

Nombrar los complejos -II- : (Solución a formular complejos I)

- 1 $[\text{Pt Cl}_2(\text{NH}_3)_2]$
- 2 $[\text{Ni}(\text{NH}_3)_6](\text{NO}_3)_2$
- 3 $[\text{Fe}(\text{H}_2\text{O})_2(\text{NH}_3)_4]^{+2}$
- 4 $[\text{Ir}(\text{AsO}_4)(\text{H}_2\text{O})_5]$
- 5 $[\text{Al}(\text{OH})_2(\text{H}_2\text{O})_4]_2\text{SO}_4$
- 6 $[\text{Pt Cl Br}(\text{CH}_3 - \text{NH}_2)_4]^{+2}$
- 7 $\text{Cs}_2[\text{Ta F}_7]$
- 8 $\text{NH}_4[\text{AlH}_4]$
- 9 $[\text{Cr O}_2(\text{O}_2)(\text{CN})_2\text{NH}_3]^{-5}$
- 10 $[\text{Pt Cl}_3\text{NO}_2]^{-2}$
- 11 $[\text{Fe}(\text{CO})_5]$
- 12 $\text{Rb}_2[\text{Zn}(\text{OH})_4]$
- 13 $[\text{Cr F Cl O}_2(\text{H}_2\text{O})_2]^{-3}$
- 14 $[\text{Rh}(\text{NO}_3)(\text{NH}_3)_5]\text{Br}_2$
- 15 $[\text{Pd Br Cl F}]$
- 16 $\text{Sn}_3[\text{Nb F}_6\text{O}]_2$
- 17 $[\text{Co}(\text{NO}_3)_6]^{-3}$
- 18 $\text{K}_3[\text{Fe}(\text{CN})_5(\text{NO})]$
- 19 $\text{Ag}[\text{Pt Br Cl}(\text{NO}_2)(\text{NH}_3)]$
- 20 $\text{Ca}_2[\text{Mo Cl}_8]$
- 21 $[\text{Cr}(\text{CrO}_4)(\text{H}_2\text{O})_4]_2\text{Cr}_2\text{O}_7$
- 22 $[\text{Zn}(\text{OH})_3(\text{H}_2\text{O})]^{-1}$
- 23 $[\text{Cr}(\text{NH}_3)_6][\text{Co}(\text{CO}_3)_3]$
- 24 $[\text{Mo}(\text{SCN})(\text{NH}_3)_5]\text{Cl}_2$
- 25 $[\text{Ru}(\text{NH}_3)_5(\text{N}_2)]\text{Br}_2$
- 26 $\text{Cd}[\text{Hg Cl}_2(\text{CN})_2]$
- 27 $\text{Na}[\text{Co}(\text{CO})_4]$
- 28 $\text{K}_4[\text{Fe}(\text{CN})_6]$
- 29 $[\text{Cr}(\text{OH})(\text{NH}_3)_2(\text{H}_2)_3]\text{NO}_3$
- 30 $[\text{Co}(\text{NH}_3)_6][\text{Cr CO}_3)_3]$
- 31 $[\text{Cr}(\text{CO})_6]$
- 32 $[\text{Cu}(\text{NH}_3)_4](\text{NO}_3^-)_2$
- 33 $\text{Al}_2[\text{Pt}(\text{CN})_4]_3$
- 34 $\text{K}_2[\text{Co}(\text{NO}_2)_5(\text{NH}_3)]$