

## *Stoichiometry Practice Answer Key*

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### **Stoichiometry Practice Answer Key**

Practice Problems (Chapter 5): Stoichiometry CHEM 30A Part I: Using the conversion factors in your tool box g A mol A mol A 1. How many moles CH<sub>3</sub>OH are in 14.8 g CH<sub>3</sub>OH? 2. What is the mass in grams of 1.5 x 10<sup>16</sup> atoms S? 3. How many molecules of CO<sub>2</sub> are in 12.0 g CO<sub>2</sub>? 2 4. What is the mass in grams of 1 atom of Au? KEY Tool Box: To ...

### **Practice Problems (Chapter 5): Stoichiometry**

Practice Problems: Stoichiometry (Answer Key) Balance the following chemical reactions: a. 2 CO + O<sub>2</sub> → 2 CO<sub>2</sub> b. 2 KNO<sub>3</sub> → 2 KNO<sub>2</sub> + O<sub>2</sub> c. 2 O<sub>3</sub> → 3 O<sub>2</sub> d. NH<sub>4</sub>NO<sub>3</sub> → N<sub>2</sub>O + 2 H<sub>2</sub>O e. 4 CH<sub>3</sub>NH<sub>2</sub> + 9 O<sub>2</sub> → 4 CO<sub>2</sub> + 10 H<sub>2</sub>O + 2 N<sub>2</sub> f. Cr(OH)<sub>3</sub> + 3 HClO<sub>4</sub> → Cr(ClO<sub>4</sub>)<sub>3</sub> + 3 H<sub>2</sub>O Write the balanced chemical equations of each reaction:

### **Practice Problems: Stoichiometry**

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### **Stoichiometry Homework Sheet With Answer Key**

Chemistry: Stoichiometry - Problem Sheet 2 KEY 9) 2 24 2 2 23 2 2 2 4.63 x 10 molecules I 1 mol I 6.02 x 10 molecules I 1 mol Cl 1 mol 71 g Cl Cl x 546 g Cl 10) 292 g Ag 1 mol Ag 108 g Ag 1 mol Cu 1 mol Ag 63.5 g Cu

### **Stoichiometry: Problem Sheet 2**

Stoichiometry: Mixed Problems (KEY) 1) N<sub>2</sub> + 3H<sub>2</sub> → 2NH<sub>3</sub> What volume of NH<sub>3</sub> at STP is produced if 25.0 of N<sub>2</sub> is reacted with an excess of H<sub>2</sub>? 3 3 3 2 3 2 2 2 40.0L NH 1mol NH 22.4L NH 1mol N 2mol NH 28.0g N 25.0g N 1mol N x x x = 2) 2KClO<sub>3</sub> → 2KCl + 3O<sub>2</sub> If 5.0g of KClO<sub>3</sub> is decomposed, what volume of O<sub>2</sub> is produced at STP? 2

### **Stoichiometry: Mixed Problems (KEY)**

Chapter 6 Balancing and Stoichiometry Worksheet and Key Topics: • Balancing Equations • Writing a chemical equation • Stoichiometry Practice: 1. In the reaction: 4Li (s) + O<sub>2</sub> (g) → 2Li<sub>2</sub>O (s) a. what is the product? b. what are the reactants? c. what does the "(s)" after the formula of lithium oxide signify?

### **chapter 6 balancing stoich worksheet and key**

Stoichiometry Practice Worksheet Solve the following stoichiometry grams-grams problems: 1) Using the following equation: 2 NaOH + H<sub>2</sub>SO<sub>4</sub> → 2 H<sub>2</sub>O + Na<sub>2</sub>SO<sub>4</sub> How many grams of sodium sulfate will be formed if you start with 200.0

### **Stoichiometry Practice Worksheet - Social Circle City Schools**

Gas Stoichiometry Practice For all of these problems, assume that the reactions are being performed at a pressure of 1.0 atm and a temperature of 298 K. 1) Calcium carbonate decomposes at high temperatures to form carbon dioxide and calcium oxide: CaCO<sub>3</sub>(s) → CaO(s) + CO<sub>2</sub>(g) How many grams of calcium carbonate will I need to form 3.45 liters of LiSO

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Answer the following stoichiometry-related questions: 12) Write the balanced equation for the reaction of acetic acid with aluminum hydroxide to form water and aluminum acetate: 13) Using the equation from problem #12, determine the mass of aluminum acetate that can be made if I do this reaction with 125 grams of acetic acid

### **Stoichiometry Practice Worksheet**

• While you should practice working as fast as possible, it is more important at this point in the course, that you practice without a calculator, even if it slows you down. Look for the "easy math"

– common factors and rough estimation – do not do “long division” to try to get exact values.  
Remember it is a MC test, use the answers

### Practice Test Ch 3 Stoichiometry Name Per

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### Stoichiometry Answer Keys - Chem I - Google Sites

Stoichiometry Worksheet #1 Answers 1. Given the following equation:  $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$ , show what the following molar ratios should be. a.  $\text{C}_4\text{H}_{10} / \text{O}_2$  b.  $\text{O}_2 / \text{CO}_2$  c.  $\text{O}_2 / \text{H}_2\text{O}$  d.  $\text{C}_4\text{H}_{10} / \text{CO}_2$  e.  $\text{C}_4\text{H}_{10} / \text{H}_2\text{O}$  2. Given the following equation:  $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$  a. How many moles of  $\text{O}_2$  can be produced by ...

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