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The atom is a ball of positive charge with electrons stuck into the ball. Electrons are arranged in concentric circular paths around the nucleus. Most of an atom's mass is concentrated in the small, positively charged nucleus. The electrons surround the nucleus and the rest of the atom is empty space. false farther It is called a quantum.

SECTION 5.1 MODELS OF THE ATOM (pages 127-132)

- spin: a quantum mechanical property of electrons; it may be thought of as clockwise or counterclockwise
- Hund's rule : electrons occupy orbitals of the same

Chapter 5

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many electrons will be involved in the formation of ions, the following suggestions may be helpful:

1. The atoms of Group 1 metals lose one electron when they form ions. The atoms of Group 2 metals lose two electrons when they form ions. The atoms of Group 3 metals lose three electrons when they form ions.
- 2.

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atoms. • electron configuration: the arrangement of various orbitals around the nuclei of atoms

5.2 Electron Arrangement in Atoms > CHEMISTRY YOU

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A Correlation of Pearson Chemistry ©2012 to the Next Generation Science Standards – DRAFT, May 2012 Grades 9-12 ... c. Develop explanations about how the patterns of electrons in the outer level of atoms, as represented in the periodic table, reflect and can predict properties of elements.

[Clarification Statement: An example of a

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This video describes light as a particle and wave. It also describes matter and quantum of energy.

Chapter 5 Electrons in Atoms Pt 1

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5.2 – Quantum Theory and the Atom. • Bohr Model of the Atom o According to his model, hydrogen's atomic emission spectrum results from electrons dropping from higher-energy atomic orbits to lower-energy atomic orbits. • The Quantum Mechanical Model of the Atom o It is based on the assumption that electrons are waves.

CHEMISTRY CHAPTER 5 OUTLINE and NOTES - Glenco TB

nuclei of atoms the ejection of electrons by metals when light shines on them the region around the nucleus of an atom where an electron is likely to be moving An atomic orbital may describe at most two electrons. the regions within which electrons have the highest probability of being found When electrons occupy orbitals of equal energy, one

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