

## *Exam Style Questions All Answers*

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### Exam Style Questions All Answers

Answers to exam-style questions for Paper 1 Section B An abbreviated answer is given for one Paper 1 Section B example question per syllabus topic. Section B questions have a maximum mark allocation of 15 marks. Marks should be allocated

### Answers to exam-style questions for Paper 1 Section B

Answers to exam-style questions Topic 5 d 1 B 2 C 3 B 4 D 5 C 6 B 7 D 8 D 9 B 10 A 11 a The average amount of energy required to break one mole of covalent bonds in a gaseous molecule under standard conditions. 'Average' refers to the fact that the bond enthalpy is different in different

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be +7 in order to cancel out all but one of the negative charges and leave an overall charge of -1. ii Fe<sup>2+</sup> is the reducing agent (because it has been oxidised). (aq) iii The equation can be separated into its two half equations: MnO<sub>4</sub><sup>-</sup>(aq) reduction reaction. + H<sup>+</sup>(aq) → Mn<sup>2+</sup>(aq) + H<sub>2</sub>O(l) d (aq) Fe<sup>2+</sup>(aq) → Fe<sup>3+</sup>(aq)

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are constant/the concentrations of all reactants and products remain constant. The rate of the forward reaction is equal to the rate of the reverse reaction. b i The position of equilibrium shifts to the right, as the forward reaction is endothermic. As the temperature is increased, the position of equilibrium shifts in the direction that

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