

## *Stoichiometry Worksheet 2 Percent Yield Answers*

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**Stoichiometry Worksheet 2 Percent Yield**

STOICHIOMETRY WORKSHEET 2 (PERCENT YIELD) Complete the following on SEPARATE paper showing all work. 1. Sodium carbonate reacts with nitric acid according to the following balanced equation: a. If 30 grams of sodium carbonate react with excess nitric acid, how many grams of sodium nitrate should be produced? b.

**STOICHIOMETRY WORKSHEET 2 (PERCENT YIELD)**

Name \_\_\_\_\_ Date \_\_\_\_\_ Pd \_\_\_\_\_ Stoichiometry Worksheet 2: Percent Yield Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get

**Name \_\_\_\_\_ Date \_\_\_\_\_ Pd \_\_\_\_\_ Stoichiometry Worksheet 2: Percent Yield ...**

Stoichiometry Worksheet 2: Percent Yield Name \_\_\_\_\_ Date \_\_\_\_\_ Pd \_\_\_\_\_ Stoichiometry Worksheet 2: Percent Yield For each of the problems below: a. Write the balanced chemical equation b. Identify the given (with units) and what you want to find (with units) c. Show set up with units. Check sig figs, give final answer with units and label.

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Name \_\_\_\_\_ Date \_\_\_\_\_ Pd \_\_\_\_\_ Stoichiometry Worksheet 2: Percent Yield For each of the problems below: a. Write the balanced chemical equation b. Identify the given (with units) and what you want to find (with units) c. Show set up with units. Check sig figs, give final answer with units and label.1.

**Name \_\_\_\_\_ Date \_\_\_\_\_ Pd \_\_\_\_\_ Stoichiometry Worksheet 2: Percent Yield ...**

Some of the worksheets displayed are Percent yield work, Work percent yield name, Chemistry percent yield, Stoichiometry work 2 percent yield, Plc resource cover understanding yield percentage, Theoretical and percent yield work, , Bakers percentages and formula weights. Once you find your worksheet, click on pop-out icon or print icon to worksheet to print or download. Worksheet will open in a new window.

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**Modeling Chemistry Stoichiometry Worksheet 2 Percent Yield ...**

2 O<sub>3</sub> is used, how many kilograms of iron can be produced? The reaction is: Fe<sub>2</sub>O<sub>3</sub> + 3 C → 2 Fe + 3 CO  
9. The average human requires 120.0 grams of glucose (C<sub>6</sub>H<sub>12</sub>O<sub>6</sub>) per day. How many grams of CO<sub>2</sub> (in the photosynthesis reaction) are required for this amount of glucose? The photosynthetic reaction is: 6 CO<sub>2</sub> + 6 H<sub>2</sub>O → C<sub>6</sub>H<sub>12</sub>O<sub>6</sub> + 6 O<sub>2</sub>

**Stoichiometry Worksheet 2 Answer Key - mrromswinckel.com**

If the percentage yield is 65.0 % calculate the actual yield. 2H<sub>2</sub> + O<sub>2</sub> → 2H<sub>2</sub>O . 100. g H<sub>2</sub> x 1 mole x 2 mole H<sub>2</sub>O x 18.02 g x 0.650 = 580 g H<sub>2</sub>O  
2.02 g 2 mole H<sub>2</sub> 1 mole . 9. Calculate the theoretical mass of iron produced from the reaction of 5.67 g of iron III oxide. If the percentage yield is 85.0 % calculate the actual yield.

**Worksheet #1 Stoichiometry - iannonechem.com**

d) If the actual yield of CO<sub>2</sub> is 69.2 g CO<sub>2</sub>, what is the percent yield? Answer: \_\_\_\_\_ 23.2 g 93.7 g 8 10 2 13 : C<sub>8</sub>H<sub>20</sub>O<sub>26</sub> charge 0 C<sub>4</sub>H<sub>10</sub> 70.3 g CO<sub>2</sub> actual yield CO<sub>2</sub> theoretical yield CO<sub>2</sub> % yield CO<sub>2</sub> = (100%) = 69.2 g CO<sub>2</sub> 70.271 g CO<sub>2</sub> 98.476% yield CO<sub>2</sub> 98.5 % yield CO<sub>2</sub> 1 mol C<sub>4</sub>H

**Practice Problems (Chapter 5): Stoichiometry**

Stoichiometry - Percent Yield Worksheet SHOW ALL WORK!!!! 0/0 Yield = Actual Yield x 10th Theoretical Yeild Theoretical Yield = answer to your stoich problem. Actuah Yield = given in the

problem or the experimental Yield. Balance the equation for the reaction of iron (III) phosphate with sodium sulfate to make iron (III) sulfate and sodium phosphate.  $2 \text{FePO}_4 + 3\text{Na}_2\text{SO}_4 \rightarrow \text{Fe}_2(\text{SO}_4)_3 + 6\text{Na}_3\text{PO}_4$

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