

Stoichiometry Mixed Problems Answer Key

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Stoichiometry: Mixed Problems (KEY)

ANSWER KEY. Mixed Stoichiometry Problems . 1. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$. a). How many moles of H_2 would be required to produce 5.0 moles of water? 5.0 moles water. b). What mass of H_2O is formed when H_2 reacts with 384 g of O_2 ? 432g H_2O . 2. $\text{H}_2\text{SO}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{SO}_4 + 2\text{H}_2\text{O}$. a). Balance this equation.

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STOICHIOMETRY: MIXED PROBLEMS Name What volume of NH_3 at STP is produced if 25.0 g of N_2 is reacted with an excess of H_2 ? 08008 2. $\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$ If 5.0 g of KClO_3 is decomposed, what volume of O_2 is produced at STP? 3. How many grams of KCl are produced in Problem 2? 4. $\text{Zn} + 2\text{HCl} \rightarrow \text{ZnCl}_2 + \text{H}_2$

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KEY Chemistry: Stoichiometry - Problem Sheet 1 Directions: Solve each of the following problems. Show your work, including proper units, to earn full credit. 1. Silver and nitric acid react according to the following balanced equation: $3\text{Ag(s)} + 4\text{HNO}_3\text{(aq)} \rightarrow 3\text{AgNO}_3\text{(aq)} + 2\text{H}_2\text{O(l)} + \text{NO(g)}$ A.

Stoichiometry: Problem Sheet 1

Chemistry: Stoichiometry - Problem Sheet 2 KEY 9) $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$, show what the following molar ratios should be. a. $\text{C}_4\text{H}_{10} / \text{O}_2$ b. O_2 / CO_2 c. $\text{O}_2 / \text{H}_2\text{O}$ d. $\text{C}_4\text{H}_{10} / \text{CO}_2$ e. $\text{C}_4\text{H}_{10} / \text{H}_2\text{O}$ 2. Given the following equation: $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

Stoichiometry: Problem Sheet 2

Stoichiometry Worksheet #1 Answers 1. Given the following equation: $2\text{C}_4\text{H}_{10} + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$, show what the following molar ratios should be. a. $\text{C}_4\text{H}_{10} / \text{O}_2$ b. O_2 / CO_2 c. $\text{O}_2 / \text{H}_2\text{O}$ d. $\text{C}_4\text{H}_{10} / \text{CO}_2$ e. $\text{C}_4\text{H}_{10} / \text{H}_2\text{O}$ 2. Given the following equation: $2\text{KClO}_3 \rightarrow 2\text{KCl} + 3\text{O}_2$

2 a. How many moles of O_2 can be produced by ...

Stoichiometry Worksheet #1 Answers

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