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pH and pOH CONTINUED 1. 0.01 M HCl $[H^+] = 0.01M$ $pH = -\log(0.01) = 2$ 2. 0.0010 M NaOH $[OH^-] = 0.0010M$ $pOH = -\log(0.010) = 3 \rightarrow pH = 11$ 3. 0.050 M NaOH

pH and pOH CONTINUED - newburyparkhighschool.net

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pH and pOH Practice Worksheet - ThoughtCo

Calculating pH and pOH worksheet W 335 Everett Community College Tutoring Center Student Support Services Program 1) What is the pH of a 0.0235 M HCl solution? 2) What is the pOH of a 0.0235 M HCl solution? 3) What is the pH of a 6.50×10^{-3} M KOH solution? (Hint: this is a basic solution -

Calculating pH and pOH worksheet

Part 2: For each of the problems below, assume 100% dissociation. 1. A. Write the equation for the dissociation of hydrochloric acid. B. Find the pH of a 0.00476 M hydrochloric acid solution. 2. A.

pH and pOH - msmogcksclassroom.com

pH, pOH, $[H^+]$, $[OH^-]$ Worksheet 1) What is the $[H^+]$ for a one liter aqueous solution containing 5.73×10^{17} H^+ ions? 2) A 0.23 M $HClO_2$ solution is 6.3% ionized. What is the $[H^+]$? 3) What is the pH of a 0.034 M $Ba(OH)_2$ solution? 4) What is the hydroxide concentration in 2.7×10^{-3} M HCl? 5) Calculate the hydrogen ion concentration for an aqueous solution whose

pH, pOH, $[H^+]$, $[OH^-]$ Worksheet - mmsphyschem.com

Solutions for the pH practice worksheet: The important thing to remember for all of these problems is that $pH = -\log [H^+]$, and that $[H^+]$ is equivalent to the molarity of acid present in a solution. When the pH is less than 7, the solution is acidic, when the $pH = 7$ it is neutral, and when it is greater than 7, it is basic.

pH Practice Worksheet - mrphysics.org

Acids/Bases & pH Worksheet (continued) Complete the following table by filling in the empty spaces. Indicate if the solution is acidic, basic or neutral. You should be able to do the starred (*) items without a calculator. (Problem #23-34) [H³⁰] + [OH][–] pH pOH Acidic/Basic/Neutral

Acids/Bases & pH Worksheet

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