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Number Number Number Atomic of Protons of Electrons of Valence Type of Ion Oxidation Element Number (+) (-) Electrons Formed Number HydrOgen 1 1 1 1 ÷ 1÷ Helium Lithium Beryllium Boron Carbon Nitrogen Oxygen Fluorine Neon Sodium Magnesium, Aluminum Silicon Phosphorus Sulfur Chlorine Argon Potassium Calcium I @ Prentice-Hall, Inc. 31

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If you need more than one polyatomic ion copy to make oxidation numbers add up to zero, use parentheses. Example: Mg(+2) N03(-1) Magnesium wants to give away two electrons, but nitrate(N03) wants to gain one. You must have two N03 groups to pick up al/ the electrons given by magnesium.

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Prentice Hall Chemistry, Chapter 20: Oxidation-Reduction Reactions. The oxidation number of hydrogen in a compound is, except in metal hydrides, such as NaH, where it is -1.

Prentice Hall Chemistry, Chapter 20: Oxidation-Reduction ...

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The oxidation number of fluorine in a compound is always -1. Oxygen has an oxidation number of -2 unless it is combined with F (when it is +2), or it is in a peroxide (such as 1-1202 or Na2O2), when it is The oxidation state of hydrogen in most of its compounds is +1 unless it is combined with a metal, in which case it is -1 In compounds, the elements of groups 1 and 2 as well as aluminum have oxidation numbers of

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Assigning Oxidation Numbers: Home: Oxidation-reduction reactions (redox reactions) are reactions in which electrons are lost by an atom or ion in one reactant and gained by an atom or ion in another reactant. Although electrons are gained and lost in these reactions, the balanced equation for a redox reaction does not show the electrons that ...

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Oxidation Numbers Chart The oxidation number (or oxidation state) of an element represents the number of electrons that are gained, transferred or shared. The oxidation number of an element is a positive or negative number assigned to an element which can be used to predict the chemical

bonds that it will form with other elements.

Oxidation Numbers Chart | Periodic Table with Oxidation ...

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The Oxidation-Reduction Reactions chapter of this Prentice Hall Chemistry Companion Course helps students learn the essential lessons associated with oxidation-reduction reactions.

Prentice Hall Chemistry Chapter 20: Oxidation-Reduction ...

Section A: Complete the chart using a periodic table to help you. Element Atomic Symbol Total # of Electrons # of Valence Electrons # of Electrons Needed to Gain or Lose (to Fill Outer Shell)

Oxidation Number Chlorine Cl 17 7 Gain 1 1- Potassium K 19 1 Lose 1 1+ Magnesium Mg 12 2 Lose 2 2+ Fluorine F 9 7 Gain 1 1-

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