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Background. The second law of thermodynamics states that \geq , where Q is the amount of energy the system gains by heating, T is the temperature of the surroundings, and ΔS is the change in entropy. The equal sign refers to a reversible process, which is an imagined idealized theoretical limit, never actually occurring in physical reality, with essentially equal temperatures of system and surroundings.

Isentropic process - Wikipedia

The first and second law of thermodynamics are the most fundamental equations of thermodynamics. They may be combined into what is known as fundamental thermodynamic relation which describes all of the changes of thermodynamic state functions of a system of uniform temperature and pressure. (Schmidt-Rohr 2014) As a simple example, consider a system composed of a number of k different types of ...

Thermodynamic equations - Wikipedia

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