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**Stoichiometry Chapter 9 Review Answers**

CHAPTER 9 REVIEW Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas:  $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$  How many moles of  $\text{O}_2$  form if 3.0 mol of  $\text{KClO}_3$  are totally consumed?

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Chapter 9 describes how to use mole ratios, molar masses, conversions, limiting reactants, and percent yield to ... Stoichiometry Review - ScienceGeek.net Homepage

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Modern Chemistry 73 Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. \_\_\_\_ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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Chapter 9: Standard Review Worksheet 1. Answers will vary. An example is included below:  $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$  This describes the decomposition reaction of hydrogen peroxide.

**Chapter 9: Standard Review Worksheet**

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