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Title:	Heteroleptic Iminophosphonamide In(III) Complexes: Source of Mild Lewis Acid Indium Centers
Authors:	Prashanth, B. (/jspui/browse?type=author&value=Prashanth%2C+B.) Bawari, D. (/jspui/browse?type=author&value=Bawari%2C+D.) Singh, Sanjay (/jspui/browse?type=author&value=Singh%2C+Sanjay)
Keywords:	Chelate Dimer Indium Lewis Acid Ligand
Issue Date:	2017
Publisher:	Wiley-VCH Verlag GmbH & Co. KGaA, Weinheim
Citation:	ChemistrySelect, 2(6), pp.2039-2043.
Abstract:	New In(III) complexes with iminophosphonamide ligand are reported for the first time. Reaction of the lithium iminophosphonamide, $\text{LLi}\cdot 2\text{OEt}_2$ with $\text{InCl}_3$ afforded the heterobimetallic compound, $\text{LiInCl}(\mu\text{-Cl})_2\text{Li}\cdot 2\text{OEt}_2$ (1) ( $\text{L}=(2,6\text{-iPr}_2\text{C}_6\text{H}_3\text{N})\text{P}(\text{Ph}_2)(\text{NtBu})$ ). To prevent the formation of $\text{LiCl}$ adduct in 1, and to obtain the targeted $\text{LiInCl}_2$ molecule, potassium salt of the ligand $\text{LK}$ was reacted with $\text{InCl}_3$ that gave the adduct $\text{LiInCl}_2(\text{THF})$ (2). Formation of compounds 1 and 2 can be considered an outcome of the In(III) centers to exhibit Lewis acidity. Reaction of $\text{LK}$ with $\text{InCl}_3$ in non coordinating solvent toluene gave the dimer $[\text{LiInCl}(\mu\text{-Cl})]_2$ (3) and its further treatment with $\text{Ph}_2\text{P}(=\text{O})\text{NHtBu}$ afforded the adduct, $\text{LiInCl}_2\cdot(\text{O}=\text{P}(\text{Ph}_2)\text{NHtBu})$ (4). Formation of the dimer 3 or the adduct 4 are again the reminiscent of the Lewis acid tendency of In(III) centers. Complexes 1–4 have been characterized using multinuclear NMR and HRMS. The single crystal X-ray structures of 1, 3 and 4 have also been elucidated.
URI:	<a href="https://chemistry-europe.onlinelibrary.wiley.com/doi/abs/10.1002/slct.201700026">https://chemistry-europe.onlinelibrary.wiley.com/doi/abs/10.1002/slct.201700026</a> ( <a href="https://chemistry-europe.onlinelibrary.wiley.com/doi/abs/10.1002/slct.201700026">https://chemistry-europe.onlinelibrary.wiley.com/doi/abs/10.1002/slct.201700026</a> ) <a href="http://hdl.handle.net/123456789/2626">http://hdl.handle.net/123456789/2626</a> ( <a href="http://hdl.handle.net/123456789/2626">http://hdl.handle.net/123456789/2626</a> )
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