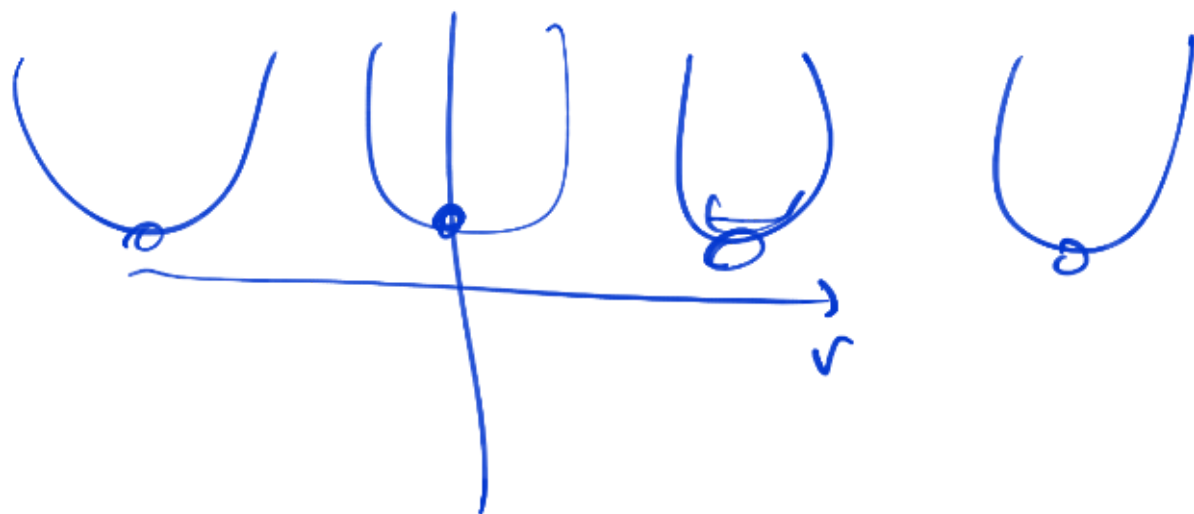


Solids

$$C_V \approx \underset{\uparrow}{25} \frac{J}{mol \cdot K}$$

$$R = 8,314 \frac{J}{mol \cdot K}$$

$$C_V \approx 3R \quad V(r)$$



$$U_i = \frac{1}{2} m v^2 + \frac{1}{2} k \Delta r^2$$

$$U_i = \frac{1}{2} m (v_x^2 + v_y^2 + v_z^2) + \frac{1}{2} k_{ef} (\Delta x_i^2 + \Delta y_i^2 + \Delta z_i^2)$$

$$U = \sum_{i=1}^N U_i = \sum_{i=1}^N \left(\frac{1}{2} m v_i^2 + \frac{1}{2} k_{ef} \Delta r_i^2 \right)$$

$$\frac{1}{2} \quad \frac{1}{2}$$

$$z = 6$$

$$U = z \left(\frac{1}{2} k T \right)$$

$$U = 3 k T \cdot N$$

$$U = 3 N k T \quad \frac{N_2}{N_4}$$

$$U = 3 n R T$$

$$\Delta U = 3 n R \Delta T \quad \text{—}$$

$$Q = C_v \cdot n \Delta T \quad \text{—}$$

$$\boxed{C_v = 3R}$$

$$C_{v, Pb} = 3R$$

Isol de Cu

$$C_v^{Cu} = 3R$$

1 mol de Pb
de 273K a 373K

$$Q = C_v n \Delta T$$

$$= 3 \cdot 8,314 \frac{J}{mol \cdot K} \cdot 1 mol \cdot 100K$$

$$= 2494,2 J$$

$$= 2,5 kJ$$

1 kg de Cu.

$$C_v = 3R$$

$$Q = C n \Delta T$$

$$= C \frac{m}{M} \Delta T$$

$$Q = \left(\frac{C}{M} \right) m \Delta T$$

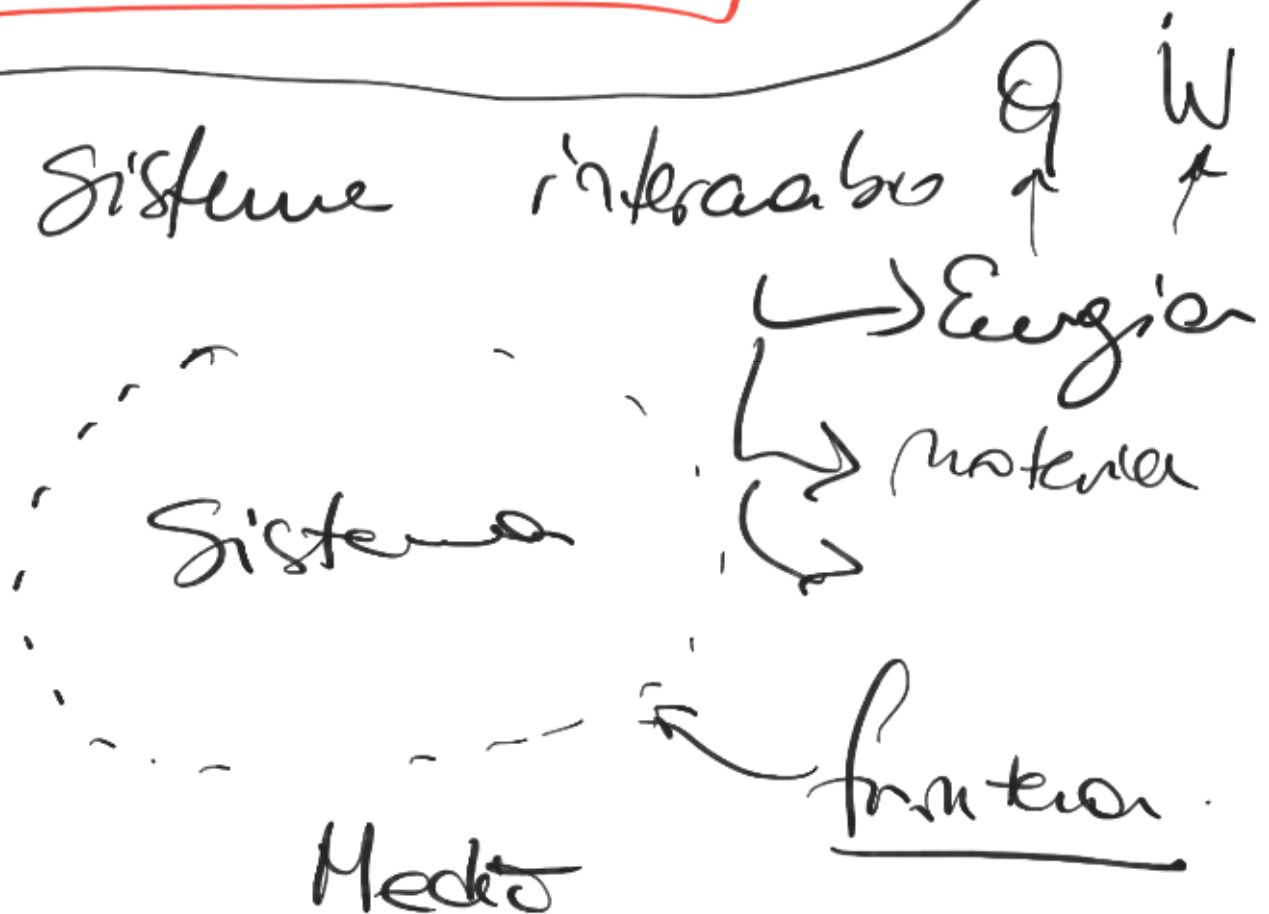
Copacabana
Colômbia

$$c = \frac{C}{M} = \frac{3R}{M} = c$$

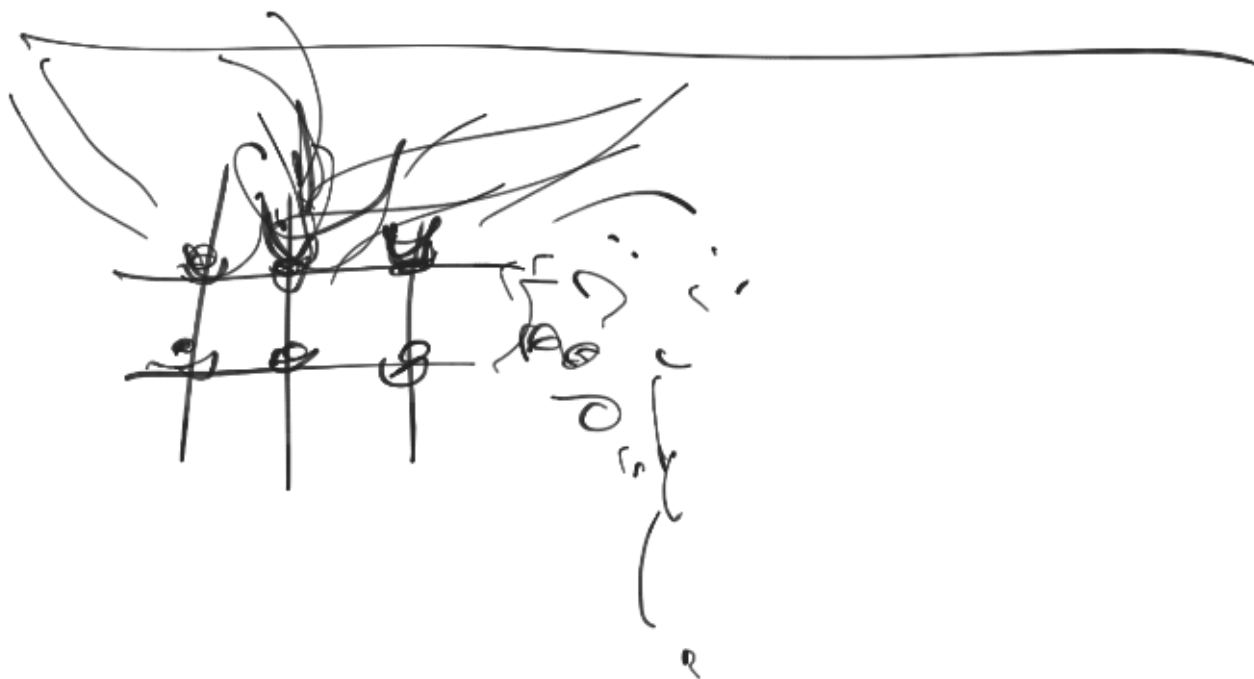
Calorimetric

$$Q_w = \frac{3.8314 \text{ J/mol K} \cdot 1000 \text{ g} \cdot 100\%}{63.5 \text{ g/mol}}$$

$$Q_w = 38.3 \text{ kJ}$$



Sistema
Aislado.



$$Q = L \mu$$
$$= L n$$

$$L = Q/\mu \quad L = Q/\mu$$