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In addition to the work on this text, Cox is a co-author of four editions of Lehninger Principles of Biochemistry. He has received awards for both his teaching and his research, including the 1989 Eli Lilly Award in Biological Chemistry, and two major teaching awards from the University of Wisconsin and the University of Wisconsin System.

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6. Calculation of pH from Concentration of Strong Acid Calculate the pH of a solution prepared by diluting 3.0 mL of 2.5 M HCl to a final volume of 100 mL with H₂O. Answer Because HCl is a strong acid, it dissociates completely to H⁺ Cl⁻. Therefore, 3.0 mL 2.5 M HCl 7.5 meq of H⁺. In 100 mL of solution, this is 0.075 M H⁺.

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Answer The strength, or force (F), of ionic interactions in a solution depends on the magnitude of the charges (Q), the distance between the charged groups (r), and the dielectric constant (ϵ , which is dimensionless) of the solvent in which the interactions occur, according to the following equation:

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