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Chapter 20 Oxidation-Reduction Reactions 221 SECTION 20.1 THE MEANING OF OXIDATION AND REDUCTION (pages 631-638) This section explains oxidation and reduction in terms of the loss or gain of electrons, and describes the characteristics of a redox reaction. It also explains how to identify oxidizing and reducing agents.

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Chapter 20 Oxidation-Reduction Reactions 523 Practice Problems In your notebook, solve the following problems. SECTION 20.1 THE MEANING OF OXIDATION AND REDUCTION Determine what is oxidized and what is reduced in each reaction. Identify the oxidizing agent and the reducing agent. 1.  $2\text{Sr} + \text{O}_2 \rightarrow 2\text{SrO}$  2.  $2\text{Li} + \text{S} \rightarrow 2\text{Li}_2\text{S}$  3.  $2\text{Cs} + \text{Br}_2 \rightarrow 2\text{CsBr}$  4.  $3\text{Mg} + \text{N}_2 \rightarrow \dots$

### SECTION 20.1 THE MEANING OF OXIDATION AND REDUCTION

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b. butanoic acid from 1-butanol 5. Classify each of the reactions in problem 4 as an oxidation-reduction, substitution, addition, or esterification reaction. SECTION 23.4 POLYMERIZATION 1. Draw the structures of propene (propylene) and tetrafluoroethene. Draw the basic repeating units when each of these compounds polymerizes. Describe

### SECTION 23.1 INTRODUCTION TO FUNCTIONAL GROUPS

636 Chapter 20 Redox Reactions Figure 20-1 The reaction of magnesium and oxygen involves a transfer of electrons from magnesium to oxygen. Therefore, this reaction is an oxidation-reduction reaction. Using the classifications given in Chapter 10, this redox reaction also is classified as a combustion reaction.  $\text{X} + \text{X} + \text{X} + \text{X} + \text{X} + \text{X} + 2\text{Mg} \rightarrow 2\text{MgO}$  Each ...

### Chapter 20: Redox Reactions - Neshaminy School District

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SECTION 20.2 OXIDATION NUMBERS 1. Give the oxidation number of each kind of atom or ion. a. Sn c. S<sub>2</sub> e. Se g. Sn<sub>4</sub> b. K d. Fe<sub>3</sub> f. Mg<sub>2</sub> h. Br 2. Calculate the oxidation number of chromium in each of the following formulas. a. Cr<sub>2</sub>O<sub>3</sub> b. H<sub>2</sub>Cr<sub>2</sub>O<sub>7</sub> c. CrSO<sub>4</sub> d. CrO<sub>4</sub><sup>2-</sup> 3. Use the changes in oxidation number to determine which elements are oxidized ...

**05 CTR ch20 7/12/04 8:17 AM Page 517 THE MEANING OF ...**

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**Chapter 2 Weathering and Soil Section 1 Summary Minerals ...**

that are introduced in this section. Each blank can be completed with a term, short phrase, or number. The oxidation number of an element in an uncombined state 1. is . 1 . The oxidation number of a monatomic ion is the same 2. . \_\_ ~ ~ \_\_\_\_ in magnitude and \_=2\_ as its ionic 3 . The sum of the . 3. oxidation numbers of the elements . in . a ...

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