

LEIPZIG UNIVERSITY  
Faculty of Physics and Earth Sciences

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## Master's Thesis

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# Filtered Cathodic Arc Plasma Processes for TiAl and TiAlN

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## Abstract

Titanium aluminum nitride (Ti–Al–N) coatings are widely used in cutting tools and wear-resistant applications, but their industrial deposition typically requires substrate heating above 400°C. Cathodic arc deposition offers room-temperature crystallization through energetic ion bombardment, but multiple process parameters are inherently coupled: external magnetic fields simultaneously increase both ion charge states (potential energy) and ion flux, while reactive gases alter plasma composition through charge-exchange collisions. This thesis systematically investigates the individual roles of ion flux, ion potential energy, and ion kinetic energy in room-temperature TiAlN film growth by varying magnetic field strength (0–0.25 T), nitrogen pressure (0–0.3 Pa), and spatial position (10–20 cm from the plasma source).

Three diagnostics were used to separate flux and energy effects: a Langmuir probe for ion current density, a quartz crystal microbalance for deposited mass, and an energy-resolving mass spectrometer for charge-state-resolved ion energy distributions. Measurements in metallic mode (vacuum) and reactive mode (nitrogen background) characterized the transition between deposition regimes and quantified the contribution of ionized versus neutral nitrogen species to film growth.

The results show that increasing the magnetic field strength can amplify ion flux by up to a factor of eight. In reactive mode, the presence of nitrogen further boosts ion flux, as charge-exchange collisions reduce the number of highly ionized species and increase the proportion of singly ionized species. Ion energy distribution functions measured by energy-resolved mass spectrometry, shift toward lower ionization levels when operating in reactive mode. The total ion energy (kinetic and potential energy combined) remains within the 30–60 eV range (down from 50–70 eV), which is beneficial for room-temperature crystalline film deposition.

This work extends the energy-flux decoupling framework established for (V,Al)N coatings to the industrially relevant TiAlN system and demonstrates that ion flux control via magnetic field strength provides a viable route to tailoring film microstructure without the need of external substrate temperature.

# CHAPTER 1

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## Introduction and Literature Review

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### 1.1 Motivation

Titanium aluminum nitride ( $\text{Ti}_{1-x}\text{Al}_x\text{N}$ ) coatings are widely used, because of their high hardness (25–35 GPa), thermal stability, and oxidation resistance [1]. These properties depend on maintaining the metastable cubic B1 crystal structure, which provides superior mechanical performance compared to the wurtzite phase [2]. Traditional deposition methods require substantial substrate heating to achieve dense, crystalline coatings. Cathodic arc deposition offers a key advantage: highly ionized metal plasmas with intrinsic high ion energies enable room-temperature crystallization through energetic condensation [3].

The challenge is that cathodic arc processes depend on the interplay of several interdependent parameters. Ion energy has two components: kinetic energy from plasma expansion and potential energy released upon neutralization. The ion flux determines how rapidly this energy is delivered to the growing film. External magnetic fields can enhance ion charge states and ion flux, the latter increasing by up to an order of magnitude [4], while adding reactive gases such as nitrogen introduces charge-exchange collisions that alter plasma composition [3]. Understanding how these parameters individually influence film properties remains a fundamental challenge for predictive process control.

## 1.2 State of the art

Recent systematic studies have made progress in understanding ion energy effects in cathodic arc deposition. Unutulmazsoy et al. showed that applying an external magnetic field at the plasma source increases ion charge states (and thus potential energy), while applying a substrate bias adjusts kinetic energy [5]. Their work on (V,Al)N films demonstrated that crystalline films can be deposited at room temperature through energetic ion bombardment. Unutulmazsoy et al. observed that “application of an external magnetic field also leads to an enhancement of the ion flux and hence the desired complete decoupling of the potential and kinetic energy effects will require further steps” [5]. In their study, the magnetic field not only modified ion charge states but also increased ion flux by up to a factor of 10, making it impossible to isolate the effect of potential energy from that of ion flux intensity.

Kalanov et al. refined this approach using detailed energy-resolved mass spectrometry and systematic distance variation [6]. Their work on (V,Al)N films showed that the enhancement of room-temperature crystallinity correlates primarily with the increase in potential energy input. Notably, applying a DC bias to increase kinetic energy by approximately 30 eV did not produce the same crystallization effect, even though it increased the total ion energy. This demonstrated that potential energy plays a distinct role in promoting film crystallinity. Crucially, Kalanov et al. addressed the flux-energy coupling by varying the source-to-substrate distance from 8 to 20 cm: moving the substrate further away reduced the ion flux while maintaining similar ion charge states and energies, enabling direct comparison of potential energy effects at matched flux levels.

Despite these advances, key questions remain for extending this framework to other material systems. Both studies explored the effects of activated nitrogen in reactive deposition, which consists of both ionized and neutral (but excited) nitrogen species [3, 7]. While Unutulmazsoy et al. [5] focused on reactive deposition with nitrogen and Kalanov et al. [6] examined both metallic and reactive modes, a key question remains unanswered: what fraction of the activated nitrogen consists of neutral species versus ionized species? This distinction is crucial for understanding the relative contributions of ions and neutrals to film growth, particularly in reactive mode where both species can participate in nitride formation.

For TiAlN specifically, which benefits from greater natural abundance and lower cost than vanadium-based nitrides, while comparative studies between cathodic arc and magnetron sputtering have shown that ion energy flux significantly affects texture

and mechanical properties [8], the flux-energy decoupling framework established for (V,Al)N has not been applied to this industrially important system.

### 1.3 Objectives

This thesis addresses above literature gaps by systematically investigating the role of ion flux in Ti–Al–N thin film growth via pulsed filtered cathodic arc deposition. The specific objectives are:

1. Characterize the ion flux and the atomic flux as functions of external magnetic field strength and distance from the macroparticle filter. Furthermore, analyze the ion energy distributions of individual ion species.
2. Investigate the sweet spot from metallic to reactive mode by varying nitrogen pressure and quantify how this affects the ion-to-neutral ratio arriving to the substrate.
3. Establish quantitative correlations between ion flux and atomic flux across the experimental parameter space of magnetic field, nitrogen pressure, and distance from the macroparticle filter.
4. Characterize the crystallinity, film thickness, film composition of deposited TiAlN films to correlate plasma process parameters with film properties.

The experimental approach combines multiple *in situ* plasma diagnostics with *ex situ* film characterization to extend the Unutulmazsoy/Kalanov framework to both reactive mode operation and the industrially relevant TiAlN system. This multi-diagnostic strategy, applied across a systematic parameter space, enables investigation of two key novel aspects: (1) the application to the TiAlN material system, and (2) the combined measurement of ion flux versus total deposited flux to distinguish the contribution of non-ionized activated nitrogen species and correlate this with film properties.

# CHAPTER 2

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## Theoretical Background

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### 2.1 Plasma generation and composition

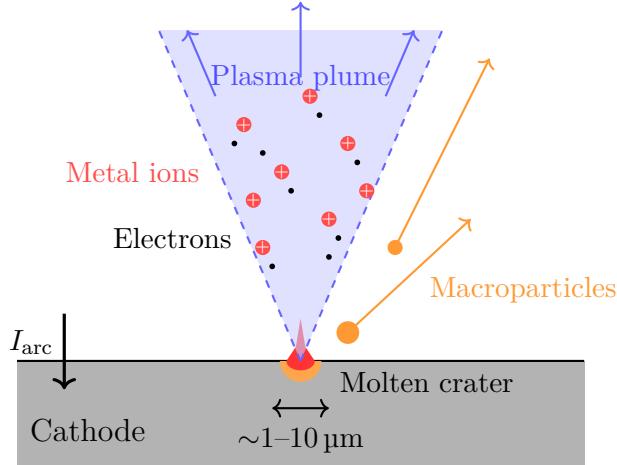
#### 2.1.1 Cathode spot plasma generation

Cathodic arc plasmas form at microscopic emission centers, known as cathode spots, on an otherwise cold metal electrode under vacuum. Spot ignition occurs when the local cathode surface, through breakdown of adsorbates or field-enhanced thermionic emission, undergoes explosive release of electrons and vaporized metal. During a single spot pulse, a few nanograms of the cathode material rapidly heat up, vaporize and ionize producing a dense, quasineutral plasma plume composed mostly of metal ions and electrons. The peak spot current densities reach  $10^{10}$ – $10^{12} \text{ A m}^{-2}$ , far above steady-state thermionic or field emission limits. These microexplosions, termed ectons (explosive electron emission centers), were first described by Mesyats [9] and produce localized nanosecond-scale plasma bursts. The arc plasma is sustained by repetitive ecton events occurring at or near the same location [3, Chap. 3.3–3.4].

The spot-generated plasma exhibits two defining characteristics. First, the extreme power density at the cathode spot produces a high degree of ionization, with more than 90% of the ejected metal atoms emerging as ions [3, Chap. 3.5]. Second, the high electron temperature and density in the spot plasma generate multiply charged ions, with charge state distributions extending to  $Q = 3$ – $4$  for refractory metals such as Ti and Al [3, Chap. 3.5].

Spot ignition and quenching occur on timescales of 10–100 ns, with each pulse eject-

ing a fully ionized burst of metal vapour. The sustained arc discharge thus consists of continuously overlapping microplasma pulses, producing a metal-rich, high-flux ion stream well-suited for energetic thin-film deposition.



**Figure 2.1.** Schematic of cathode spot operation. The arc current  $I_{\text{arc}}$  concentrates at a microscopic spot ( $1-10 \mu\text{m}$ ), creating a molten crater from which a plasma plume of metal ions, electrons and macroparticles expand. Adapted from [10]

### 2.1.2 Pulsed vs. continuous arc operation

Cathodic arcs can operate in either continuous (DC) or pulsed mode, with fundamental differences in plasma generation dynamics. In DC operation the cathode spot moves continuously across the surface, maintaining a steady-state plasma density determined by the balance between plasma generation at the spot and losses through expansion. The time-averaged plasma properties remain constant, and the ion flux to the substrate is continuous.

In pulsed operation, the arc is periodically initiated and extinguished, creating discrete plasma bursts separated by periods with no plasma generation. During the active phase of each pulse, the instantaneous plasma density can be significantly higher than in DC arcs operating at the same average power, because the energy is concentrated in short time intervals. The peak plasma density scales with the instantaneous arc current. Between pulses the plasma expands and dissipates, allowing the cathode surface to cool. This temporal modulation affects both the spot dynamics and the resulting plasma composition.

The ion charge state distributions in pulsed arcs are typically similar to or slightly enhanced compared to DC arcs, as the higher instantaneous power density can promote additional ionization events in the cathode spot region [3, Chap. 10].

### 2.1.3 Plasma expansion and macroparticle filtering

Following their generation at the cathode spots, the plasma bursts expand into the vacuum chamber. This expansion is supersonic, with ions carrying directed kinetic energy away from the cathode. In many industrial and research systems, the expanding plasma is guided through a magnetic filter that removes macroparticles while allowing plasma to pass along curved magnetic field lines.

In the region near the cathode (within a few centimetres of the spot), plasma densities are on the order of  $10^{18} \text{ cm}^{-3}$  and electron temperatures  $T_e \approx 5\text{--}10 \text{ eV}$ . As the plume propagates, its density decreases according to

$$n(r) = \frac{C I_{\text{arc}}}{r^2} \quad (2.1)$$

where  $I_{\text{arc}}$  is the arc current,  $r$  the distance, and  $C$  a constant related to the ion erosion rate of the cathode material. This  $1/r^2$  scaling assumes free expansion, but deviations can occur due to magnetic fields or reactive gases, which may alter the plasma trajectory or cause recombination [3, Chap. 4.3, Eq. 4.3, p. 178].

In cathodic arc discharges from titanium cathodes, whether pure Ti or Ti-Al compounds, ions generally carry an average charge state  $\langle Q \rangle \approx 2.1\text{--}2.2$  at the source [3, Chap. 4.1, App. B.8]. This high degree of ionization reflects the extreme power density of the spot and follows the cohesive energy rule, which links  $\langle Q \rangle$  to the cohesive energy of the cathode material [3, App. B.8].

In the present work, a  $90^\circ$  curved magnetic filter guides the expanding plasma toward the substrate region while removing macroparticles. After passing through the filter, the plasma has evolved from its initial state at the cathode spot, having undergone expansion, potential collisions with background gas and interaction with guiding magnetic fields. The properties of this filtered plasma in the substrate region determine the energy and flux delivered to the growing film, and are the focus of the following sections.

## 2.2 Ion energies and flux in the substrate region

This section focuses on the properties of ions in the plasma after expansion and filtering, in the region where they reach the substrate and form the growing film.

### 2.2.1 Ion energies: origins and implications

Ions in cathodic arc plasmas carry both kinetic and potential energy. The kinetic energy  $E_{\text{kin}}$  arises from the supersonic expansion of plasma from the cathode spot, while the potential energy  $E_{\text{pot}}$  is released upon neutralization at the substrate surface and is determined by the ionization states of the ion.

The total energy delivered by an ion to the growing film is

$$E_{\text{tot}} = E_{\text{kin}} + E_{\text{pot}}. \quad (2.2)$$

For cathodic arc plasmas, the kinetic energy is closely linked to the arc burning voltage, which remains nearly constant at 30–35 V [3, Chap. 4.2]. This voltage accelerates ions away from the cathode region, giving them characteristic drift velocities. As ions traverse the expanding plasma, they may undergo collisions that modify their energy distribution.

Table 2.1 summarizes characteristic ion properties for Ti and Al cathodic arc plasmas in vacuum, measured near the cathode.

**Table 2.1.** Characteristic ion properties for Ti and Al cathodic arc plasmas in vacuum, near the cathode [3, App. B, Table B.8].

Species	$\langle Q \rangle$	$E_{\text{kin}}$ (eV)	$E_{\text{pot}}$ (eV)	$E_{\text{tot}}$ (eV)
Ti <sup>2+</sup>	2.1	59	21	80
Al <sup>2+</sup>	1.7	28	24	52

These values, measured near the cathode spot, serve as reference for understanding the energy budget of ions. In this work, the plasma is characterized after expansion and filtering, where ion energies may differ from these initial values. Nonetheless for the materials used in this work, the total ion energies are expected to exceed the approximately 30 eV threshold for subplantation, enabling densification and improved crystallinity in Ti–Al–N films without requiring external substrate heating [11].

## 2.2.2 Effect of distance on ion properties

As ions travel from the cathode through the filter and toward the substrate, their properties evolve due to geometric expansion and potential collisions. The ion flux decreases with the square of the distance due to the expanding plasma front as seen in Eq. 2.1.

In vacuum (metallic mode), the ion energy distributions remain relatively narrow and well-defined. At increased distances, the plasma density decreases but the relative composition and charge states are largely preserved, therefore the ion flux at the substrate position becomes an important parameter, as it determines the rate at which energy is delivered to the growing film. Understanding how distance affects the ion flux therefore provides insight into the spatial uniformity of the deposition process and allows optimization of substrate positioning.

## 2.2.3 Effect of external magnetic fields

Applying an external axial magnetic field at the arc source modifies the plasma properties in several ways. The magnetic field confines electrons near the cathode spot through their gyration around field lines, while the heavier ions are less affected and can still reach the anode. This "magnetic insulation" effect reduces electron current to the anode, requiring a higher voltage to maintain the arc current. The higher voltage accelerates ions more strongly and increases the plasma temperature in the cathode spot region, prolonging the interaction time between electrons and ions [4]. These changes lead to:

- Increased average ion charge state  $\langle Q \rangle$ , which increases the ion potential energy,
- Simultaneously increased ion kinetic energy,
- Enhanced ion flux, which can increase by up to an order of magnitude [4].

These effects are inherently coupled: the external magnetic field simultaneously increases both the ion charge states (and thus potential energy) and the total ion flux. This coupling presents a challenge for isolating the individual contributions of ion energy and ion flux to film growth. Decoupling these parameters requires additional experimental approaches, such as varying the source-to-substrate distance to modulate flux while maintaining similar ion energies, or applying bias voltages to shift ion kinetic energies independently [5, 6].

## 2.2.4 Ion flux

The ion flux  $\Gamma$  represents the number of ions arriving per unit area per unit time, expressed in ions  $\text{cm}^{-2} \text{s}^{-1}$ . In a cathodic arc plasma, the total ion current density  $J_i$  ( $\text{A cm}^{-2}$ ) relates to  $\Gamma$  via

$$\Gamma = \frac{J_i}{e \langle Q \rangle}, \quad (2.3)$$

where  $e$  is the elementary charge and  $\langle Q \rangle$  the average ion charge state.

In vacuum cathodic arcs, the burning voltage remains nearly constant at 30–35 V for arc currents, so the plasma generation rate and thus  $\Gamma$  increases approximately linearly with  $I_{\text{arc}}$  [3, Chap. 6.5]. The absolute ion flux at the substrate depends not only on the arc current but also on the distance from the source and the presence of magnetic fields or reactive gases. These factors determine the fraction of generated plasma that reaches the substrate and the composition of that plasma. The transition from metallic to reactive operation introduces additional complexity, as discussed in the following section.

## 2.3 Reactive mode and nitrogen activation

### 2.3.1 Transition from metallic to reactive mode

Cathodic arc deposition operates in two distinct regimes. In metallic mode, the cathode surface remains metallic and the plasma consists exclusively of metal ions, characterized by high ionization degrees. In reactive mode, a background gas such as  $\text{N}_2$  adsorbs onto the cathode surface, forming a compound layer that poisons the cathode and alters both spot behaviour and plasma composition [3, Chap. 9.2].

When  $\text{N}_2$  is introduced, a dynamic equilibrium develops between compound formation (through adsorption and reaction at the cathode surface) and compound removal (via explosive ejection events that eject both metal and nitrogen containing particles) [3, Chap. 9.3]. The equilibrium position depends on gas pressure, arc current, and cathode composition. At low  $\text{N}_2$  pressures or high power densities, type-2 (metal-rich) spots prevail maintaining predominantly metal ion flux. At higher pressures, type-1 (poisoned) spots dominate producing a mixed plasma of metal and nitrogen ions [3, Chap. 9.4]. This transition affects not only the chemical composition of the deposited film but also the energy distribution and charge state distribution of the plasma, as compound formation at the cathode alters the electron emission and plasma generation mechanisms.

### 2.3.2 Activated nitrogen species

In reactive mode, the plasma contains not only metal ions but also activated nitrogen species. These include:

- Ionized nitrogen:  $\text{N}^+$  and  $\text{N}_2^+$  ions formed by electron-impact ionization,
- Neutral but excited nitrogen: metastable  $\text{N}_2$  and atomic N species that carry internal energy but no net charge.

The term “activated nitrogen” encompasses both ionic and neutral excited species that participate in film growth. While ionized species can be detected directly by mass spectrometry, the contribution of neutral activated species is more difficult to quantify. This distinction is important because the total deposited flux includes both ionic and neutral components, whereas ion current measurements detect only the charged fraction.

Charge exchange with  $\text{N}_2$  reduces the average charge state of metal ions and introduces gas-ion species, altering the potential energy delivered to the film [3, Chap. 9.4]. Collisions during plasma expansion also reduce ion drift velocities, lowering kinetic energy before substrate impact. These effects collectively modify the energy budget available for film growth in reactive mode compared to metallic mode. The interplay between metal ion flux, activated nitrogen flux, and their respective energies determines the resulting film composition, structure, and properties, as discussed in the following section.

## 2.4 Plasma–surface interactions and film growth

### 2.4.1 Energetic condensation and subplantation

When metal ions with sufficient energy strike the growing film, they penetrate below the surface and deposit energy through a shallow collision cascade. This subplantation process produces two key effects:

- **Localized densification:** Ions with energies above approximately 30 eV implant beneath the surface, occupying interstitial sites and displacing near-surface atoms through knock-on collisions. This reduces porosity and increases film density, which is particularly important for transition-metal nitride coatings such as Ti–Al–N [11, 3].
- **Atomic-scale heating:** The deposition of kinetic energy and release of potential energy (ionization enthalpy) generate localized, nanosecond-scale tempera-

ture spikes. These enhance adatom mobility and promote crystallite coalescence without requiring global substrate heating [3, Chap. 8.2].

As the energetic input from ions increases, films transition from porous, amorphous structures to dense, crystalline coatings. This densification introduces compressive stresses of several GPa through atomic peening [3, Chap. 8.1–8.4]. For example, TiN films grown with total ion energies of approximately 60 eV develop a preferred cubic (111) texture and hardness exceeding 30 GPa.

The relationship between ion flux and film growth rate is dependent on

$$R = \frac{m_{\text{ion}} \Gamma S}{\rho_{\text{film}}}, \quad (2.4)$$

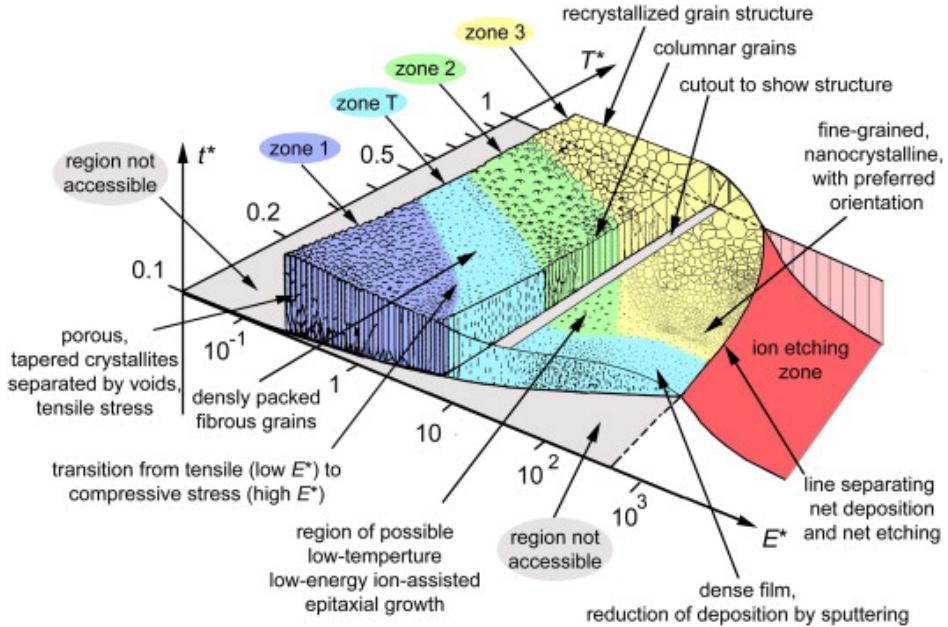
where  $m_{\text{ion}}$  is the average ion mass,  $\Gamma$  the ion flux,  $S$  the sticking coefficient, and  $\rho_{\text{film}}$  the film density. The sticking coefficient  $S$  represents the probability that an arriving ion incorporates into the growing film rather than being reflected or resputtered; for metal ions at moderate energies  $S \approx 1$  (below the resputter threshold of approximately 100–200 eV). In contrast the film density  $\rho_{\text{film}}$  itself depends on the ion energy and flux, as higher energies promote densification through subplantation. This interdependence between flux, energy, and resulting film structure motivates the systematic study of these parameters, which is the focus of this work.

#### 2.4.2 Structure-zone models

The microstructure of thin films deposited by physical vapour deposition depends strongly on the energy and flux of incident species. Thornton's structure-zone model was originally developed for magnetron sputtering and relates film morphology to the homologous temperature  $T/T_m$  (substrate temperature normalized to the melting point) and working gas pressure [12]. At low  $T/T_m$  and high pressures, films exhibit porous, columnar structures (Zone 1) due to limited adatom mobility. As  $T/T_m$  increases denser columnar (Zone T) and eventually equiaxed crystalline structures (Zone 2 and Zone 3) develop.

Anders extended this framework to account for the energetic ion bombardment characteristic of cathodic arc deposition [13]. In the revised model, ion energy  $E^*$  (normalized to a displacement energy) replaces gas pressure as the second axis, reflecting the dominant role of ion bombardment in densification. High-energy ions can induce subplantation and atomic peening even at low substrate temperatures, enabling dense and crystalline films without external heating, which is a key advantage of cathodic arc

processes. However, excessive ion energy leads to lattice damage, defect accumulation and eventually amorphization or resputtering, defining an optimal energy window for film growth [3, Chap. 8.3].



**Figure 2.2.** Structure-zone diagram for plasma based thin film deposition, showing film microstructure as a function of generalized temperature  $T^*$  and normalized ion energy  $E^*$ . From Anders [13].

### 2.4.3 TiAlN crystal structures

Titanium aluminium nitride ( $\text{Ti}_{1-x}\text{Al}_x\text{N}$ ) coatings are widely used for wear protection and cutting tools due to their high hardness, oxidation resistance, and thermal stability. The crystal structure depends primarily on the aluminium content  $x$ :

- For  $x \lesssim 0.6\text{--}0.7$ ,  $\text{Ti}_x\text{Al}_{1-x}\text{N}$  crystallizes in the metastable cubic B1 structure, where Al atoms substitute for Ti on the metal sublattice. This cubic phase exhibits hardness values of 25–35 GPa and is the preferred structure for most industrial applications [1].
- For  $x \gtrsim 0.7$ , the stable wurtzite (B4) structure becomes dominant. The wurtzite phase has lower hardness (typically 15–20 GPa) and is generally undesirable for hard coating applications [2].
- At intermediate compositions, mixed cubic-wurtzite structures or nanocomposite arrangements may form depending on deposition conditions.

The metastable cubic phase is retained at high Ti contents through kinetic limitations during low-temperature deposition. Energetic ion bombardment in cathodic arc

processes can extend the solubility limit of Ti in the cubic phase by providing additional energy for atomic rearrangement without the diffusion lengths associated with thermal equilibration [14].

The cathode composition used in this work (75 wt.% Ti – 25 wt.% Al) is expected to produce cubic-phase  $\text{Ti}_x\text{Al}_{1-x}\text{N}$  films under typical cathodic arc conditions.

# CHAPTER 3

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## Experimental Methodology

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This chapter describes the experimental apparatus, diagnostic techniques, and data processing methods used to investigate the correlation between ion flux and mass deposition rate in pulsed filtered cathodic arc deposition of Ti-Al-N thin films. The approach combines three complementary in-situ diagnostics: a Langmuir probe for ion current measurements, a quartz crystal microbalance (QCM) for real-time mass deposition monitoring, and an energy-resolving mass spectrometer (ERMS) for charge-state-resolved ion energy distributions.

### 3.1 Experimental setup

#### 3.1.1 Vacuum & gas control

The vacuum chamber was evacuated using a two-stage pumping system consisting of a dry rotor vacuum pump (Leybold ECODRY+) for roughing and a cryogenic pump (Leybold COOLVAC) for the main pumping line, achieving a base pressure on the order of  $1 \times 10^{-5}$  Pa with the cryopump fully open. To control the working pressure during reactive deposition, the cryopump gate valve was partially closed to position 345 (approximately 34.5% open), which reduced the effective pumping speed and increased the base pressure to approximately  $1 \times 10^{-4}$  Pa. This configuration allowed for the nitrogen pressures.

Nitrogen gas ( $\text{N}_2$ , 99.999% purity) was introduced via a mass flow controller (MKS IO Type B), with chamber pressures monitored using a Pfeiffer PKR 361. The nitrogen flow rate of 1 to 12.8 sccm (standard cubic centimeters per minute) leads to nitrogen

partial pressures between 0.025 and 0.3 Pa.

Plasma was generated using a water-cooled cylindrical anode and a rod cathode with composition 75 wt.% Ti; 25 wt.% Al (corresponding to 62.8 at.% Ti; 37.2 at.% Al). The cathode had a diameter of 6.35 mm and length of 38.1 mm. The expanding plasma was guided through a 90° curved magnetic macroparticle filter connected in series with the arc source.

### 3.1.2 Power circuits and arc operation

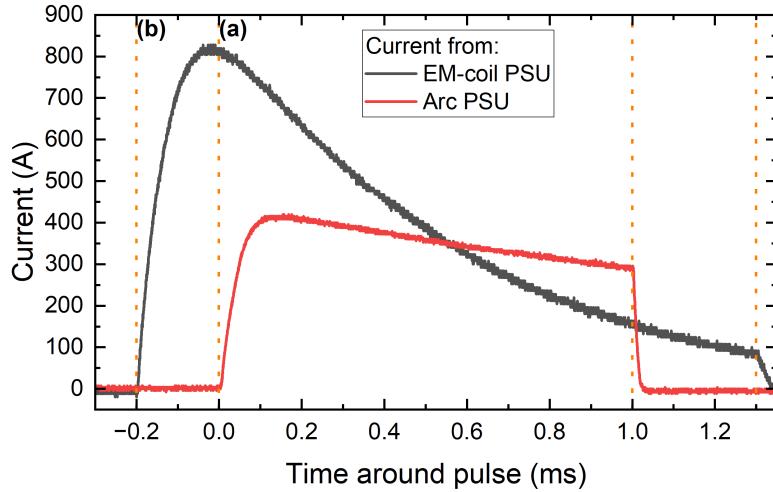
The arc power supply (Melec SPIK3000A) operated in pulsed DC mode, delivering arc currents up to 450 A at pulse frequencies ranging from 0.2 to 5 Hz, with a typical pulse width of 1 ms. The arc current was monitored using a Chauvin Arnoux MA200 connected to a Tektronix MSO64 oscilloscope. The cathode voltage relative to ground and anode voltage relative to ground were measured on additional channels to determine the burning voltage during the pulse.

A coil wound around the cylindrical anode (EM-coil) was used to enhance the ion charge states leaving the source through magnetic confinement (see Section 2.2.3). This coil was connected to a separate pulsing unit (Melec SPIK3000A) capable of delivering currents up to 850 A. The EM-coil was triggered 200  $\mu$ s before arc ignition to ensure the magnetic field reached stable conditions before plasma generation. The coil current was monitored using another Chauvin Arnoux MA200, with a typical pulse width of 1.5 ms.

The magnetic field strength within the EM-coil solenoid was estimated using:

$$B = \frac{\mu_0 N I}{L} \quad (3.1)$$

where  $L = 0.02$  m is the solenoid length,  $N = 5$  is the number of turns,  $\mu_0 = 1.256 \times 10^{-6}$   $\frac{\text{T}\cdot\text{m}}{\text{A}}$  is the vacuum permeability, and  $I$  is the peak coil current. The peak current was determined from oscilloscope measurements at the beginning of the pulse, as this represents the maximum field strength before resistive losses cause the current to decay. The current waveform drops within a pulse due to the design of the pulsing unit in the power supplies. Figure 3.1 shows example current waveforms for both the arc and EM-coil for a 250 V input to the coil power supply. An additional example for 100 V input is provided in Appendix B.1.



**Figure 3.1.** Example pulse waveforms showing the arc current (red) and EM-coil current (black) as functions of time. The orange dashed lines indicate the trigger timing: (a) EM-coil trigger at  $t = -0.2$  ms, (b) arc trigger at  $t = 0$  ms. The waveforms shown are pulses at 250 V EM-coil input voltage. Distance: 10 cm, pressure: 0 Pa, magnetic field: 0.25 T.

### 3.1.3 Diagnostic positioning and synchronization

A delay generator (Stanford Research Systems DG645) served as the master clock for the experimental system, providing precisely timed trigger signals to:

- the arc power supply (channel **a**),
- the EM-coil power supply (channel **b**),
- the energy-resolving mass spectrometer (channel **c**).

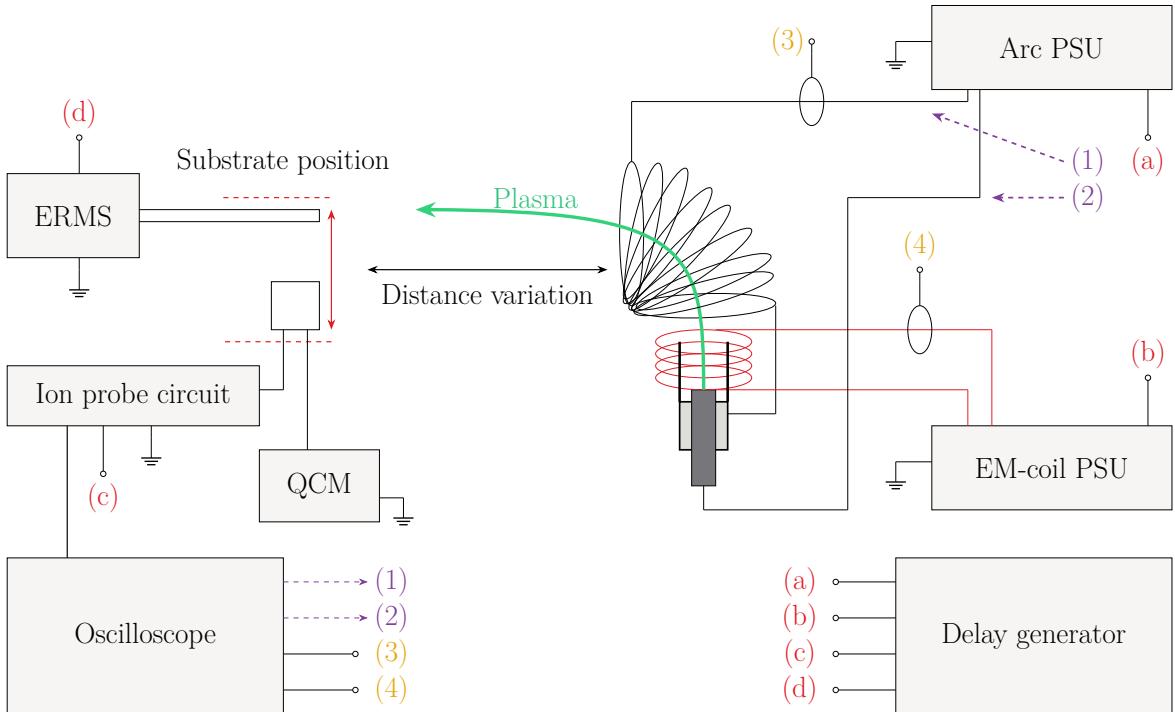
The oscilloscope was triggered on the rising edge of the cathode voltage (channel 1), which marks the onset of the arc pulse. Channels 1 and 2 of the oscilloscope measured the cathode-to-ground voltage and anode-to-ground voltage, respectively. Channels 3 and 4 recorded the arc current and EM-coil current via current probes. The ion current was recorded separately, more on that in Section 3.2.1.

The vacuum chamber setup for cathodic arc plasma diagnostics and thin film deposition is shown in Figure 3.2. The green arrow illustrates the trajectory of the plasma plume as it expands from the cathode, passes through the 90° macroparticle filter, and reaches the target. This experimental configuration enabled systematic investigation of plasma properties and film growth as functions of distance, magnetic field strength, and nitrogen pressure, as detailed in the following sections.

The ion current probe and quartz crystal microbalance (QCM) were not triggered by the delay generator. The ion probe signal was recorded directly by the oscilloscope

(triggered on the cathode voltage), automatically synchronizing ion current measurements with the arc pulses. The QCM operated continuously, with frequency measurements recorded before and after each sputtering run to determine the accumulated mass change over a fixed number of pulses. The BNC cable connector was detached during sputtering, this was done to protect the sensitive components in the oscillator from the highly energetic plasma.

For in situ plasma diagnostics, the Langmuir probe and QCM were mounted on a custom movable assembly positioned at distances ranging from 10 to 20 cm from the macroparticle filter exit. The energy-resolving mass spectrometer (ERMS) was mounted on a separate linear feedthrough allowing independent adjustment of its position (distance from the filter exit) for spatial characterization of the plasma. Film depositions were performed using silicon substrates mounted on a third movable holder, also positioned at controlled distances from the filter exit.



**Figure 3.2.** Schematic of the vacuum chamber setup for cathodic arc plasma diagnostics and thin film deposition with the plasma trajectory through the EM-coil and the macroparticle filter shown in green. At the substrate position the ERMS could be pulled back allowing for the substrate holder or a custom holder assembly integrated the ion current probe and QCM A.1 to be placed. The delay generator synchronizes both PSUs, the ERMS (when used) and the biasing for the ion current probe, and the oscilloscope records all the data as described in Sec.3.1.3.

## 3.2 Plasma diagnostics

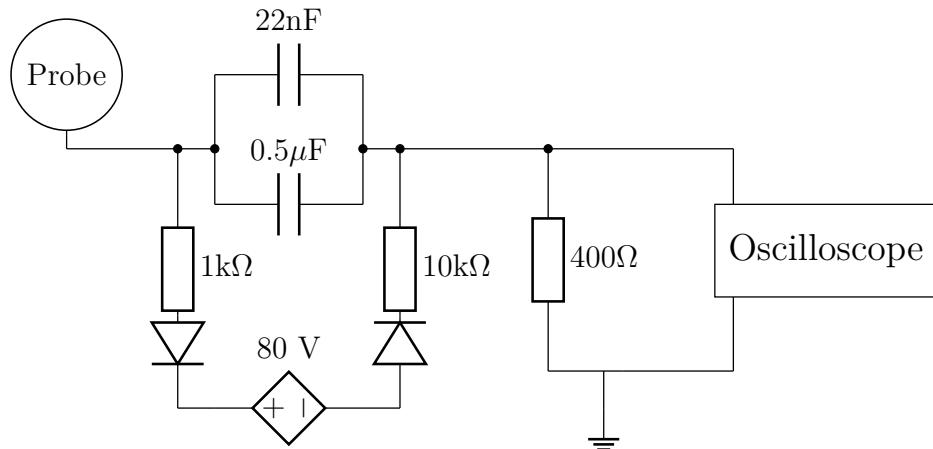
### 3.2.1 Ion current probe

An in-house-built ion collector probe was designed to measure the ion current density in ion saturation mode. The probe consisted of a 5 mm diameter copper rod machined to form a cylindrical collection surface with a nail-head geometry. This design prevented short circuits between the probe and the grounded mount due to metallic deposition. The probe body was wrapped in Kapton tape for electrical insulation from the aluminum mounting assembly (Figure A.1).



**Figure 3.3.** In-house built ion collector probe wrapped in Kapton tape for electrical insulation from the aluminum assembly holder. The probe includes an attachment point for a screw terminal connector, enabling connection to the ion probe circuit.

To ensure full ion collection and suppress electron current, the probe was negatively biased at  $V_b = -80$  V to operate in the ion saturation regime [15], as determined through bias voltage characterization measurements (Appendix A.5).



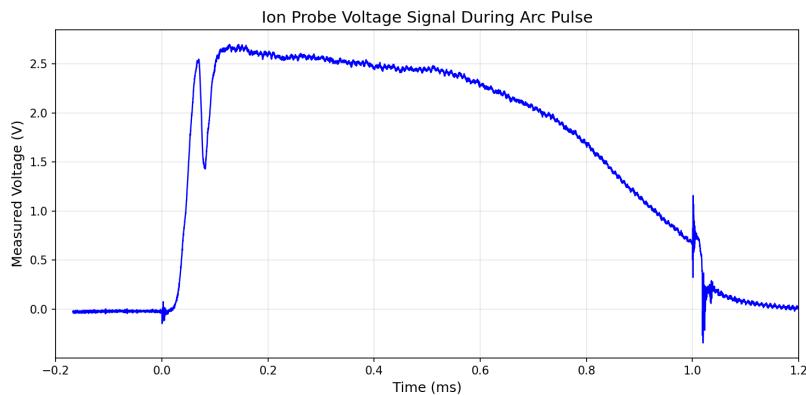
**Figure 3.4.** Schematic of the ion-flux probe circuit. The  $400\Omega$  resistor converts ion current to voltage, while the  $0.5\mu\text{F}$  capacitor and  $400\Omega$  resistor form a high-pass filter with a 795 Hz cutoff.

The probe circuit (Figure 3.4) converts the collected ion current to a voltage signal. The ion current passes through a  $0.5\mu\text{F}$  coupling capacitor and develops a voltage

across the  $400 \Omega$  measurement resistor. This RC configuration forms a high-pass filter with cutoff frequency:

$$f_c = \frac{1}{2\pi RC} = \frac{1}{2\pi \cdot 400 \Omega \cdot 0.5 \times 10^{-6} \text{ F}} \approx 795 \text{ Hz} \quad (3.2)$$

The high-pass filter blocks DC offsets while passing the pulsed ion current signal. The RC time constant  $\tau = 200 \mu\text{s}$  is much shorter than the 2 s interval between pulses, ensuring complete capacitor discharge between measurements. The bias supply line incorporates a  $22 \text{ nF}$  capacitor and  $1 \text{ k}\Omega$  resistor as a low-pass filter to smooth the  $-80 \text{ V}$  bias voltage.



**Figure 3.5.** Example voltage waveform measured across the  $400 \Omega$  resistor during a single arc pulse. The signal shows transients at pulse ignition (0 ms) and termination (1 ms), with gradual decay during the pulse due to the high-pass filter characteristics.

The ion current is calculated from the measured voltage as:

$$I_{\text{ion}} = \frac{V_{\text{measured}}}{R} \quad (3.3)$$

The voltage signal was recorded using a Tektronix MSO64 oscilloscope. Figure 3.5 shows an example waveform of 64 pulses averaged. The ion current density  $J_i$  is calculated as:

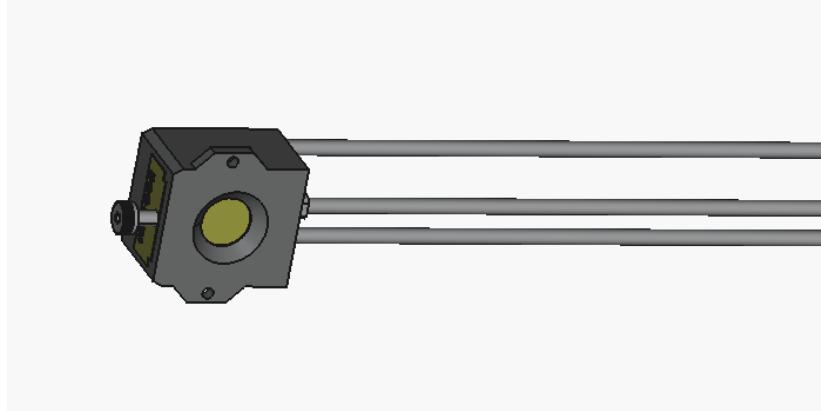
$$J_i = \frac{V_{\text{measured}}}{A \cdot R} \quad (3.4)$$

where  $A = 0.196 \text{ cm}^2$  is the probe collection area and  $R = 400 \Omega$ .

### 3.2.2 Quartz crystal microbalance

A quartz crystal microbalance (QCM) was used to measure the deposited mass during pulsed cathodic arc operation. The system consisted of an INFICON Cool Drawer<sup>TM</sup>

with a single sensor in standard orientation, water-cooled to ensure thermal stability. A 14 mm diameter, 6 MHz AT-cut quartz crystal was operated with an SQM-160 controller for electronic readout.



**Figure 3.6.** Sensor head of the INFICON Cool Drawer™ Quartz Crystal Microbalance (QCM) used for in-situ mass operation monitoring during cathodic arc sputtering. The assembly includes a water-cooled housing, a 14 mm diameter AT-cut quartz crystal (6 MHz), and electrode leads for connection to the SQM-160 controller. (Schematic adapted from INFICON STP file, available at <https://www.inficon.com/en/products/thin-film-technology/cool-drawer-single-sensor>).

The measurement principle follows the Sauerbrey equation [16], which relates the change in resonance frequency of the quartz crystal to the deposited mass:

$$\Delta m = \frac{N_{\text{AT}} \rho_q \pi r^2}{F_q^2} \Delta F = 18.8146023 \times 10^{-9} \frac{\text{g}}{\text{Hz}} \cdot \Delta F \quad (3.5)$$

Here  $\rho_q = 2.649 \text{ g cm}^{-3}$  is the quartz density,  $r = 0.7 \text{ cm}$  is the radius of the crystal area,  $N_{\text{AT}} = 166 \text{ 100 Hz cm}$  is the frequency constant of the AT-cut quartz,  $F_q = 6 \text{ MHz}$  is the uncoated resonance frequency, and  $\Delta F$  is the measured frequency shift.

The Sauerbrey relation is accurate as long as  $\Delta F \lesssim 0.05 F_q$  (approximately 0.3 MHz for a 6 MHz crystal), for more details on the correction above this limit look in Appendix A.1. In the present experiments, the observed frequency shifts ranged from approximately 1 Hz to 50 Hz over 64 pulses. These shifts are well above the SQM-160 resolution of approximately 0.03 Hz at 6 MHz, yet orders of magnitude below the Sauerbrey breakdown limit. The Sauerbrey approximation was therefore fully sufficient for all measurements in this work.

### 3.2.3 Simultaneous ion probe and QCM measurements

To enable direct comparison between ion current and deposited mass, the Langmuir probe and QCM were mounted on the same movable aluminum assembly. The probe was positioned through a precision-milled pass-through hole, while the QCM was secured in a dedicated cutout and fixed via screws (see Appendix A.2 for assembly details). This configuration ensured rigid mechanical alignment between the two diagnostics throughout all measurements.

The probe collection area ( $A_{\text{probe}} = 0.196 \text{ cm}^2$ , diameter 5 mm) was smaller than but contained within the QCM crystal area ( $A_{\text{QCM}} = 1.54 \text{ cm}^2$ , diameter 14 mm). Both diagnostics were positioned as close as possible to each other to minimize spatial gradients between measurement locations. This design is meant to mitigate radial plasma density gradients inherent to expanding cathodic arc plasmas [3, Chap. 6.2], though even with close proximity, gradients cannot be completely eliminated and remained evident in the experimental results.

### 3.2.4 Energy-resolving mass spectrometer (ERMS)

An energy-resolving mass spectrometer (ERMS, Hiden EQP 1000) was used to measure ion energy distribution functions (IEDFs) and charge-state-resolved fluxes of plasma species. The system combines an electrostatic energy analyzer with a quadrupole mass filter to measure distributions of energy-to-charge ratios (E/Q) at fixed mass-to-charge ratios (M/Q), and vice versa.

Ions enter the ERMS through a 50  $\mu\text{m}$  sampling orifice and are transported to the energy analyzer, where their kinetic energy  $E_i$  is selected according to:

$$E_i = \left( V_{\text{ENERGY}} + \frac{R}{d} V_{\text{PLATES}} - V_{\text{AXIS}} \right) Q \cdot e \quad (3.6)$$

Here,  $V_{\text{ENERGY}}$  and  $V_{\text{AXIS}}$  are opposing potentials applied to the analyzer,  $R$  is the mean radius of the cylindrical sector,  $d$  is the plate separation,  $V_{\text{PLATES}}$  is the potential difference across the sector plates,  $Q$  is the ion charge state, and  $e$  is the elementary charge [17].

The energy-selected ions are then injected into the quadrupole mass filter, where a combination of AC and DC electric fields creates a stability region dependent on  $M/Q$ , described by the Mathieu equations [18]. The potential in the quadrupole is:

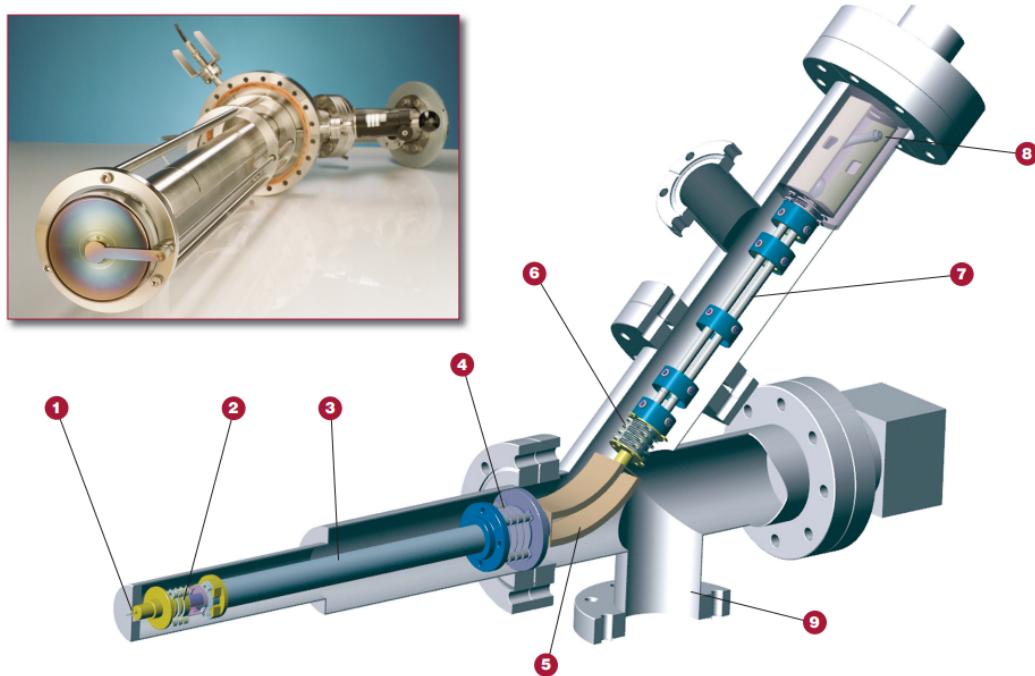
$$V(x, y, t) = \frac{U_0 \cos(\omega t)}{r_0^2} (x^2 - y^2) \quad (3.7)$$

where  $U_0$  is the amplitude of the AC voltage,  $\omega$  is the angular frequency, and  $r_0$  is the field radius. The stability of ion motion is determined by the dimensionless parameters:

$$a = \frac{8eU_{DC}}{Mr_0^2\omega^2}, \quad q = \frac{4eU_0}{Mr_0^2\omega^2} \quad (3.8)$$

with  $U_{DC}$  as the superimposed DC voltage and  $M$  the ion mass. For a given  $M/Q$ , stable transmission occurs only within specific  $(a, q)$  regions, enabling mass separation [18, 19]. Only ions with trajectories stable in both radial and axial directions reach the detector.

To reduce interference from the arc's magnetic field, the ERMS was equipped with a grounded mu-metal shield [4].



**Figure 3.7.** ERMS, Hiden EQP HE 1000: (1) Sampling Orifice, (2) Electron Impact Ion Source, (3) Transfer Ion Optics, (4) Quadrupole Lens, (5) Energy Filter, (6) Decelerating Lens, (7) Quadrupole Mass Filter, (8) Detector, (9) Differential Pump Port [20]

IEDFs were measured using a double-trigger acquisition scheme synchronized with the arc pulses (1 ms duration, 5 Hz repetition rate). For each  $M/Q$  value, two 20 ms acquisition windows were recorded, initiated 10 ms before the onset of the pulse. The combined 40 ms of data for each point were averaged to obtain the final IEDF. Measurements were performed for charge states  $1^+$ ,  $2^+$ , and  $3^+$  of aluminum ions, and for charge states  $1^+$ ,  $2^+$ ,  $3^+$ , and  $4^+$  of titanium ions. For nitrogen ion species ( $N$  and  $N_2$ ), only the  $1^+$  ionization level was measured. This was achieved by scanning  $V_{ENERGY}$  while fixing the quadrupole mass filter to the corresponding  $M/Q$  values (Table 3.1).

**Table 3.1.** Mass-to-charge ratios measured for each ion species and charge state.

Charge state	Al	Ti	N	N <sub>2</sub>
	(M/Q, amu)			
1+	27	47.867	14	28
2+	13.5	23.933	–	–
3+	9	15.955	–	–
4+	–	11.966	–	–

### Energy Conversion and Mass Transmission Correction

The ERMS system measures energy-per-charge ( $E/Q$ ) distributions for each mass-to-charge ratio ( $M/Q$ ). To derive ion energy distributions for different charge states, the measured  $E/Q$  distributions are converted to actual ion energies by multiplication with the corresponding charge state number  $Q$ :

$$E = Q \cdot (E/Q) \quad (3.9)$$

This conversion accounts for the charge-dependent scaling of ion energies and is necessary because the electrostatic energy analyzer resolves ion energy per unit charge, not total ion energy [17, 18].

A critical aspect of quantitative mass spectrometry is accounting for the mass-dependent transmission efficiency of the quadrupole mass filter and detector assembly. The transmission function  $T(M/Q)$  varies with mass-to-charge ratio due to the quadrupole's RF/DC field configuration and detector efficiency [7, 19]. Correction factors were determined through calibration measurements and applied to the raw ion signal intensities  $I_{\text{raw}}$  to obtain corrected intensities:

$$I_{\text{corrected}}(M/Q) = \frac{I_{\text{raw}}(M/Q)}{T(M/Q)} \quad (3.10)$$

Without this correction, lighter ions would be systematically underestimated relative to heavier ions, leading to errors in charge state distributions and composition analysis. From the corrected ion signals, statistical parameters are extracted for each charge state and species: mean ion energy  $\langle E \rangle$ , standard deviation  $\sigma_E$  characterizing the energy spread, and peak energy  $E_{\text{peak}}$  [3].

## Kinetic and Potential Energy Determination

The measured ion energies represent the kinetic energy of the ions. For each element, the mean kinetic energy is calculated as a weighted average over all charge states:

$$\langle E_{\text{kin}} \rangle = \frac{\sum_Q I_Q \cdot \langle E \rangle_Q}{\sum_Q I_Q} \quad (3.11)$$

where  $I_Q$  is the integrated ion current for charge state  $Q$  (after mass transmission correction) and  $\langle E \rangle_Q$  is the mean energy measured for that charge state.

The potential energy associated with each charge state is calculated from the cohesive energy and ionization potentials of the cathode material, following the methodology of Kalanov et al. [6]:

$$E_{\text{pot},Q} = E_{\text{coh}} + C \sum_{i=1}^Q E_{\text{ion},i} \quad (3.12)$$

where  $E_{\text{coh}}$  is the cohesive energy (4.85 eV for Ti, 3.39 eV for Al),  $E_{\text{ion},i}$  are the successive ionization energies, and  $C$  is a factor that accounts for the fraction of ionization energy deposited in the solid, with the remainder lost through radiation and Auger electron emission. Following Kalanov et al.,  $C = 0.8$  is used, representing the energy fraction contributing to atomic-scale heating during film growth. The mean potential energy for each element is then:

$$\langle E_{\text{pot}} \rangle = \frac{\sum_Q I_Q \cdot E_{\text{pot},Q}}{\sum_Q I_Q} \quad (3.13)$$

These kinetic and potential energy components are calculated separately, with total ion energy  $E_{\text{tot}} = E_{\text{kin}} + E_{\text{pot}}$  used in the analysis of film growth mechanisms (Chapter 5). Implementation details of the data processing workflow are provided in Appendix A.3.

## 3.3 Thin film deposition and characterization

Thin films were deposited on Si (100) substrates mounted on a movable substrate holder positioned at controlled distances from the macroparticle filter exit. Prior to deposition, the 2x2 cm substrates were cut into four pieces and cleaned with compressed nitrogen gas. For thickness measurements, a masking technique was employed: a marker line was drawn near the edge of each substrate before deposition, creating a well-defined step edge. After deposition, this marker was removed by ultrasonic cleaning in isopropanol, leaving a sharp boundary between the coated and uncoated regions for profilometry analysis.

The deposition parameters (distance, magnetic field strength, nitrogen pressure) were selected to match a subset of the plasma diagnostic measurements, enabling direct correlation between plasma properties and film characteristics. All depositions were performed with 6000–8000 pulses at 5 Hz repetition rate (1 ms pulse duration) to ensure sufficient film thickness for ex situ characterization. The substrate holder was not heated, allowing the investigation of room-temperature film growth under energetic ion bombardment conditions characteristic of cathodic arc deposition.

### 3.3.1 Profilometry

Stylus profilometry was used to measure film thickness by mechanically tracing the surface topography using a diamond-tipped stylus. The technique provides direct measurement of step heights between masked and deposited regions, making it particularly useful for verifying film thickness values obtained by QCM [21].

In profilometry, a stylus with a small tip radius is dragged across the sample surface with a controlled force of 3 mg while its vertical displacement is monitored electromagnetically. The resulting trace provides a profile of the surface from which the step height (film thickness) with vertical resolution down to  $\sim 1$  nm can be extracted.

The Bruker DektakXT stylus profilometer used in this study. For thickness measurements, silicon substrates were partially masked with a marker line, which created a well-defined step edge during deposition. This marker line was subsequently removed with Isopropanol in a ultrasonic bath. Multiple scans across each step were performed to ensure reproducibility, and the average thickness was calculated from at least three different positions on each sample.

### 3.3.2 X-ray diffraction (XRD)

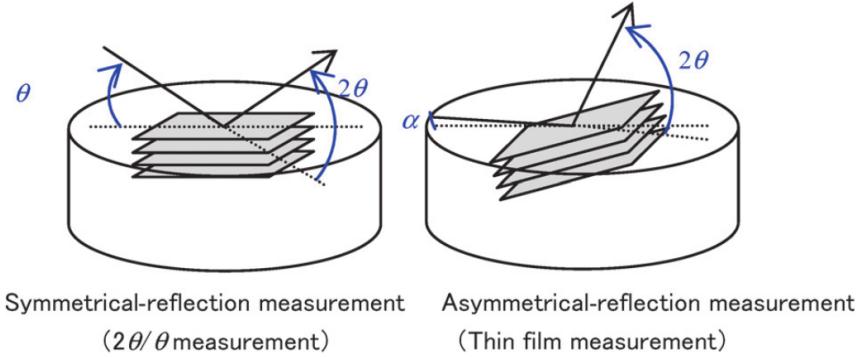
X-ray diffraction (XRD) was employed to analyze the crystallographic structure of the thin films deposited during the experiments. XRD is a non-destructive technique that provides detailed information about the crystalline phases present in the material, as well as their lattice parameters, crystallite size, and strain.

Due to the small film thickness, out-of-plane diffraction techniques were used to enhance the film signal relative to the substrate. Two measurement geometries are employed [22]:

- **Symmetrical reflection ( $2\theta/\theta$  scan):** Both incident and diffracted beams make equal angles with the sample surface. This geometry probes lattice planes parallel

to the substrate and is suitable for textured films, but substrate peaks can obscure weak film signals.

- **Asymmetrical reflection (thin-film method):** The incident beam is fixed at a small grazing angle  $\alpha$ , while the detector scans in  $2\theta$ . This reduces the X-ray penetration depth from tens of micrometers to a few micrometers, greatly enhancing sensitivity to thin films [22].



**Figure 3.8.** Schematic of out-of-plane diffraction geometries: symmetrical reflection (left) and asymmetrical reflection (right) for thin film analysis. Taken from [22].

The XRD measurements were performed using a Rigaku Ultima IV system equipped with a non monochromatic  $Cu_\alpha$  source. The crystallographic structure was determined by analyzing the diffraction patterns. The Bragg equation was used to identify the crystalline phases:

$$2d \sin(\theta) = n\lambda, \quad (3.14)$$

where  $d$  is the spacing between atomic planes,  $\theta$  is the diffraction angle,  $n$  is an integer, and  $\lambda$  is the X-ray wavelength.

### 3.3.3 X-ray reflectometry (XRR)

X-ray reflectometry (XRR) is a non-destructive technique used to determine film thickness, density, and surface or interface roughness by measuring the intensity of X-rays reflected at grazing incidence angles ( $0.01\text{--}5^\circ$ ) [23].

When X-rays strike a flat surface at shallow angles below the critical angle  $\theta_c$ , total external reflection occurs. The critical angle is related to the electron density of the material through:

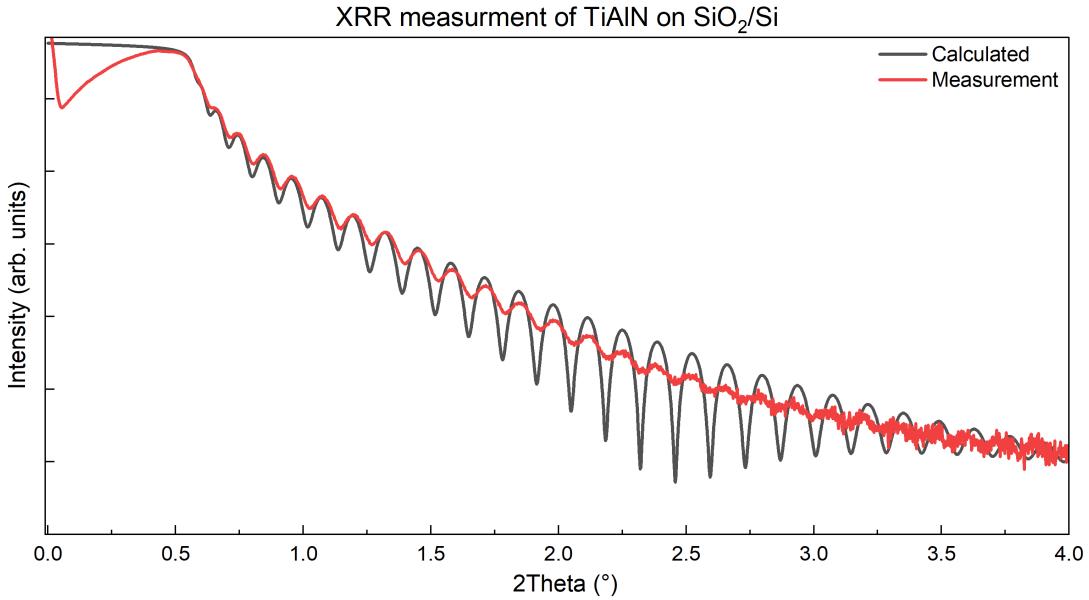
$$\theta_c = \sqrt{2\delta}, \quad (3.15)$$

where  $\delta$  depends on the material density and composition [23]. Above  $\theta_c$ , X-rays penetrate the film and reflect from interfaces, creating interference patterns known as

Kiessig fringes. The period of these oscillations is directly related to film thickness [24]:

$$d \approx \frac{\lambda}{2\Delta\theta}, \quad (3.16)$$

where  $\lambda$  is the X-ray wavelength and  $\Delta\theta$  is the angular spacing between fringes.



**Figure 3.9.** Example XRR measurement and fit for a TiAlN thin film deposited on Si substrate. The oscillations (Kiessig fringes) arise from interference between reflections from the film surface and the film-substrate interface.

The XRR technique provides three key parameters [23]:

- **Film thickness:** Determined from the oscillation period, with accuracy of 0.1–0.2 nm for films in the 1–150 nm range.
- **Density:** Extracted from the critical angle position and oscillation amplitude. Larger density contrast between film and substrate produces higher amplitude oscillations.
- **Surface and interface roughness:** Surface roughness causes faster decay of reflectivity at higher angles, while interface roughness reduces oscillation amplitude.

The measured reflectivity curves were fitted using GlobalFit software, which implements the Parratt recursive formalism [25] to account for multiple reflections at each interface. Film thickness, density, and roughness parameters were refined by fitting the calculated reflectivity curve to the experimental data. Initial thickness estimates from Profilometry measurements (Section 4.3.1) were used as starting values for the fitting procedure. XRR is particularly suited for verifying thicknesses and characterizing film density and surface quality for the deposited TiAl/TiAlN films.

### 3.3.4 Scanning electron microscopy (SEM)

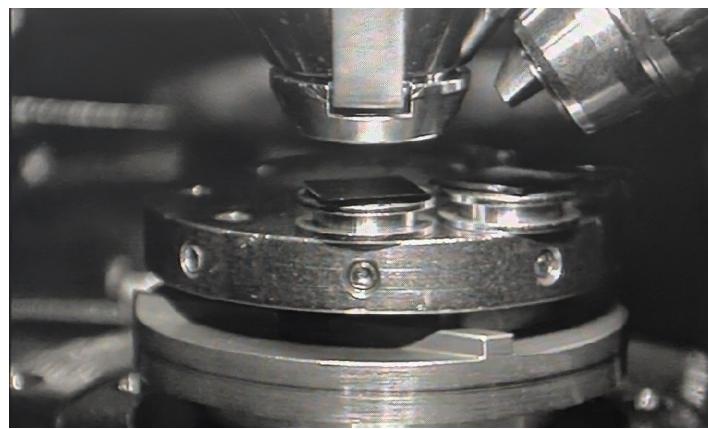
Scanning electron microscopy (SEM) was employed to characterize the surface morphology and microstructure of the deposited films. SEM uses a focused electron beam to scan the sample surface, generating secondary electrons (SE) and backscattered electrons (BSE) that provide information about surface topography and compositional contrast, respectively [26, 27].

SEM imaging was performed using a Zeiss Ultra 55 field emission scanning electron microscope equipped with an in-lens detector for high-resolution secondary electron imaging. Samples were imaged at an accelerating voltage of 15 kV with magnifications ranging from  $30,000\times$  to  $100,000\times$ . The conductive TiAlN films required no additional coating preparation.

The primary objectives of SEM characterization were to:

- Examine surface morphology and the presence of any macroparticles typical of cathodic arc deposition
- Assess film uniformity across the substrate
- Identify microstructural features resulting from different deposition conditions

Representative SEM micrographs and analysis of surface features are presented in Section 4.3.5.



**Figure 3.10.** Interior view of the SEM sample chamber showing the sample holder stage and detector configuration used for imaging and EDX analysis.

### 3.3.5 Energy-dispersive X-ray spectroscopy (EDX)

Energy-dispersive X-ray spectroscopy (EDX) relies on the ionization of inner-shell electrons by the incident electron beam. When an inner-shell electron is ejected, an electron

from a higher energy level fills the vacancy, releasing energy in the form of a characteristic X-ray. The energy of this X-ray is unique to each element, allowing identification and quantification of the sample composition. The intensity of the characteristic X-ray peaks is proportional to the concentration of each element, enabling quantitative analysis through comparison with standards or standardless quantification algorithms [26].

EDX measurements were performed using a silicon drift detector (SDD) integrated with the SEM system, operating at an accelerating voltage of 10 kV. To minimize substrate contribution from the thin films, samples were tilted to 55° relative to the electron beam increasing the effective path length by  $\approx 1.74$ .

Prior to analysis, samples were cleaned ultrasonically in isopropanol to remove loose particles. While profilometry provides straightforward thickness measurements of the deposited films (see Section 4.3.1), EDX analysis requires careful consideration of potential contamination sources. Organic solvent residues and hydrocarbon contamination from the vacuum system can lead to carbon deposition under electron beam irradiation [26, 27]. A small carbon peak was therefore typically observed in EDX spectra and was excluded from compositional quantification of the TiAlN films. Multiple area measurements were taken on each sample to verify that all surface features, including any macroparticles or morphological structures, had the same metal-nitride composition as the underlying thin film and to identify potential impurities beyond the surface carbon layer.

The primary elements analyzed were titanium (Ti), aluminum (Al), and nitrogen (N). Oxygen (O), silicon (Si) and carbon (C) were also monitored to assess surface oxidation, contribution from the Si substrate and contamination from Isopropanol during the ultrasonic bath cleaning steps. The measured film compositions are used to calculate the effective molar mass  $M_{\text{eff}}$  required for determining total deposited flux from QCM measurements (Section 3.2.4). Compositional results and representative EDX spectra are presented in Section 4.3.4.

## 3.4 Calculation of ion flux and total deposited flux

The plasma diagnostic measurements described in Section 3.2.1–3.2.4 provide electrical currents, deposited masses, and ion energy distributions. To quantify the actual particle arrival rates and relate these to film growth, these raw measurements must be converted to particle fluxes through the conversion factors and procedures described below. Uncertainty estimates for the calculated fluxes, accounting for instrument un-

certainties and measurement statistics, are provided in Appendix A.4.

### 3.4.1 Ion flux from probe and ERMS measurements

The ion flux  $\Gamma_{\text{ion}}$  ( $\text{ions cm}^{-2} \text{s}^{-1}$ ) represents the number of charged metal ions arriving per unit area per unit time. It is calculated from the ion current measured by the Langmuir probe (Section 3.2.1) combined with the mean charge state determined from ERMS measurements (Section 3.2.4):

$$\Gamma_{\text{ion}} = \frac{I_{\text{ion}}}{e\langle Q \rangle A_{\text{probe}}} \quad (3.17)$$

where  $\langle Q \rangle$  is the average charge state determined from ERMS measurements (Section 3.2.4),  $e = 1.602 \times 10^{-19} \text{ C}$  is the elementary charge,  $A_{\text{probe}} = 0.196 \text{ cm}^2$  is the probe collection area, and  $R = 400 \Omega$  is the measurement resistor. This combination of probe and mass spectrometer data accounts for the multiply charged nature of the cathodic arc plasma and enables quantitative determination of the particle flux from the electrical current measurement.

$$\langle Q \rangle = \frac{\sum_Q Q \cdot I_Q}{\sum_Q I_Q} \quad (3.18)$$

where  $I_Q$  is the integrated ion current for charge state  $Q$  after applying the mass transmission correction (Equation 3.10) [28, 5]. This combination of probe and mass spectrometer data accounts for the multiply charged nature of the cathodic arc plasma and enables accurate determination of the particle flux from the electrical current measurement.

### 3.4.2 Total deposited flux from QCM and film characterization

The total deposited flux  $\Phi_{\text{total}}$  ( $\text{atoms cm}^{-2} \text{s}^{-1}$ ) quantifies the arrival rate of all species contributing to film growth, including both ions and neutrals. It is calculated from QCM mass measurements (Section 3.2.2) combined with film composition from EDX (Section 3.3.5):

$$\Phi_{\text{total}} = \frac{\Delta m \cdot N_A}{A_{\text{QCM}} \cdot \Delta t \cdot M_{\text{eff}}} \quad (3.19)$$

where  $\Delta m$  is the mass change measured by QCM over time interval  $\Delta t$ ,  $A_{\text{QCM}} = 1.54 \text{ cm}^2$  is the crystal area (14 mm diameter),  $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$  is Avogadro's constant, and  $M_{\text{eff}}$  is the effective molar mass calculated from film composition:

$$M_{\text{eff}} = \sum_i x_i M_i = x_{\text{Ti}} \cdot 47.867 + x_{\text{Al}} \cdot 26.982 + x_{\text{N}} \cdot 14.007 \quad (3.20)$$

where  $x_i$  are the atomic fractions from EDX measurements (Table 4.3) and  $M_i$  are the atomic masses.

The effective molar mass varies significantly between metallic and reactive mode:  $M_{\text{eff}} \approx 43.5 \text{ g}\cdot\text{mol}^{-1}$  for Ti–Al films versus  $M_{\text{eff}} \approx 31.8\text{--}33.2 \text{ g}\cdot\text{mol}^{-1}$  for TiAlN films. This 25% reduction reflects the incorporation of light nitrogen atoms (14.0 g·mol<sup>-1</sup>) displacing heavier metal atoms.

# CHAPTER 4

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## Results

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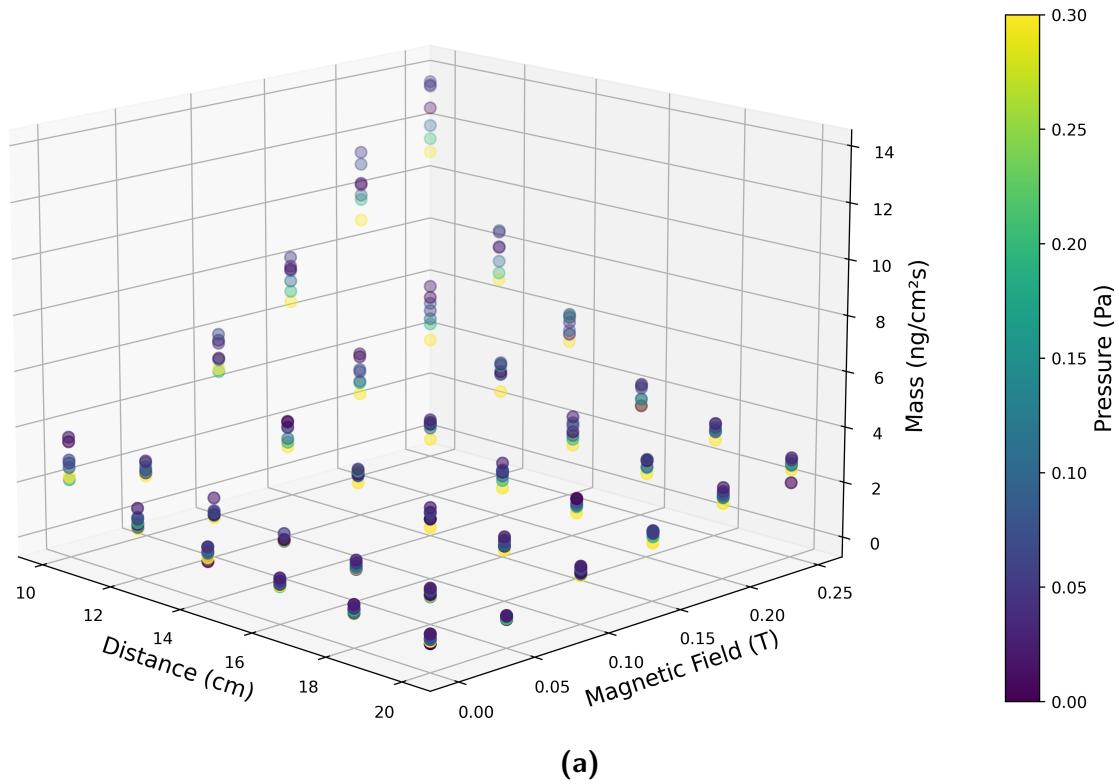
### 4.1 Mass deposition rate and ion current measurements

To systematically characterize the plasma dynamics and deposition behavior, measurements were conducted using a quartz crystal microbalance (QCM) for mass deposition rate and a biased ion collector probe for ion current density.

The experimental parameter space spanned three variables: distance from the macroparticle filter (10–20 cm), applied magnetic field strength (0–0.25 T), and nitrogen background pressure (0–0.3 Pa), with all permutations measured as listed in Table B.1. At this stage, the raw quantities mass deposition rate (in  $\text{ng cm}^{-2} \text{ s}^{-1}$ ) and ion current (in mA) are shown to establish trends before flux calculations in Section 3.4.

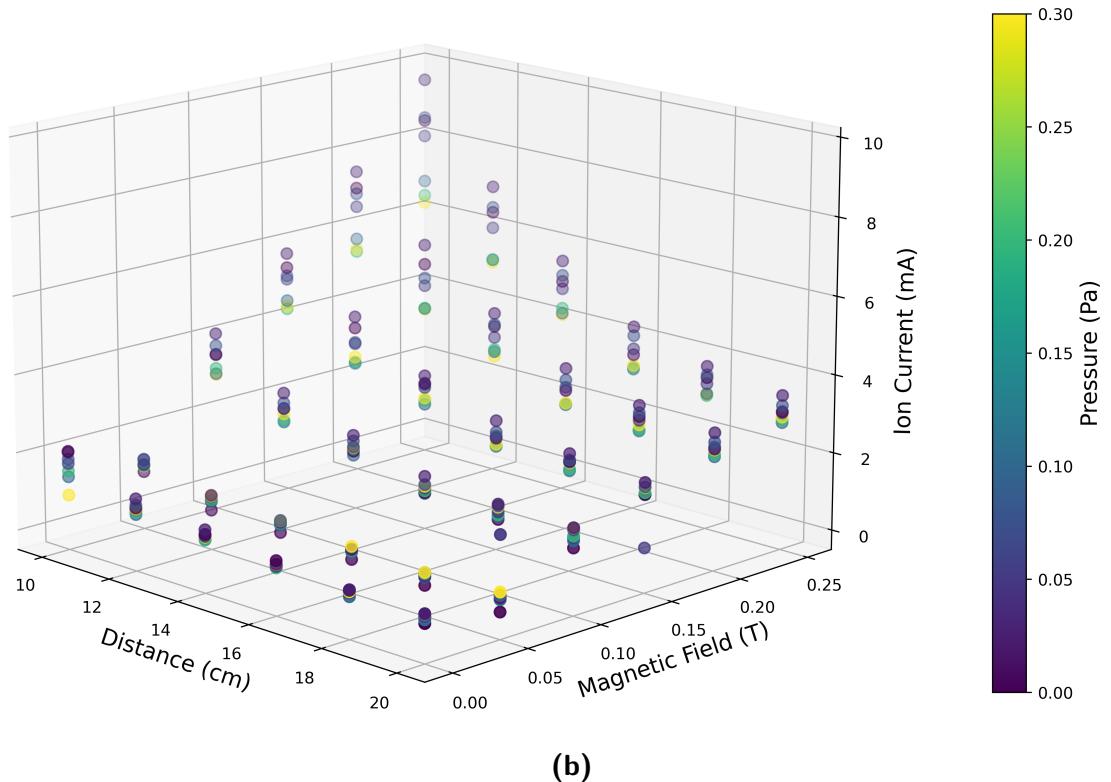
Figure 4.1 provides a three-dimensional visualization of the complete dataset, illustrating how mass deposition rate and ion current vary simultaneously with all three control parameters. Several global trends are immediately apparent: both quantities decrease with increasing distance due to plasma expansion, increase with applied magnetic field due to enhanced plasma enhancement and ionization, and exhibit complex pressure dependence that warrants detailed investigation.

Mass ( $\text{ng}/\text{cm}^2\text{s}$ ) vs. Distance (cm), Magnetic Field (T) and Pressure (Pa)



(a)

Ion Current (mA) vs. Distance (cm), Magnetic Field (T) and Pressure (Pa)



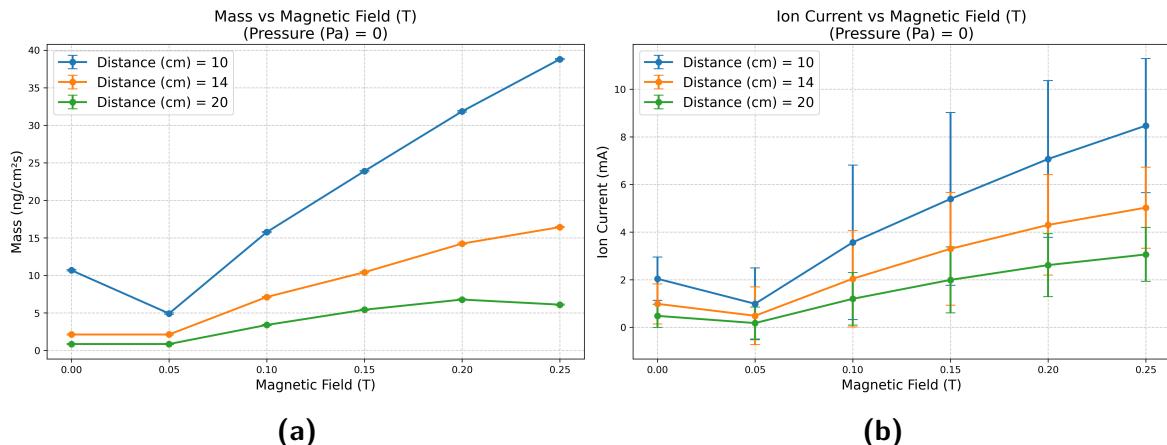
(b)

**Figure 4.1.** Three-dimensional visualization of (a) mass deposition rate and (b) ion current as functions of magnetic field strength, source-to-substrate distance, and nitrogen pressure. Each data point represents an average over 64 pulses at 5 Hz repetition rate, normalized to a single pulse. Surface colors indicate the magnitude of the nitrogen pressure.

To systematically dissect these multidimensional trends, the following subsections examine cross-sections through the parameter space. We begin with the metallic case (no nitrogen, Section 4.1.1) to characterize the plasma behaviour without reactive gas present. Sections 4.1.2–4.1.4 then examine the individual effects of distance, magnetic field, and nitrogen pressure by varying each parameter while holding the others constant. The parameter combinations analyzed in detail—distances of 10, 14, and 20 cm; magnetic fields of 0, 0.15, and 0.25 T; and nitrogen pressures of 0.1 and 0.3 Pa—were selected because they correspond to conditions used for energy-resolved mass spectrometry (ERMS) measurements and thin film deposition, enabling direct correlation between plasma diagnostics and film properties in later sections.

#### 4.1.1 Metallic case (no nitrogen)

In the absence of reactive gas, plasma expansion and deposition dynamics are governed solely by geometric spreading and magnetic enhancement. Figure 4.2 presents mass deposition rate and ion current as functions of magnetic field strength for three representative distances.



**Figure 4.2.** Magnetic field dependence of (a) mass deposition rate and (b) ion current in metallic mode (0 Pa). Both quantities increase monotonically above 0.1 T due to improved plasma guiding, with the strongest enhancement at 10 cm where radial expansion is minimal. The anomalous dip at 0.05 T (discussed in Sec. 5.2) appears across all distances. Error bars represent intra-pulse variability.

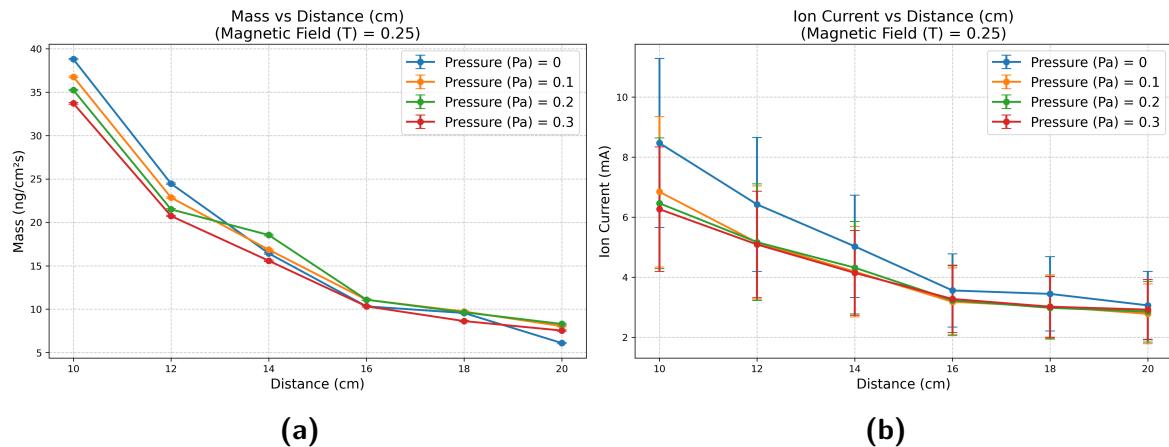
As shown in Figure 4.2, both mass deposition rate and ion current increase with applied magnetic field. The enhancement is most pronounced at the shortest distance (10 cm), where the magnetic field can effectively guide ions before significant radial expansion occurs. At 20 cm, the plasma has already expanded substantially, reducing the relative impact of magnetic enhancement on the collected flux.

A surprising anomaly appears at 0.05 T, where both quantities drop below their

zero-field values before recovering and increasing at higher fields (0.1 T and above). Similar behavior is observed across all experimental conditions involving increasing magnetic field strength, more on that in section 5.2.

#### 4.1.2 Distance as a variable

Figure 4.3 examines the effect of source-to-substrate distance under fixed magnetic enhancement (0.25 T) for both metallic and reactive conditions. Four nitrogen pressures are compared: 0 Pa (metallic), 0.1 Pa, 0.2 Pa, and 0.3 Pa.



**Figure 4.3.** Mass deposition rate (a) and ion current (b) as functions of distance from the macroparticle filter exit at 0.25 T magnetic field strength, measured at nitrogen pressures of 0, 0.1, 0.2, and 0.3 Pa.

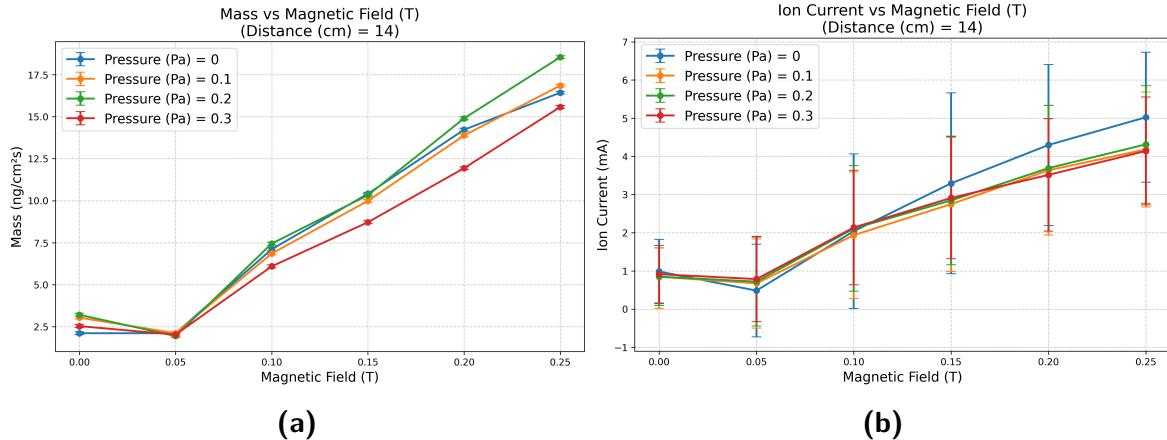
Both mass deposition rate and ion current decrease with distance, reflecting plasma expansion and the corresponding decrease in particle density. In metallic mode (0 Pa), mass deposition rate follows an approximate  $1/r^2$  dependence expected for free expansion [3, Chap. 4.3], with 20 cm values approximately 25–30% of those measured at 10 cm.

Ion current exhibits a more pronounced pressure dependence than mass deposition rate. When nitrogen is introduced, the measured current drops steeply, while the QCM mass deposition rate shows more modest changes. The contrast between the two measurements is most pronounced at short distances and diminishes at 14–20 cm.

Additional measurements at constant magnetic field (0 T) and varying distance show similar trends and are presented in Appendix B.2.

### 4.1.3 Magnetic field as a variable

Figure 4.4 presents mass flux and ion current as functions of magnetic field strength at a fixed intermediate distance (14 cm) for the same set of pressures.



**Figure 4.4.** (a) Mass deposition rate and (b) ion current as functions of magnetic field strength (0–0.25 T) at 14 cm distance, shown for metallic mode (0 Pa) and reactive mode (0.1, 0.2, 0.3 Pa nitrogen pressure).

The magnetic mirror anomaly at 0.05 T, clearly visible in the 10 cm metallic data (Figure 4.2), is significantly attenuated at 14 cm. Both mass deposition rate and ion current remain approximately constant between 0 and 0.05 T, suggesting that the adverse mirror effect is either weaker after 4 cm of additional expansion or is masked by increased statistical noise at this intermediate distance.

Above 0.1 T, both quantities show a clear, consistent increase for all conditions, confirming that magnetic enhancement enhances plasma density and flux at the substrate. The enhancement is approximately linear with field strength in the 0.1–0.25 T range. Mass deposition rate (Figure 4.4a) increases by a factor of approximately 5–6× from zero field to 0.25 T, with all pressure conditions following similar trajectories. This similarity indicates that the magnetic field effect on total mass flux is largely independent of nitrogen pressure at this distance.

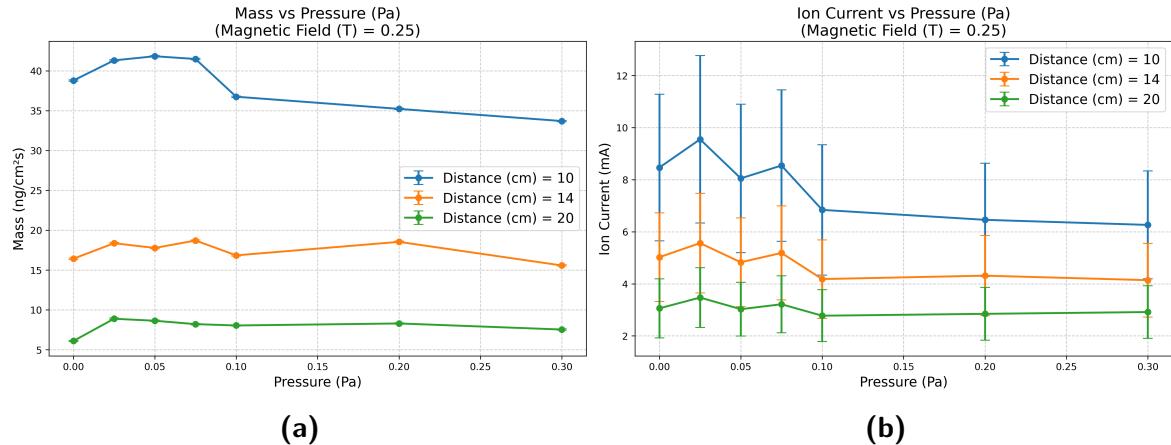
Ion current (Figure 4.4b) also increases with magnetic field, but exhibits greater variability, particularly in reactive mode. At 0.25 T, the ion current in metallic mode (0 Pa) reaches approximately 5 mA, while reactive mode conditions (0.1–0.3 Pa) show currents in the range of 4 mA with substantial error bars. The overlap between metallic and reactive conditions at high fields, combined with the large variation of ion current during a pulse, makes it difficult to draw definitive conclusions about pressure effects on ion current at 14 cm distance.

The similar behavior of mass deposition rate across all pressures, combined with the more variable ion current measurements, suggests that at 14 cm the magnetic field primarily affects the total particle flux rather than selectively enhancing ionization. The reduced contrast between metallic and reactive mode compared to the 10 cm measurements (Section 4.1.2) indicates that charge-exchange collisions have already reduced the ion charge states by this distance.

To isolate the effect of nitrogen pressure more clearly, the next section examines pressure as the primary variable under fixed magnetic field conditions.

#### 4.1.4 Nitrogen pressure as a variable

Figure 4.5 presents mass flux and ion current as functions of nitrogen pressure at maximum magnetic enhancement (0.25 T) for three distances (10, 14, and 20 cm).



**Figure 4.5.** (a) Mass deposition rate and (b) ion current as functions of nitrogen pressure (0–0.3 Pa) at a magnetic field strength of 0.25 T, shown for the distances 10, 14 and 20 cm.

At 10 cm, the mass deposition rate exhibits non-monotonic pressure dependence, with a local maximum near 0.05–0.1 Pa before decreasing at higher pressures. The ion current shows substantial scatter across the low pressure ranges this is an artefact from the experimental setup. A general decreasing trend is evident despite the variability.

At 14 cm and 20 cm, both mass deposition rate and ion current become largely independent of nitrogen pressure. Within experimental uncertainty, both quantities remain essentially constant across the 0–0.3 Pa range investigated.

The measurements presented in this section establish the general trends in mass deposition rate and ion current across the experimental parameter space. However, these raw quantities do not directly reveal the underlying plasma composition. To

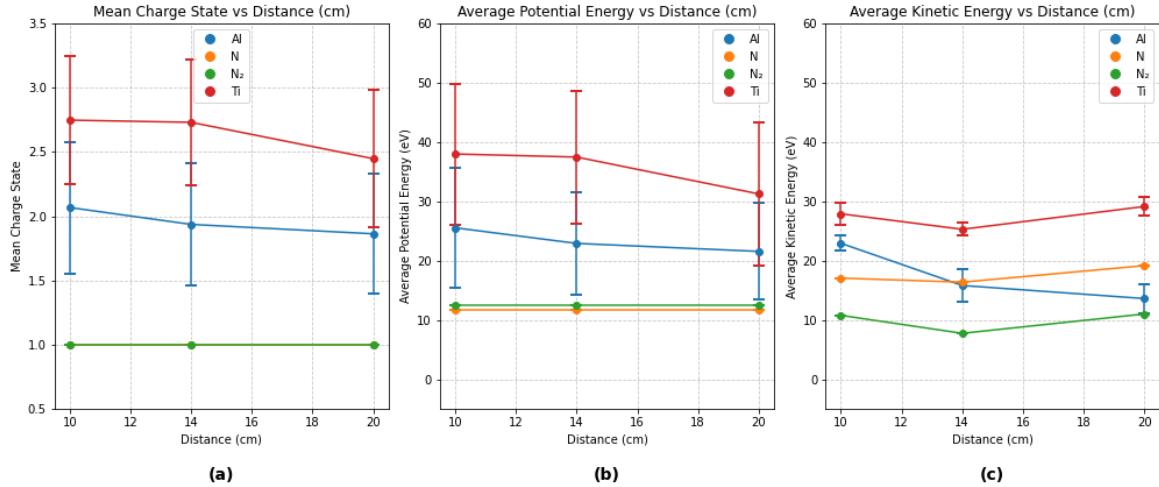
address these questions, energy-resolved mass spectrometry was performed under selected conditions to characterize the ion energy distributions and mean charge states, as presented in the following section.

## 4.2 Energy-resolved mass spectrometry results

Energy-resolved mass spectrometry (ERMS) measurements were performed to characterize charge-state distributions and ion energies under selected conditions corresponding to film depositions. The measurements provide charge-state-resolved ion energy distribution functions (IEDFs) for  $\text{Al}^{1+,2+,3+}$ ,  $\text{Ti}^{1+,2+,3+,4+}$  and  $\text{N}/\text{N}_2^{1+}$ , from which mean charge states  $\langle Q \rangle$  (3.18), kinetic energies  $E_{\text{kin}}$  (3.11), and potential energies  $E_{\text{pot}}$  (3.12) are derived. All measurements were performed at 0.25 T magnetic field strength, with particular emphasis on the 14 cm distance selected for systematic film depositions, as well as for all the above mentioned species.

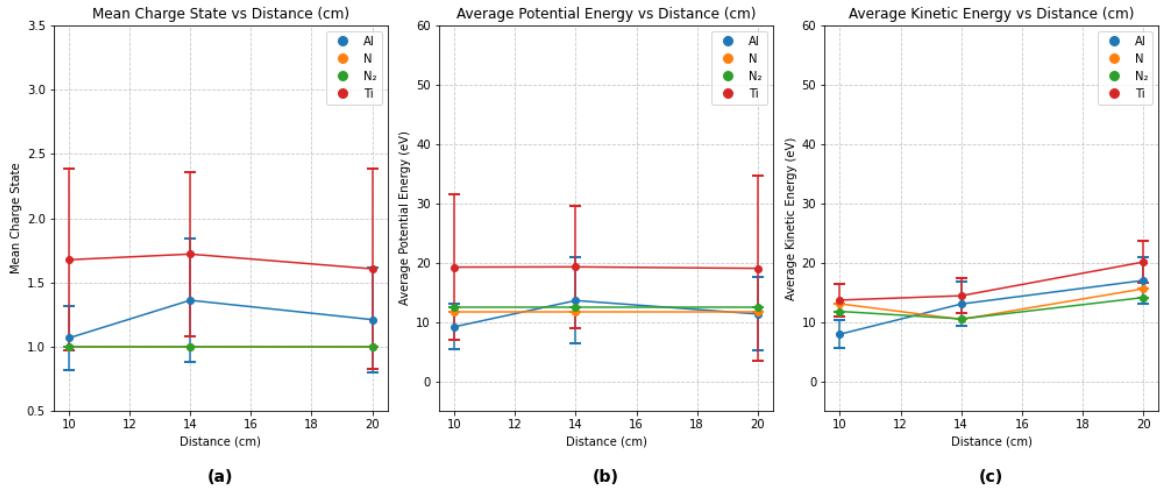
### 4.2.1 Distance dependence

Figures 4.6 and 4.7 present mean charge states, potential energies, and kinetic energies as functions of distance for metallic mode (0 Pa) and reactive mode (0.3 Pa  $\text{N}_2$ ), respectively.



**Figure 4.6.** Mean charge state (a), average potential energy (b), and average kinetic energy (c) as functions of distance in metallic mode (0 Pa nitrogen) at 0.25 T magnetic field. Data shown for average values of Al, Ti, and N ion species.

In metallic mode (Figure 4.6), titanium ions exhibit  $\langle Q \rangle \approx 2.6$  and aluminum ions  $\langle Q \rangle \approx 2.0$  across all distances. These charge states reflect the high electron temperatures ( $T_e \approx 5\text{--}10$  eV) characteristic of cathode spot plasmas [3, Chap. 4.3]. The mean

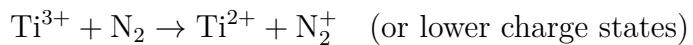


**Figure 4.7.** Mean charge state (a), average potential energy (b), and average kinetic energy (c) as functions of distance in reactive mode at 0.25 T magnetic field. Data shown for average values of Al, Ti, and N ion species at nitrogen pressures of 0.1, 0.2, and 0.3 Pa.

charge states remain remarkably constant with distance, varying by less than 10% between 10 and 20 cm, indicating that recombination is negligible during the transit time through this region on the order of  $\sim 100 \mu\text{s}$  scale.

Potential energies scale directly with charge state, with Ti showing  $E_{\text{pot}} \approx 38 \text{ eV}$  at 10 cm decreasing to  $\approx 31 \text{ eV}$  at 20 cm, and Al showing  $E_{\text{pot}} \approx 26 \text{ eV}$  remaining nearly constant across the measurement range. Kinetic energies are in the range of 20–30 eV for both species and remain approximately constant with distance within measurement uncertainty. Importantly, total ion energies ( $E_{\text{kin}} + E_{\text{pot}}$ ) exceed 40 eV at all distances, well above the  $\approx 30 \text{ eV}$  threshold for subplantation-driven densification [11]. This confirms that energetic condensation remains accessible across the entire 10–20 cm range without substrate heating.

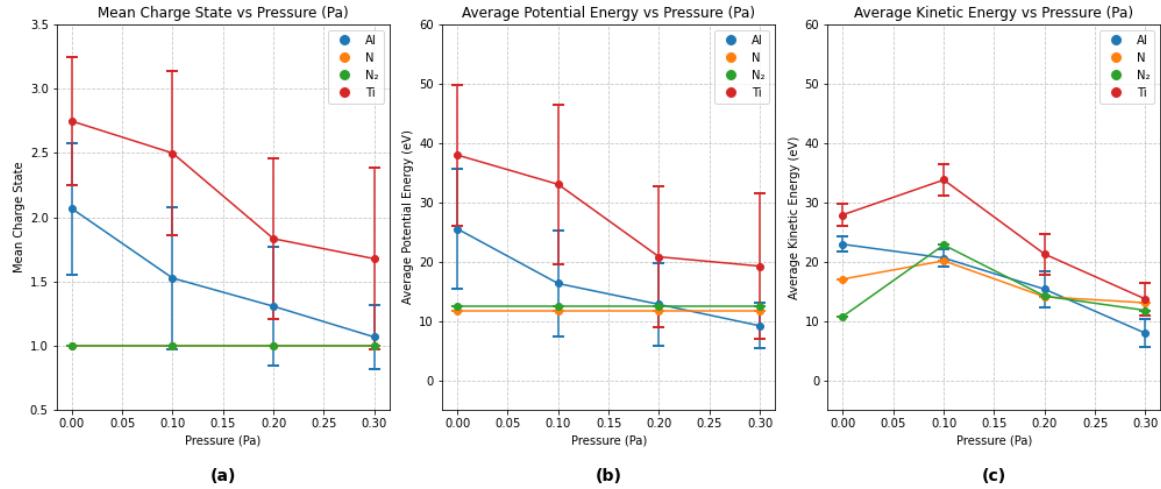
Introducing nitrogen at 0.3 Pa substantially alters the plasma composition (Figure 4.7). Metal ion charge states decrease by up to 50%, coinciding with charge-exchange conditions where multiply charged metal ions interact with nitrogen molecules [7]:



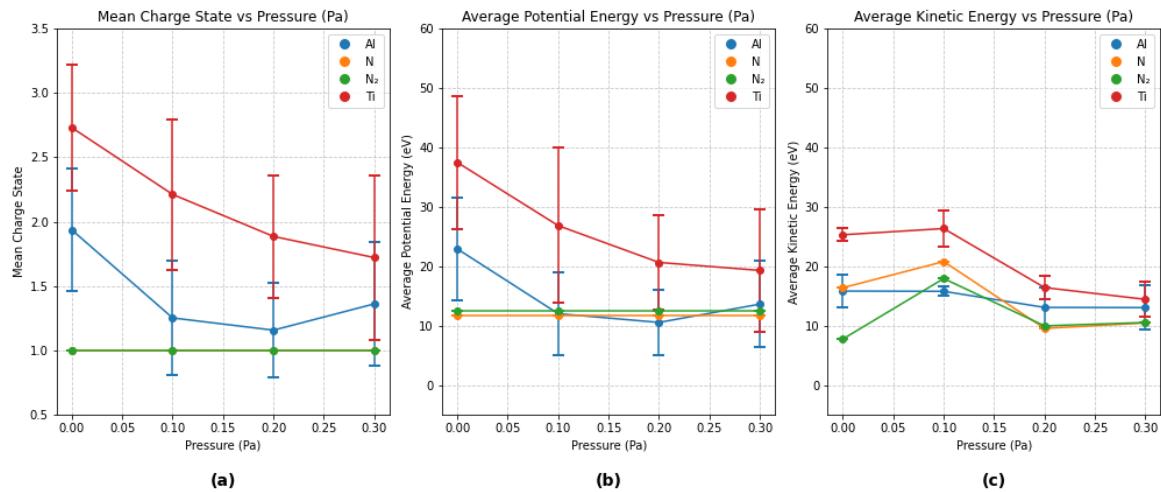
Potential energies decrease proportionally with the reduced charge states, falling from  $\approx 35\text{--}40 \text{ eV}$  (metallic) to  $\approx 20\text{--}30 \text{ eV}$  in reactive mode. Kinetic energies show a slight increase in reactive mode. Despite these changes, total ion energies remain in the range of 30 eV at 10 cm and 35 eV at 20 cm. At the 14 cm deposition distance, energies are approximately 25–35 eV.

## 4.2.2 Pressure dependence

Figures 4.8 and 4.9 present the pressure dependence of charge states and ion energies at 10 cm and 14 cm, respectively. The 14 cm measurements are emphasized as this distance was selected for film depositions.



**Figure 4.8.** Mean charge state (a), average potential energy (b), and average kinetic energy (c) as functions of nitrogen pressure for Al, Ti, and N ions at 10 cm distance and 0.25 T magnetic field.



**Figure 4.9.** Mean charge state (a), average potential energy (b), and average kinetic energy (c) as functions of nitrogen pressure for Al, Ti, and N ions at 14 cm distance and 0.25 T magnetic field.

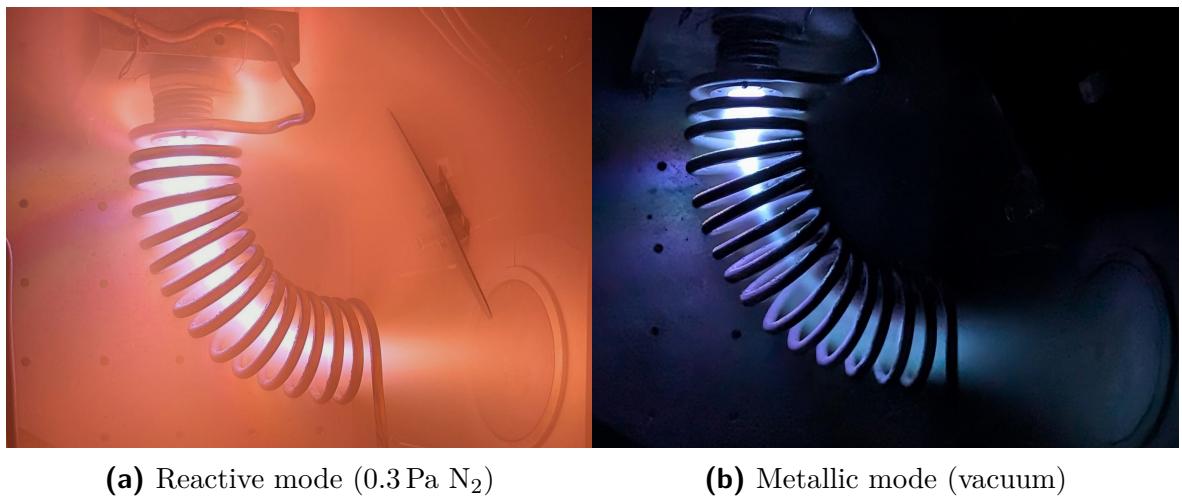
At both distances, metal ion charge states generally decrease with increasing nitrogen pressure as charge-exchange collisions accumulate. At 14 cm, Ti decreases from

$\langle Q \rangle \approx 2.7$  (0 Pa) to  $\approx 1.7$  (0.3 Pa), and Al from  $\approx 1.9$  to  $\approx 1.3$ . This trend is non-monotonic, with a slight increase observed between 0.2 and 0.3 Pa at 14 cm. This deviation from systematic reduction suggests that charge-exchange collisions are not the sole process affecting charge-state distributions. The overall charge-state reduction is qualitatively similar at both distances.

Kinetic energies show more complex behavior. At both 10 cm and 14 cm, Ti and Al exhibit non-monotonic trends with a local maximum near 0.1 Pa. The kinetic energy measurements show increased scatter at 0.2–0.3 Pa compared to metallic mode. The origin of the non-monotonic behavior requires further investigation, but the overall effect is modest compared to the charge state reduction.

### 4.3 Thin film analysis

Thin films were deposited on Si(100) substrates positioned at 14 cm from the macroparticle filter exit under conditions listed in Table B.1. The substrates were mounted on the substrate holder with the same positions in space used for QCM and ion probe measurements to ensure identical positioning. Films were deposited under four representative conditions spanning the parameter space: metallic mode (0 Pa) and reactive mode (0.1, 0.2, 0.3 Pa) and at high magnetic field strength (0.25 T). All depositions used 850 A arc current, 1 ms pulse width, and 5 Hz repetition rate.



**Figure 4.10.** Photographs of filtered cathodic arc plasma in (a) reactive mode at 0.3 Pa nitrogen pressure and (b) metallic mode in vacuum. Images taken during deposition at 14 cm distance with 0.25 T magnetic field. Camera settings: f/1.7, 1/100 s, ISO 90.

### 4.3.1 Profilometry

Samples were prepared by placing a straight marker line near the edge of each substrate prior to deposition. This marker could be cleanly removed by ultrasonic washing after film growth, revealing the deposition step for thickness analysis. Profilometry measurements were performed at three positions on each sample: one near the center and one near each edge. The three measurements were averaged to obtain the mean film thickness.

In reactive conditions, the deposition rate decreased compared to metallic mode. To achieve comparable film thicknesses for characterization, the pulse count was increased from 6000 to 8000 pulses for nitrogen containing depositions. The resulting thickness values are summarized in Table 4.1.

**Table 4.1.** Film thickness measured by stylus profilometry at two substrate positions (QCM and ion probe locations) for films deposited at 14 cm distance with 0.25 T magnetic field. All films deposited with 850 A arc current, 1 ms pulses at 5 Hz.

Film ID	Pressure (Pa)	Pulses	Thickness QCM (nm)	Thickness Ion probe (nm)
10	0.0	6000	65	57
9	0.3	6000	42	35
13	0.1	8000	59	41
14	0.2	8000	46	38

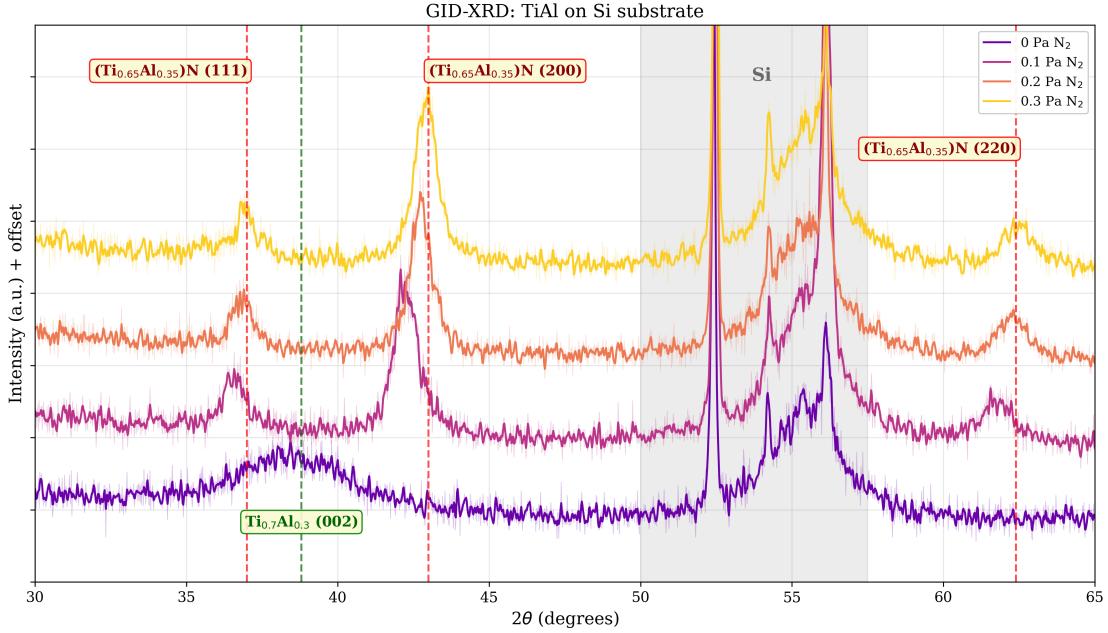
The thickness differences between the QCM and ion probe positions range from 14–44% and were used to correct the final flux values. Films deposited in reactive mode show deposition rates reduced by 36–45% compared to metallic mode, despite the increased pulse count used to partially compensate for this reduction.

### 4.3.2 X-ray diffraction (XRD)

X-ray diffraction measurements were performed to characterize the crystallographic structure and phase composition of the deposited films as a function of nitrogen pressure. Figure 4.11 shows the diffraction patterns for films deposited at 14 cm distance with 0.25 T magnetic field under varying N<sub>2</sub> pressures from 0 Pa (metallic mode) to 0.3 Pa (reactive mode).

The diffraction patterns reveal a clear transition from metallic to nitride phases with increasing nitrogen pressure:

**Metallic mode (0 Pa N<sub>2</sub>):** At zero nitrogen pressure, the film exhibits a weak, broad peak at approximately 38.8°, consistent with the (002) reflection of hexagonal



**Figure 4.11.** X-ray diffraction patterns showing hexagonal TiAl phase at 0 Pa and cubic  $(\text{Ti},\text{Al})\text{N}$  phase at 0.1–0.3 Pa nitrogen pressure. Films deposited at 14 cm with 0.25 T magnetic field. Si substrate peaks at 52–57°.

$\text{Ti}_{0.7}\text{Al}_{0.3}$  (space group  $\text{P}6_3/\text{mmc}$ , PDF 04-004-9157).

**Reactive mode (0.1–0.3 Pa  $\text{N}_2$ ):** Introduction of nitrogen fundamentally alters the film structure. At 0.1 Pa  $\text{N}_2$ , three distinct peaks emerge at approximately 37°, 43°, and 62°, corresponding to the (111), (200), and (220) reflections of the cubic B1 ( $\text{NaCl}$ -type)  $(\text{Ti}_{0.65}\text{Al}_{0.35})\text{N}$  phase (space group  $\text{Fm}\bar{3}\text{m}$ , PDF 04-017-5094). The hexagonal TiAl peak disappears completely, indicating full conversion to the nitride phase.

As the nitrogen pressure increases from 0.1 Pa to 0.3 Pa, the peak positions shift progressively closer to the reference values. Small deviations at 0.1 Pa correlate with the lower nitrogen content measured by EDX (33 at.% vs 39 at.%). The reference cubic nitride exhibits a lattice parameter of  $a = 4.202 \text{ \AA}$  and a calculated density of  $4.8 \text{ g cm}^{-3}$  [29], which will be compared with the film densities determined by XRR (Section 4.3.3).

The sharp, well-defined (111), (200), and (220) reflections in all nitride films indicate crystalline growth in Zone T or Zone 1 of Anders' structure zone model [13], where ion energies of 30–50 eV (measured by ERMS, Section 4.2) enable room-temperature crystallization through energetic condensation.

### 4.3.3 X-ray reflectometry (XRR)

X-ray reflectometry measurements were performed to determine film thickness and mass density for selected samples. Table 4.2 presents measurements at both substrate positions (QCM and ion current probe locations) for films 10, 11, 13, and 14. These measurements provide independent verification of the profilometry results and additionally yield quantitative mass density values.

**Table 4.2.** X-ray reflectometry provides independent thickness validation and reveals density variations. Mass densities extracted from critical angle and Kiessig fringe spacing range from 4.00 (for TiAl) to 4.8 g cm<sup>-3</sup>, reaching theoretical TiAlN bulk density (4.8 g cm<sup>-3</sup>).

Film ID	Pressure (Pa)	Thickness (nm)		Mass density (g/cm <sup>3</sup> )
		QCM	Ion current	
10	0.0	63	58	4.00
13	0.1	46.5	38	4.80 / 4.75
14	0.2	42	34.5	4.80 / 4.75
11	0.3	45.5	39	4.65

A systematic thickness gradient is observed between the two substrate positions, with the ion current location consistently showing 5–8.5 nm lower thickness compared to the QCM position, corresponding to a reduction of 8–18%. This spatial non-uniformity is consistent with the profilometry measurements presented in Table 4.1, where thickness differences of 5–18 nm between the two positions were observed across all samples.

The XRR thickness values are on average 10% lower than profilometry measurements. This agreement validates both measurement techniques and confirms that the observed gradient is a real spatial variation in the amount of deposited particles at each measurement device and can therefore be accounted and corrected for.

The introduction of nitrogen substantially increases the film mass density from 4.00 g/cm<sup>3</sup> in metallic mode to 4.65–4.80 g/cm<sup>3</sup> in reactive mode. This density increase reflects the transition from metallic TiAl to a dense cubic TiAlN nitride phase. The measured densities in reactive mode closely approach the reference value of 4.8 g/cm<sup>3</sup> for stoichiometric cubic TiAlN [29], indicating nearly complete densification and crystallization.

#### 4.3.4 Energy dispersive X-ray spectroscopy (EDX)

Energy dispersive X-ray spectroscopy was performed to determine the elemental composition of the deposited films as a function of nitrogen pressure. All films were deposited at 14 cm distance with a magnetic field of 0.25 T. Three measurement points were taken on each sample and averaged to obtain representative composition values.

**Table 4.3.** Elemental composition (at.%) and effective molar mass of films deposited at varying nitrogen pressures. Data obtained from EDX measurements averaged over three positions per sample. All films deposited at 14 cm distance with 0.25 T magnetic field.

Nitrogen (Pa)	N (at.%)	Al (at.%)	Ti (at.%)	$M_{\text{eff}}$ (g · mol <sup>-1</sup> )
0.0	0	21	79	43.5
0.1	33	17	50	33.2
0.2	38	15	47	31.9
0.3	39	14	47	31.8

The results in Table 4.3 reveal several observations:

**Metallic mode (0 Pa):** In vacuum deposition, the film composition shows a Ti:Al ratio of approximately 79:21, compared to the cathode composition (62.8 at.% Ti; 37.2 at.% Al). The effective molar mass is 43.5 g·mol<sup>-1</sup>, between pure aluminum (27.0 g·mol<sup>-1</sup>) and pure titanium (47.9 g·mol<sup>-1</sup>).

**Reactive mode (0.1–0.3 Pa N<sub>2</sub>):** Nitrogen incorporation reaches approximately 33 at.% at 0.1 Pa and saturates near 39 at.% above 0.2 Pa. The final composition at 0.3 Pa is Ti<sub>0.47</sub>Al<sub>0.14</sub>N<sub>0.39</sub>, close to the reference material composition used for XRD analysis [29]. The Ti:Al ratio shifts from approximately 3.75:1 in metallic mode to approximately 3.4:1 in reactive mode.

**Effective molar mass trends:** The effective molar mass decreases from 43.5 g·mol<sup>-1</sup> (metallic mode) to approximately 32 g·mol<sup>-1</sup> (reactive mode), a reduction of approximately 25%. The molar mass stabilizes above 0.2 Pa at  $M_{\text{eff}} \approx 31.8$  g·mol<sup>-1</sup>, consistent with the compositional saturation observed in nitrogen content.

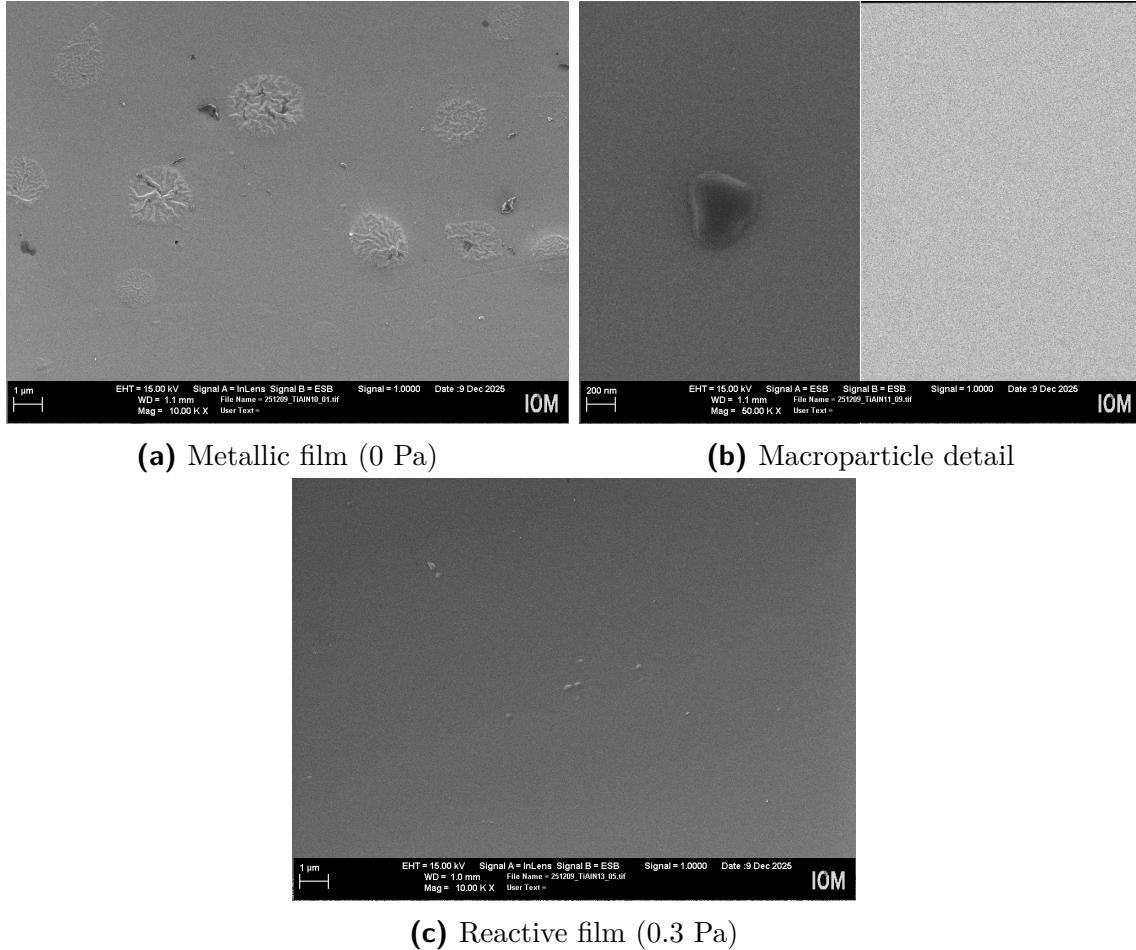
The effective molar mass is calculated from the atomic fractions  $x_i$  and atomic masses  $M_i$  according to:

$$M_{\text{eff}} = \sum_i x_i M_i = x_{\text{Ti}} \cdot 47.867 + x_{\text{Al}} \cdot 26.982 + x_{\text{N}} \cdot 14.007 \quad (4.1)$$

where the atomic fractions sum to unity ( $x_{\text{Ti}} + x_{\text{Al}} + x_{\text{N}} = 1$ ).

### 4.3.5 Scanning electron microscopy (SEM)

Scanning electron microscopy was performed to examine the surface morphology of selected TiAl and TiAlN films. Figure 4.12 shows representative images for films deposited under different conditions.



**Figure 4.12.** SEM images of (a) metallic TiAl film deposited at 0 Pa, (b) macroparticle on TiAlN film showing secondary electron (left) and backscattered electron (right) imaging at 50,000 $\times$  magnification, and (c) reactive TiAlN film deposited at 0.3 Pa nitrogen pressure. All films deposited at 14 cm distance with 0.25 T magnetic field.

The metallic TiAl film [Fig. 4.12(a)] exhibits localized surface swelling features, likely resulting from stress relaxation induced by energetic ion bombardment ( $E_{\text{tot}} \approx 60$  eV). Scattered macroparticles are visible as bright spots across the surface, consistent with typical cathodic arc deposition despite magnetic filtering.

The high-magnification image [Fig. 4.12(b)] shows a direct comparison between secondary electron imaging (left, sensitive to topography) and backscattered electron imaging (right, sensitive to atomic number contrast) of the same macroparticle on a TiAlN film. The absence of compositional contrast in the BSE image confirms that the

macroparticle has the same elemental composition as the surrounding film, indicating it originated from molten cathode material rather than external contamination.

The reactive TiAlN film [Fig. 4.12(c)] macroparticle are present, but lower quantities than for the metallic case. The overall surface remains relatively smooth between defects, with no evidence of large-scale roughening or columnar grain boundaries at this magnification.

These observations confirm that the magnetic macroparticle filter effectively reduces but does not completely eliminate macroparticle contamination, consistent with the filtered cathodic arc deposition process described in Chapter 3.

## 4.4 Particle flux calculations

The ion current and mass deposition rate measurements represent electrical and mass quantities that must be converted to particle fluxes using charge-state data from ERMS (Section 4.2) and compositional data from EDX (Section 4.3.4). This section presents the derived ion flux  $\Gamma_{\text{ion}}$  and total atomic flux  $\Phi_{\text{total}}$  for the 14 cm deposition distance.

### 4.4.1 Calculation methods

The ion flux ( $\text{ions}\cdot\text{cm}^{-2}\cdot\text{s}^{-1}$ ) is calculated from the ion current measurements:

$$\Gamma_{\text{ion}} = \frac{I_{\text{ion}}}{e\langle Q \rangle A_{\text{probe}}} \quad (4.2)$$

where  $I_{\text{ion}}$  is the measured ion current,  $e = 1.602 \times 10^{-19} \text{ C}$ ,  $A_{\text{probe}} = 0.196 \text{ cm}^2$ , and  $\langle Q \rangle$  is the weighted mean charge state from ERMS:

$$\langle Q \rangle = \frac{\sum_i Q_i \cdot I_i^{\text{ERMS}}}{\sum_i I_i^{\text{ERMS}}} \quad (4.3)$$

where the sum includes all charge states of Ti and Al.

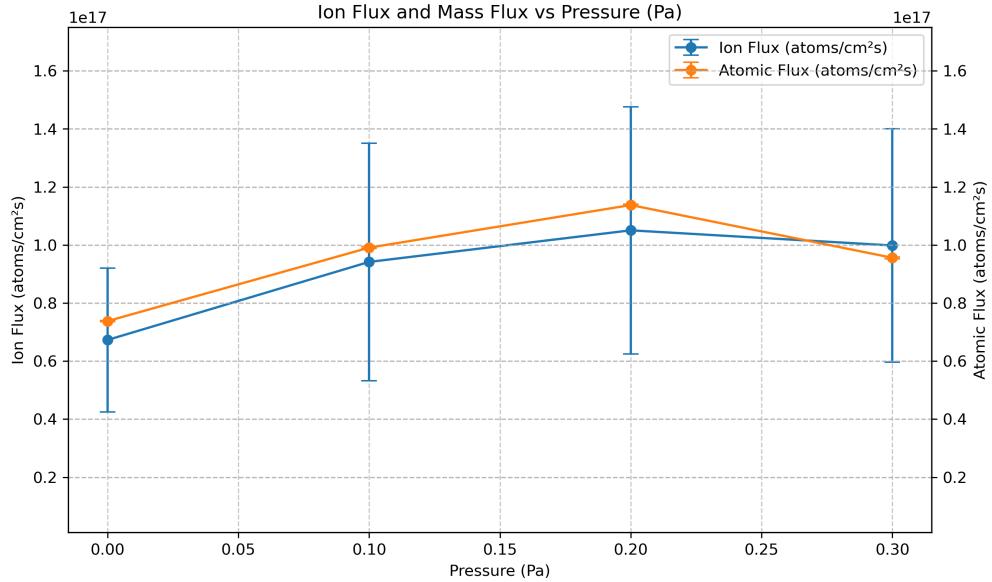
The total atomic flux ( $\text{atoms}\cdot\text{cm}^{-2}\cdot\text{s}^{-1}$ ) is calculated from QCM measurements:

$$\Phi_{\text{total}} = \frac{\Delta m \cdot N_A}{A_{\text{QCM}} \cdot \Delta t \cdot M_{\text{eff}}} \quad (4.4)$$

where  $\Delta m$  is the mass change over time interval  $\Delta t$ ,  $A_{\text{QCM}} = 1.54 \text{ cm}^2$ ,  $N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$ , and  $M_{\text{eff}}$  is the effective molar mass from Table 4.3.

## 4.4.2 Flux measurements at 14 cm distance

Figure 4.13 presents the calculated ion flux and atomic flux as functions of nitrogen pressure at 14 cm distance with 0.25 T magnetic field.



**Figure 4.13.** Ion flux and atomic flux as functions of nitrogen pressure at 14 cm distance with 0.25 T magnetic field. Error bars represent combined uncertainties from current measurements and charge-state determination (ERMS).

The ion flux shows an increase with nitrogen pressure, from  $(0.7 \pm 0.2) \times 10^{17}$  ions·cm $^{-2}$ ·s $^{-1}$  in metallic mode to  $(1.0 \pm 0.3) \times 10^{17}$  ions·cm $^{-2}$ ·s $^{-1}$  at 0.1 Pa, remaining relatively constant at higher pressures. The error bars ( $\pm 30\text{--}40\%$ ) reflect combined uncertainties from ion flux variation within a pulse from the ion current probe measurements and ERMS charge-state determination A.1. The overall trend mirrors the ion current behavior observed in Figure 4.5b, where current initially increases from 0 to 0.1 Pa due to additional nitrogen ion production before saturating at higher pressures. The reduction in atomic flux relative to ion flux at higher reactive gas pressures may be partially explained by reduced sticking coefficients under these conditions.

The atomic flux exhibits clear non-monotonic pressure dependence, increasing from  $0.75 \times 10^{17}$  atoms·cm $^{-2}$ ·s $^{-1}$  in metallic mode to a maximum of  $1.15 \times 10^{17}$  atoms·cm $^{-2}$ ·s $^{-1}$  at 0.2 Pa (a 53% increase), before decreasing to  $0.95 \times 10^{17}$  atoms·cm $^{-2}$ ·s $^{-1}$  at 0.3 Pa.

# CHAPTER 5

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## Discussion of Results

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This chapter interprets the experimental findings presented in Chapter 4 and connects them to the theoretical framework established in Chapter 2. The discussion is organized around two central themes: film growth mechanisms and plasma transport.

### 5.1 Film Growth and Microstructure

XRD results demonstrate that nitrogen incorporation enables room-temperature crystallization of TiAlN. The metallic TiAl film (0 Pa) exhibits only weak hexagonal reflections, while all reactive mode films (0.1–0.3 Pa nitrogen) develop sharp cubic TiAlN (111) and (200) reflections.

XRR measurements show the measured densities are in good agreement with reference values for their respective crystal structures (Table 4.2). The metallic TiAl film shows  $4.00 \text{ g}\cdot\text{cm}^{-3}$ , while reactive mode films reach  $4.65\text{--}4.80 \text{ g}\cdot\text{cm}^{-3}$ , similar to the cubic TiAlN reference value of  $4.8 \text{ g}\cdot\text{cm}^{-3}$  [29].

EDX measurements show systematic deviations from cathode stoichiometry (Table 4.3). In metallic mode, the Ti:Al ratio shifts from 62.8:37.2 (cathode) to 79:21 (film), likely arising from differential erosion rates, ionization efficiencies, or transport through the magnetic filter. In reactive mode, the ratio shifts to approximately 3.4:1, closer to cathode composition. Nitrogen saturation at 39 at.% (0.2–0.3 Pa) approaches stoichiometric  $(\text{Ti},\text{Al})\text{N}$  (50 at.% N).

XRR and profilometry reveal 5–8.5 nm thickness gradients between measurement

positions (Tables 4.1–4.2), attributable to plasma transport through the curved magnetic filter. The 90 degree filter guides electrons along field lines while ions follow electrostatically. Different species follow slightly different trajectories, leading to potential spatial variations in flux and composition [31].

## 5.2 Plasma transport and magnetic field effects

The anomalous decrease in both mass deposition rate and ion current at 0.05 T magnetic field (Figure 4.2) reflects a magnetic mirror effect at the entrance of the EM-coil, where plasma transitions from the weak fringing field (approximately 0.01 T) outside into the stronger field (0.05 T) inside. Electrons with sufficient perpendicular velocity components are reflected, creating a localized space-charge layer that temporarily retards ion flow [3]. At higher fields (0.1 T and above), the beneficial effects of plasma compression and enhanced ionization overcome this mirror loss, restoring monotonic increase. This effect persists across all distances and pressures, indicating it is a fundamental feature of the plasma-coil interaction.

In metallic mode, mass deposition rate follows approximate  $1/r^2$  dependence (Figure 4.3a), consistent with free plasma expansion. Ion current decreases more steeply in reactive mode due to charge-exchange neutralization. The magnetic field enhancement increases both mass flux and ion current by factors of 5–8 from zero field to 0.25 T (Figure 4.4), demonstrating effective plasma enhancement.

## 5.3 Reactive mode plasma chemistry

The pressure dependence of mass flux exhibits complex behavior at 10 cm (Figure 4.5b), with an initial decrease from 0 to 0.05 Pa before decreasing at higher pressures. At low nitrogen pressures (0–0.1 Pa), enhanced plasma generation at the cathode from altered spot dynamics and nitrogen ion production contribute to increased measured mass flux. The same cannot be said for the ion current, which is gradually decreasing with increasing nitrogen pressure. This is due to the many reactions possible with the background gas, reduction of the metallic species can lead to ionization, dissociation and excitations.

ERMS measurements confirm this mechanism: mean charge states decrease from  $\langle Q \rangle \approx 2.4$  (metallic) to  $\approx 1.6$  (0.3 Pa reactive), with charge-state distributions shifting toward lower ionization levels (Figures 4.8–4.9). This represents a 35% reduction in average charge state. Importantly, total ion energy (kinetic plus potential) remains

above 30 eV across all conditions, sufficient for subplantation processes.

Table 5.1 summarizes the measured plasma parameters and resulting film properties across the pressure range at 14 cm distance with 0.25 T magnetic field.

**Table 5.1.** Plasma parameters and film properties correlation at 14 cm, 0.25 T

Pressure (Pa)	Ion Flux ( $10^{17}$ cm $^{-2}$ s $^{-1}$ )	Atomic Flux ( $10^{17}$ cm $^{-2}$ s $^{-1}$ )	$\langle Q \rangle$	Density (g cm $^{-3}$ )	XRD Quality
0.0	0.7	0.75	2.4	4.00	weak hex
0.1	0.95	1.0	2.0	4.80	strong cubic
0.2	1.05	1.15	1.8	4.80	strong cubic
0.3	1.0	0.95	1.6	4.65	strong cubic

The ionization fraction  $f_{\text{ion}} = \Gamma_{\text{ion}}/\Phi_{\text{total}}$  remains between 0.9 and 1.1 across all pressures. This high ionization persists even as mean charge states decrease by 35% from metallic to reactive mode, confirming that charge-exchange reduces individual ion charges without substantial neutralization of the metal plasma.

Atomic flux exhibits non-monotonic pressure dependence with maximum at 0.2 Pa (54% increase over metallic mode), while ion flux remains approximately constant. This indicates that reactive film formation enhances mass deposition through incorporation of nitrogen atoms from the gas phase, while the metal ion flux remains determined primarily by arc current and magnetic field. The flux minimum at 0.3 Pa may relate to pressure-dependent sticking coefficients, as observed by Mahieu et al. for sputtered atoms [32]. Further systematic pressure variation studies are needed to clarify this effect.

### 5.3.1 Activated nitrogen species

EDX measurements (Table 4.3) show substantial nitrogen incorporation in reactive mode films (33–39 at.%), while ERMS detects only modest N $^+$ /N $_2^+$  signals (10–28 at.%) compared to metal ions. The imbalance suggests that nitrogen has some excited species but since ERMS estimates the nitrogen contents via the overall species counts, therefore the imbalance can only be used as a hint towards activated nitrogen.

## 5.4 Summary

This work established a comprehensive methodology for characterizing magnetically filtered cathodic arc plasmas using combined QCM and ion probe measurements, supplemented by ERMS charge-state analysis and film characterization via XRD, XRR, and EDX. The approach was validated for the TiAl cathode system across metallic and reactive nitrogen modes, demonstrating near-unity ionization fractions ( $f_{\text{ion}} = 0.92\text{--}1.05$ ) and revealing the dominant role of ions in mass transport.

The methodology enabled quantification of ion and atomic fluxes under varying magnetic field (0–0.25 T), distance (10–14 cm), and nitrogen pressure (0–0.3 Pa) conditions. Magnetic field enhancement factors of 5–8 were demonstrated, while charge-exchange reactions in reactive mode reduced mean charge states by 35% without equally decreasing the ion flux. Film analysis confirmed that high ionization fractions and ion energies exceeding 30 eV enable room-temperature crystallization of dense cubic TiAlN films.

## 5.5 Outlook

Further work should focus on characterizing neutral species contributions to identify potential gaps in the flux balance. Improved measurement stability with more consistent ion currents would help resolve subtle trends obscured by current error bars. Systematic investigation of pressure-dependent sticking coefficients is needed to explain the flux behavior at higher pressures. These studies should begin with TiN as a simpler binary system before extending to multi-component materials like TiAlN.

# CHAPTER 6

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Secondly I want to thank

Lastly I want to thank my supporting family

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# APPENDIX A

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## Appendix: Experimental Methods supplementary

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### A.1 Longer QCM depositions

For larger mass loadings of the QCM, the linear approximation fails and the Z-match™ technique is used. This method, introduced by Lu and Lewis in 1972 on the basis of Miller and Bolef's theoretical treatment [33, 34], incorporates the acoustic properties of both the quartz and the deposited film via the acoustic impedance ratio

$$Z = \left( \frac{d_q \mu_q}{d_f \mu_f} \right)^{1/2}, \quad (\text{A.1})$$

with  $d$  and  $\mu$  denoting the density and shear modulus of quartz ( $q$ ) and film ( $f$ ), respectively [35]. In practice, the controller applies a correction function  $f(Z)$  to the Sauerbrey relation,

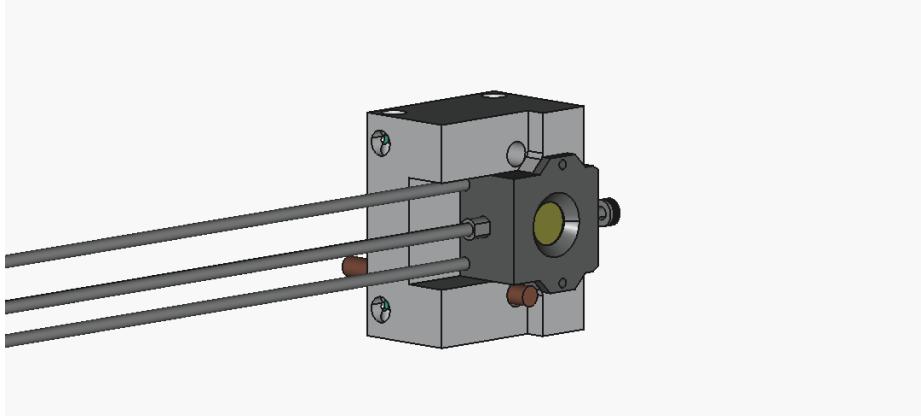
$$m_f = \frac{N_{\text{AT}} d_q \pi r^2}{F_q^2} \cdot \Delta F \cdot f(Z), \quad (\text{A.2})$$

which compensates for the acoustic mismatch and extends the validity of thickness determination up to  $\sim 0.4F_q$ .

### A.2 Holder assembly design

The integrated holder assembly (Figure A.1) enabled simultaneous QCM and ion probe measurements with minimal spatial separation. The aluminum mount positioned both diagnostics as close as practicable to bring the flux and plasma conditions as close as

possible between measurement locations.



**Figure A.1.** Holder assembly for In-Situ plasma Diagnostics: Integrated QCMs and Ion Collector Probe (interactive 3D model; static preview shown in non-Adobe viewers).

## A.3 Data processing workflow

Experimental data were processed using custom Python scripts to ensure consistency and reproducibility across all measurements. This section describes the data handling procedures and processing workflows used throughout this work.

### A.3.1 Data organization and logbook system

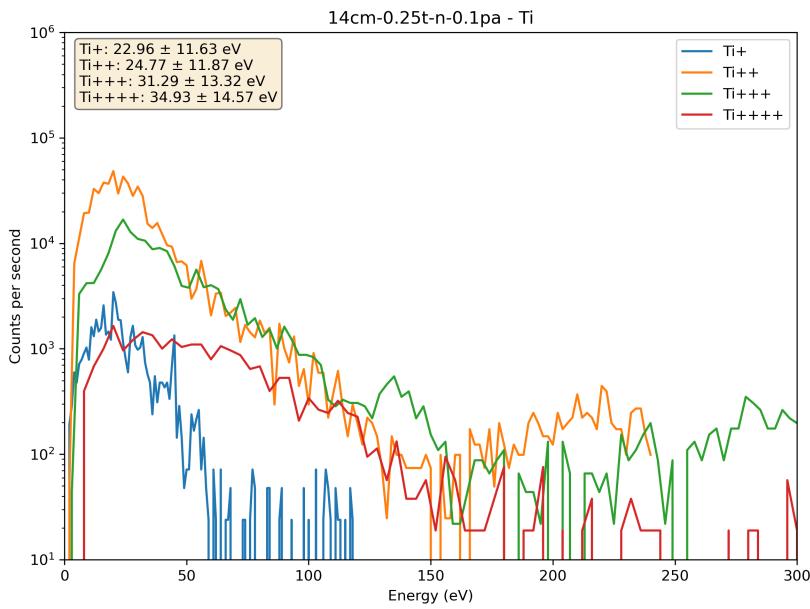
A central Excel logbook served as the reference for all measurements, with each measurement identified by a unique suffix and linked to its corresponding data files. The logbook recorded:

- Date and time of measurement
- Spatial parameters: distance from macroparticle filter (10–20 cm)
- Gas parameters: N<sub>2</sub> flow rate (MFC setting in sccm), working pressure (Pa)
- Vacuum system parameters: cryopump gate valve position, base pressure
- Power supply settings: arc voltage and current, EM-coil voltage and current
- Pulse characteristics: frequency (Hz), pulse width (ms), number of pulses
- QCM frequencies: initial ( $f_0$ ) and final ( $f_1$ ) values for mass determination

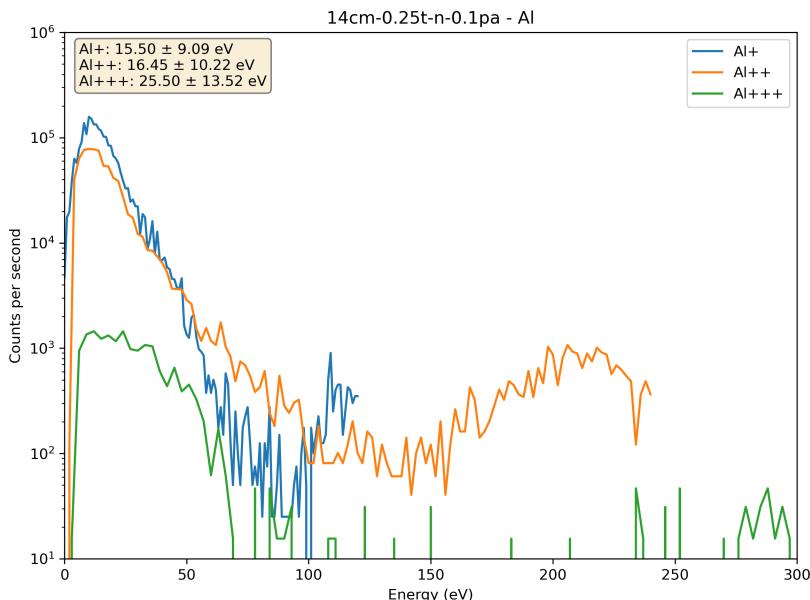
Associated oscilloscope waveforms were stored as CSV files for ion current analysis, with filenames linked to the logbook suffix for traceability.

### A.3.2 Ion current data processing

Ion current waveforms were recorded using a Tektronix MSO64 oscilloscope. For each measurement condition, multiple pulses were recorded and processed as follows:



**Figure A.2.** Charge-state-resolved ion energy distribution functions for  $\text{Ti}^+$  through  $\text{Ti}^{4+}$  measured at 14 cm distance with 0.25 T magnetic field and 0.1 Pa  $\text{N}_2$  pressure. Mean kinetic energies and standard deviations are shown in the legend.



**Figure A.3.** Charge-state-resolved ion energy distribution functions for  $\text{Al}^+$  through  $\text{Al}^{3+}$  measured at 14 cm distance with 0.25 T magnetic field and 0.1 Pa  $\text{N}_2$  pressure. Mean kinetic energies and standard deviations are shown in the legend.

- Pulse averaging:** Individual pulse waveforms were averaged to obtain the mean ion current waveform.
- Time-integrated current:** The mean ion current over the pulse duration (0–1 ms) was calculated by integrating the averaged waveform and dividing by the pulse width.

The Python script automatically matched logbook entries to oscilloscope CSV files and compiled all parameters into CSV files for analysis.

### A.3.3 QCM data processing

QCM frequency measurements were recorded before and after each set of pulses. The deposited mass was calculated using the Sauerbrey equation (Equation 3.5). The mass deposition rate was obtained by dividing by the number of pulses and the crystal area.

### A.3.4 ERMS data processing

The ERMS data processing was implemented in Python and performed the following operations on the raw spectral data. The underlying measurement principles and correction methodologies are described in Section 3.2.4, while this section focuses on the computational implementation.

- Spectral integration:** Raw energy distribution functions (EDFs) for each  $M/Q$  value were averaged over two measurement windows of 20 ms duration each. This averaging improves signal-to-noise ratio while maintaining synchronization with the pulsed arc operation (1 ms pulse duration, 5 Hz repetition rate).
- Mass transmission correction:** Correction factors  $T(M/Q)$  were applied to the raw ion signal intensities according to Equation 3.10. The correction function accounts for the mass-dependent detection efficiency of the quadrupole mass filter, which systematically affects lighter ions (Al) differently than heavier ions (Ti). The transmission function was determined through calibration measurements and validated against the known cathode stoichiometry ( $\text{Ti}_{0.75}\text{Al}_{0.25}$ ).
- Energy conversion:** The ERMS measures energy-per-charge ( $E/Q$ ) distributions. For multiply charged ions, the measured  $E/Q$  values were multiplied by the charge state  $Q$  to obtain the actual ion energy in eV (Equation 3.9). This conversion was performed for charge states  $Q = 1^+, 2^+, 3^+$  (Al) and  $Q = 1^+, 2^+, 3^+, 4^+$  (Ti).
- Statistical extraction:** For each charge-state-resolved energy distribution, the following statistical parameters were calculated:

- Mean ion energy  $\langle E \rangle$  - weighted average over the distribution
- Standard deviation  $\sigma_E$  - characterizing the energy spread
- Peak energy  $E_{\text{peak}}$  - the most probable energy (distribution maximum)

The mean ion energy represents the kinetic energy of ions in that charge state.

5. **Mean charge state calculation:** The mean charge state  $\langle Q \rangle$  for each element was calculated as a weighted average over all detected charge states using Equation 3.18. This parameter is essential for converting the ion current measurements to particle flux (Section 3.4).
6. **Potential energy calculation:** The potential energy  $E_{\text{pot}}$  for each charge state was calculated separately from the cohesive energy and ionization potentials of the cathode material, following Kalanov et al. [6]:

$$E_{\text{pot}}(Q) = E_{\text{coh}} + C \sum_{i=1}^Q E_{\text{ion},i} \quad (\text{A.3})$$

where  $E_{\text{coh}}$  is the cohesive energy (4.85 eV for Ti, 3.39 eV for Al),  $E_{\text{ion},i}$  are the successive ionization energies, and  $C = 0.6$  is a factor accounting for the fraction of ionization energy deposited in the solid. The mean potential energy was then calculated as a weighted average over charge states.

7. **Energy analysis output:** The kinetic energy (measured) and potential energy (calculated) for each element were exported along with their standard deviations. These independent energy components enable investigation of their respective effects on film structure and properties.

The processed results were exported to CSV files containing: pattern identification, mean charge states, mean kinetic energies, mean potential energies, and charge-state-resolved energies for each element. These data were correlated with ion probe and QCM measurements for comprehensive flux and energy analysis. Example charge-state-resolved ion energy distribution functions for titanium and aluminum ions are shown in Figures A.2 and the following figure, measured at 14 cm distance with 0.25 T magnetic field and 0.1 Pa N<sub>2</sub> pressure.

## A.4 Uncertainty analysis

This section documents the uncertainty analysis for measurements reported in this work. Uncertainties were determined from three primary sources: instrument specifications, statistical variations in repeated measurements, and propagation through derived quantities.

### A.4.1 Primary measurements

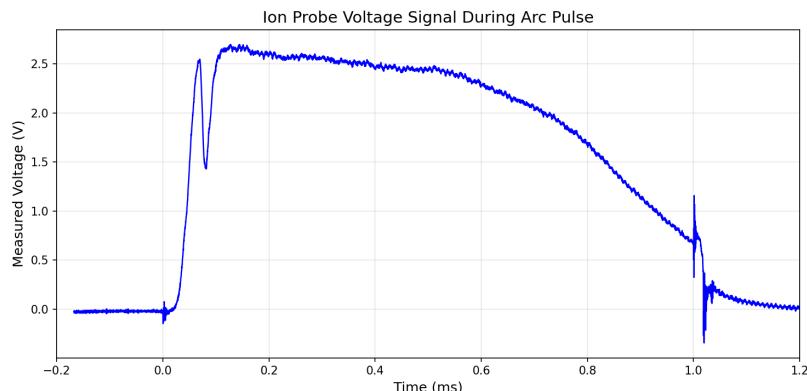
#### QCM mass measurements

The QCM frequency measurements have a manufacturer-specified resolution that translates to a mass uncertainty of approximately  $\pm 5.64 \times 10^{-10}$  g per measurement. This corresponds to a relative uncertainty of approximately 0.1% for typical deposited masses. This instrumental precision is negligible compared to other uncertainty sources in this work.

The number of pulses (64 in most cases) was controlled digitally and has negligible counting uncertainty. The crystal area ( $1.54 \text{ cm}^2$ ) is specified by the manufacturer.

#### Ion current measurements

Ion current measurements exhibited significant variability due to the non-stationary nature of cathodic arc plasmas. The standard deviation of time-averaged current within individual pulses ranged from 33% to 68% of the mean value (relative uncertainty), with a mean of 45% across all conditions measured.



**Figure A.4.** Example ion current waveform showing variability within a pulse. The voltage signal measured across the  $400 \Omega$  resistor is proportional to the ion current. Rapid transients at pulse ignition (0 ms) and termination (1 ms) are excluded from the time-averaged current calculation.

This large variability arises from cathode spot motion and splitting or extinction events, plasma potential fluctuations during pulse ignition and termination, and charge-exchange collisions with background gas that depend on pressure. The oscilloscope (Tektronix MSO64) has a specified voltage measurement accuracy of less than 1%, which is negligible compared to the plasma variability. Error bars on ion current measurements represent the within-pulse standard deviation and characterize the inherent variability of the arc discharge under nominally constant operating conditions.

## ERMS measurements

Mean charge states determined from energy-resolved mass spectrometry (Equation 3.18) have uncertainties ranging from 14% to 24% (relative), with a mean of 20%. This value is calculated as the weighted standard deviation of the charge state distribution:

$$\sigma_Q = \sqrt{\sum_Q w_Q(Q - \langle Q \rangle)^2} \quad (\text{A.4})$$

where  $w_Q = I_Q / \sum I_Q$  are the normalized signal intensities after mass transmission correction. This represents the spread of the charge state distribution and is used in error propagation for flux calculations.

### A.4.2 Derived quantities

Ion flux (calculated from ion current and mean charge state) has significantly larger uncertainties. For ion flux  $\Phi_i = I_{\text{ion}}/(Q \cdot e \cdot A)$ , the relative uncertainty is

$$\frac{\sigma_{\Phi_i}}{\Phi_i} = \sqrt{\left(\frac{\sigma_I}{I}\right)^2 + \left(\frac{\sigma_Q}{Q}\right)^2}, \quad (\text{A.5})$$

where uncertainties were propagated using standard error propagation for uncorrelated variables. This quantity is dominated by the ion current uncertainty (33–68%) and charge state uncertainty (14–24%), resulting in combined uncertainties of 36–72% depending on conditions.

Atomic flux was calculated by converting mass flux to atomic flux using the effective molar mass (determined from film composition measured by ex-situ EDX analysis) and Avogadro’s number. Since composition uncertainties were not formally characterized, the reported atomic flux values do not include propagated uncertainties.

### A.4.3 Uncertainties not quantified

Several systematic uncertainty sources were not explicitly quantified in this work. Spatial plasma uniformity across the measurement area is estimated to contribute less than 10% based on probe size relative to plasma diameter. Long-term calibration drift is expected to be less than 5% as all measurements were completed within 3 months. Temperature effects on QCM sensitivity are minimal since chamber temperature remained stable within  $\pm 2^\circ\text{C}$  and the QCM temperature coefficient is less than 0.1% per degree. Geometric alignment uncertainties in probe positioning of  $\pm 2$  cm correspond to less than 25% flux variation at typical distances.

These effects are expected to be small compared to the dominant uncertainty sources identified above but were not independently characterized.

#### A.4.4 Summary

Table A.1 summarizes the typical uncertainties for key measured and derived quantities. All error bars in figures represent these uncertainties unless otherwise noted. The large uncertainties in ion-based measurements reflect the intrinsic variability of cathodic arc plasmas rather than instrumental limitations.

**Table A.1.** Summary of measurement uncertainties

Quantity	Relative Uncertainty	Dominant Source
QCM mass	0.1%	Instrument precision
Ion current	33–68% (mean: 45%)	Plasma variability
Mean charge state	14–24% (mean: 20%)	Charge state spread
Mass flux	0.1%	QCM precision
Ion flux	36–72%	Ion current + charge state
Atomic flux	Not reported	Composition error not characterized

## A.5 Langmuir probe bias voltage characterization

Before plasma investigations were performed, the functionality of the custom-built ion current probe was verified by characterizing its voltage-current relationship and determining the appropriate operating voltage within the ion saturation regime. This characterization ensures that the probe operates in a regime where the collected current is dominated by ions while effectively repelling electrons.

The bias voltage measurements were conducted without nitrogen in the chamber, at a fixed distance of 10 cm away from the filter, and with an EM-coil strength of 0.25 T.

The relationship between the collected current  $I$  and the bias voltage  $V$  was analyzed using the modified Langmuir equation [15]:

$$I = I_{\text{sat}} \left( 1 - e^{-V/V_0} \right) + m \cdot V \quad (\text{A.6})$$

where  $I_{\text{sat}}$  is the ion saturation current,  $V_0$  is a characteristic voltage, and the linear term  $m \cdot V$  accounts for sheath expansion and collisional effects at higher bias voltages.

Figure A.5 displays the measured ion current as a function of bias voltage. The blue data points represent experimental measurements, while the green curve shows the fit to Equation A.6 and the red dashed line shows the limit between the transition and the saturation regimes of the ion probe.

As the bias voltage increases, the sheath surrounding the probe expands. At higher voltages the sheath becomes non-planar, causing the probe to attract ions from a broader solid angle beyond the immediate frontal area. This geometric effect, combined with ion-neutral collisions within the extended sheath region, produces a residual linear increase in current even in the saturation regime. The linear correction term  $m \cdot V$  in Equation A.6 captures these effects, ensuring accurate modeling across the full voltage range [15, Chap. 7].

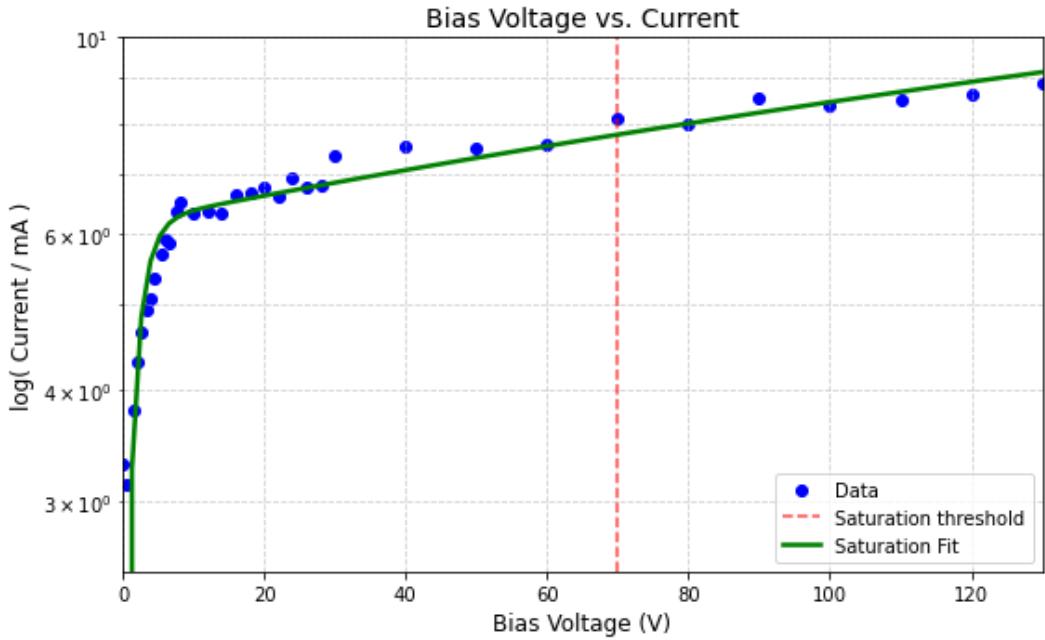
The experimental data were fitted using a nonlinear least-squares method in Python, yielding:

$$I_{\text{sat}} = 6.17 \text{ mA}, \quad V_0 = 1.8 \text{ V}, \quad m = 0.023 \text{ mA/V}.$$

The voltage-current characteristic exhibits two distinct regimes:

### 1. Transition Regime (0–70 V):

At lower bias voltages, the probe collects both ions and electrons. As the negative bias increases, electrons are progressively repelled, reducing their contribution to



**Figure A.5.** Measured ion current vs. bias voltage, showing fit of eq: A.6. The vertical dashed line at 70 V marks the transition to full ion saturation. Process conditions: no gas, 10 cm from the macroparticle filter, 0.25 T EM-coil strength.

the measured current. This produces a rapid rise in the net ion current as electron repulsion becomes increasingly effective. The exponential term in Equation A.6 dominates this regime, reflecting the transition from mixed electron-ion collection to predominantly ion collection.

## 2. Ion Saturation Regime (70–130 V):

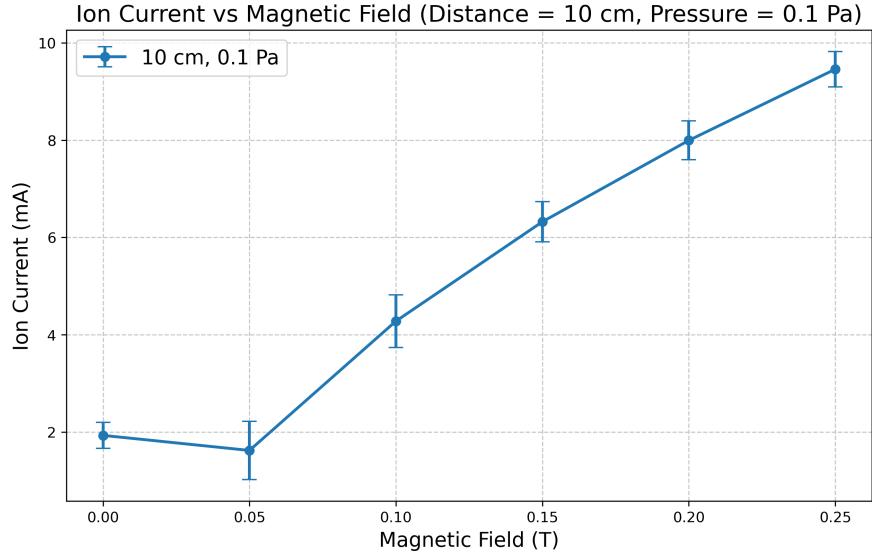
Beyond approximately 70 V, the current curve resembles more a plateau, indicating that the probe has entered the ion saturation regime. At these voltages, the negative bias effectively repels all electrons, and the measured current represents the ion flux to the probe surface. The slight residual increase in current with voltage is captured by the linear term  $m = 0.023 \text{ mA/V}$ , which accounts for sheath expansion and collisional effects as described above.

Based on this characterization, a bias voltage of  $-80 \text{ V}$  was selected for all subsequent plasma measurements. This voltage lies well within the ion saturation regime, ensuring complete electron repulsion.

## A.6 Ion current variation over different pulses

To analyze the pulse-to-pulse variation, approximately 30 single pulses were recorded for each magnetic field strength. The average ion current of each pulse was calculated over the 0–1 ms pulse interval. The final mean and standard deviation were then

determined by combining these averages, allowing for an assessment of the differences between pulses.



**Figure A.6.** Pulse to pulse variation in mean ion current at varying magnetic fields (0.1 Pa N<sub>2</sub>, 10 cm distance)

EM-coil field (T)	Mean $I_{\text{ion}}$ (mA)	Std (%)	Range (mA)	# Pulses
0.00	1.931	13.8	[1.13, 2.26]	30
0.05	1.622	37.1	[1.16, 4.85]	32
0.10	4.279	12.7	[2.72, 4.88]	31
0.15	6.323	6.6	[4.99, 6.92]	27
0.20	7.996	5.0	[7.26, 8.78]	22
0.25	9.460	3.8	[8.61, 10.02]	30

**Table A.2.** Summary of pulse measurement statistics for varying magnetic field strengths. Note: The 0.05 T data includes variability from unstable arc behavior at this field strength with a large outlier of 4.85 mA.

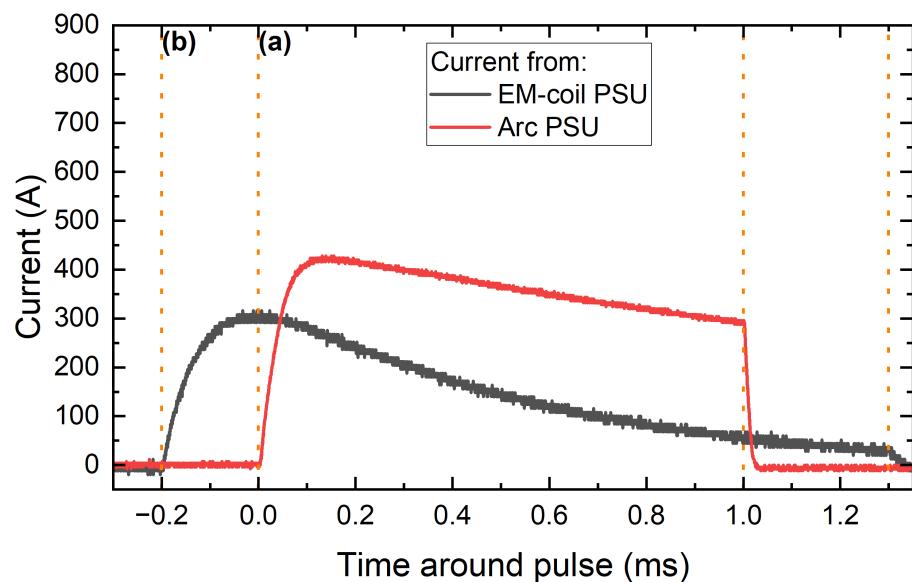
In all other measurements, the oscilloscope averaged the voltage signal over multiple pulses to improve signal-to-noise ratio. Therefore, the error reported in subsequent sections primarily reflects variations in ion current within individual pulses, rather than differences between distinct pulses.

## APPENDIX B

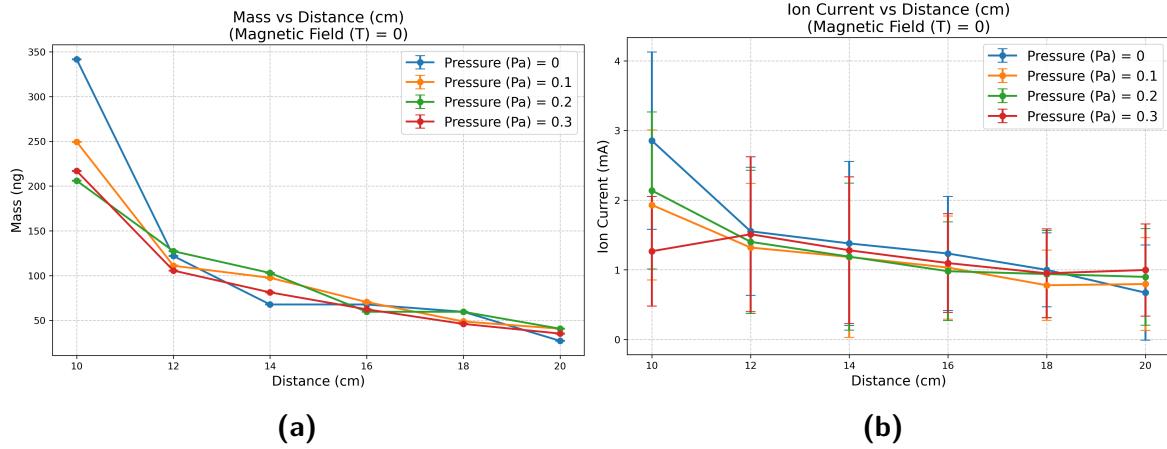
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### Appendix: supplementary material

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**Figure B.1.** Additional Pulse waveform with the triggering timings (a) and (b) for the Arc-PSU and the EM-coil PSU marked with the orange dashed line for 100V input



**Figure B.2.** Additional plot for Section 4.1.2. QCM and Ion Probe measurements showing (a) the deposited mass after 64 pulses and (b) the ion current averaged over a single pulse, each plotted as a function of distance. Data shown for representative pressures, with error bars representing variation within pulses.

## B.1 Data availability

All raw data files, processed datasets, and analysis scripts are archived and available upon request. The main Python scripts used for data processing include:

- [Ion current and mass csv generator.py](#): Ion current waveform analysis (mean and std) and qcm mass inclusion if provided.
- [Massspec final.py](#): ERMS data integration, mass transmission correction and energy states, mean charge states and potential aswell as kinetic energy extraction.
- [XRD plot stacked.py](#): plotting of stacked XRD plots from .txt files
- Data sets created and used for the Thesis

dunno if i will modify the csv merger script and add it since it is too specific maybe

**Table B.1.** Complete overview of experimental measurements. All distance/field/pressure combinations not explicitly listed were measured with ion current probe and QCM only. In total: 252 Ion+QCM measurements, 17 ERMS measurements, 4 deposited films

Distance	Magnetic Field	Pressure	Measurements Performed			Film ID
(cm)	(T)	(Pa)	Ion/QCM	Mass Spec	Ex-situ	
<i>Systematic parameter scan:</i>						
12, 16, 18	0, 0.05, 0.1, 0.15, 0.2, 0.25	0, 0.025, 0.05, 0.075, 0.1, 0.2, 0.3	✓	—	—	—
<i>Distance 10 cm — standard conditions:</i>						
10	0.05, 0.1, 0.2	all pressures <sup>†</sup>	✓	—	—	—
10	0	0.025, 0.05, 0.075, 0.2	✓	—	—	—
10	0.15	0.025, 0.05, 0.075, 0.2	✓	—	—	—
10	0.25	0.025, 0.05, 0.075	✓	—	—	—
<i>Distance 10 cm — with mass spectrometry:</i>						
10	0, 0.15	0, 0.1, 0.3	✓	✓	—	—
10	0.25	0	✓	✓	✓	003
10	0.25	0.1	✓	✓	✓	004
10	0.25	0.2	✓	✓	—	—
10	0.25	0.3	✓	✓	✓	002
<i>Distance 14 cm — standard conditions:</i>						
14	0, 0.05, 0.1, 0.2	all pressures <sup>†</sup>	✓	—	—	—
14	0.15	0, 0.025, 0.05, 0.075, 0.2, 0.3	✓	—	—	—
14	0.25	0.025, 0.05, 0.075, 0.2	✓	—	—	—
<i>Distance 14 cm — with mass spectrometry:</i>						
14	0.15	0.1	✓	✓	—	—
14	0.25	0	✓	✓	✓	008/010
14	0.25	0.1	✓	✓	✓	011/013
14	0.25	0.2	✓	✓	✓	014
14	0.25	0.3	✓	✓	✓	009/011
<i>Distance 20 cm — standard conditions:</i>						
20	0, 0.05, 0.1, 0.15, 0.2	all pressures <sup>†</sup>	✓	—	—	—
20	0.25	0.025, 0.05, 0.075, 0.1, 0.2	✓	—	—	—
<i>Distance 20 cm — with mass spectrometry:</i>						
20	0.25	0, 0.3	✓	✓	—	—

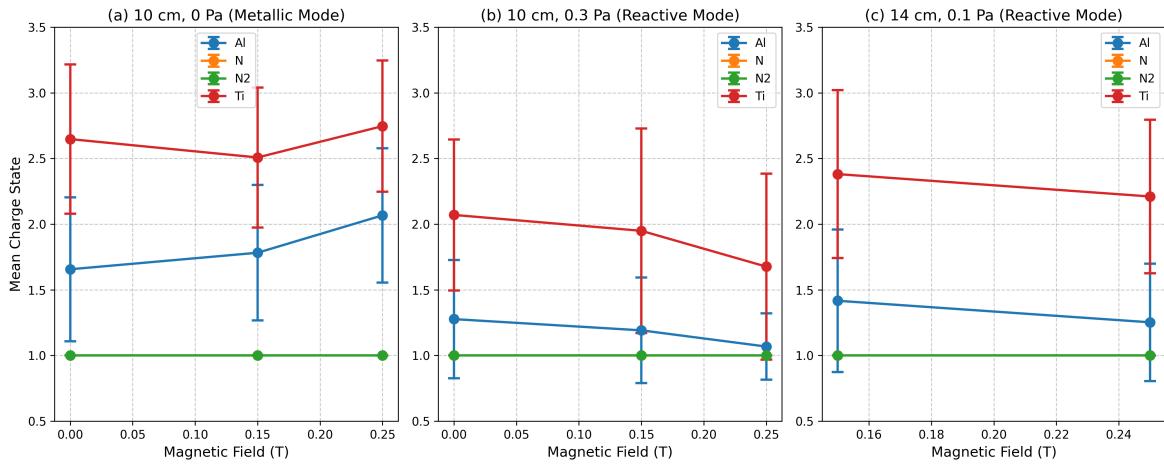
<sup>†</sup>all pressures = 0, 0.025, 0.05, 0.075, 0.1, 0.2, 0.3 Pa

Ion/QCM: Ion current probe and quartz crystal microbalance measured at the same time

Mass Spec: Energy-resolved mass spectrometry (ERMS)

Ex-situ: XRD, SEM, EDX, profilometry characterization of deposited films

Magnetic Field Enhancement of Ion Charge States



**Figure B.3.** Mean charge state as a function of magnetic field strength for Ti and Al ions under representative conditions: (a) 10 cm distance in metallic mode (0 Pa), (b) 10 cm distance in reactive mode (0.3 Pa), and (c) 14 cm distance in reactive mode (0.1 Pa). The systematic increase in charge states with magnetic field strength occurs across all conditions, demonstrating the inherent coupling between magnetic confinement and ion charge states. This coupling prevents independent control of ion flux and potential energy, as discussed in Section 5.