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You do this by reacting 50.0 L of nitrogen gas with excess hydrogen gas at STP to create ammonia gas using the Haber process given by the following reaction:  $\text{N}_2(\text{g}) + 3\text{H}_2(\text{g}) \rightarrow 2\text{NH}_3(\text{g})$ . Assuming the reaction had 100% yield, you then bubble the ammonia into 5.0 L of water. What is the molarity of the resulting ammonia solution? 0.89 M

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