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REDOX AND ELECTROCHEMISTRY Date ____ Period ____ Analyzing Oxidation-Reduction Reactions When chemical bonds form, electrons are either lost, gained or shared. Metals lose electrons. This is what happens when iron rusts. When the iron, a metal, combines with oxygen, a non metal, to form rust, it loses electrons. This

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Balance a redox equation, by balancing electrons lost with electrons gained . Use Table J to predict whether a particular reaction will occur or not. In an electrochemical (or galvanic or voltaic) ... ©Evan P. Silberstein, 2003 '~1 I I ASSIGNING OXIDATION NUMBERS Name -----

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Example of using table of standard reduction potentials to calculate standard cell potential. Watch the next lesson: <https://www.khanacademy.org/science/chem...>

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