \text{H} & & \text{CH}_3 & & \text{H} & \end{array}$</td></tr></table> <p>Alcohols can be in any order</p> <p>Allow any combination of structural / displayed formulae / skeletal formulae</p> <p>Allow (1) for any two or three alcohols correct</p> <p>Allow (1) for at least two alcohols as TE from R group in (f)(i)</p> <p>Penalise missing H once only</p> <p>Penalise incorrect connectivity of horizontal OH groups once only e.g. OH-C on left</p> <p>Penalise omission of OH once only i.e. 4 correct carbon skeletons for R scores (1)</p>	$\begin{array}{ccccccc} & \text{H} & & \text{H} & & \text{H} & & \text{H} \\ & & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - & \text{C} - \text{OH} \\ & & & & & & & \\ & \text{H} & & \text{H} & & \text{H} & & \text{H} \end{array}$	$\begin{array}{ccccccc} & \text{H} & & \text{H} & & \text{H} & \\ & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - \text{OH} \\ & & & & & & \\ & \text{H} & & \text{CH}_3 & & \text{H} & \end{array}$			$\begin{array}{ccccccc} & \text{H} & & \text{H} & & \text{OH} & & \text{H} \\ & & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - & \text{C} - \text{H} \\ & & & & & & & \\ & \text{H} & & \text{H} & & \text{H} & & \text{H} \end{array}$	$\begin{array}{ccccccc} & \text{H} & & \text{OH} & & \text{H} & \\ & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - \text{H} \\ & & & & & & \\ & \text{H} & & \text{CH}_3 & & \text{H} & \end{array}$	(2)
$\begin{array}{ccccccc} & \text{H} & & \text{H} & & \text{H} & & \text{H} \\ & & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - & \text{C} - \text{OH} \\ & & & & & & & \\ & \text{H} & & \text{H} & & \text{H} & & \text{H} \end{array}$	$\begin{array}{ccccccc} & \text{H} & & \text{H} & & \text{H} & \\ & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - \text{OH} \\ & & & & & & \\ & \text{H} & & \text{CH}_3 & & \text{H} & \end{array}$								
		$\begin{array}{ccccccc} & \text{H} & & \text{H} & & \text{OH} & & \text{H} \\ & & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - & \text{C} - \text{H} \\ & & & & & & & \\ & \text{H} & & \text{H} & & \text{H} & & \text{H} \end{array}$	$\begin{array}{ccccccc} & \text{H} & & \text{OH} & & \text{H} & \\ & & & & & & \\ \text{H} & - \text{C} & - & \text{C} & - & \text{C} & - \text{H} \\ & & & & & & \\ & \text{H} & & \text{CH}_3 & & \text{H} & \end{array}$						

Question Number	Answer	Additional Guidance	Mark
4(f)(iii)	<ul style="list-style-type: none"> • structure of any butyl benzoate (1) • tertiary butyl R group (1) 	<p>Example of structure:</p>  <p>Allow any combination of structural / displayed formulae / skeletal formula</p> <p>Allow (1) for structure as TE from R group in (f)(i) Allow another mark if the R group would give 2 peaks on ^{13}C NMR spectrum</p>	(2)

(Total for Question 4 = 15 marks)