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ΔS = change in entropy or randomness and Complete the table for the sign of ΔG ; +, or undetermined, When conditions allow for an undetermined sign of ΔG , temperature will decide spontaneity. Answer the questions below. The conditions in which ΔG is always negative is when ΔH is and ΔS is ΔH ΔS 9. 3. ... Chemistry IF8766 ...

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Name _____ Date _____ Period ____ ENTROPY and GIBBS FREE ENERGY . Entropy is the degree of randomness in a substance. The symbol for change in entropy is ΔS . Solids are very ordered and have low entropy. Liquids and aqueous ions have more entropy because they move

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Entropy, S , is a measure of how dispersed or spread out a system's energy is among the available energy levels. The number of energy levels and how many are populated depend on a number of factors. As temperature is increased more of the excited state energy levels will be populated Increasing temperature always increases entropy

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CH302 Spring 2009 Worksheet 1 Answer Key: A Little Thermo Review 1. What two processes (a.k.a. path functions) can transfer internal energy between a system and its surroundings? What symbols are used for these variables? Heat and work. q for heat and w for work. 2. Write a good definition for a state function.

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