

## ***Oxidation Number Exercise Answers***

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### Oxidation Number Exercise Answers

Oxidation Number Exercise - answers Page 57 Oxidation Number Exercise Do not hand in this work sheet. When you are ready, you will be given an examination over this material. Complete the examination by yourself and hand it in to receive credit. Purpose: This exercise is designed to teach the student how to assign oxidation numbers.

### Oxidation Number Exercise - Adobe Media Server

An oxidation number is a positive or negative number assigned to an atom according to a set of rules. Redox reactions can be balanced by the use of oxidation numbers. A simple way to remember a monatomic ion's oxidation number is to recall the number of electrons it gains or loses, which is based on its group number.

### Oxidation Numbers Quiz - Softschools.com

Calculating Oxidation Numbers for Carbon. Calculate the oxidation number for iodine in each of the following compounds: (a) iodine, I<sub>2</sub> (b) potassium iodide, KI (c) silver periodate, AgIO<sub>4</sub>. 4 (d) zinc iodate, Zn(IO<sub>3</sub>)<sub>2</sub>. Answers: (a) 0; (b) - 1; (c) +7; (d) +5. Practice Exercise

### Example Exercise 17.1 Calculating Oxidation Numbers for Carbon

Oxidation Showing top 8 worksheets in the category - Oxidation . Some of the worksheets displayed are Work 7, Oxidation state work and answers, Oxidation number exercise, Work 1 determination of oxidation number or valence, Work oxidation numbers name, Oxidation number work 1 write the rule next to your, Chapter 20 work redox, Key review work on balancing redox equations.

### Oxidation Worksheets - Printable Worksheets

Oxidation Numbers Worksheet Directions: Use the Rules for Assigning Oxidation Numbers to determine the oxidation number assigned to each element in each of the given chemical formulas. Formula Element and Oxidation Number Formula Element and Oxidation Number 1. Cl<sub>2</sub> Cl 16. Na<sub>2</sub>O 2 Na O 2. -Cl Cl 17. SiO<sub>2</sub> Si O 3. Na<sub>2</sub>Na 18. CaCl<sub>2</sub>

### Oxidation Numbers Worksheet - brookville.k12.oh.us

WORKSHEET - ASSIGNING OXIDATION NUMBERS Name \_\_\_\_\_ Period \_\_\_\_\_ Oxidation Number Rules: 1. A pure element has an oxidation number of 0. 2. The oxidation number of an element in a monatomic ion equals the charge of the ion. 3. The sum of the oxidation number of all the elements in a compound equals 0.

### WORKSHEET - ASSIGNING OXIDATION NUMBERS

Exercise - Balance the following redox equations using either the inspection technique or the oxidation number method. Be sure to check that the atoms and the charge are balanced for each. (The answers are at the bottom of this page.) a. Al(s) + MnO<sub>2</sub>(s) → Al<sub>2</sub>O<sub>3</sub>(s) + Mn(s)

### Exercise for Balancing Redox Equations Using the Oxidation ...

1. Give the oxidation numbers of all the elements in the following molecules and ions: a. SO<sub>2</sub>, SO<sub>3</sub>, SO<sub>3</sub><sup>2-</sup>, SO<sub>4</sub><sup>2-</sup> b. ClO<sub>2</sub>, ClO<sub>2</sub><sup>-</sup>, ClO<sub>2</sub><sup>2-</sup>, ClO<sub>3</sub><sup>-</sup>, ClO<sub>4</sub><sup>-</sup> c. N<sub>2</sub>O, NO, NO<sub>2</sub>, N<sub>2</sub>O<sub>4</sub>, N<sub>2</sub>O<sub>5</sub>, NO<sub>2</sub><sup>-</sup>, NO<sub>3</sub><sup>2-</sup> 2. Determine the oxidation number of the sulfur atom: a. H<sub>2</sub>S b. S c. H<sub>2</sub>SO<sub>4</sub> d. S<sub>2</sub><sup>2-</sup> e. HS<sup>-</sup> f. SO<sub>2</sub> g. SO<sub>3</sub>

### Worksheet: Oxidation Numbers Name - MARRIC

oxidation numbers of element X are A. +1 and +2 B. +2 and +3 C. +1 and +3 D. +2 and +4 8. Oxygen has a positive oxidation number in the compound A. H<sub>2</sub>O B. H<sub>2</sub>O<sub>2</sub> C. OF<sub>2</sub> D. IO<sub>2</sub> 9. What is the oxidation number of sulfur in H<sub>2</sub>SO<sub>4</sub>? A. 0 B. 2 C. +6 D. +4 10. In the equation Cu(s) + 2Ag<sup>+</sup>(aq) → Cu<sup>2+</sup>(aq) + 2Ag(s), the oxidizing agent is A. Cu<sub>0</sub> B. Ag<sup>+</sup> C. Cu<sup>2+</sup> D. Ag<sub>0</sub> 11.

### Redox practice worksheet - Imghs.org

Practice Problems: Redox Reactions (Answer Key) Determine the oxidation number of the elements in each of the following compounds: a. H<sub>2</sub>CO<sub>3</sub> H: +1, O: -2, C: +4 b. N<sub>2</sub>N: 0 c. Zn(OH)<sub>2</sub> Zn:

2+, H: +1, O: -2 d. NO<sub>2</sub>-N: +3, O: -2 e. LiH Li: +1, H: -1 f. Fe<sub>3</sub>O<sub>4</sub> Fe: +8/3, O: -2; Identify the species being oxidized and reduced in each of the following reactions:

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