

Electron Configuration Yahoo Answers

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Best Answer: As you will have been told s-orbitals are spherical, p-orbitals lie along the three major axes (x,y and z) and d orbitals have more than one type. The electrons in these orbitals are attracted by the protons in the nucleus. The attraction depends on the distance from the nucleus and also the type of orbital.

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digital configuration of Carbon (C), Nitrogen(N), Oxygen(O) are as follows: C(6)- 1s² 2s² 2p^{x1} 2p^{y1} N(7)- 1s² 2s² 2p^{x1} 2p^{y1} 2p^{z1} O(8)- 1s² 2s² 2p^{x2} 2p^{y1} 2p^{z1} the following, Carbon has 2 electrons in its 2p subshell. So, there is an chance to make good 1/2-crammed P orbital contained in the perfect shell of Carbon by gaining a million electron.

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Best Answer: I just had to do this for an online homework assignment. I hope it helps. (a) Co (Z = 27) electron configuration: 1s² 2s² 2p⁶ 3s² 3p⁶ 4s² 3d⁷ condensed electron configuration: [Ar] 4s² 3d⁷

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I don't understand electron configuration can some explain to me. I have a test tomorrow on it. [Rn] 7s² 5f¹² i have a chart off this to do but i dont understand electron configuration! what about when there is 3??? like: 4s² 3d¹⁰ 4p⁶? its confusing and i dont understand how the element name for that is aluminum. i need an explanation please.

10 points ELECTRON CONFIGURATION ... - id.answers.yahoo.com

Best Answer: count to the left one for each e- you took away so Ir+4 is isoelectronic with Ta and thus has configuration of 1s² 2s² 2p⁶ 3s² 3p⁶ 4s² 3d¹⁰ 4p⁶ 5s² 4d¹⁰ 5p⁶ 6s² 4f¹⁴ 5d³

CHEMISTRY HELP: Electron Configurations? | Yahoo Answers

what are the long & shorthand configurations for ; Fluorine , Copper , Oxygen ion , Nickel , Iron & Sodium ion. thankss. (:

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Best Answer: Copper exists with the electron configuration of [Ar] 4s¹ 3d¹⁰ because it is more stable. This is because Copper will want to have a very stable d orbital, but it only has 9 out of 10.

Which is the more stable form of ... - in.answers.yahoo.com

Bromine is element 35, so 1s² 2s² 2p⁶ 3s² 3p⁶ 4s² 3d¹⁰ 4p⁵ is the correct electron configuration In short hand, this would be [Ar] 4s² 3d¹⁰ 4p⁵.

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Yahoo UK & Ireland Answers ... What would be the electron configuration of the following elements: a) Calcium b) Nitrogen c)Bromine d) Gallium. Follow . 5 answers 5. Report Abuse. Are you sure that you want to delete this answer? ... Electron configurations and electron affinity of nitrogen ?

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Electron configuration of Cu²⁺. STEP 1 Write down the electron configuration of the atom, Cu. [Ar] 3d¹⁰ 4s¹ STEP 2 The ion has a 2+ charge so two electrons are removed. One is taken from the 4s sub-shell, the other from the 3d sub-shell. So the electron configuration of Cu²⁺ is: 1s² 2s² 2p⁶ 3s² 3p⁶ 3d⁹.

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