Experimental Methods	Measured Quantities	Applications	References	
Calorimetry • Microcalorimetrc	Heats of combustion, vaporization, and formation	Species in solid, solution, or gas-solid	(1) 1995LEB/CHI (2) 2004SIL/MAT	
Solution titration	surface Reaction heats		(3) 1992BUR (3) 1992HOF (4) 1995NOL	
Gas-solid surface	Femtomole adsorption heats on single-crystal surface		(5) 1998BRO/KOS (6) 1998CER	
Thermal desorption spectroscopy (TDS) Temperature programmed-desorption (TPD)	An adsorbed surface is heated and desorbing species are detected with MS, IR or others	Gas-solid interface	(1) 1986WOO/DEL (2) 1990JON/NIE	
Helium atom scattering (HAS)	Monoenergetic beam of helium atoms are scattered from ordered surfaces and detected	Gas-solid interface	1996HOF/TOE	

The readers may find more methods for determining the BDEs in e-publication (2002SIL).

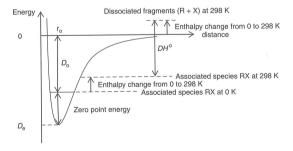
2.2 The conversion between D_0 , D_e , and DH^0

Related energies are often encountered when searching for experimental BDE data.

Do the bond dissociation energy at 0 K

 $D_{\rm e}^{}$ the electronic energy, i.e., the minimum on the potential energy curve or surface or the well depth.

 $D_0(R-X)$ may be determined from D_e , as shown in figure:



$$D_0 = D_e - \Delta(ZPE) \tag{2.1}$$

Here Δ (ZPE) is the change of the zero point vibrational energies in the species involved in the bond dissociation processes, the dissociating state as the energy reference.

For diatomic molecules and Morse potential, the approximate relationship follows:

$$D_0 \simeq \frac{\omega_e^2}{4\omega_e x_e} - \frac{\omega_e}{2} + \frac{\omega_e x_e}{4}$$

$$D_e \simeq \frac{\omega_e^2}{4\omega_e x_e}$$

$$\Delta(ZPE) = \frac{\omega_e}{2} - \frac{\omega_e x_e}{4}$$
(2.2)

Here ω_e and x_e are the spectroscopic parameters.

In general, the relationship between the dissociation enthalpy change and temperature must be determined to make the conversion from D_0 to DH^0 . There are four methods for this conversion.

- 1. Thermochemical calculations (2005RUS/BOG).
- Calculations by quantum chemistry (1997CUR/RAG; 2000CIO/SCH; 2002NIST/COM).
- 3. Using NIST-JANAF Thermochemical Tables (1998CHA).
- 4. Simple approximation. This is good for biatomic, triatomic, and tetraatomic species. This method is reviewed below.

2.2.1 Diatomic molecules

From thermodynamics, the enthalpy change, ΔH , and internal energy change, ΔE , for the bond dissociation processes is described as

$$\Delta H = \Delta E + \Delta (PV) \tag{2.3}$$

Where $\Delta(PV)$ is the molar work that is equal to $\Delta(RT)$, assuming 1 molar ideal gas. For diatomic species, a vibrational free degree becomes a translational one because of the bond dissociation. An approximate equation is obtained as

$$\Delta H = \Delta E + RT \approx D_0 + \frac{3}{2}RT \tag{2.4}$$

For T = 298.15 K, the BDE is derived based on the experimental D_0 at 0 K

$$BDE \equiv DH^{0} \approx D_{0} + \frac{3}{2}RT$$
 (2.5)
$$= D_{0} + 0.8886 \text{ kcal/mol} = D_{0} + 3.7181 \text{ kJ/mol}$$

Here the R is the gas constant; the H in the nomenclature DH^0 emphasizes that the BDE is the enthalpy change of the dissociation processes. The superscript '0' denotes the standard conditions of thermodynamics.

Table 4.5.3 (continued) C-C BDEs in Organic Compounds Containing Sulfur Atom(s)

The broken bonds (boldface =dissociated	recomme	BDEs (boldface= recommended data; reference in parentheses)		OVOR
group)	kcal/mol	kJ/mol	parentheses)	References
3-Methyl-1-butanethiol iso-C ₄ H ₉ -CH ₂ SH	80.5±2.5	336.8 ± 10.5	Derived from $\Delta_f H^0$ in ref.	1986PED/NAY
2,2-Dimethyl-1-propanethic $tert$ - C_4H_9 - CH_2 SH	ol 78.7±2.5	329.3±10.5	Derived from $\Delta_f H^0$ in ref.	1986PED/NAY
1-Hexanethiol nC₅H ₁₁ –CH ₂ SH	80.3±2.5	336.0±10.5	Derived from $\Delta_f H^0$ in ref.	1986PED/NAY
1-Heptanethiol nC ₆ H ₁₃ –CH ₂ SH	80.2±2.5	335.6 ± 10.5	Derived from $\Delta_f H^0$ in ref.	1986PED/NAY
Phenylmethane-1-thiol C_6H_5 – CH_2SH	93.2±2.5	389.9 ± 10.5	Derived from $\Delta_f H^0$ in ref.	1986PED/NAY
1,2-Ethanethiol HSCH ₂ -CH ₂ SH	74.9±2.8	313.4±11.7	Derived from $\Delta_f H^0$ in ref.	1986PED/NAY
Ethyl methyl sulphide CH ₃ -CH ₂ SCH ₃	82.0±2.0	343.1±8.4	Derived from $\Delta_f H^0$ in ref.	1986PED/NAY
Methyl propyl sulphide C_2H_5 – CH_2SCH_3	80.8±2.0	338.1±8.4	Derived from $\Delta_f H^0$ in ref.	1986PED/NAY
Ethyl phenyl sulphide CH ₃ -CH ₂ SC ₆ H ₅	80.9±2.0	338.5±8.4	Derived from $\Delta_{\rm f} H^0$ in ref.	1986PED/NA
Dibenzyl sulphide C ₆ H ₅ -CH ₂ SC ₆ H ₅	97.1±2.0	406.3±8.4	Derived from $\Delta_f H^0$ in ref.	1986PED/NA

4.6 C-C BDEs in radicals

Table 4.6 C-C BDEs in Radicals

The broken bonds (boldface=dissociated	recomme	ooldface = ended data; a parentheses)	Methods (reference in	
group)	kcal/mol	kJ/mol	parentheses)	References
C-C	(1) 142.36	595.63	(1) Derived from $\Delta_t H^0$ in ref.	(1) 1998CHA
Man Select (I)	(2) 147.8±3.7	618.3±15.4	(2) Spectroscopy	(2) 2000LU/TOS
C-C _x				
x=2	(1) 175.5	734	 Derived from Δ_tH⁰ in ref. 	(1) 1998CHA
	(2) 167.9	703	(2) Derived from $\Delta_t H^0$ in ref.	(2) 1994GIN/FIN
3	(1) 135.3 (2) 118.5	566 496	$\Delta_{\rm f}H^{\circ}$ in ref.	
	(3) 108.6 ± 3.5	454.4 ± 14.5	(3) PD	(3) 2000CHO/BIS
4	(1) 169.3	708		
(1) 2006(210)	(2) 164.4	688		
5	(2) 116.1 (2) 168.2	486 704		
	(2) 168.2	704		
C ⁺ -C (also see Chapter 24)	143.9±4.6	601.9±19.3	GIB MS	1999LU/TOS
C ⁻ -C (also see Chapter 24)	188.7±0.2	789.6±1.0	Derived from EA in ref.	1996NIST
C ₂ -C ₂	(1) 139.1	582	Derived from $\Delta_t H^0$ in ref.	(1) 1994GIN/FIN
	(2) 168.4	705		(2) 1998CHA
C ₂ -C ₃	(1) 135.5	567	(1) Derived from $\Delta_t H^0$ in ref.	(1) 1994GIN/FIN
	(2) 162.2	679	(2) Derived from $\Delta_t H^0$ in ref.	(2) 1998CHA
	(3) 137.4 ± 3.5	575.0 ± 14.5	(3) PD	(3) 2000CHO/BIS
				(continued

1998GER/GRA

	, and the Energies
1997YAN/LIN 1997YEO/VAT 1997YEO/VAT(b) 1997ZAN/TAY 1997ZHA/BAR 1997ZHA/BOR 1997ZHA/CHE 1997ZHA/CHE 1997ZHA/LIN 1997ZHA/LIN 1997ZHU/ZHA 1997ZOL/BLA	Yang, Y., Linnert, H. V. et al., J. Phys. Chem. A, 101, 2371–2378, 1997. Yeo, Y. Y., Vattuone, L., and King, D. A., J. Chem. Phys., 106, 1990–1996, 1997. Yeo, Y. Y., Vattuone, L., and King, D. A., J. Chem. Phys., 106, 392–401, 1997. Zanni, M. T., Taylor, T. R. et al., J. Chem. Phys., 107, 7613–7619, 1997. Zhang, C. M., Bartelt, M. C. et al., J. Crys. Growth, 174, 851–857, 1997. Zhao, Y., Bordwell, F. D. et al., J. Am. Chem. Soc., 119, 9125–9129, 1997. Zhao, H. Q., Cheung, Y. S. et al., J. Chem. Phys., 107, 7230–7241, 1997. Zhao, H. Q., Cheung, Y. S. et al., J. Chem. Phys., 106, 86–93, 1997. Zhao, R., Lind, J. et al., J. Chem. Soc. Perkin Trans. 2, 569–574, 1997. Zhu, Q., Zhang, X. M., and Fry, A. J., Polymer Degrad Stability, 57, 43–50, 1997. Zolfaghari, A., Blais, S., and Jerkiewicz, G., Proceedings-Electrochem. Soc., 97–16, Electrochemical Surface Science of Hydrogen Adsorption and Absorption.
1997ZOL/CHA	Zolfaghri, A., Chayer, M., and Jerkiewicz, G., J. Electrochem. Soc. 144
1998ALI/KLO 1998ALI/WAL 1998ALL/VER	3034–3041, 1997. Ali, T., Klotzer, B. et al., <i>J. Chem. Phys.</i> , 109, 10996–11009, 1998. Ali, T., Walker, A. V. et al., <i>Surf. Sci.</i> , 414, 304–314, 1998. Allouche, A., Verlaque, P., and Pourcin, J., <i>J. Phys. Chem. B</i> , 102, 89–98, 1998.
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2000DAH/TOR	Dahl, S., Tornqvis, E., and Chorkendorff, I., J. Catal., 192, 381–390, 2000.	2000HOP/SAU	Hop, C. E. C. A., Saulys, D. A. et al., J. Mass Spectrom., 35, 1003–1010, 2000.
2000DAV/GOO	Davis, K. K. and Goodman, D. W., J. Phys. Chem. B, 104, 8557–8562, 2000.	2000HOV/KOL	Hovel, S., Kolczewski, C. et al., J. Chem. Phys., 112, 3906–3816, 2000.
2000DEN/DEN	Denisov, E. T. and Denisova, T. G., Handbook of Antioxidants, CRC Press, Boca	2000HUS/AGU	Husband, L. Aguirre, F. et al., I. Phys. Chem. A, 104, 2020–2024, 2000.
	Raton, 2000.	2000ING/BUS	Ingolfsson, O., Busolt, U., and Sugawara, K., J. Chem. Phys., 112, 4613–4620,
2000DEN/KHU	Denisov, E. T. and Khursan, S. L., Russ. J. Phys. Chem., 74, Suppl. 3, S491–	2000ING/ DOS	2000.
=0002211/1410	S497, 2000.	2000JAR/MIL	Jarek, R. L., Miles, T. D. et al., J. Phys. Chem. A, 104, 2230–2237, 2000
2000DIS/BOE	Distelrath, V. and Boesl, U., Faraday Discuss., 115, 161–174, 2000.	2000JAK/WILL 2000KIM	Kimura, K., In Photoionization and Photodetachment, Part 1, Ng, C. Y. Ed.,
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2000EIC/ZIM	Eicher, B., Zimmermann, H. P., and Gaggeler, H. W., J. Phys. Chem. A, 104,	2000KOD/PET	Kodambaka, S., Petrova, V. et al., Surf. Rev. Lett., 7, 589–593, 2000.
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2000EVA/GER	Evans, C. J. and Gerry, M. C. L., J. Mol. Spectrosc., 203, 105–117, 2000.	2000KOF/KIE	56, 1613–1626, 2000.
2000EVII/GER 2000FAN/CHE	Fang, L., Chen, X. et al., Low Tem. Phys., 26, 752–755, 2000.	2000KOS/MIR	Kostikova, L. M., Miroshnichemko, E. A., and Matyushin, Y. N., Inter Ann
2000FAN/CHE(b)	Fang, L., Chen, X. et al., J. Phys. Chem. A, 104, 9153–9155, 2000.	2000KO3/WIIK	Conf. of ICT, Karlsruhe, Germany, 31st. pp. 50/1–50/11, 2000.
2000FAN/CHE(c)	Fang, L., Chen, X. et al., J. Chem. Phys., 113, 10202–10206, 2000.	2000KRA/CIR	Kranenburg, M., Ciriano, M. V. et al., J. Phys. Chem. A, 104, 915–921, 2000.
2000FAR/BUR	Farrell, G. L. and Burkey, T. J., J. Photochem, Photobiol. A Chem., 137, 135–139,	2000KRA/CIK 2000KRE/SCH	Kretzschmar, I., Schroder, D. et al., J. Phys. Chem. A, 104, 5046–5058, 2000.
	2000.	2000LAF/SZA	Lafleur, R. D., Szatary, B., and Baer, T., J. Phys. Chem. A, 104, 1450–1455, 2000.
2000FER/ALC	Fernandez-Morata, F., Alcami, M. et al., J. Phys. Chem. A, 104, 8075–8080,	2000LAT/SAT	Latini, A., Satta, M. et al., Chem. Eur. J., 6, 1042–10949, 2000.
	2000.	2000LEE/AHN	Lee, J. G., Ahner, J. et al., J. Chem. Phys., 112, 3351–3357, 2000.
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