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62 Review Module / Chapters 13-16 ____11. The electron configuration of a fluoride ion, F²⁻, is: a. 1s² 2s² 2p⁵. b. the same as that of the neon atom. c. 1s² 2s² 2p⁶ 3s¹. d. the same as that of a

potassium ion. ____12. Metals are good conductors of electricity because they: a. form crystal lattices. b. contain positive ions. c. contain mobile valence electrons. d. form ionic bonds.

15 Ionic Bonding and Ionic Compounds Chapter Test A - LPS

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1.1 Chemistry Section Review - LPS

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