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solutions with bromthymol blue (pH = 6.0-7.6). • Forensic analysis of DNA by electrophoresis requires a buffer that will keep the charge on the DNA molecules relatively constant so that their migration in an electric field will depend only on their size.

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An ideal acetic acid-sodium acetate buffer system has a pH of 4.75 and its buffer range is 3.75-5.75. Equation 8 shows the calculation for the lower pH limit of an acetic acid-sodium acetate buffer solution (when the concentration ratio of the weak acid component to the conjugate base component is equal to 10:1)

Lab #16 - Properties of Buffer Solutions - LHS AP Chemistry

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1. pH=pKa + log (base/acid), best with equimolar concentrations 2. C6H8O7 + NaOH = NaC6H7O7 + H2O C6H7O7 + NaOH = NaC6H6O7 + H2O C6H6O7 + NaOH = NaC6H5O7 + H2O 3. a. Equal molar concentrations of C6H8O7 and NaC6H7O7 b. Equal molar concentrations of C6H6O7 and NaC6H5O7 4. Ideal

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Hydrolysis of salts will be used to study the acid-base properties of dissolved ions in aqueous solutions. The approximate pH of these solutions will be determined using acid-base indicators. A buffer solution will be prepared, and its ability to moderate pH will be investigated alongside solutions that cannot function as buffers.

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