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$P_{ideal} = 1.0004$, $P_1 = nRT/(V - nb)$ $P_1 = 1.0029$, $P_2 = (a \cdot n^2)/V^2$ $P_2 = 0.0129$, $P_{waals} = P_1 +$
 P_2 $P_{waals} = 1.0158$ The ideal gas law predicts a pressure of 1.0004 atmospheres, while the van
der Waals model predicts 1.0158 atmospheres.

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