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PROPERTIES OF SOLUTIONS SECTION REVIEW

a solution containing the maximum amount of solute for a given.... amount of a substance that dissolves in a given quantity of so.... solution: a solution that contains less solute than a saturate.... describes liquids that dissolve in one another in all proporti....

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that are introduced in this section. Each blank can be completed with a term, short phrase, or number. Changes in the temperature of a system and of a 1. solute alter the at which a solute dissolves. The extent 2. to which a gas dissolves in a liquid is proportional to the 3. of the gas in accordance with law. The solubility of a gas 4.

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Chapter 16 Solutions 167. SECTION 16.1 PROPERTIES OF SOLUTIONS (pages 471–477) This section identifies the factors that affect the solubility of a substance and determine the rate at which a solute dissolves. Solution Formation (pages 471–472) Look at Figure 16.1 on page 471 to help you answer Questions 1 and 2.

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SECTION 16.2 CONCENTRATIONS OF SOLUTIONS 1. Calculate the molarity of each of the following solutions. a. 0.40 mol of NaCl dissolved in 1.6 L of solution b. 20.2 g of potassium nitrate, KNO₃, in enough water to make 250.0 mL of solution 2. Calculate the number of grams of solute needed to prepare each of the following solutions.

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Concepts. 2. Most matter exists as compounds because most pure elements are chemically unstable. Atoms of most elements quickly form chemical bonds with other atoms to make more stable compounds. 3. A covalent bond holds a water molecule together. 4. A molecule 5. The subscript tells you the ratio of atoms of each element in a compound.

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Solution Properties Review. Choose the correct answer for each question. Show all questions <=> The substance dissolved in a solution is the: ? solvent ? solution ? solute ? mixture; The dissolving medium in a solution is the: ... Some solutions will boil at a temperature below 100°C, and some of the solutions will boil at a temperature ...

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