

The reconciliation and validation of a combined interatomic potential for the description of Xe in γ U-Mo

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Abstract

A U-Mo alloy has been selected as the fuel design for the conversion of high-performance research reactors in the United States. Efforts are ongoing to describe the fuel evolution as a function of time, for a variety of different reactor conditions. The accurate prediction of fuel evolution under irradiation requires the implementation of correct thermodynamic properties into mesoscale and continuum-level fuel performance modeling codes. Molecular dynamics has proven to be a valuable tool to parameterize or inform these higher-length scale models. However, there are currently inaccuracies in the only available U-Mo-Xe potential, which limits the predictive capabilities of molecular dynamics to inform critical phenomena in these fuel systems such as fission gas swelling. This work provides an updated U-Mo-Xe ternary interatomic potential which combines existing potentials in a reconciled format. The validation of the interatomic potential is performed by analyzing the phase stability and vacancy formation energies. Subsequently, Xe solution energies and an equation of state to describe Xe bubbles in U-Mo are calculated, providing 1) evidence of the significant differences between the prior ternary potential and the currently presented potential, and 2) updated data/tools for implementation into mesoscale simulation methodologies to study fission gas bubble evolution.

1. Introduction

The United States High-Performance Research Reactor (USHPRR) program plans to replace current highly enriched uranium (HEU) fuel in high-power research reactors with low enriched uranium (LEU) fuel [1]. In order to achieve a reduced enrichment in these fuel types while maintaining the power density, there is 5 a requirement for increased uranium density. The design for conversion includes a monolithic foil of uranium with 10 wt% molybdenum, with a zirconium diffusion barrier in aluminum cladding.

Recently, substantial effort has been made on mesoscale models to predict the swelling behavior of U-Mo fuels [2, 3, 4, 5, 6, 7, 8, 9]. These models rely on phase-field and/or rate theory descriptions of the microstructure evolution of material systems in order to model swelling on realistic timescales. These

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10 simulation methodologies include a number of parameters that are either fit to limited experimental data, calculated from lower length scale modeling methodologies, or assumed based on other material systems. Classical molecular dynamics (CMD) [..¹] has proven valuable for the parametrization of such mesoscale models [10, 11, 12, 13, 14]. However, there are specific limitations to the present interatomic potentials (which are required for CMD simulations) that inhibit the scope of research that is currently capable.

15 The two interatomic potentials that have been used for the majority of the cited research are from Smirnova [15] and Starikov [16]. The Smirnova potential is an embedded-atom method (EAM) [17, 18] interatomic potential that includes descriptions for U, Mo, and Xe. At the time of its development, to the authors' knowledge, it was the only potential capable of describing any of the pair systems included in this ternary. Thus, it presented a dramatic advancement in terms of available tools for the exploration of the
20 U-Mo system. In particular, the inclusion of Xe in this potential allowed for the analysis of gas bubble behavior, which is a critical phenomenon related to the performance of U-Mo fuels [19, 20]. However, certain properties of the γ phase of U-Mo were found to not be well captured by this EAM potential, such as point defect diffusion coefficients and phase stabilities [15, 16]. In particular, as will be shown later, the body-centered-cubic (bcc) solid solution phase was found unstable with Mo precipitate formation in the U matrix.
25 In the literature, efforts have been put forth by Starikov et al. [16] to improve upon the EAM potential and develop a U-Mo potential in the angular dependent potential (ADP) formalism [21]. The ADP-type potential can correctly predict the stability of the bcc solid solution phase of U-Mo. Although there is a dramatically improved performance of the ADP compared with the EAM potential, the interactions for Xe were excluded in the ADP parametrization, and this limits the applicability of the ADP-type potential for
30 Xe in U-Mo.

Motivated by the need of simulating Xe with a reliable potential for U-Mo, this work bridges prior [..²] separate Xe-U, Xe-Mo, and Xe-U-Mo interactions, with the U-Mo ADP developed by Starikov et al. [16] to achieve an interatomic potential for the U-Mo-Xe system. The interatomic potential is validated, and prior results investigating the U-Mo-Xe system are investigated to demonstrate the differences from prior ternary
35 descriptions.

2. Computational Details

2.1. [..³] Interatomic potential formalism

The ADP [21] formalism is a generalization of the EAM type potential, where the total energy E_i (commonly in eV) of an atom i is given by:

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$$E_i = F_\alpha \left(\sum_{j \neq i} \rho_\beta(r_{ij}) \right) + \frac{1}{2} \sum_{j \neq i} \phi_{\alpha\beta}(r_{ij}) + \frac{1}{2} \sum_s (\mu_i^s)^2 + \frac{1}{2} \sum_{s,t} (\lambda_i^{st})^2 - \frac{1}{6} \nu_i^2 \quad (1)$$

40

$$\mu_i^s = \sum_{j \neq i} u_{\alpha\beta}(r_{ij}) r_{ij}^s \quad (2)$$

$$\lambda_i^{st} = \sum_{j \neq i} w_{\alpha\beta}(r_{ij}) r_{ij}^s r_{ij}^t \quad (3)$$

$$\nu_i = \sum_s \lambda_i^{ss} \quad (4)$$

where α and β are the element types of atoms i and j , F is the embedding energy (the energy required to place an atom of type α into the electron cloud) and is a function of the electron density (ρ), ϕ is a pair potential interaction, [..⁴] and [..⁵] s and $t = 1, 2, 3$ and refer to the Cartesian coordinates. The μ and λ terms represent the dipole and quadrupole distortions of the local atomic environment. This formalism extends the original EAM (the first two terms in eq. (1)) by introducing angular forces as dipole and quadrupole moments (third, fourth, and fifth terms). A similar approach that involves the modification of the EAM formalism to include angular forces is the modified-EAM (MEAM) [22, 23].

50 A given ternary potential requires interactions for each of the three species themselves, and interactions between each species, for a total of six pair descriptions. The U-U, Mo-Mo, and U-Mo interactions are taken from the 2018 version of the Starikov U-Mo ADP [16]. The descriptions of U-Xe, Mo-Xe, and Xe-Xe interactions have been previously parameterized, but for an EAM interatomic potential [15]. Thanks to the shared underlying physics between the EAM and the ADP-type interatomic potentials, the EAM-based Xe pair interactions can be implemented into an ADP formalism, provided that appropriate scaling is performed. Therefore, the construction of a ternary interatomic potential in this work should not necessarily be considered a development step, but a reconciliation step, where multiple models have been unified under a single potential formalism, but the underlying descriptions have remained unchanged.

2.2. [..⁶] Interatomic potential adjustment

60 This work makes use of the assumption that for Xe interactions no angular forces are required for the accurate representation of forces. Given the noble nature of Xe [..⁷] and the lack of unpaired valence electrons, we believe that this is a reasonable assumption. Therefore, the EAM descriptions from Smirnova [15] can be taken as complete, and the third, fourth, and fifth terms in eq. (1) can be set to zero.

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The first step in the reconciliation procedure involves an invariant scaling process. The quantity $\bar{\rho}$ is
65 given by the sum:

$$\bar{\rho} = \sum_{j \neq i} \rho(r_{ij}) \quad (5)$$

where ρ is the electron density. The energy and forces in the system are invariant to the scaling of the electron density and the embedding energy,

$$\rho(r_{ij}) = \alpha \rho(r_{ij}) \quad (6)$$

$$F(\bar{\rho}) = F\left(\frac{\bar{\rho}}{\alpha}\right) \quad (7)$$

where F is the embedding energy and α is a scale factor. This scaling does not change the properties of the pure systems but does^[..⁸], in fact, change the multi-component system's behavior. The scale factor was
70 applied only to Xe^[..⁹] and in such a fashion as to lead to the same electron density range as that employed for the U-Mo ADP. Thus, the scaling factor α is simply the ratio between the ρ discretization in the EAM and ADP parameterizations (approximately 0.936). This is considered minor scaling and should yield minimal impacts on ternary interactions. The invariant scaling process was validated by studying the pressure as a function of volume for pure Xe gas prior to and after the scaling process, which showed an exact match within the statistical
75 fluctuations of the molecular dynamics simulations. No further scaling is required, as the^[..¹⁰] discretization in space (r) is identical for both the ADP and EAM potentials utilized. For further detail on the fitting procedures to obtain the original potentials, please refer to the original manuscripts for the EAM [15] and the ADP [16].

2.3. Molecular^[..¹¹] dynamics simulations

Simulations were performed utilizing the LAMMPS [24] software package and the U-Mo-Xe potential
80 constructed in this work. General simulation details are listed here, and further information, where necessary, is included in the results section.

2.3.1. Phase stability

The γ phase of U-Mo is a bcc random solid substitutional solution phase with a wide compositional range. As the primary purpose of the ADP-type potential is to study the γ phase of U-Mo, its capability of stabilizing
85 the bcc phase in the desired temperature and compositional regions needs to be justified. The stability of the bcc γ phase was examined by calculating the c/a ratio as a function of temperature and Mo molar fraction, c_{Mo} , and the stability of the bcc U-10Mo solid solution against phase separation.

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A bi-crystal simulation cell with a $\langle 100 \rangle \Sigma 5$ symmetrical tilt grain boundary was used to study possible phase separation and grain boundary segregation. The simulation cell was 15.2 nm (x) by 1.4 nm (y) by 30.4 nm (z) in size and contained 21.6 atomic percent Mo for the purpose of representing U-10Mo. The grain boundary was inserted in the x-y plane. After the relaxation of the initial simulation cell, hybrid molecular dynamics / Monte Carlo (MDMC) simulations were carried out at 300, 600, and 900 K. Both the EAM and ADP-type potentials have been utilized to compare their performance. In the MDMC simulations, a randomly selected Mo atom was swapped with a randomly selected U atom at every MD step. The change in total potential energy, ΔE , was calculated. The swap was accepted if $\xi < \exp(-\frac{\Delta E}{K_B T})$ and rejected otherwise. Here K_B is the Boltzmann constant, and ξ is a random number in the range of 0 to 1. Each simulation lasted until the total potential energy converged.

To compute the c/a ratio, a $20 \times 20 \times 20$ bcc simulation cell was used with 16000 atoms. The Mo molar fraction was varied from 0 to 1, with an incremental step of 0.1. The point $c_{Mo} = 0.216$ was also inserted to represent U-10Mo. The temperature was varied from 0 to 1000 K, with one data point every 100 K. At each temperature and for each Mo fraction, the length of the cell along the three principal axes was relaxed using a zero-stress boundary condition. The cells were found to evolve into either a body-centered-tetragonal (bct) or bcc structure, i.e., the γ phase. Here, c is taken as the smallest lattice parameter in an orthorhombic unit cell and a as the average of the two other lattice parameters. The γ phase is indicated by $c/a \approx 1$, while $c/a < 1$ represents the bct phase. It should be noted that at high temperatures, c/a is not exactly 1 because of thermal fluctuations in the simulations, and we took $c/a > 0.99$ as the criteria for the appearance of the γ phase.

2.3.2. Vacancy formation energy and Xe solution energy

The vacancy formation energy, E_v^f , and Xe solution energy, E_{Xe}^s , were also computed at 300 K as a further test of the potential. The vacancy formation energy is defined as

$$E_v^f = E(N_U N_{Mo} V_1) - N_U \mu_U - N_{Mo} \mu_{Mo} \quad (8)$$

[..¹²] and the Xe solution energy is defined as

$$E_{Xe}^s = E(N_U N_{Mo} Xe_1) - N_U \mu_U - N_{Mo} \mu_{Mo} - \mu_{Xe} \quad (9)$$

[..¹³]

where $E(N_U N_{Mo} V_1)$ is the total energy of simulation cells with N_U U atoms, N_{Mo} Mo atoms, and one vacancy, and $E(N_U N_{Mo} Xe_1)$ is the corresponding total energy when the vacancy is occupied by a Xe atom, forming a Xe substitutional. μ_U and μ_{Mo} are the chemical potentials of U and Mo, respectively. The calculations were performed using a $10 \times 10 \times 10$ cell with 1568 U and 432 Mo atoms, representing U-10Mo. An NPT (constant pressure and constant temperature) ensemble with zero external pressure and a

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temperature at 300 K was used for the [..¹⁴]calculations. A random atom was either removed or replaced with Xe to create a vacancy or Xe substitutional, respectively. The statistical approach recently proposed 120 [25] was used to compute E_v^f , μ_U , and μ_{Mo} , which were then used to computed E_{Xe}^s using [..¹⁵]eq. (9). The process is briefly described below, and the detailed derivation can be reached in the original reference.

2.3.3. [..¹⁶]

Starting from a perfect bcc simulation cell containing $N = 2000$ atoms, with a total energy of E^0 , 2000 defected simulation cells were created, each with a vacancy created by removing a different atom (denoted as k) 125 and a total energy of E^k computed by MD. The mean vacancy formation energy, $\langle E_v^f \rangle$, was obtained, as,

$$\langle E_v^f \rangle = \frac{1}{N} \sum_{k=1}^N E^k - \frac{N-1}{N} E^0 \quad (10)$$

The chemical potential of element j (U or Mo), μ_j , was obtained as,

$$\mu_j = E^0 + \langle E_v^f \rangle - \frac{1}{N_j} \sum_{k=1}^{N_j} E^k \quad (11)$$

In eq. (11), N_j is the number of type j atoms, and the summation term in the right-hand side enumerates all simulations in which a type j atom was removed to create the vacancy. The chemical potentials for U and Mo 130 were obtained using the above equations, and that of Xe was taken as 0 because of its inert nature. The chemical potentials were then used to compute the vacancy formation energies and Xe solution energies with 2000 different atomic configurations to achieve reasonable statistics.

2.3.3. Xe bubble equation of state

A supercell of 40x40x40 bcc unit cells (128,000 U atoms) was generated, and 21.6 percent of U atoms 135 [..¹⁷]were randomly switched to Mo atoms, yielding a U-10Mo alloy in the bcc structure. Relaxation of the bulk system was performed in an NPT ensemble, relaxing each x, y, and z component individually, with a damping parameter of 0.1 ps. Temperatures of interest were 300 - 700 K, in increments of 100 K, which spans the realistic operating temperatures for U-Mo fuels [26]. The system was allowed to equilibrate for 200 ps at a given temperature, and subsequently, a void was constructed by deleting a sphere of atoms from 140 the center of the supercell. This void was relaxed for 200 ps under the same simulation conditions described above. Different sizes of voids are prescribed to ensure that there were no simulation artifacts present due to the existence of a single bubble size. Voids were ensured to be sufficiently large that the surface energy had converged [14].

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In order to analyze bubbles, simulations were performed in an NVT ensemble to mimic a bubble in a very large system that effectively exerts a resistive pressure on the bubble. This allows for the calculation of a Xe bubble pressure and a subsequent equation of state (EOS) based on the density of the bubble. The generation of bubbles was performed by inserting Xe atoms into the void one at a time, while the relaxation of the system is ongoing. The insertion rate was less than one Xe atom per 8 ps for all systems. The rate of insertion was modified according to the bubble size in order to ensure a similar rate of Xe to vacancy ratio change as a function of time. **This gradual insertion allows for on-the-fly relaxation and avoids pressure/energy spikes that would accompany a rapid insertion of a large number of Xe atoms.** In order to track the bubble size, two atoms (one on either side of the void) were tracked throughout the simulation and the distance between the two atoms is classified as the diameter of the bubble. The pressure of the bubble was determined by computing the stress per atom on each of the Xe atoms in the system, summing the individual components of the stress tensor over all Xe atoms, and finally dividing by the degrees of freedom (three) and the volume of the bubble.

A minimization script was utilized to fit the EOS for a Virial functional form to the determined pressure and molar volume data from the molecular dynamics simulations. The data was input into the script, and the relative error was summed and utilized to optimize the EOS, iterating by providing a random step to each of the fitting coefficients and only accepting the iteration if the total relative error is reduced. This methodology is in accordance with a prior study [14] developing a Xe bubble EOS utilizing the U-Mo-Xe EAM potential.

It should be noted that this approach assumes the nucleation of bubbles, but not does explicitly explore the nucleation phenomenon. Negative/attractive binding energies of Xe to voids/bubbles have been observed previously [14], and bubbles are known to form experimentally. Thus, the inclusion of a nucleation stage is not expected to alter any of the results or conclusions of this manuscript.

3. Results

3.1. Interatomic potential validation

3.1.1. Phase stabilities and structural constants

The MDMC simulations with the EAM and ADP-type potentials predicted different stabilities for the bcc solid solution phases. As shown in [..¹⁸]fig. 1, with the EAM potential, U-10 Mo solid solution was not stable and went through a phase separation with the formation of Mo precipitates. The same results were obtained at all three simulation temperatures, 300 K, 600 K, and 900 K. In contrast, no such phase separation was observed in the simulations with the ADP-type potential at all three temperatures, where the solid solution U-10Mo remained stable, as shown in [..¹⁹]fig. 1. This again indicates the benefit of achieving

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¹⁷⁵ an ADP-type potential with Xe. In all simulations with both potentials, no substantial grain boundary segregation was identified.

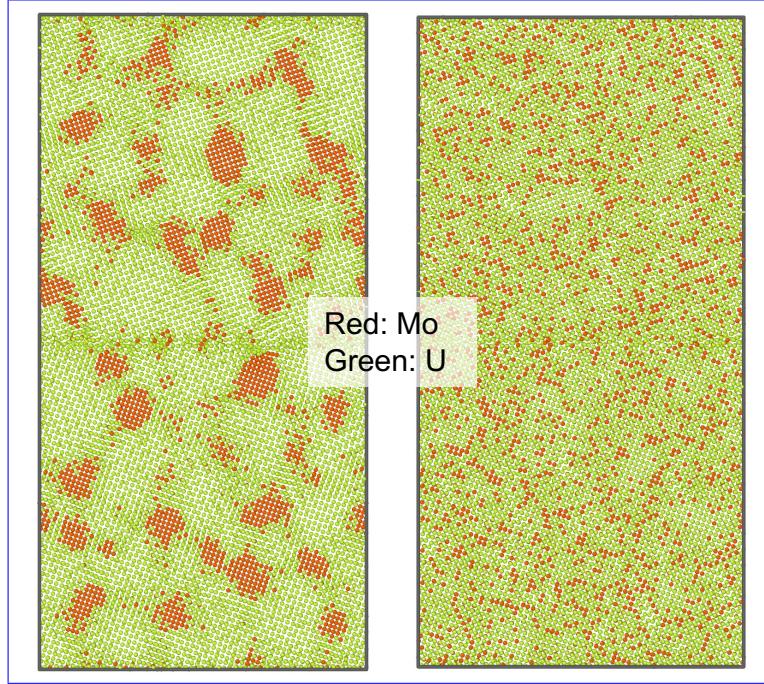


Figure 1: Atomic configurations from ^{[..²⁰]hybrid} MDMC simulations at 900 K showing (^{[..²¹]left}) Mo precipitation with the EAM potential and (^{[..²²]right}) stable solid solution with the ADP-type potential.

The stable zone of the γ phase is shown by the c/a ratio plotted in ^{[..²³]fig. 2}. The bct phase is stable when $T < 650$ K and the molar fraction of Mo $c_{Mo} < 0.2$. In this zone, the c/a ratio increases with increasing c_{Mo} and is insensitive to temperature with a weak U-shaped dependence. In contrast, the γ phase, indicated by $c/a \approx 1$, is stable when $T > 650$ K or $c_{Mo} \geq 0.2$. Although the bct phase is not exactly the same as the α -U structure, and the transition temperature is lower than the actual $\alpha - \beta$ transition point of U, ^{[..²⁴]fig. 2} indicates that the potential captures the overall behavior that the γ phase is stabilized by increasing temperature and Mo concentration. It is thereby applicable for studying the properties of U-Mo systems either with high Mo content or at high temperatures. Note that this potential was primarily developed for ¹⁸⁵ simulating bcc U-Mo solid solution alloys.

The atomic pressure and volume were computed to investigate their dependence on the local atomic environment. The atomic pressure is computed as $\frac{1}{3}(\sigma_{xx} + \sigma_{yy} + \sigma_{zz})$, which are the normal atomic stress components computed in LAMMPS. To compute atomic volume, a Voronoi tessellation was carried out by taking the atoms as cell centers, and the cell volumes were taken as the atomic volumes. In the solid solution phase, by average, U atoms are under compression while Mo atoms are under tension, as shown in fig. 3(a), while the overall simulation ¹⁹⁰

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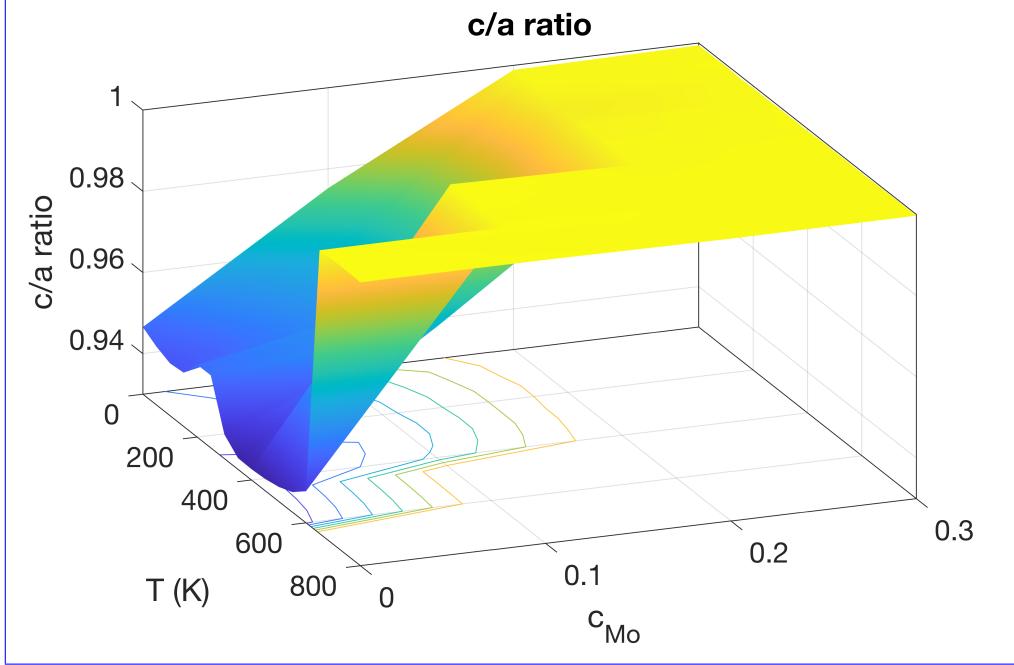


Figure 2: The c/a ratio as a function of Mo molar fraction and temperature. The bcc solid solution phase is stable when c_{Mo} is above 0.2 or T is above 650 K.

cell exhibits zero pressure at the boundaries. This is consistent with the fact that U has a larger atomic size than Mo. At room temperature, bcc U has a lattice constant of 3.47 Å, versus 3.15 Å for bcc Mo [27]. The atomic pressure of U atoms increases slightly in magnitude with an increasing number of 1st nearest neighbor (1NN) Mo atoms, showing the dependence on the local atomic environment. In U-10Mo solid solution, U atoms are found to have a larger average atomic volume than Mo, as shown in fig. 3(b). For both, the atomic volumes decrease with an increasing number of 1NN Mo atoms almost linearly until five or six 1NN Mo atoms, for which the few data points may not warrant good statistics. The results imply that Mo-rich regions will have smaller atomic volumes than U-rich regions and are under tension.

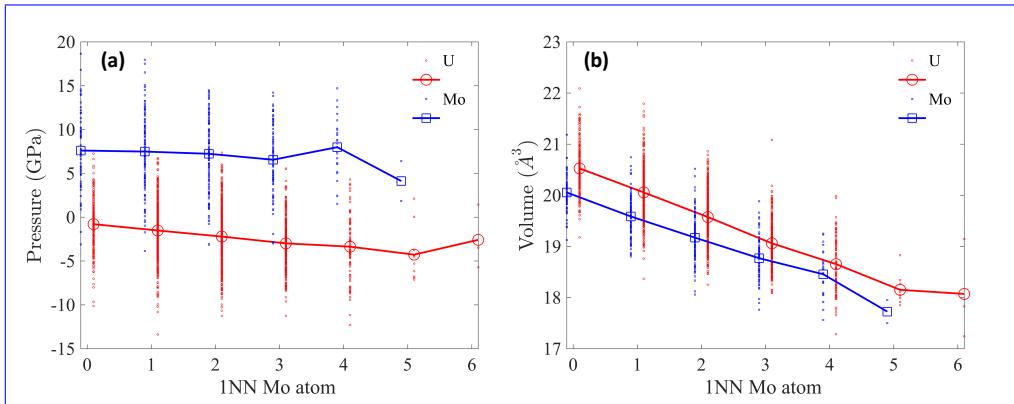


Figure 3: (a) Atomic pressure and (b) volume of U and Mo atoms as functions of 1NN Mo atoms. Positive (negative) pressure means tension (compression). Data points for Mo are shifted to the left and U to the right by 0.1 for clarity.

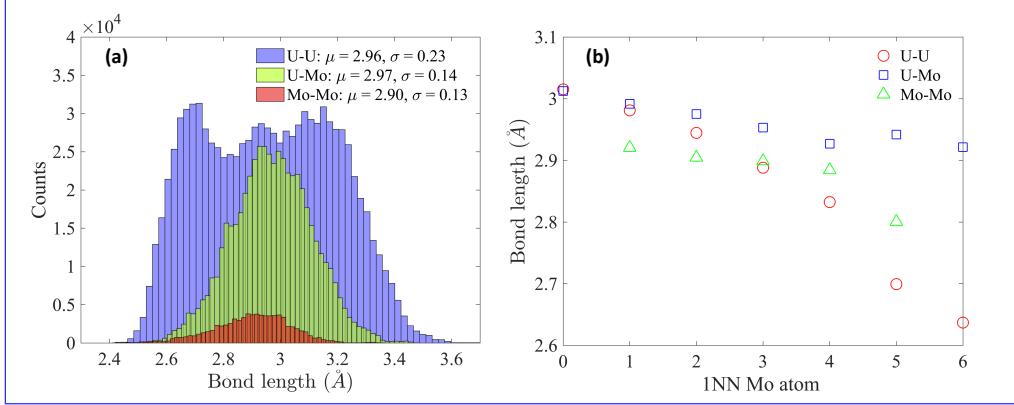


Figure 4: (a) Distributions of U-U, U-Mo, and Mo-Mo bond lengths, and (b) the dependence of average U-U, U-Mo, and Mo-Mo bond lengths on 1NN Mo atoms.

The bcc lattice of U-10Mo solid solution is highly distorted, as indicated by the bond length distribution in fig. 4(a). Here, 1NN bonds in a 2000 atom supercell equilibrated at 300 K were characterized. The U-U bond has the broadest distribution, with a standard deviation of $\sigma = 0.23 \text{ \AA}$, which is about 8% of the mean bond length, $\mu=2.96 \text{ \AA}$, while the range of variation spreads from 2.4 \AA to about 3.6 \AA . The U-Mo bond has the largest mean length ($\mu=2.97 \text{ \AA}$), and the Mo-Mo bond has the smallest mean length ($\mu=2.90 \text{ \AA}$). For the three types of bonds, the lengths decrease with an increasing number of 1NN Mo atoms, as shown in fig. 4(b). In all atomic environments, U-Mo bonds have the largest average lengths. In the U-rich environment (with less than three 1NN Mo atoms), U-U bonds have larger average lengths than Mo-Mo but become shorter than Mo-Mo bonds when there are more than three 1NN Mo atoms. The U-U bonds experience the most reduction in length from U-rich to Mo-rich environments. The change in bond length is consistent with the reduction in atomic volume in fig. 3(b).

210 3.2. Xe substitution in U-Mo

The vacancy formation energy E_v^f and the Xe solution energy E_{Xe}^s are shown in [..²⁵]fig. 5, both exhibiting [..²⁶]probability distributions, which indicates their dependence on the local atomic environment. Such a dependence is typical in concentrated alloys and implies the importance of sampling different atomic environments. The distributions were fitted using [..²⁷]a Gaussian distribution function based on their shapes. The mean and standard deviations obtained from the fitting are listed in [..²⁸]table 1 along with the chemical potentials of U and Mo[..²⁹]. The mean vacancy formation energy $\langle E_v^f \rangle$ computed is 1.67 eV. This value is higher than the average value obtained by Jin et al. in U-20Mo (in atomic percent) at 0 K using the same ADP-type potential [28], where the formation energies for U-vacancy (vacancy created by removing

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a U atom) and Mo-vacancy were computed to be 1.07 eV and 1.30 eV, respectively. The discrepancy may
220 come from the different methods used for deriving the chemical potentials. In Jin et al., the [..³⁰]average over all atoms was used for the chemical potential for both U and Mo, and in this work, the chemical potentials [..³¹]of U and Mo were derived separately with a notable difference, as given in [..³²]table 1. The mean Xe solution energy is 10.29 eV, much higher than that calculated using the EAM potential, 6.1 eV [14].

It should be noted that vacancy formation energy and Xe solution energy as defined in eq. (8) and eq. (9) do
225 not depend on the types of atoms that are removed to create the vacancies, because the defected and reference cells have the same numbers of U and Mo atoms. As discussed in [25, 29], the formation of a vacancy in a real material system does not involve removing an atom, although it is prescribed that way in atomistic calculations. Rather, the total mass should be conserved, as eq. (8) and eq. (9) imply. Vacancies can be generated from sinks such as free surfaces and grain boundaries without removing any atoms, and they do not feature chemical
230 types. In atomistic calculations, as the reference cell usually has one more atom than the defected cell, the chemical potential of the removed atom is added into the defected cell to conserve the total mass and remove the dependence of the result on the type of the removed atom. As such, there is no distinction between "U vacancies" and "Mo vacancies". There is a slight difference in composition between cells having a U atom removed and those having a Mo atom removed. Here, we assume such a negligible change in composition won't cause a substantial
235 change in vacancy formation energy. The insertion of Xe follows the creation of a vacancy, and the resulting Xe solution energies are independent of the type of the removed atom as well.

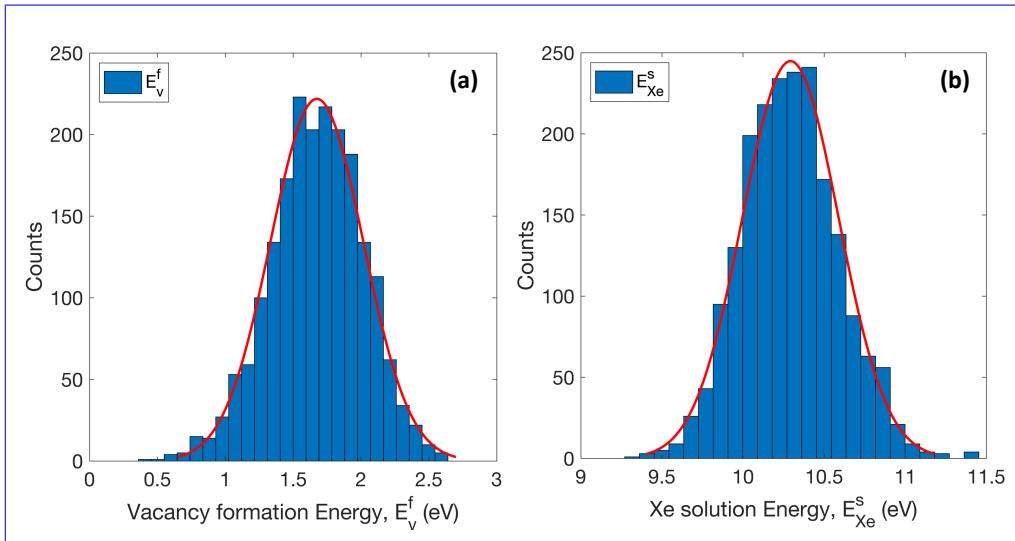


Figure 5: Distributions of (a) vacancy formation energy and (b) Xe solution energy in U-10Mo.

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Quantity	μ_U	μ_{Mo}	$\langle E_v^f \rangle$	$\langle \sigma_v^f \rangle$	$\langle E_{Xe}^s \rangle$	$\langle \sigma_{Xe}^s \rangle$
Value	-4.1125	-6.6934	1.6748	0.3415	10.2924	0.2965

Table 1: Chemical potentials of U and Mo, μ_U and μ_{Mo} , mean vacancy formation energy $\langle E_v^f \rangle$, and the standard deviation $\langle \sigma_v^f \rangle$, mean Xe solution energy, $\langle E_{Xe}^s \rangle$, and the standard deviation, $\langle \sigma_{Xe}^s \rangle$, in U-10Mo calculated at 300 K using molecular dynamics simulations, all in units of eV.

The **vacancy formation energy (E_v^f)**, Xe solution energy [..³³] (E_{Xe}^s), and their difference are plotted against the numbers of [..³⁴] Mo atoms in the 1NN [..³⁵] and 2NN [..³⁶] distances in fig. 6. The E_v^f is found to increase very slightly with the number of 1NN Mo atoms, as shown in fig. 6(a), while a clear positive correlation between E_{Xe}^s and the number of 1NN Mo atoms is identified, fig. 6(b). Note that E_{Xe}^s as defined in eq. (9) contains E_v^f . Their difference was computed and also found to increase with the number of 1NN Mo atoms, [..³⁷] as shown in fig. 6(c). The lower E_{Xe}^s in a U-rich environment than in a Mo-rich environment indicates that Xe prefers to be surrounded by U atoms, and the preference is not from the environmental dependence of the E_v^f . Instead, the preference of Xe for a U-rich environment [..³⁸] can be explained by the atomic volume shown in fig. 3(b).

245 The U-rich environment has a larger atomic volume and is therefore preferred by inert gas atoms such as Xe. In comparison, E_v^f , E_{Xe}^s , and their difference show no clear dependence on the number of 2NN Mo atoms[..³⁹], as shown in fig. 6(d)-(f), indicating that the dependence is very local within the 1NN distance. Note that for the environments with many Mo neighbors, only a few data points were collected from MD simulations in each case, and thus the averaged values are subject to large uncertainty but are included for completeness.

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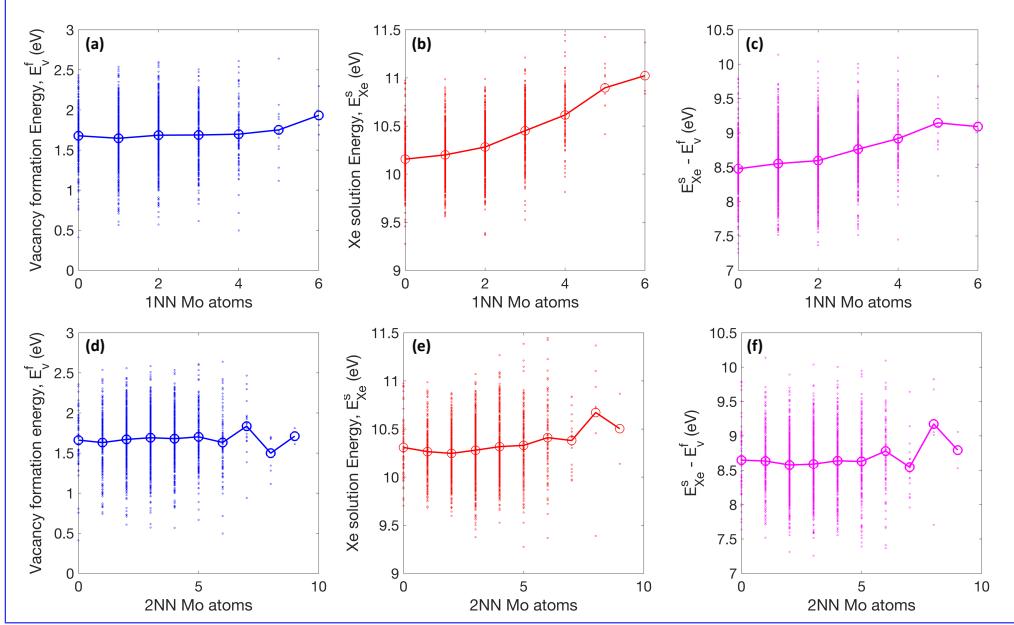


Figure 6: Dependence of (a) vacancy formation energy, (b) Xe solution energy, and (c) their difference on the number of [..⁴⁰] 1 [..⁴¹] NN Mo atoms. In ([..⁴²] d)[..⁴³] -([..⁴⁴] f) [..⁴⁵] the corresponding dependence on the number of Mo atoms within the 2NN distance are shown. The averaged values are shown by large markers connected by lines to guide the visualization.

250 3.2.1. Voids and bubbles in U-Mo

3.3. Surface energies of voids

In order to generate bubbles, voids of varying size must be generated. This allows for the calculation of a void surface area as a function of radius, and the surface energy can be determined from [..⁴⁶] eq. (12)

$$E_{surf} = \frac{(E^* - E)}{A} \times N \quad (12)$$

where E^* is the potential energy per atom of the system with a void, E is the potential energy per atom of the perfect crystal of U-Mo, A is the total surface area of the void, and N is the number of atoms in the system with a void. The void surface energy is shown in [..⁴⁷] fig. 7 as a function of the void radius.

It can be observed that the void surface energy converges above a radius of 1.5 nm for all temperatures to a value of approximately 1.5 J/m². Also, the void surface energy tends to increase slightly with increasing temperature. This is in accordance with the average surface energy for U-Mo planar surfaces as determined in [30], which utilized the Starikov ADP [16]. It should be noted that the EAM potential was previously utilized to analyze void surface energy, and predicted a value 0.4 J/m² lower than the value reported here. This provides confidence in the accurate implementation of the potential [..⁴⁸] with a validation of different

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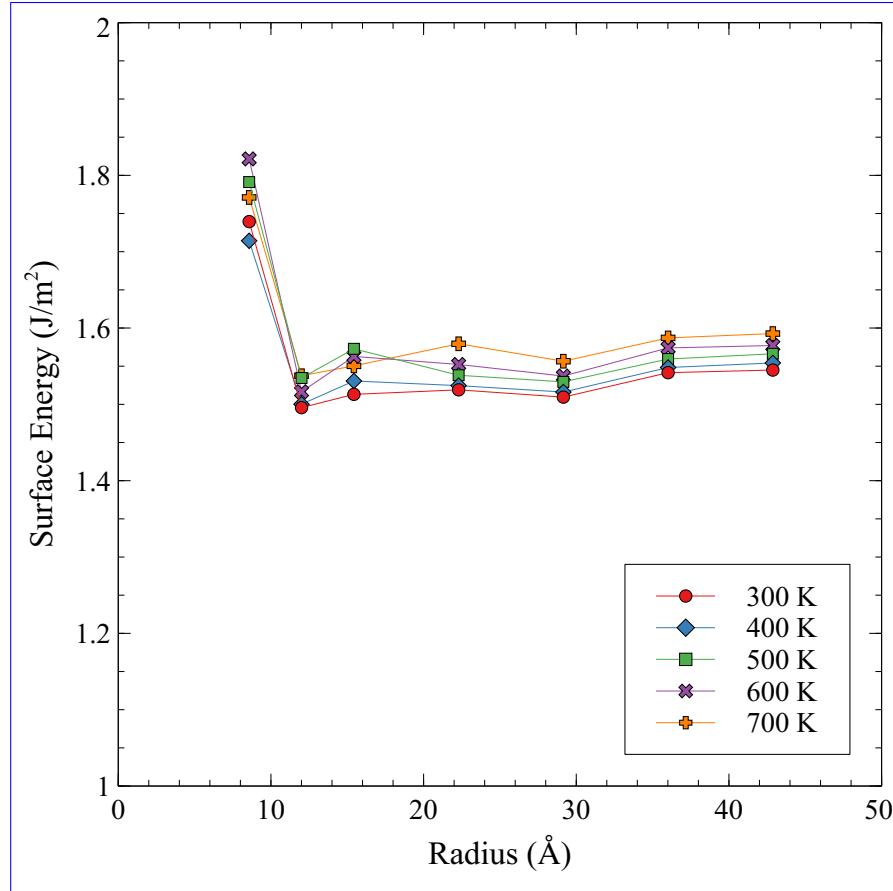


Figure 7: The void surface energy as a function of radius and temperature in U-10Mo.

methodologies for obtaining average surface energies, and underlines the importance of this work to incorporate a more accurate U-Mo potential which can also describe Xe interactions. [Statistical errors for the determination of surface energies are on the order of 0.04 J/m².](#)

3.4. Xe bubble equation of state in U-Mo

The [..⁴⁹]EOS can be determined by tracking both the pressure inside the bubble and the bubble size as a function of the number of Xe atoms present in the bubble while the system is equilibrated in an NVT ensemble, which provides a pressure versus density relationship. This data can be fit to an EOS that provides a generalized function predicting the relationship between pressure, temperature, and molar volume. [..⁵⁰]Temperatures from 300 K to 700 K are analyzed [..⁵¹]for all bubble sizes previously mentioned. [..⁵²]

In line with the previous study exploring the EOS of Xe bubbles in U-Mo [14], a Virial [..⁵³]EOS is utilized,

⁴⁹removed: equation of state (EOS)

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⁵³removed: equation of state

expanded to the third order with respect to volume and the second order with respect to temperature, as is shown in equation [..⁵⁴] eq. (13),

$$P = \frac{RT}{v} \left(A + \frac{B}{v} + \frac{C}{v^2} + \frac{D}{v^3} \right) \quad (13)$$

²⁷⁵ where $A=1$, and B , C , and D are temperature-dependent Taylor series of $1/T$ ($B=b_0 + b_1/T + b_2/T^2$, $C=c_0 + c_1/T + c_2/T^2$, and $D=d_0 + d_1/T + d_2/T^2$), leading to nine unique fitting parameters. The term R is the gas constant (8.3145 J/mol-K) and v is the molar volume (cm³/mol). The Virial equation is a general function relating pressure, molar volume, and temperature that can be directly derived from statistical mechanics [31]. **No restrictions are imposed on the fitting parameters to obtain an optimized EOS.**

²⁸⁰ The subsequent fit, with and without molecular dynamics data, is shown in [..⁵⁵] fig. 8 and fig. 9, respectively. As expected, with higher temperatures, the pressure at a given molar volume is higher, resulting in an effective rightward shift of the data as the temperature is increased. All data with molar volumes below 500 cm³/mol are displayed in fig. 8. The MD data is removed in [..⁵⁶] fig. 9 and the optimized EOS is displayed on a log/linear scale to emphasize the differences between the individual isotherms. As stated, temperature shifts the pressure to a larger value at a given molar volume, and the optimized Virial EOS accurately captures this behavior. The optimized coefficients are shown in [..⁵⁷] table 2.

Table 2: Virial [..⁶⁷]EOS ([..⁶⁸]eq. (13)) parameters for Xe bubbles in U-Mo.

Parameter	Value
b_0	14.625 cm ³ /mol
b_1	54597.728 cm ³ -K/mol
b_2	-386.344 cm ³ -K ² /mol
c_0	2968.616 cm ⁶ /mol ²
c_1	2938.010 cm ⁶ -K/mol ²
c_2	-84.545 cm ⁶ -K ² /mol ²
d_0	705527.001 cm ⁹ /mol ³
d_1	53.609 cm ⁹ -K/mol ³
d_2	421.138 cm ⁹ -K ² /mol ³

While 4th order optimizations both including and excluding the 3rd order term were pursued, as has recently been performed by Yang [32] for Xe in UO₂, the accuracy of the fit to the MD data was not improved.

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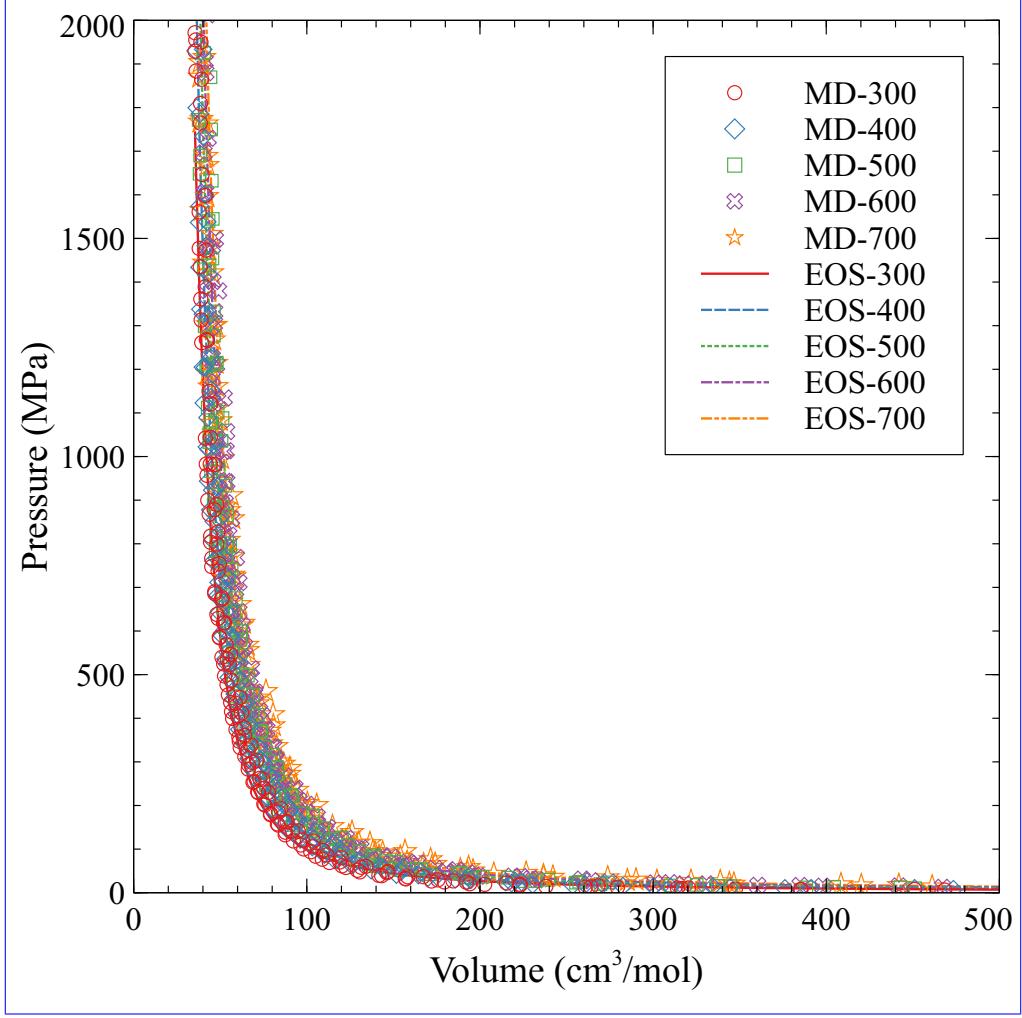


Figure 8: An [..⁵⁸]EOS [..⁵⁹]based on a virial expansion for Xe bubbles in U-10Mo from [..⁶⁰]300 K to 700 K [..⁶¹]compared to molecular dynamics data[..⁶²].

A comparison to the Kaplun EOS [33] is shown in [..⁶⁹]fig. 10 for 400 K and 700 K. Kaplun developed a unified
 290 EOS for liquid, gas, and fluid defined from thermodynamic equations which shows excellent agreement with
 experimental data for a number of species. [..⁷⁰]The comparison in fig. 10 shows that at low densities,
 the Kaplun data and the fit to the computational data converge, with the computational EOS predicting a
 very slightly higher pressure [..⁷¹](<0.2 MPa). The differences observed in fig. 10 are exaggerated due to the
 log/linear scale utilized in the graphical representation. However, at high densities and high pressures, there is a
 295 dramatic difference between the Kaplun EOS and the Virial EOS from this work, indicating the importance
 of taking into account the U-Mo matrix for the properties of Xe bubbles. Thus, this updated Virial EOS in

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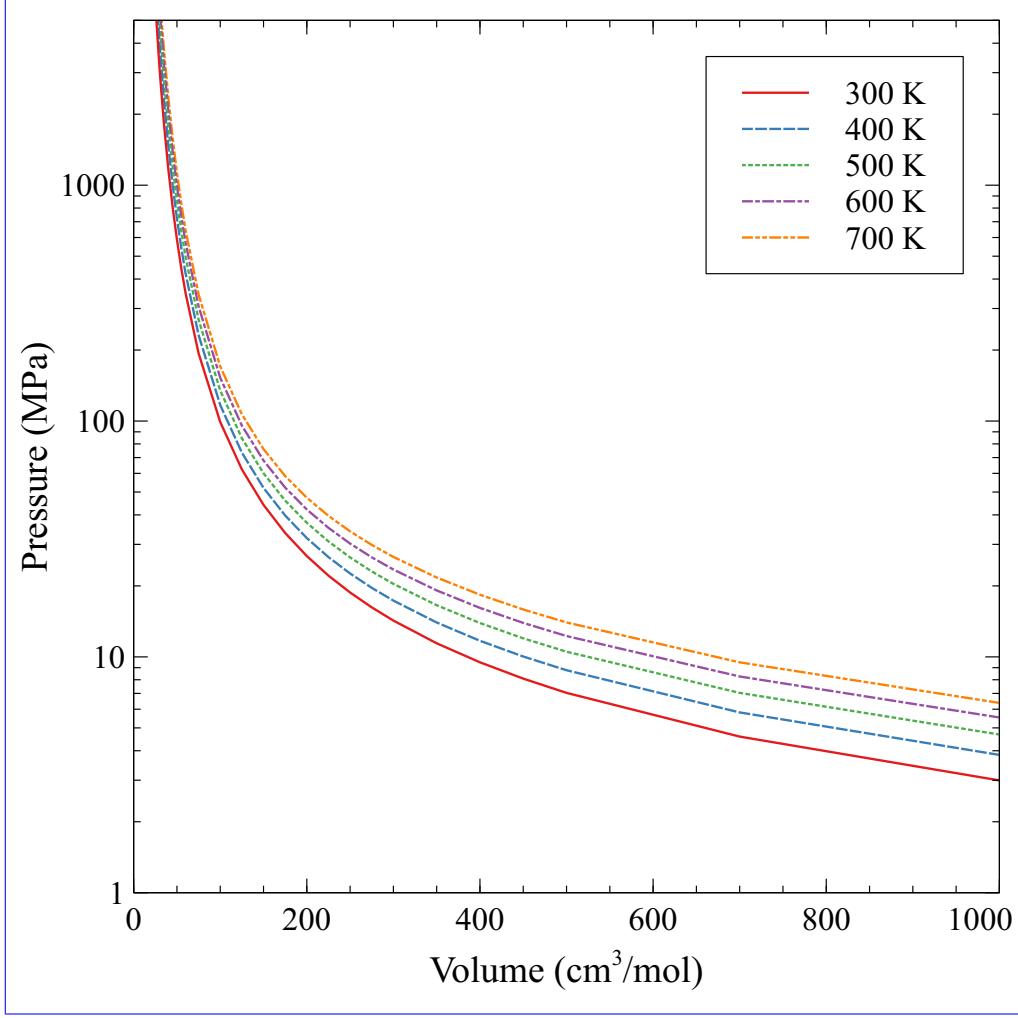


Figure 9: An [..⁶³]EOS [..⁶⁴]based on a Virial expansion for Xe bubbles in U-10Mo from [..⁶⁵]300 K to 700 K [..⁶⁶]without molecular dynamics data.

[..⁷²]eq. (13) and table 2 can be utilized to describe both high and low density/pressure bubbles in U-Mo over the prescribed temperature ranges. Unfortunately, there is no experimental data to base a comparison or to establish the ultimate accuracy of the EOS of Xe bubbles in U-Mo at high pressures and densities. Thus,
 300 we argue that the convergence to a pure Xe gas at low densities, and the specific incorporation of the accurate interaction between U, Mo, and Xe at high densities, provides a reasonable prediction for Xe bubble behavior in the U-Mo system.

⁷²removed: equation 13 and Table 2

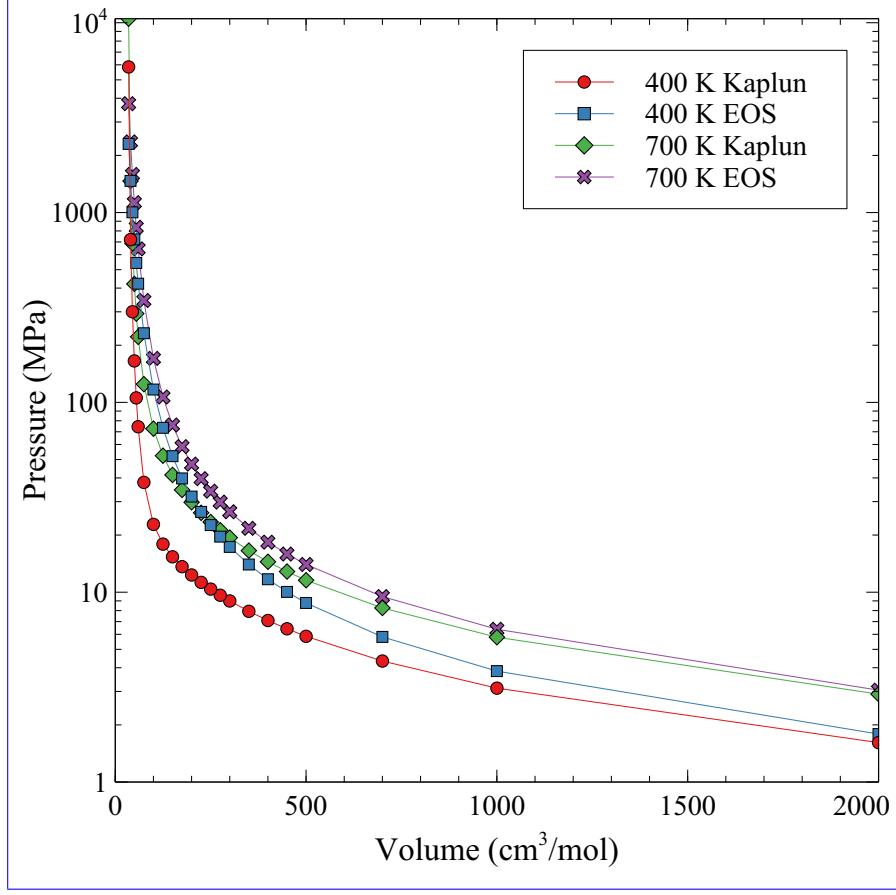


Figure 10: Comparison between the Kaplun unified EOS (Kaplun) and the Virial EOS from this work (EOS) over a wide range of molar volumes at 400 K and 700 K.

4. Conclusions

In this work, prior interatomic potentials for the U-Mo and the U-Mo-Xe system were reconciled under a single ADP framework, generating a novel ternary potential that utilizes the most accurate descriptions of the individual pair systems which exist. The phase stability of the γ phase of U-Mo was explored via MDMC simulations, demonstrating the inability of prior potential forms to correctly predict the stable solid solution behavior of the alloy. The distortion of the γ phase into a bct structure was analyzed as a function of temperature and composition, agreeing with experimental observations of bcc instability of U-rich systems at lower temperatures. A rigorous approach to vacancy and Xe substitutional energies was applied to this system for the first time, identifying key inaccuracies in previous potential formalisms and providing information on the effect of the local environment on the behavior of Xe substitutionals. Finally, voids and bubbles were explored in U-10Mo, verifying previous studies of void surface energy and providing a

new equation of state for Xe bubbles specific to the U-Mo system. [..⁷³] While further experimental validations
315 are needed to justify the accuracy of the reconciled potential, this work enables studying Xe gas behavior and gas
bubble properties in U-Mo solid solutions using molecular dynamics simulations. It allows for the exploration of
critical phenomena related to Xe gas atoms and bubbles in U-Mo fuel and will serve to inform mesoscale and
engineering scale fuel performance simulations [..⁷⁴] with better reliability than existing interatomic potentials.

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⁷³removed: This work has served to demonstrate the need for an updated interatomic potential, and provided that tool, which

⁷⁴removed: , providing increased accuracy and reducing known errors, allowing for the exploration of critical phenomena in U-Mo fuels

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