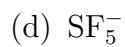
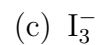
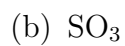


Lewis Structures and VSEPR Model

Nov 13, 2023

Draw the Lewis structure showing the formal charges. At each central atom, determine the electronic arrangement, and molecular geometry for the following compounds. Include all resonance structures and show the resonance hybrid structure:



(e) $\text{CH}_2\text{FCH}_2\text{COOH}$

(f) XeF_4

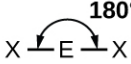
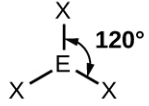
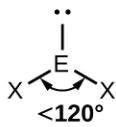
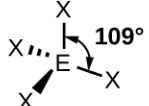
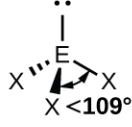

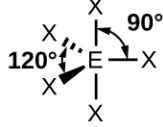
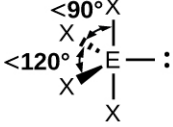
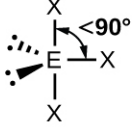
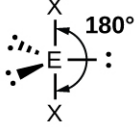

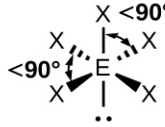
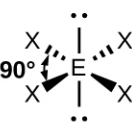
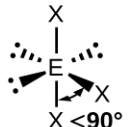

(g) CHF_3

(h) O_3

(i) HCCH

(j) CHONH_2

(k) PF_6

Number of electron pairs	Electron pair geometries: 0 lone pair	1 lone pair	2 lone pairs	3 lone pairs	4 lone pairs
2	 Linear				
3	 Trigonal planar	 Bent or angular			
4	 Tetrahedral	 Trigonal pyramid	 Bent or angular		
5	 Trigonal bipyramid	 Sawhorse or seesaw	 T-shape	 Linear	
6	 Octahedral	 Square pyramid	 Square planar	 T-shape	 Linear