

Reaction Rates

Pre-lesson assignment- textbook page 142-144

Define the following terms

- Rate of a chemical reaction
- Collision theory

Make notes on reaction rates

Use the following questions as guidance

1. Show the equation for the rate of reaction and show how the units of rate can be derived.
2. Sketch a graph showing how the concentration of reactants changes over time.
 - a. How is the gradient of the line affected by the rate of reaction?
 - b. Explain why the rate slows as the reaction progresses.
 - c. Explain why the rate of reactions eventually reaches zero.
3. State four ways to alter the rate of a chemical reaction.
4. In order for a reaction to progress, what must be true of the particles involved as they collide?
5. Explain how rate of reaction is affected by
 - a. Pressure
 - b. Concentration