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Pearson Chemistry 8 Covalent Bonding

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Covalent Bond Strength • Most simply, the strength of a bond is measured by determining how much energy is required to break the bond. • This is the bond enthalpy. • The bond enthalpy for a Cl-Cl bond, $D(\text{Cl-Cl})$, is measured to be 242 kJ/mol.

Chapter 8 Concepts of Chemical Bonding

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8.2 The Nature of Covalent Bonding > 23 Copyright © Pearson Education, Inc., or its affiliates. All Rights Reserved. • Experimental evidence, however, indicates ...

Chapter 8

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Pearson Chemistry Foundations Edition Chapter 8 1 Covalent ...

•covalent bond: a bond formed by the sharing of electrons between atoms •molecule: a neutral group of atoms joined by covalent bonds •diatomic molecule: a molecule consisting of two atoms
Glossary Terms

Covalent Bonding - Pittsfield High School

Chapter 8 Covalent Bonding and Molecular Structure 8-6 Example: PF₃ Example: CO₂ Step 1: Count Valence Electrons Count the total number of valence electrons in the molecule or ion. Anions have extra electrons, so add 1 electron for each negative charge. Cations have a deficiency of electrons, so subtract 1 electron for each positive charge.

Chapter 8: Covalent Bonding and Molecular Structure

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