CHEM101 Laboratory Exercise #2 Laboratory Notebook
Moles, Concentration, Acid-Base Reactions and Quantitative Analysis by Titration

Using Microsoft Word, students must write their in-lab notes below <u>while</u> completing the Laboratory exercise. The Laboratory notebook must be uploaded to the CHEM 101 Lab Brightspace site as a readable .pdf by the end of the lab period. Please see page 7 of the 202405 CHEM 101 Lab manual for all the information required in the in-lab notes. The documentation below must reflect the student's work without assistance from others.

Name: ____Arfaz Hossain____ **Lab Section:** _B12_**Quad:** _2_ **Date:** _May 28, 2024_

In-lab Notes:

EXPERIMENT 1:
Experimental Procedure for standardising a solution of NaOH using a titration:
Initial Without NaOH Solution Volume: 200mL
Final with added NaOH Solution: 212mL

1st
KHP Amount: 0.502g
Burette Reading Initial: 45.7mL
Burette Reading Final: 32.8mL
Difference: 12.9mL

2nd
KHP Amount: 0.507g
Burette Reading Initial: 32.8mL
Burette Reading Final: 20.6mL
Difference: 12.2

3rd
KHP Amount: 0.501g
Burette Reading Initial: 20.6mL
Burette Reading Final: 8.4mL
Difference: 12.

EXPERIMENT 2 (Vineger):
1st

Vineger Amount: 5.00mL
Burette Reading Initial: 48.5mL
Burette Reading Final: 27.9mL

2nd
Vineger Amount: 5.00mL
Burette Reading Initial: 27.9mL
Burette Reading Final: 6.8mL

3rd
Vineger Amount: 5.00mL
Burette Reading Initial: 37.8mL
Burette Reading Final: 20.8mL

Experimental Procedure: We put NaOH in a beaker and KHC in 3 flasks, titrated and mixed and took note of the data. We also did the same thing with Vineger as well.