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6 The answer is B. The important step that determines the rate of reaction is that between the halogen radical and the alkane. e Option A is not correct because the reaction is between chlorine radicals and the alkane not chlorine molecules. Options C and D are incorrect — the Cl-Cl bond is the stronger. 7 The answer is B. The initial attack is by a δ^+ Br atom.

6 The answer is B The important step that determines the ...

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As the iodine is produced it reacts with the sodium thiosulfate, until all the sodium thiosulfate has been reacted. The next iodine produced turns the starch dark blue. volume of sodium thiosulfate A measure of the rate is time taken. b In experiments 1 and 2, $[I^-]$ remains constant and $[S_2O_8^{2-}]$ is doubled.

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This teacher resource pack follows on from the earlier Philip Allan Updates publication Organic Chemistry 2 (ISBN: 978-0-86003-223-6). It contains new material that gives teachers alternative activities to help in their teaching of A2 chemistry.

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