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Because they affect both the charge and the structure of molecules, redox reactions are said to be coupled. If an electron is transferred between molecules, one molecule is oxidized while the other molecule is reduced.

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Redox (short for reduction-oxidation reaction) (pronunciation is a chemical reaction in which the oxidation states of atoms are changed. Any such reaction involves both a reduction process and a complementary oxidation process, two key concepts involved with electron transfer processes.

Oxidation and Reduction | Exams Daily

oxidation-reduction reaction redox reaction oxidation reduction oxidizing agent reducing agent. Section 20.1 Oxidation and Reduction. In Chapter 10, you learned that a chemical reaction can usually be classified as one of five types—synthesis, decomposition, combustion, single-replacement, or double-replacement.

Chapter 20: Redox Reactions - Neshaminy School District

Oxidation Reduction Reactions- Answer Key. 4.51. If nitric acid is a strong oxidizing agent and zinc is a strong reducing agent, then zinc metal will probably reduce nitric acid when the two react; that is, N will gain electrons and the oxidation number of N must decrease.

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OXIDATION REDUCTION REACTION PEARSON ANSWER KEY

Chapter 20 Oxidation-Reduction Reactions 523 Practice Problems In your notebook, solve the following problems. SECTION 20.1 THE MEANING OF OXIDATION AND REDUCTION Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent. 1. $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$ 2. $2\text{Li} + \text{S} \rightarrow 2\text{Li}_2\text{S}$ 3. $2\text{Cs} + \text{Br}_2 \rightarrow 2\text{CsBr}$ 4. $3\text{Mg} + \text{N}_2 \rightarrow \text{Mg}_3\text{N}_2$ 5. $4\text{Fe} + 3\text{O}_2 \rightarrow 2\text{Fe}_2\text{O}_3$

05 CTR ch20 7/12/04 8:17 AM Page 517 THE MEANING OF ...

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5. In redox reactions, how are oxygen atoms generally balanced? Oxidation is when the total number of electrons increases in a reaction; reduction is when the total number of electrons decreases in a reaction. Reduction is a reaction that removes an electron from a substance; oxidation is a reaction that adds electrons to a substance.

Prentice Hall Chemistry Chapter 20: Oxidation-Reduction ...

Practice Problems: Redox Reactions (Answer Key) Determine the oxidation number of the elements in each of the following compounds: a. H_2CO_3 H: +1, O: -2, C: +4 b. N_2 N: 0 c. $\text{Zn}(\text{OH})_2$ Zn: 2+, H: +1, O: -2 d. NO_2 N: +3, O: -2 e. LiH Li: +1, H: -1 f. Fe_3O_4 Fe: +8/3, O: -2; Identify the species being oxidized and reduced in each of the following reactions:

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