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Acids, Bases, and pH A water molecule (H_2O) can split apart to form a hydrogen ion (H^+) and a hydroxide ion (OH^-). The pH scale measures the concentration of hydrogen ions in a solution. The scale ranges from 0 to 14. Pure water has a pH of 7. An acid is any compound that forms H^+ ions in solution. Acidic solutions have pH values below 7.

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Slide 9 of 35. Acid-Base Theories >Arrhenius Acids and Bases. Arrhenius said that acids are hydrogen- containing compounds that ionize to yield hydrogen ions (H^+) in aqueous solution. He also said that bases are compounds that ionize to yield hydroxide ions (OH^-) in aqueous solution.

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Chemistry - Chapter 19 - Acids, Bases, and Salts. K_a , The ratio of the concentration of the dissociated (or ionized) form of an acid to the concentration of the undissociated (nonionized) form. $K_{eq} \text{X} [\text{H}_2\text{O}] = K_a = \frac{[\text{H}_3\text{O}^+][\text{CH}_3\text{COO}^-]}{[\text{CH}_3\text{COOH}]}$

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