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atomic number (Z) the number of protons in the nucleus of an atom; unique to each element, determines the identity of the element. cathode ray. a stream of electrons produced at the negative electrode of a tube containing a gas at low pressure; JJ Thomson used a cathode ray in an experiment to discover electrons.

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Dalton's Atomic Theory. 1) elements are composed of atoms. 2) atoms of same element are identical, but differ from other elements. 3) elements can mix together 4) atoms only change when mixed with other elements.

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Ch. 4 Atomic Structure Review Day - CALVERT CHEM

Materials: Prentice-Hall textbook; Chemistry by Wilbraham, Staley, Marta, and Waterman
Objectives: All students will: Be proficient in the fundamentals of chemistry, based on the two great influences; the theory of the atom and the law of conservation of mass. Core Concepts: Chemistry is the study of matter and the changes it undergoes.

Prentice-Hall textbook; Chemistry by Wilbraham, Staley ...

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Chemistry I - Mr. Benjamin's Classroom

Chapter 4 Atomic Structure Section 4.1 Studying Atoms (pages 100-105) This section discusses the development of atomic models. Reading Strategy(page 100) Summarizing As you read, complete the table about atomic models. For more information on this Reading Strategy, see the Reading and Study Skills in the Skills and Reference Handbook at the end of

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Hydrogen is a chemical element with symbol H and atomic number 1 Chapter 7 ionic and metallic bonding answer key prentice hall. With a standard atomic weight of 1. 008, hydrogen is the lightest element on the periodic table.

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