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ACKNOWLEDGEMENT

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SUMMARY

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1 INTRODUCTION

Most of the newly discovered active pharmaceutical ingredients (APIs) are poorly water-soluble in crystalline forms, which limit their bioavailability, dissolution and then their distribution through the organism. This fact limits their wider oral use as a solid drug in medical treatment. Nowadays, the techniques of combinatorial chemistry and high throughput screening have led to a sharp rise in the quantity of these molecules, so the oral administration of poorly soluble drugs became the biggest challenge for formulation scientists in the pharmaceutical industry [1]. There are strategies to overcome this issue such as forming amorphous dispersion, self-emulsifying drug delivery systems, nanocrystals or complexing with cyclodextrin [2].

In 1961, Sekiguchi and Obi provided the earliest account of solid dispersion, when they discovered that the creation of eutectic mixtures enhances the rate of drug release and, bioavailability [3].

Our aim is to overcome the poor solubility by using amorphous solid phases of APIs and to avert the rearrangement of molecules into a crystal lattice. However, crystalline forms of APIs are advantageous due to their better stability during long-term storage and better predictions of changes at the molecular level under defined conditions [4].

The better solubility of APIs in amorphous forms originates from a higher Gibbs energy of the amorphous form in comparison with the crystalline forms. During processing, storage, and after contact with water or humidity, the thermodynamically metastable amorphous forms tend to crystallise. Solid mixtures of API and excipients (e.g. polymeric excipients) create amorphous solid dispersions (ASD), and offer a way to inhibit crystallisation of the API before and after oral administration of the dose [5].

The addition of an excipient to an amorphous API can generally have a twofold effect on the rate of solid-state crystallisation, affecting both thermodynamic and kinetic aspects. Thermodynamically, it reduces the Gibbs energy due to strong intermolecular interactions between API and its excipient, as well as it increases kinetic barriers to recrystallisation. On the atomic scale of individual interactions, hydrogen bonding has the biggest contribution [6].

2 Studied compounds

2.1 Polylactic acid

2.2 Active pharmaceutical ingredients

The second selected API is **naproxen** a non-steroidal anti-inflammatory drug, used as a painkiller. Naproxen contains three oxygen atoms (one carboxyl group and one ether bond), the structure is shown in Figure ?? on the right. According to its structure, naproxen can donate one hydrogen bond and accept up to three hydrogen bonds. Naproxen is a white crystalline powder, with a molar weight of $M_w = 230.263 \text{ g mol}^{-1}$ and melting point 429.3 K [7].

Carbamazepine is a representative anticonvulsant, which is used for the treatment of seizures and neuropathic pain. Carbamazepine contains two nitrogen atoms (amide group) and one oxygen in the carboxyl group; its structure is shown in Figure ?? on the left side. According to its structure, carbamazepine can accept and donate one hydrogen bond. Carbamazepine is a white crystalline powder, with a molar weight of $M_w = 236.273 \text{ g mol}^{-1}$ and melting temperature of 463.6 K [7].

The selected API was **sulfathiazole** as a representative antibiotic drug from the sulfonamides group, which is used in the treatment of pyogenic cutaneous infections [?]. Sulfathiazole is a white crystalline powder, with molar weight (M_w) = 255.3 g mol^{-1} [?], which is highly polymorphic, there are 5 polymorphs discovered so far [?]. All known polymorphs of sulfathiazole crystallize in the $P2_1/c$ space group, but there are differences in intermolecular bonding and structural properties [?]. For our research, we chose the II polymorph, the structure of which is pictured in Figure 1, there are four molecules of sulfathiazole in the crystal monoclinic unit cell. Our goal is to use quantum computing methods to determine the charges on atoms to complete the force field file model, and then to use MD to validate the model through comparison of simulated crystallographic parameters and density with experimental data.

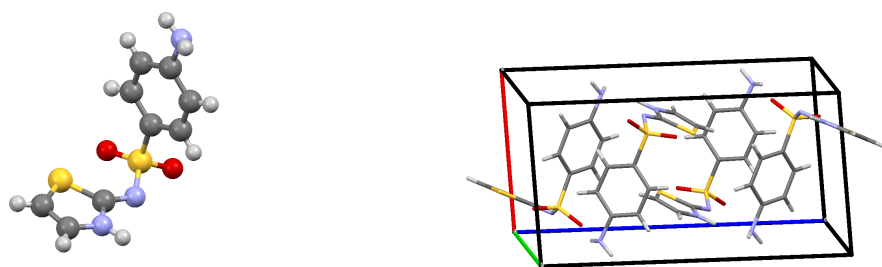


Figure 1: Sulfathiazole - molecular structure on the left and a unit cell of its II polymorph on the right.

3 THEORETICAL PART

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4 METHODS

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5 RESULTS AND DISCUSSION

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6 CONCLUSION

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List of Abbreviations

x position

v velocity

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List of Symbols

x position

v velocity

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Appendix A Headline