

## ***Gizmo Solutions Titration***

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**Gizmo Solutions Titration**

Measure the quantity of a known solution needed to neutralize an acid or base of unknown concentration. Use this information to calculate the unknown concentration. A variety of indicators can be used to show the pH of the solution.

**Titration Gizmo : ExploreLearning**

Use this information to calculate the unknown concentration. A variety of indicators can be used to show the pH of the solution. Launch Gizmo. Titration. Measure the quantity of a known solution needed to neutralize an acid or base of unknown concentration. Use this information to calculate the unknown concentration.

**Titration Gizmo : Lesson Info : ExploreLearning**

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In the Titration Gizmo™, you will use indicators to show how acids are neutralized by bases, and vice versa. To begin, check that 1.00 M NaOH is selected for the Burette, ... Solutions with a pH below 7.0 are acidic, while those with a pH above 7.0 are basic. 3. Move the slider on the burette to the top to add about 25 mL of NaOH to the flask.

**Student Exploration: Titration**

Measure: A titration can be used to determine the concentration of an acid or base by measuring the amount of a solution with a known concentration, called the titrant, which reacts completely with a solution of unknown concentration, called the analyte. The point at which this occurs is called the equivalence point.

**Student Exploration: Titration - ScienceGeek.net**

Titration is a common lab method used to determine the concentration of one material that is dissolved or mixed within another. During the process of titration, a chemist slowly adds an acid (or base) solution of known concentration to a base (or acid) solution of unknown concentration, called the analyte.

**Gizmo of the Week: Titration | ExploreLearning News**

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**Titration - Chem + French Notes 2016**

Practice: Perform the following titrations and determine the concentrations of the following solutions. In each experiment, list the volume of titrant needed to neutralize the analyte and the indicator used. Use the Worksheet tab of the Gizmo to calculate each analyte concentration. Include all units.

**Titration - About**

\_\_\_\_ Turn Blue \_\_\_\_ Gizmo Warm-up Litmus is an example of an indicator, a substance that changes color depending on its pH (pH is a measure of the concentration of protons, or  $H^+$  ions). In the Titration Gizmo™, you will use indicators to show how acids are neutralized by bases, and vice versa.

**M10L3M1TitrationGizmo - Name Date Student Exploration ...**

Titration Practice Worksheet Find the requested quantities in the following problems: 1) If it takes 54 mL of 0.1 M NaOH to neutralize 125 mL of an HCl solution, what is the concentration of the HCl?

2) If it takes 25 mL of 0.05 M HCl to neutralize 345 mL of NaOH solution, what is the concentration of the NaOH solution?

**Titration Practice Worksheet - chemunlimited.com**

Activity B (continued from previous page) 4. Calculate: Select the Worksheet tab. This tab helps you calculate the analyte concentration. Fill in the first set of boxes ("moles H<sub>2</sub>SO<sub>4</sub>" and "moles NaOH") based on the coefficients in the balanced equation.

**Activity B continued from previous page 4 Calculate Select ...**

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Gizmo Warm-up. Litmus is an example of an indicator, a substance that changes color depending on its pH (pH is a measure of the concentration of protons, or H<sup>+</sup> ions). In the Titration Gizmo™, you will use indicators to show how acids are neutralized by bases, and vice versa. To begin, check that 1.00 M NaOH is selected for the Burette,

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