		WAS THE THE PARTY OF THE PARTY
	SAP ID- 60004200132	
	Nome - Ayush Jain	
22/07/21	Engineering Chemistry	DATE:
7.00	2 3 9	
	Corrosion - Tutorial 1	
	+ 7	
1	OCH CONTRACTOR AND A CONTRACTOR OF CONTRACTO	
	The standard emf of the cell is 0.462 V. Cu/cut	2/1m) 11 Agt (1m) /A
(7)	the standard ent of the cell is used. with	
	If the standard potential of Cu is 0.337	V, what is
	the standard potential of Ag electrode.	
		· · · · · · · · · · · · · · · · · · ·
Calvison	V.OBSTUT :	
20174,00	1. 1. 1> C. 2+ +2e-	
•	and anode At anode: Cu -> Cu2+ +2e-	
	At cathode: 2Agt + 2e -> 2Ag	note - Stan
	Vaccio-	E. E. M. Committee of the Committee of t
	Net cell reaction: Cu + 2 Ag +	Cu2+
	Net cert reaction	
	Ecell = Ecathode - Eanode	
ant in	0.462 = Eng - Ecul	and the
	are is has in a statement material has	
	· co	634
EGMHON	.: E'Ag = 0.462 + 0.337	
	: E'Ag = 0:799 V	n () en
	The standard potential of Ag electrode	is 0.799 V.
	·· Me danied I	Se 1/2 14
	- 5-40	CONTRACTOR OF THE PARTY OF THE
	and the state of the	
2>	Calculate standard electrode potential of lead	electrode when
	it is in contact with 0.0096 M Pb2+ solution	
	The E value of lead is -0.18025 V.	
	Co 242	16.3
Solution	$[Pb^{2+}] = 0.0096M$	
	EPB2+1Pb = -0.18025 V	
	Water Control of the	
	EP62+/Pb = EP62+/Pb - 2.303 RT 10910 1 NF [P62+]	
	UE [LPst]	

	Page 2
	DATE:
	:. E Pb2+ /Pb = EPb2+ /Pb - 2-303RT logio [Pb2+]
	= -0.18025 - 2.303 x8.314 x 301 x 10g (0.0096)
3'	a to the same of t
	= -0.18025 - 0.02986 x (-2.0177)
	= -0.1200 V
	Total Control of State of Stat
	is -0.1200V.
	Total pace - the interior that say
	Share 3 - shares 3 = usis
3>	For the cell reaction Ni/Ni+2 (0.01M) (u2+ (0.5 m) /cu. The
	Standard reduction potentials of Ni and Cu are -0.25
	and calculate the EMF of the cell at 298k.
	and costant in the critical and an area of the critical and area of the
Ves	to a shortesta pla yo without hockanie of
Solution	At anode: $Ni \longrightarrow Ni^{+2} + 2e^{-}$
	At cathode: $Cu^{24} + 2e^{-} \rightarrow Cu$
Name of the last o	Net cell reaction: Ni + Cu2+ -> Ni+2 + Cu
	Versus as the land and the
	Fiell = Ecathode - Eanode
	= 0.34 - (-0.25)
	: E'cell = 0.59 V

	Page 3		
		DATE:	
	NOW,		
	Ecol = E'coll - 2-303 RT log [N; 2+]	0.1	
	= 0.59 - 2.363 ×8.314 × 298 × log 0.01 2 × 96500 0.5		
0	= 0.59 - 0.0592 x (-1.6989)		
	= 0.64029		
	:. Fcen = 0.6403 V		
	The EMF of the cell is 0.6403 voHs.		
•			