Enthusiast, Leader & Achiever Course

PHASE : ALL PHASE

TARGET : PRE-MEDICAL 2020

Test Type : **MAJOR** Test Pattern : **NEET (UG)**

TEST DATE : 03-09-2020

Q.	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	19	20	21	22	23	24	25	26	27	28	29	30
A.	1	2	4	4	3	1	2	2	2	1	1	2	4	4	3	2	2	2	4	1	3	3	2	1	1	2	4	4	1	1
Q.	31	32	33	34	35	36	37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54	55	56	57	58	59	60
A.	2	4	2	3	1	2	2	2	2	4	4	1	1	4	1	3	3	3	1	1	2	4	3	4	1	3	4	1	3	1
Q.	61	62	63	64	65	66	67	68	69	70	71	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86	87	88	89	90
A.	2	4	3	1	3	3	1	1	1	4	2	3	4	2	4	2	4	2	4	1	1	4	2	2	4	2	3	2	4	1
Q.	91	92	93	94	95	96	97	98	99	100	101	102	103	104	105	106	107	108	109	110	111	112	113	114	115	116	117	118	119	120
A.	2	4	2	4	3	3	3	1	4	2	1	3	3	4	3	1	1	3	2	3	2	2	2	2	3	4	2	4	1	2
Q.	121	122	123	124	125	126	127	128	129	130	131	132	133	134	135	136	137	138	139	140	141	142	143	144	145	146	147	148	149	150
A.	3	2	4	4	4	2	1	1	2	4	3	3	3	4	3	3	2	3	2	2	1	1	4	3	2	3	1	4	4	1
Q.	151	152	153	154	155	156	157	158	159	160	161	162	163	164	165	166	167	168	169	170	171	172	173	174	175	176	177	178	179	180
A.	1	3	2	4	1	2	2	4	4	3	2	3	1	2	1	1	3	1	2	3	2	1	3	1	2	1	3	3	2	3

HINT – SHEET

1. Ans (1)

Weight of cylinder = upthrust due to both liquids

$V \times D \times g$

$= \left(\frac{A}{5} \times \frac{3}{4}L\right) \times d \times g + \left(\frac{A}{5} \times \frac{L}{4}\right) \times 2d \times g$

$\Rightarrow \left(\frac{A}{5} \times L\right) \times D \times g = \frac{A \times L \times d \times g}{4}$

$\Rightarrow \frac{D}{5} = \frac{d}{4}$

$\therefore D = \frac{5}{4}d$

2. Ans (2)

$60 \times 1(40 - T) = 10 \times 80 + 10 \times T$

$2400 - 60\,T = 800 + 10\,T$

$70\,T = 1600$

$T = \frac{160}{7}\,^{o}C$

$\approx 23°C$

3. Ans (4)

Angular frequency of simple pendulum,

$\omega = \sqrt{\frac{g}{\ell}}$

As the support oscillate up and down, effective g changes

$\Delta g = 2.\omega_0^2 A = 2 \times 1^2 \times 10^{-2} ms^{-2}$

$= 0.02\ ms^{-1}$

Change in angular frequency $d\omega$ of the pendulum is given by

$\frac{d\omega}{\omega} = \frac{-1}{2} \cdot \frac{dg}{g}$

$\frac{|\Delta\omega|}{\omega} = \frac{1}{2}\frac{|\Delta g|}{g}$

$= \frac{1}{2} \times \frac{0.02}{10}$

$= 0.001\ rad\ s^{-1}$

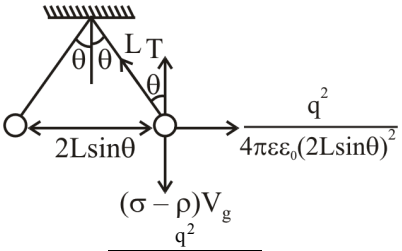
4. **Ans (4)**

By comparing given equation of progressive wave with standard equation y = acos (kx – ωt), we get

$$k = \frac{2\pi}{\lambda} = \alpha \Rightarrow \alpha = \frac{2\pi}{0.08} = 25\pi$$

$$\text{and } \omega = \frac{2\pi}{T} = \beta \Rightarrow \beta = \frac{2\pi}{2} = \pi$$

5. **Ans (3)**



$$\tan\theta = \frac{\frac{q^2}{4\pi\varepsilon\varepsilon_0(2L\sin\theta)^2}}{(\sigma-\rho)V_g}$$

$$= \frac{\frac{q^2}{4\pi\varepsilon_0(2L\sin\theta)^2}}{\sigma V_g}$$

$$\Rightarrow \varepsilon(\sigma-\rho) = \sigma$$

$$\varepsilon = \frac{\sigma}{\sigma-\rho} = \frac{1.6}{1.6-0.8}$$

$$\varepsilon = 2$$

6. **Ans (1)**

As condenser is isolated, dielectric is placed, charge constant, capacity increase by

$$C = \frac{\in . KA}{d}$$

$$\text{\& potential } V = \frac{q}{C} \Rightarrow C\uparrow, V\downarrow$$

$$\text{potential energy } U = \frac{q^2}{2C}\downarrow$$

7. **Ans (2)**

Here, i = 4A, V = 20 volt, so,

$$R = \frac{V}{I} = \frac{20}{4} = 5\Omega$$

Since, voltmeter is connected in parallel with resistance R, the effective resistance of this combination is 5 Ω only if the resistance R is greater than 5 Ω, since total resistance in parallel combination becomes less than individual resistance.

8. **Ans (2)**

Given: Energy level of ground state of hydrogen $(n_1) = 1$.

We know that atomic number of hydrogen atom $(Z_1) = 1$ and atomic number of beryllium atom $(Z_2) = 4$.

We also know that orbital radius,

$$r = 4\pi\varepsilon_0 \times \frac{n^2h^2}{4\pi me^2Z} \text{ or } n \propto \sqrt{Z}$$

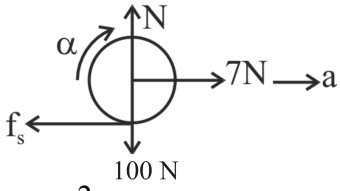
Since, orbital radius of ionised beryllium is equal to the orbital radius of hydrogen atom in ground state, therefore

$$\frac{n_1}{n_2} = \sqrt{\frac{Z_1}{Z_2}} = \sqrt{\frac{1}{4}} = \frac{1}{2}$$

or $n_2 = 2n_1 = 2 \times 1 = 2$

9. **Ans (2)**

$$F = 7 - f_s = 10a$$



$$f_s = \frac{2}{5}ma = 4a$$

$$f_s R = \frac{2}{5}(MR^2)\left(\frac{a}{R}\right)$$

$$7 = 14a$$

$$\Rightarrow a = 0.5 \text{ m/s}^2$$

$$f_s = 2N$$

$$f_L = \mu_s N = 0.1 \times 100 = 10N$$

$$f_s < f_L$$

10. **Ans (1)**

AB → constant pressure compression

ΔW = –ve, ΔU = –ve (V↓ T↓)

ΔQ = –ve

BC → constant volume

ΔW = 0, ΔU = –ve (P↓ T↓)

CD → constant pressure

ΔW = +ve, ΔU = +ve (V↑ T↑)

DA → isothermal

ΔW = –ve

ΔU = 0

ΔQ = –ve

12. Ans (2)

Magnetic field = B_H

T = 60/12 = 5 sec

Magnetic field = $B_H + B$

T = 60/15 = 4 sec

Let after reversing magnetic needle makes x oscillation per minute.

Magnetic field = $B_H - B$, T = 60/x sec

$$T \propto \frac{1}{\sqrt{\text{Magnetic field}}} \Rightarrow \frac{5}{4} = \sqrt{\frac{B_H + B}{B_H}}$$

$$\Rightarrow \frac{B}{B_H} = \frac{9}{16}$$

$$\Rightarrow \frac{5}{\left(\frac{60}{x}\right)} = \sqrt{\frac{B_H - B}{B_H}} = \frac{x}{10} = \sqrt{\frac{7}{16}}$$

$$\Rightarrow x = \sqrt{63}$$

13. Ans (4)

Let $I = I_0 \sin \omega t$

Where $I_0 = 10$, $\omega = 100$

Then $\varepsilon = M \frac{dI}{dt}$

$\therefore \varepsilon_{max} = M \frac{d}{dt} I_0 \sin \omega t$

$= M I_0 \omega \cos \omega t$

$= M \times 10 \times 100\pi$

M = 5 mH

14. Ans (4)

Current will be maximum in the condition of resonance.

So, $I_{max} = \frac{E}{R} = \frac{E}{10} A$

Energy stored in the coil

$$U_L = \frac{1}{2} L I_{max}^2 = \frac{1}{2} \times 10^{-5} E^2$$

energy stored in the capacitor

$$U_C = \frac{1}{2} C E^2 = 10^{-6} E^2$$

$$\therefore \frac{U_C}{U_L} = \frac{1}{5}$$

16. Ans (2)

For TIR $i > i_C$

$60° > i_C \Rightarrow \sin 60 > \sin i_c$

$$\frac{\sqrt{3}}{2} > \frac{\mu}{(3/2)} \quad \Rightarrow \quad \mu < \frac{3\sqrt{3}}{4}$$

17. Ans (2)

$$I^{11} = \frac{I}{2} \cos^2 45$$

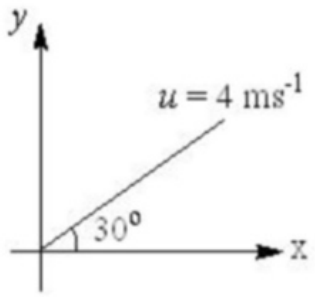
$$I^{11} = \frac{I}{4}$$

$$\text{Transmitted light} = \frac{I^{11}}{I} = \frac{\frac{I}{4} \times 100}{I}$$

$$= \frac{I}{\frac{4}{I}} \times 100 = 25\%$$

18. Ans (2)

Components of velocity of ball relative to lift are



$u_x = 4 \cos 30° = 2\sqrt{3} ms^{-1}$

and $u_y = 4 \sin 30° = 2 ms^{-1}$

and acceleration of ball relative to lift is 12 ms^{-2}

in negative y-direction or vertically downwards.

Hence, time of flight

$$T = \frac{2u_y}{12} = \frac{u_y}{6} = \frac{2}{6} = \frac{1}{3} s$$

21. Ans (3)

$$T = \frac{F}{2\ell} = \frac{728}{2 \times 5}$$

$\Rightarrow$ T = 72.8 dyne/cm

22. Ans (3)

$$\frac{R_1}{R_2} = \frac{R_3}{R_4} \text{ and } R \propto \frac{1}{K}$$

$$\text{so, } \frac{K_2}{K_1} = \frac{K_4}{K_3}$$

$$\Rightarrow K_1K_4 = K_2K_3$$

24. Ans (1)

In case of spherical metal conductor the charge quickly spreads uniformly overthe entire surface because of which charges stay for longer time on the spherical surface. While in case of non-spherical surface, the charge concentration is different at different points due to which the charges do not stay on the surface or longer time.

25. Ans (1)

$$\text{Power dissipated} = i^2R = \left(\frac{E}{R+r}\right)^2 R$$

$$\therefore \left(\frac{E}{R_1+r}\right)^2 R_1 = \left(\frac{E}{R_2+r}\right)^2 R_2$$

$$\Rightarrow$$

$$R_1(R_2^2 + r^2 + 2R_2r) = R_2(R_1^2 + r^2 + 2R_1r)$$

$$\Rightarrow R_1R_2^2 + R_1r^2 + 2R_1R_2r$$

$$= +R_2R_1^2 + R_2r^2 + 2R_1R_2r$$

$$\Rightarrow (R_1 - R_2)r^2 = R_1R_2(R_1 - R_2)$$

$$\Rightarrow r = \sqrt{R_1R_2}$$

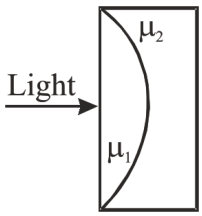
26. Ans (2)

$B = B_0n^2$

$= (5)^2 B_0$

$= 25 B_0$

27. Ans (4)



Equivalent focal length is given by

$$= \frac{1}{f_{eq}} = \frac{1}{f_1} + \frac{1}{f_2}$$

$$\frac{1}{f_{eq}} = (\mu_1 - 1)\left(\frac{1}{\infty} - \frac{1}{-R}\right) + (\mu_2 - 1)\left(\frac{1}{-R} - \frac{1}{\infty}\right)$$

$$\Rightarrow f_{eq} = \frac{R}{\mu_1 - \mu_2}$$

28. Ans (4)

$I_C = 10$ mA

$$i_C = i_E \times \frac{95}{100}$$

$$i_E = \frac{100\, i_C}{95} = 10.53 \text{ mA}$$

$i_B = i_E - i_C = 10.53 - 10 = 0.53$ mA

30. Ans (1)

At equator

$g' = g - \omega^2R$

ω = increase

g' = decrease

31. Ans (2)

$$\eta = 1 - \frac{T_2}{T_1} = 1 - \frac{300}{500} = \frac{2}{5}$$

$$\eta = \frac{W}{Q} \qquad Q = \frac{W}{\eta} = \frac{1000J}{2/5}$$

$Q = 2500$ J

33. Ans (2)

Resultant magnetic force on current carrying loop in uniform magnetic field is always zero.

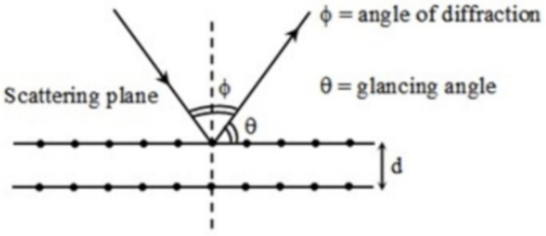
$$|\vec{F}_1 + \vec{F}_2 + \vec{F}_3 + \vec{F}_{QP}| = 0$$

$$\vec{F}_{QP} = |-\vec{F}_1 + \vec{F}_2 + \vec{F}_3|$$

$$= \sqrt{(F_3 - F_1)^2 + F_2^2},$$

where $F_3 > F_1$ and $(F_3 - F_1) \perp F_2$

35. Ans (1)



From the diagram,

$$\theta = 90^\circ - \frac{\phi}{2}$$

40. Ans (4)

Using the relation

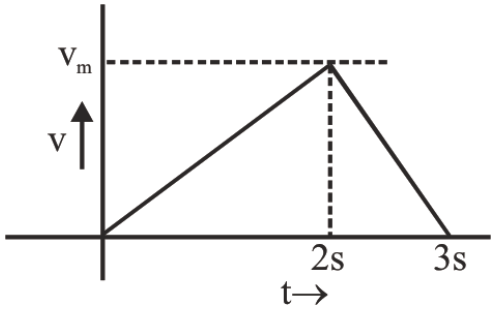
$$s = \frac{\alpha\beta t^2}{2(\alpha+\beta)} = \frac{2\times4\times3\times3}{2(2+4)}$$

$$\left(\text{Given : } \alpha = 2m/s^2,\ \beta = 4m/s^2\right)$$

$$= \frac{2\times4\times3\times3}{2\times6} = 6m$$

OR

$v_m = 2 \times t_1$, $v_m = 4 \times t_2$



$\Rightarrow 2t_1 = 4t_2 \Rightarrow t_1 = 2t_2$

also $t_1 + t_2 = 3s \Rightarrow t_1 = 2s$ and $t_2 = 1s$

and $v_m = 2 \times 2 = 4$ m/s^2

distance = area $= \frac{1}{2} \times 4 \times 3 = 6m$

41. Ans (4)

$$m_{total} = \frac{m}{\left(1 - \frac{1}{2}\right)} = 2m$$

$$\Rightarrow a = \frac{F}{2m} \Rightarrow F_3 = \left(\frac{m}{4}\right)a = \frac{F}{8}$$

42. Ans (1)

$$w = \frac{1}{2} \times 6 \times 10 - 4 \times 5 + 4 \times 5 - 5 \times 2$$

$$= 30 - 20 + 20 - 10 = 20J$$

$w = k_f - k_i$

$20 = k_f - 25$

$k_f = 45J$

43. Ans (1)

For $\theta < 5°$ $\tan\theta \approx \theta \approx 2 = 2 \times \frac{\pi}{180}$

$$\tan\theta = \frac{v^2}{rg} = \frac{h}{b}$$

$$\frac{2 \times 3.14}{180} = \frac{h}{1800} = 62.8 \text{ mm}$$

44. Ans (4)

Angular momentum = momentum × perpendicular distance.

For motion along BC, perpendicular distance does not change.

46. Ans (3)

$\Delta G = (2 \times 60) - (2 \times 90 + 70)$

$= 120 - (180 + 70)$

$= 120 - 250 = -130$ kJ/mol

$\therefore \Delta G < 0$, so the reaction is spontaneous

47. Ans (3)

$$(OH^-) = \frac{(0.6 \times 100) - (0.5 \times 100)}{200}$$

$$[OH^-] = \frac{60 - 50}{200} = 0.05 \text{ M}$$

$pOH = \log (5 \times 10^{-2})$

$= 2 - \log 5 = 1.3$

$pH = 12.7$

48. Ans (3)

$T^o_b = 99.28^oC$

$T_b = ?, \Delta T_b = K_{b.m}$

$\Rightarrow T_b - 99.28 = 0.513 \times 0.7$

$\Rightarrow T_b = 99.28 + 0.3591$

$= 99.64^oC$

49. Ans (1)

If the experiment is set such that only the colloidal particles are allowed to move towards anode and net the DM it is known as anaphoresis.

50. Ans (1)

SC $2r = a$

fcc $4r = \sqrt{2}\, a$

bcc $4r = \sqrt{3}\, a$

51. Ans (2)

$P_1V_1 = nRT_1 \quad ; P_2V_2 = nRT_2$

$P_1 = P_2 \; ; V_1 = V_2$

$\Rightarrow n_1T_1 = n_2T_2$

$\Rightarrow 1 \times 300 = n_2 \times 500$

$\Rightarrow n_2 = 0.6$

$\Rightarrow$ 0.4 moles are withdrawn

52. Ans (4)

$$Mn^{+7} + X^{th} \rightarrow \overset{+5}{XO_3^-} + Mn^{+2}$$

eq of MnO_4^- = eq of X^{th}

$n_1 \times v.f_1 = n_2 \times v.f_2$

$\Rightarrow 6 \times 10^{-3} \times 5 = 9 \times 10^{-3}(5 - n)$

$$\Rightarrow \frac{10}{3} = 5 - n$$

$$\Rightarrow n = \frac{5}{3}$$

54. Ans (4)

Two rates are related as

$$\frac{-d[BrO_3^-]}{dt} = \frac{1}{3}\frac{d[Br_2]}{dt}$$

$$\frac{-d[BrO_3^-]}{dt} = \frac{1}{3}(0.025)$$

$\Rightarrow (0.0083)$

55. Ans (1)

$2H_2O_2(l) \rightarrow 2H_2O(l) + O_2(g)$

$\Delta H^{o}_{reaction} = \Delta H_f^{o}(products) - \Delta H_f^{o}(reactants)$

$-196 \Rightarrow 2\Delta H_f^{o}[H_2O(l)] + \Delta H_f^{o}(O_2) - 2\Delta H_f^{o}2H_2O_2(l)$

$-196 = 2x + 0 - 2(-188)$

$x = \Delta H_f^{o}[H_2O(l)] = -286\ kJ/mol.$

56. Ans (3)

$$\frac{\text{wt of Al deposited}}{\text{wt of Cu deposited}} \Rightarrow \frac{\text{eq. wt of Al}}{\text{eq. wt of Cu}}$$

$$\frac{W(Al)}{6.35} \Rightarrow \frac{27/3}{63.5/2}$$

$$W(Al) = \frac{9 \times 2}{63.5} \times 6.35$$

$= 1.8\ g$

57. Ans (4)

For isotonic solution

$(Na_2SO_4)\pi_1 = \pi_2(G)$

$i\ C_1\ RT = C_2.\ RT$

$(1 + 2\alpha)\ 0.008 = 0.01$

$$1 + 2\alpha = \frac{0.01}{0.008} = \frac{10}{8} = 1.25$$

$$\alpha = \frac{0.25}{2} = 0.125$$

$\% = 12.5\ \%$

58. Ans (1)

$$k_1 = \frac{[NH_3]^2}{[N_2][H_2]^3} \quad k_2 = \frac{[NO]^2}{[N_2][O_2]} \quad k_3 \frac{[H_2O]}{[H_2][O_2]^{1/2}}$$

$$2NH_3 + \frac{5}{2}O_2 \rightleftharpoons 2NO + 3H_2O$$

$$k_{eq} = \frac{[NO]^2[H_2O]^3}{[NH_3]^2[O_2]^{5/2}} \Rightarrow \frac{k_2 \times k_3^3}{k_1}$$

Required reaction

$$NH_3 + \frac{5}{4}O_2 \rightleftharpoons NO + \frac{3}{2}H_2O$$

$$k_{eq} = \frac{k_2^{1/2}\, k_3^{3/2}}{k_1^{1/2}}$$

59. Ans (3)

$P_1 = 620$ mm $\quad P_2 = 640$ mm

$V_1 = 300$ cc $\quad V_2 = 310$ cc

$T_1 = 27° + 273 = 300$ k $\quad T_2 = ?$

Moles are constant

$PV = nRT$

$$\frac{P_1V_1}{RT_1} = n = \frac{P_2V_2}{RT_2}$$

$$\frac{P_1V_1}{T_1} = \frac{P_2V_2}{T_2}$$

$$\frac{620 \times 300}{300} = \frac{640 \times 310}{T_2}$$

$$T_2 = \frac{640 \times 310}{620}$$

$= 320$ k

60. Ans (1)

For isothermal reversible compression

$\Delta E = 0,\ q + w$

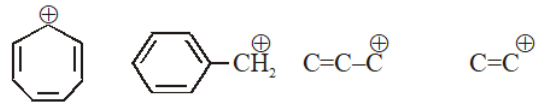
$w = -q$

$q = -208J$ (Heat released so –ve)

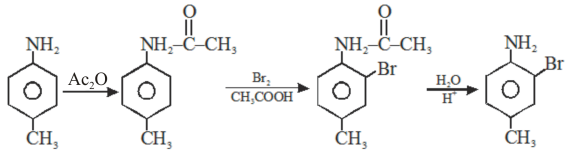
$w = +208J$

(Compression so work done on the system +ve)

61. Ans (2)



62. Ans (4)



63. Ans (3)

$CH_3–CH_2–CH_3$ $CH_3–CH_2–CH_2–CH_3$

Propane Butane

$CH_3–\underset{\displaystyle CH_3}{\underset{|}{CH}}–CH_3$ $CH_3–CH_3$

Iso butane Ethane

64. Ans (1)

SN^2 rate $\propto$ less streic crowding

65. Ans (3)

Find R & S

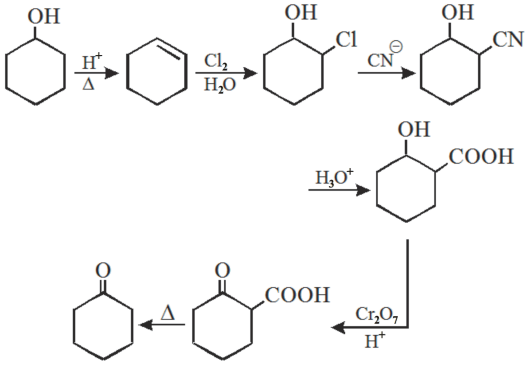
Enantiomer had opposite R, S.

66. Ans (3)

$CH_3–CH_2–CH_2–O–\overset{1}{CH_2}–\overset{2}{\underset{\displaystyle CH_3}{\underset{|}{CH}}}–\overset{3}{CH_3}$

1-propoxy-2-methyl propane

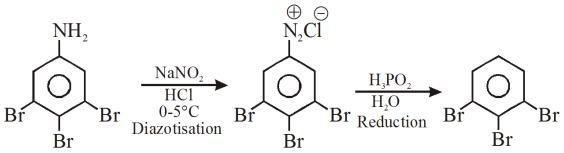
67. Ans (1)



68. Ans (1)

Soframicine is antiseptic.

69. Ans (1)

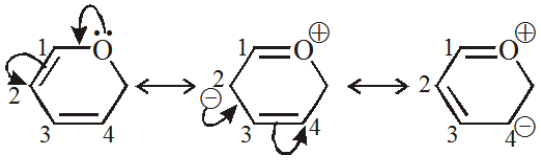


70. Ans (4)

3° Amine is most sterically crowded does not able to form H-bonding that's why lesat basic order of boiling point for isomeric anines is 1° > 2° > 3° due to H-bonding

71. Ans (2)

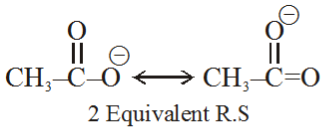
πe^- density is identified by resonating structure.



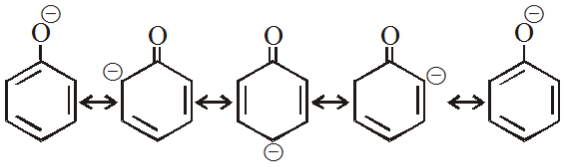
–ve charge is at 2 and 4 position. So C_2 and C_4 have highest electron density.

72. Ans (3)

Due to equivalent resonating structure formed by acetate ion it is more stable than phenoxide ion which form non equivalent resonating structure.



2 Equivalent R.S



Non equivalent R.S

73. Ans (4)

According to CIP priority is assigned according to atomic number of directly attached atom

$–OH > –COOH > –\underset{\displaystyle OH}{\underset{|}{CH}}–CH_3 > CH_2OH$

74. Ans (2)

ESR $\propto$ electron density of benzene

$\propto$ Resonance

$\propto$ Hyperconjugation

75. Ans (4)

Reactivity of alcohol with HBr $\propto$ stability of carbcation.

76. Ans (2)

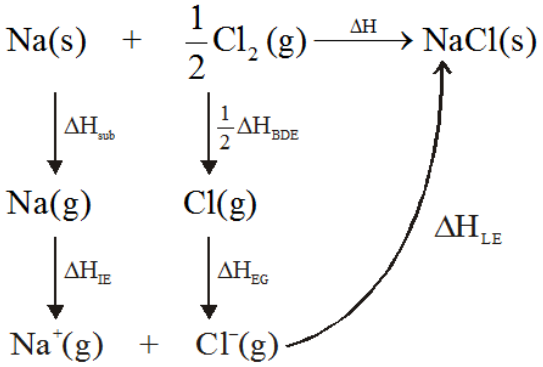
When CO_2 is passed through brine solution saturated with NH_3, it gives sodium bicarbonate which on drying and heating gives sodium carbonate.

77. Ans (4)

(1) $Zn(NH_3)_2Cl_2$ → Tetrahedral

(4) $[Cr(H_2O)_3Cl_3]$ → Ma_3b_3 shows 2 GI (fac-Mer)

79. Ans (4)



80. Ans (1)

$CaOCl_2 \longrightarrow CaCl_2 + [O]$

81. Ans (1)

$$I.E. \propto Z_{eff}$$

$O^+ > F^+ > N^+ > C^+$

$2s^22p^3 \quad 2s^22p^4 \quad 2s^22p^2 \quad 2s^22p^1$

82. Ans (4)

H_2Se and H_2S are polar, but molecular weight of H_2Se is higher so higher dispersion forces in H_2Se.

83. Ans (2)

$H_2SO_4 + PbO_2 \longrightarrow PbSO_4 + ½O_2 + H_2O$

84. Ans (2)

$K_3[Co(CO_3)_3] \Rightarrow Co^{3+}(d^6)$ (WFL) $\Rightarrow sp^3d^2$

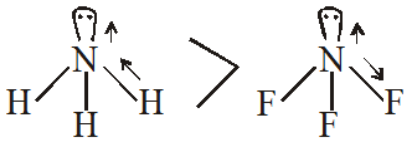
octahedral no. of unpaired electron = 4

$$\mu = \sqrt{n(n+2)} = \sqrt{4(4+2)} = 4.9 \text{ BM}$$

86. Ans (2)

(1) BF_3 and CCl_4 → Dipole moment = 0

(2) NH_3 and NF_3 → Non-zero dipole moment



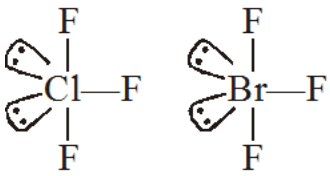
87. Ans (3)

Solubility : $MgSO_4 > CaSO_4 > SrSO_4 > BaSO_4$

Stability : $BeCO_3 < MgCO_3 < CaCO_3 < SrCO_3$

Bond energy – $Cl_2 > Br_2 > F_2 > I_2$

88. Ans (2)



89. Ans (4)

All d-block elements are transition elements except Zn, Cd and Hg due their full filled electronic configuration in ground state as well as stable oxidation state.

90. Ans (1)

CuO, Cu_2O and $CuSO_4$ gives Cu and SO_2 on heating with Cu_2S during Bessemerisation.

95. Ans (3)

NCERT Pg # 117

101. Ans (1)

NCERT (XII) Pg # 30(E), 32(H)

105. Ans (3)

NCERT (XI) Pg # 265

107. Ans (1)

NCERT (XI) Pg. # 272, Ist para

111. Ans (2)

NCERT XI, Page # 294

112. Ans (2)
NCERT (XIth) (E), Para-6, Pg. # 297

113. Ans (2)
NCERT Pg # 84, 85

115. Ans (3)
NCERT-XI Page No. 72

118. Ans (4)
NCERT XIIth, Pg. no # 147

123. Ans (4)
NCERT (E) Pg. # 322

129. Ans (2)
NCERT Pg. # 49

130. Ans (4)
NCERT- Pg # 56, Para- 4.2.11.2, Fig # 4.19 (a)

132. Ans (3)
NCERT XI, Pg. # 132, Para-04

140. Ans (2)
NCERT XII Pg.# 177 (E), 193 (H)

142. Ans (1)
NCERT Pg. # 6

145. Ans (2)
NCERT-XI, Pg # 31, Fig.3.1(a-i) & (b-ii)

147. Ans (1)
NCERT-XI Pg # 36

151. Ans (1)
NCERT-XI Pg. # 145

154. Ans (4)
NCERT Pg. # 133

160. Ans (3)
NCERT XI Pg # 101

164. Ans (2)
NCERT Pg. # 188

171. Ans (2)
NCERT-XIth Pg# 250,251

177. Ans (3)
NCERT Pg # 266

178. Ans (3)
NCERT- Pg. # 280, 282, 283