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Modern Chemistry 71 Chapter Test Name Class Date Chapter Test B, continued 17. A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19. A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.

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Modern Chemistry 35 Chapter Test Name Class Date Chapter Test B, continued Write the orbital notation for the following elements in the space provided. 37. lithium, atomic number 3 38. carbon, atomic number 6 39. neon, atomic number 10 PART VI Write the answers to the questions on the line to the left, and show your work in the space provided.

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Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test A Chapter: The Periodic Law Use the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1.

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Modern Chemistry 137 Chapter Test Name Class Date Chapter Test A, continued ____ 19. In the figure on the previous page, the pH at the equivalence point a. is equal to 7.0. b. is greater than 7.0. c. is less than 7.0. d. cannot be determined from the data given. ____ 20. In the figure on the previous page, the volume of titration standard

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Modern Chemistry 47 Chapter Test Name Class Date Chapter Test A, continued ____ 7. The melting points of ionic compounds are higher than the melting points of molecular compounds because a. ionic substances tend to vaporize at room temperature. b. ionic substances are brittle. c. attractive forces between ions are greater than the attractive ...

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CHAPTER 7 • HOMEWORK 7-1 VOCABULARY 1. false 2. true 3. false 4. false 5. true SKILL BUILDER AgCl—In the compound, there is one atom of silver for each atom of chlorine. Ba(OH) 2 —The subscript 2 refers to everything inside the parentheses, so there are two hydroxide ions for

HOMEWORK 7-2 - Kettering City School District

Modern Chemistry: Chapter Tests with Answer Key, 2006, Holt Rinehart & Winston, Holt, Rinehart and Winston Staff, Harcourt School Publishers, 2006 ... Michigan Holt Chemistry and Modern Chemistry Test Preparation Workbook: Help for the MME , Holt, Holt Rinehart & Winston, Jun 1, 2006, Juvenile Nonfiction, 126 pages. .

Modern Chemistry: Chapter Tests with Answer Key, 2006 ...

Modern Chemistry . Name Section Quiz, continued Class Date 6. How many total moles of atoms are there in one mole of silver(I) phosphate, Ag₃P₀₄? d. 11 7. What is the correct molar mass of Ag₃P₀₄? ... Assessment Class Date Chemical Formulas and Chemical Compounds Section Q Iz: Determining Chemical Formula

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CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

mc06se cFMsr i-vi

Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test B Chapter: The Periodic Law PART I In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1. In his periodic table, Mendeleev did not list all of the elements in order

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Modern Chemistry 101 Chapter Test Name Class Date Chapter Test B, continued 18. The pressure exerted by each gas in a mixture is called the of that gas. 19. If the temperature and number of moles of a gas remain constant but the volume increases, the pressure of the gas will . 20. The lowest possible temperature, corresponding to zero on the ...

Assessment Chapter Test B

Modern Chemistry 141 Chapter Test Name Class Date Chapter Test B, continued 18. When a strong acid is titrated with a weak base, the pH of the solution at the equivalence point is than 7. 19. A is a

highly purified solid used to check the concentration of a standard solution. 20. A 1 M solution of NaOH will have a pH that is

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