

Stephen Murray Naming Compounds Answers

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Naming Compounds - Mrs. McCutchen's Class

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compound covalent compound ionic compound USE “- IDE” ENDING (NO PREFIXES!) Name the metal and non-metal and change the ending to “ide”. USE GREEK PREFIXES Put prefixes in front of element names to tell how many atoms are there. Don’t use “mono” for first name, but always for second name. CHECK THE CHART BELOW (NO PREFIXES!)

chap 19 no 4 - Mr Murray's Science and Music

two non-metals form a _____ compound. Fall Semester Review - IPC Calcium oxide (CaO) Water Find the molecular mass of the following compounds Classify (ionic, covalent or polyatomic) and name the following compounds. 1. ... C. Stephen Murray 1. “If I __ I full” is an easy way to remember the _____

Fall Semester Review - IPC - winterschemistry.com

Answers - Naming Chemical Compounds . Name the following chemical compounds: 1) NaBr sodium bromide. 2) Ca(C₂H₃O₂)₂ 2 calcium acetate. 3) P₂O₅ diphosphorus pentoxide. 4) Ti(SO₄)₂ titanium(IV) sulfate. 5) FePO₄ iron(III) phosphate. 6) K₃N potassium nitride. 7) SO₂ sulfur dioxide. 8) CuOH copper(I) hydroxide. 9) Zn(NO

Answers - Naming Chemical Compounds

write the compound as a formula. Way 2 Cl⁻¹ - You know it is a balanced compound because $2(1) + 1(-2) = 0$. Balanced ionic compounds have a neutral charge. Again, you know it is a balanced compound because $1(2) + 2(-1) = 0$. Balanced ionic compounds have a neutral charge. Make the ionic compound of magnesium oxide. Make lithium chloride.

Naming and Covalent Compounds - Mr Murray's Science and Music

naming compounds and ionic.covalent compounds -answer key.jnt Author: jrtowers508 Created Date: 20130926120810Z ...

naming compounds and ionic.covalent compounds -answer key

Name These Covalent Compounds Penta = Mono Tetra = Hexa = Hepta = Deca = N₃C₁₄ P₀₅
Disilicon oxide nitrogen tetrachloride Sulfur oxide Phosphorous sulfur ox fluor Ionic, co K₂S₀₄ MgBr₂
P₂₀₄ Name these Polyatomic Compounds (Remember — no prefixes!) Classify and Name These Compounds Covalent, Ionic Name Barium. chloride bon

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CSTEPHENMURRAY NAMING COMPOUNDS REVIEW ANSWER KEY PDF

Name the metal and non-metal and change the ending to "ide". Li₂S Metal and non-metal— ionic Lithium Sulfide (22! dilithium sulfide— no prefixes for ionic compounds) Why are ionic compounds so easy to name? Because most ionic com- pounds can only form one way, using the oxidation numbers. In covalent compounds, though, non-metals can

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Name These Ionic Compounds Magnesium Fluor-ide I ,ithium Ox- ide Sodium Chlor- Potassium Ox- Brom- the Polyatomic Chart on the front of the worksheet to name these Polyatomic Ions: carbonate NH₄ crop- OH⁻ P₀₄ coe COVälent Disilicon oxide nitrogen tetrachloride Sulfur oxlde Phosphorous 1. 5. 6. LiC₁ coz B. c. D E. dioxide Thcarbon monoxide

www.npenn.org

Covalent Compound Naming Worksheet 1 covalentname.sxw. Naming Covalent Compounds Solutions Write the formulas for the following covalent compounds: 1) antimony tribromide SbBr₃ 2) hexaboron silicide B₆Si 3) chlorine dioxide ClO₂ 4) hydrogen iodide HI 5) iodine pentafluoride IF₅

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