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Empirical Formula. a chemical formula that shows the composition of a compound in terms of the relative numbers and kinds of atoms in the simplest ratio.

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Chapter 10 - Chemical Quantities

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Percent Composition and Chemical Formulas 11) A compound analyzed in a chemistry laboratory consists of 5.34 g of carbon, 0.42 g of hydrogen, and 47.08 g of chlorine.

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Textbook: Nivaldo J. Tro, ^Chemistry: A Molecular Approach __, Pearson/Prentice Hall, 2008.
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Chapter 10 Chemical Quantities 241 Section Review Objectives • Relate Avogadro's number to a mole of a substance • Calculate the mass of a mole of any substance • Describe methods of measuring the amount of something • Compare and contrast the atomic mass of an element and its molar mass Vocabulary Key Equations • moles representative particles • representative particles moles

Objectives Vocabulary Key Equations

1. Introduction to Chemistry 2. Matter and Change 3. Scientific Measurement 4. Atomic Structure 5. Electrons in Atoms 6. The Periodic Table and Chemical Periodicity 7. Ionic Bonding and Ionic Compounds 8. Covalent Bonding 9. Chemical Names and Formulas 10. Chemical Quantities 11. Chemical Reactions 12. Stoichiometry 13. States of Matter 14. The ...

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