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In the case of carbon dioxide, we have two dipoles. According to an electronegativity chart, oxygen is more electronegative than carbon, so we draw our dipoles pointing from carbon to oxygen.

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CS₂ and CO₃²⁻ follow the octet rule. The octet rule is a chemical rule of thumb that reflects observation that atoms of main-group elements tend to combine in such a way that each atom has eight electrons in its valence shell, giving it the same electronic configuration as a noble gas.

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