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Chemistry stoichiometry. 1. MOLES AND MASS Name Determine the number of moles in each of the quantities below. . 25g of NaCl 2. 125g of H_2SO_4 3. 74g of $KMnO_4$ 4. 74g of KCl 5. 35 g of CuSCVShLO Determine the number of grams in each of the quantities below. 1. 2.5 moles of NaCl 2. 0.50 moles of H_2SO_4 3.

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Stoichiometry: Mixed Problems (KEY) 1) $N_2 + 3H_2 \rightarrow 2NH_3$ What volume of NH_3 at STP is produced if 25.0 of N_2 is reacted with an excess of H_2 ? 3 3 2 3 2 2 40.0L NH_3 1mol NH_3 22.4L NH_3 1mol N_2 2mol NH_3 28.0g N_2 25.0g N_2 1mol N_2 $x \times x =$ 2) $2KClO_3 \rightarrow 2KCl + 3O_2$ If 5.0g of $KClO_3$ is decomposed, what volume of O_2 is produced at STP? 2

Stoichiometry: Mixed Problems (KEY)

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Chemistry IF8766 Page 66 Instructional Fair, Inc. Title: Microsoft Word - pg 66 - Stoichiometry-limiting reagent.doc Author: Nathan Created Date: Chemistry I Worksheet. Chemistry I Worksheet NAME Formula Mass, Percent Composition, and Empirical and Molecular Formulas INSTRUCTIONS: Work the following problems.

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Stoichiometry – Limiting Reagent Laboratory NAME _____ SECTION _____ 2 Because CaCl_2 contains one mole of calcium (II) ions per mole of calcium chloride and Na_2CO_3 contains one mole of carbonate ions per mole of sodium carbonate, the reagent with the fewest number of moles will be limiting.

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