

Stoichiometry Vocabulary Review Answers

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Chapter 12 (Stoichiometry) Vocabulary. STUDY. PLAY. actual yield. the amount of product that forms when a reaction is carried out in the laboratory. excess reagent. a reagent present in a quantity that is more than sufficient to react with a limiting reagent; any reactant that remains after the limiting reagent is used up in a chemical reaction.

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Chapter 12 Stoichiometry Vocabulary Review Answer Key

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Stoichiometry Test Review 1. How many moles of oxygen are made if 12.0 moles of potassium chlorate react? $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ 18.00 mol O_2 2. Copper(II) chloride reacts w/sodium nitrate to produce copper(II) nitrate and sodium chloride. a) Write the balanced equation for the reaction. $\text{CuCl}_2 + 2\text{NaNO}_3 \rightarrow \text{Cu}(\text{NO}_3)_2 + 2\text{NaCl}$

Stoichiometry Test Review - Hillsboro City Schools

Chapter 12 Review Stoichiometry Important Vocabulary • Stoichiometry- The calculations of quantities in chemical reactions in a subject of chemistry • Mole ratio- a conversion factor derived from the coefficients of a balanced chemical equation interpreted in terms of moles

Chapter 12 Review - Scarsdale Public Schools / Overview

Practice Problems, Vocabulary Review, Quiz, Chapter Test A, Chapter Test B Technology • Computer Test Bank, Chapter 12 Test • Interactive Textbook with ChemASAP, Chapter 12 • Virtual Chem Labs, Lab 14 378 Chapter 12 12CHAPTER Study Guide Key Concepts 12.1 The Arithmetic of Equations • A balanced chemical equation provides the

CHAPTER 12 Study Guide - Quia

368 Chapter 11 • Stoichiometry Section 11.11.1 Objectives Describe the types of relationships indicated by a balanced chemical equation. State the mole ratios from a balanced chemical equation. Review Vocabulary reactant: the starting substance in a chemical reaction New Vocabulary stoichiometry mole ratio Defining Stoichiometry

Chapter 11: Stoichiometry

Stoichiometry Worksheet Review Name _____. 1. Given the following equation: $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$, show what the following molar ratios should be.

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Answer the following items in the space provided. 11. When 3.00 g of Mg is ignited in 2.20 g of pure oxygen, what is the limiting reactant? What is the theoretical yield of MgO? ... Concept Review: Stoichiometry and Cars 1. c 2. b 3. d 4. b 5. d 6. If there is too much oxygen and not enough gasoline, the engine will stall.

Skills Worksheet Concept Review - Marian High School

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