

## Lecture 20 Problems

### Problem 1

$\text{PbSO}_4$  has solubility  $4.25 \cdot 10^{-2} \text{ g/L}$ ; then,

$$K_{sp} = (4.25 \cdot 10^{-2})^2 = 1.8 \cdot 10^{-3}$$

## Problem 2

(a): The concentration/solubility doubles.

(b): The partial pressure of the  $\text{CO}_2$  did not change, thus the solubility did not change.