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PERCENTAGE COMPOSITION ine the percentage composition of each of the compounds below. Name. O = 950(051- 070. 2. /Hé?b. H = Cl = V5 '. 3.9M9(N03)2. SCIENCE FAIR RESOURCES equipment worksheet ANSWER KEY TO UNIT G PRE-TEST and study guide J1_Unit J Outline_ Moles, Percentage Composition, Empirical and

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2S: $2(23.0) + 32.1 = 78.1$ g/mol 9H. 2O: $9(18.0) = 162.0$ g/mol Na. $2S \cdot 9H$. 2O (total): 240.1 g/mol %H. $2O = 162.0/240.1 \cdot 100 = 67.5\%$ 2. A 5.0 g sample of a hydrate of BaCl. 2 was heated, and only 4.3 g of the anhydrous salt remained.

KEY - Composition of Hydrates

Percent Composition Directions: Determine the percentage composition of each of the compounds below. ... Chemistry IF8766 Page 54 Instructional Fair, Inc. Title: Microsoft Word - pg 54 - Percent Composition.doc Author: Nathan

pg 54 - Percent Composition

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Percent Composition Worksheet - Solutions. Find the percent compositions of all of the elements in the following compounds: 1) CuBr 2 Copper (II) Bromide Cu: 28.45% Br: 71.55% 2) NaOH Sodium Hydroxide Na: 57.48% O: 39.99% H: 2.53% 3) (NH 4)2S Ammonium Sulfide N: 41.1% H: 11.8% S: 47.1% 4) N2S2 Dinitrogen disulfide.

Percent Composition Worksheet II

Percentage Composition Worksheet - Answer Key Give the % composition of all elements in these compounds. Show all work! _____ 1) ammonium sulfite % N 24.1 % H 6.9 % S 27.6 % O 41.3 _____ 2) aluminum acetate % Al 13.2 % C 35.3

Percentage Composition Worksheet - Auburn School District

PERCENTAGE COMPOSITION Name _____
Find the percentage composition of each of the compounds below.
1. MnO_2 2. $(\text{NH}_4)_3\text{PO}_4$ 3. C_2H_6 4. Zn 5. Al Solve the following problems.
6. How many grams of oxygen can be produced from the decomposition of 100. g of KClO_3 ?
7. How much iron can be recovered from 25.0 g of Fe_2O_3 ?

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Chemistry IF8766 Page 27 Instructional Fair, Inc. Title: Microsoft Word - pg 27 - Atomic Structure.doc Author: Nathan Created Date: 3/28/2007 10:37:14 PM ... Percent Composition

Instructional Fair Inc Chemistry IF8766 27

Write the correct formula for calcium acetate and then answer 21 - 23 based on it.
21. What is the mass of exactly one mole of calcium acetate? $\text{Ca}(\text{C}_2\text{H}_3\text{O}_2)_2$ 158g/mol
22. How many moles are contained in 1.58 g of the substance in #21? 0.01mol
23. How much does 0.400 mole of #21 weigh? 63.3g
24. Write the formula for oxygen gas. O_2
25.

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