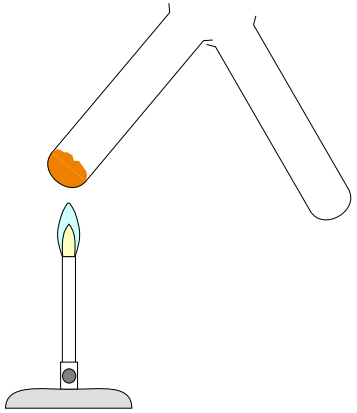
Mark Scheme (Results)

Summer 2019

Pearson International Advanced Subsidiary Level
In Chemistry (WCH13) Paper 01 Practical Skills in Chemistry I

Question number	Answer	Additional guidance	Mark
1(a)	• correct balanced equation	Example of correct equation: $(NH_4)_2CO_3 \rightarrow 2NH_3 + H_2O + CO_2$ Allow multiples H_2CO_3 for $H_2O + CO_2$ Ignore state symbols even if incorrect	**(1)**

Question Number	Answer	Additional guidance	Mark
1(b)		For all the tests ignore indicators If name and formula given both must be correct Observation marks are dependent on test	(6)
	For ammonia • test: reaction with hydrogen chloride / HCl(g) **(1)**	Allow (add /introduce / place next to) HCl If HCl(aq) / conc HCl is used a suitable method is needed e.g. dipping a glass rod into HCl(aq) or opening a bottle of HCl(aq) close to the ammonia. Do not award 'add hydrochloric acid' / HCl(aq) / other hydrogen halides but allow the result mark	
	• result: white smoke **(1)**	Allow white fumes / white solid Do not award steamy / misty fumes / precipitate /cloud	
	For water • test: add (anhydrous) copper((II)) sulfate or cobalt((II)) chloride **(1)**	Accept $CuSO_4$ / $CoCl_2$	
	• result: white to blue or blue to pink **(1)**	If start & finish colours are given both must be correct Allow just $CuSO_4$ turns blue or $CoCl_2$ turns pink Allow observation mark if $CuSO_4$ / $CoCl_2$ solutions are used Do not award $CoCl_2$ turns red	
	For carbon dioxide • test: (add / add to) lime water or (saturated) solution of calcium hydroxide **(1)**	Ignore boiling temperature measurement Accept $Ca(OH)_2$((aq))	
	• result: any indication that a white suspension is formed **(1)**	turns cloudy / turns milky / white precipitate forms	

Question number	Answer	Additional guidance	Mark
1(c)	Diagram showing collecting test tube angled down with mouth of the tube close to and below that of the heated test tube	Example of diagram:  ALLOW angles to the vertical 0—75° Ignore lime water in collecting tube Do not award if additional apparatus used e.g. delivery tube. Do not award if horizontal distance between test tube lips >1cm	**(1)**

Question number	Answer	Additional guidance	Mark
1(d)(i)	An answer that makes reference to the following points: • white **and** precipitate (forms) **(1)** • identifies the precipitate as barium carbonate **(1)**	Ignore subsequent tests in (i) and (ii) Allow white solid / crystals Accept formula $BaCO_3$ If name and formula are given, both must be correct Ignore ammonium chloride (and water) if the precipitate is clearly identified	**(2)**

Question number	Answer	Additional guidance	Mark
1(d)(ii)	An answer that makes reference to the following points: • effervescence (precipitate dissolves) **(1)** • carbon dioxide (is evolved) **(1)**	 Accept bubbling / bubbles / fizzing Ignore gas evolves Accept formula CO_2 Ignore barium chloride / $BaCl_2$ (product) ammonium chloride / NH_4Cl water / H_2O	**(2)**

(Total for Question 1= 12 marks)

Question number	Answer	Additional guidance	Mark
2(a)	An answer that makes reference to the following points: • suitable reagent **(1)** • observation **(1)**	Phosphorus(V) chloride / phosphorus pentachloride / PCl_5 (solid) Allow thionyl chloride / $SOCl_2$ Do not award PCl_5 / $SOCl_2$ **solution** but allow the result mark Steamy fumes / (dense) white fumes / misty fumes Do not award white smoke Allow add sodium (1) and effervescence / fizzing / bubbles (1) add named carboxylic acid **and** strong acid catalyst (1) gives fruity smell (1) Do not award acidified dichromate and orange to green	**(2)**

Question number	Answer	Additional guidance	Mark
2(b)(i)	• potassium dichromate(VI) / $K_2Cr_2O_7$ / sodium dichromate(VI) / $Na_2Cr_2O_7$ **and** sulfuric acid / H_2SO_4	Allow omission of the oxidation number Just 'acid / acidified' $Cr_2O_7^{2-}/H^+$ Ignore heat / reflux / concentrated Do not award Potassium manganate (VII) potassium chromate((VI)) Incorrect oxidation number e.g. potassium dichromate(IV) hydrochloric acid / HCl / Nitric acid / HNO_3	**(1)**

Question number	Answer	Additional guidance	Mark
2(b)(ii)	An answer that makes reference to the following points: • choice of apparatus 1 **(1)** • the ease of oxidation of the aldehyde **(1)**	Example of a justification: The aldehyde is easily oxidised (to a carboxylic acid) / more easily oxidised than the alcohol Allow To prevent further oxidation Partial oxidation occurs Use of reflux (apparatus 2) results in further oxidation M1 and M2 are standalone	**(2)**

Question number	Answer	Additional guidance	Mark
2(b)(iii)	An answer that makes reference to the following points: • suitable reagent **(1)** • result of the selected test **(1)**	**Route 1** (warm with) (blue) Fehling's / (blue) Benedict's reagent Red / brown **and** precipitate / solid **Route 2** (warm with) Tollens' reagent Silver mirror or grey/ black precipitate Ignore Brady's reagent Do not award potassium dichromate(VI) No observation TE on incorrect reagent	**(2)**

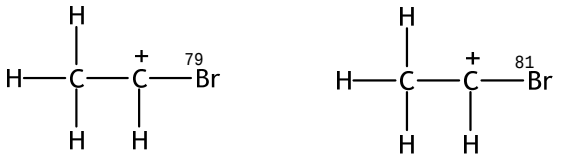
Question number	Answer	Additional guidance	Mark
2(b)(iv)	An answer that makes reference to the following points: • The alcohol cannot be identified **and** because there are two primary alcohols with the molecular formula $C_4H_{10}O$	Accept Alcohol could be butan-1-ol / $CH_3CH_2CH_2CH_2OH$ or 2-methylpropan-1-ol / $(CH_3)_2CHCH_2OH$ both alcohols needed Allow any clear structural / displayed / skeletal formulae Ignore just 'carbon chain could be straight or branched' just 'there are isomers'	**(1)**

Question number	Answer	Additional guidance	Mark
2(b)(v)	An answer that makes reference to the following points: • choice of apparatus 2 **(1)** • ensuring complete reaction / oxidation **(1)**	M2 dependent on M1 Ignore subsequent distillation Ignore reference to preventing loss of volatile reagents or products. Just 'because the ketone does not oxidise further' Just 'reaction is slow'	**(2)**

Question number	Answer	Additional guidance	Mark
2(c)(i)	An answer that makes reference to any **two** of the following points: • the mineral wool holds the alcohol in place (at the end of the tube) **(1)** • the alcohol vapour would not pass over the catalyst slowly enough to react (without the mineral wool) **(1)** • the mineral wool is chemically inert / does not react with the alcohol **(1)**	Ignore large surface area / high melting temperature / good absorbant / prevents evaporation (of the alcohol)/ slow reaction Allow prevents the alcohol mixing with the aluminium oxide / Al_2O_3 / catalyst Allow so the alcohol is not heated directly (by the Bunsen) Ignore Any reference to alcohol burning Allow mineral wool does not burn	**2** **exp**

Question number	Answer	Additional guidance	Mark
2(c)(ii)	An explanation that makes reference to the following points: • the possibility of suck-back **(1)** EITHER • explanation of the cause of suck-back **(1)** OR • description of the consequences of suck-back **(1)**	Examples of correct responses: Suck-back will occur / Water will be drawn up into the reaction tube (from the water bath) Do not award suck-back of anything other than water (On cooling) the pressure in the tube drops and atmospheric pressure acting on the water in the water bath which causes a pressure difference (resulting in suck-back) Allow just drop in pressure / vacuum formed in the reaction tube. Do not award just 'cooling causes suck-back' just 'due to pressure differences' Cold water causes hot tube to crack Allow just test tube cracks/shatters Do not award water will react with the aluminium oxide / tube explodes	**(2)**

Question number	Answer	Additional guidance	Mark
2(d)(i)	Red-brown / brown to colourless	Both needed Allow orange to colourless Allow orange-brown to colourless Allow yellow to colourless Ignore 'clear' Do not award red	**(1)**

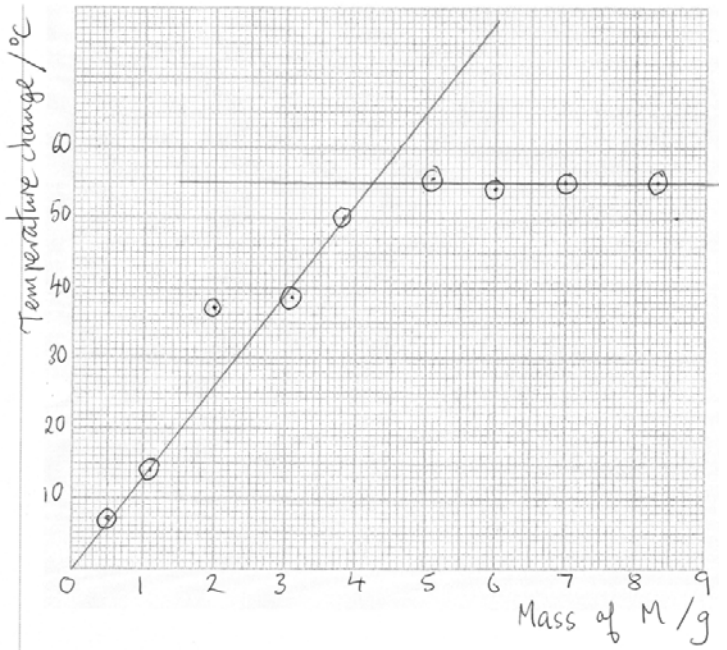
Question number	Answer	Additional guidance	Mark
2(d)(ii)	An explanation that makes reference to the following points: • identification of the peaks by molecular formula or structure **(1)** • only 2,3-dibromobutane can produce the fragments at *m* / *z* = 107 and *m* / *z* = 109 **(1)** • Identifies butan-2-ol as the **only** alcohol that can form but-2-ene (as a product of dehydration and only but-2-ene can form 2,3-dibromobutane) **(1)**	Do **not** penalise omission of charges $C_2H_4{}^{79}Br^+$ and $C_2H_4{}^{81}Br^+$ OR  Allow peaks due to $C_2H_4Br^{(+)}$ Allow identifies $C_4H_8Br_2$ as 2,3-dibromobutane Do not award Just 'alcohol must be butan-2-ol' Just a sequence of structures	**(3)**

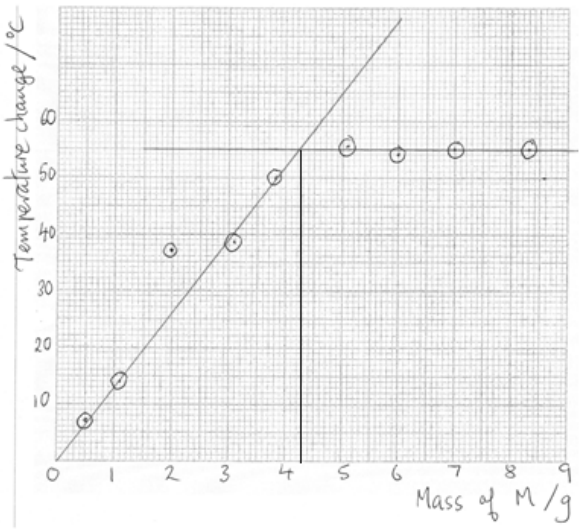
(Total for Question 2= 18 marks)

Question number	Answer	Additional guidance	Mark
3(a)(i)	An answer that makes reference to the following: • the stability of the polystyrene cup	To ensure that the polystyrene cup does not tip over Because the cup is so light, it tips over easily Allow Just 'to provide support' So if the polystyrene cup is damaged the reaction mixture will go into the beaker Do not award To prevent heat loss To provide insulation Because the polystyrene cup gets hot Ignore Just 'to prevent spillage'	**(1)**

Question number	Answer	Additional guidance	Mark
3(a)(ii)	• pipette	Accept 50 cm^3 pipette 25 cm^3 pipette (twice) graduated pipette Allow burette / measuring cylinder / volumetric flask Do not award beaker / flask	**(1)**

Question number	Answer	Additional guidance	Mark
3(a)(iii)	• heat loss is similar for all the experiment runs	Allow to minimise heat loss Allow no heat loss Allow reverse argument e.g. heat loss greater with filings Do not award so reaction goes to completion Ignore References to reaction rate References to temperature	**(1)** **exp**

Question number	Answer	Additional guidance	Mark
3(b)(i)	• suitable choice of scale and correct choice of axes **(1)** • axes labelled, with units **(1)** • all points plotted correctly **(1)**	Example of graph below  Points plotted should cover at least 50% of the graph in both directions. Allow 2 g per large square on x-axis with y-axis scale as shown. Allow 'temperature' and 'T' for 'temperature change' Ignore punctuation errors e.g. (g) instead of /g Ignore scale errors that lie outside the range of the points plotted	**(3)**

Question number	Answer	Additional guidance	Mark
3(b)(ii)	• two best fit lines **(1)** • correct reading of the mass at the intersection of the best fit lines **(1)**	 Allow reasonable best fit line close to points 1, 2, 4 and 5 If the line is continued to the axis it must be within one small square of the origin Do not award a BFL that includes point 3 Second line should be horizontal at about 55°C Do not award if horizontal line deviates 2 small squares from extreme left to extreme right 4.1—4.5 (g) TE on lines drawn but do not award unless value is obtained from the intersection of a **straight**, angled best fit line and a straight best fit line through the last four points	**(2)**

<table>
<tr><th>Question number</th><th>Answer</th><th>Additional guidance</th><th>Mark</th></tr>
<tr><td>3(b)(iii)</td><td>• calculation of amount (moles) of copper(II) sulfate (1)

• equating of moles of copper(II) sulfate and of M via equation (1)

• calculation of A_r of M to 2 or 3 SF (1)</td><td>Example of calculation

Do not penalise intermediate rounding unless incorrect or 1 SF

TE on mass from 3(b)(ii) and at each stage

amount of $CuSO_4$
$= 50.0 \times 1.35 \times 10^{-3}$
$(= 6.75 \times 10^{-2} / 0.0675$ mol)

From equation

4.3 g of M ≡ mol $CuSO_4$ $= 50.0 \times 1.35 \times 10^{-3}$

A_r of M $= 4.3 / 6.75 \times 10^{-2} = 64 / 63.7$

If no working, correct answer to 3 SF using data from (b)(ii) scores (3)
<table><tr><th>Mass</th><th>A_r</th></tr><tr><td>4.1</td><td>61 / 60.7</td></tr><tr><td>4.2</td><td>62 / 62.2</td></tr><tr><td>4.4</td><td>65 / 65.2</td></tr><tr><td>4.5</td><td>67 / 66,7</td></tr></table>
Ignore units of g mol^{-1}</td><td>(3)</td></tr>
</table>

Question number	Answer	Additional guidance	Mark
3(b)(iv)	• Mixture was not stirred (effectively) Or • local heating occurred	Allow Temperature of the solution was not uniform Do not award incorrect quantities used temperature / mass measured incorrectly heat loss incomplete reaction reactants not mixed	**(1)**

(Total for Question 3= 12 marks)

Question number	Answer	Additional guidance	Mark
4(a)	An answer that makes reference to the following: • results that are within 0.2 cm^3	Allow within 0.1 cm^3 ±0.1 cm^3 of the mean results 0.2 / 0.1 (cm^3) apart Do not award ±0.2 cm^3 Ignore Omission of units Reference to 'good agreement' / similar values / same values	**(1)**

Question number	Answer	Additional guidance	Mark
4(b)	((24.10 + 24.30)/2) = 24.2(0) (cm^3)		**(1)**

Question number	Answer	Additional guidance	Mark
4(c)	• from yellow **(1)** • to orange **(1)**	Do not award red Correct colours in reverse order scores (1)	**(2)**

Question number	Answer	Additional guidance	Mark
4(d)	Possible route through the calculation • calculation of amount of hydrochloric acid **(1)** • calculation of amount of NaOH in 250 cm^3 (1:1 ratio) **(1)** • calculation of mass of NaOH **(1)** • calculation of percentage purity of NaOH **(1)**	TE on mean titre from (b) and TE at each stage mol (HCl) = $24.20 \times 0.095 \times 10^{-3}$ $= 2.299 \times 10^{-3} / 0.002299$ NaOH (in 250) = $10 \times 2.299 \times 10^{-3}$ (mol) $= 2.299 \times 10^{-2} / 0.02299$ (mol) Mass NaOH (in 250) $= 40 \times 2.299 \times 10^{-2} = 0.9196$ (g) Purity NaOH = 100 x 0.9196 / 0.95 = 96.8% Ignore SF except 1 SF Do not award purity >100% or any value based on an uncalculated mass of NaOH Correct answer no working scores (4) If mean titre calculated using all four titres (24.28 cm^3) purity = 97.12% If calculated using all last three titres (24.03 cm^3) purity = 96.13%	**(4)**

(Total for Question 4 = 8 marks)
Total for Paper = 50 marks