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macroscopic thermodynamics of equilibrium (MTE).¹ In order to obtain a preliminary impression of the features in which this modernized Gibbs theory differs from that of CKC, it is convenient to compare the geometrical methods utilized by each. The application of geometrical methods to thermodynamics is based on a

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The equilibrium state of a system is completely specified by a certain number (this number is called the number of (thermodynamic) degrees of freedom) of independent variables. In equilibrium these specify the other variables via equations of state. Classic example: ideal gas, with a fixed number of particles N . Equation of state (EoS): $PV = Nk_B T$.

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Equilibrium thermodynamics. Here, typically a system, as cylinder of gas, initially in its own state of internal thermodynamic equilibrium, is set out of balance via heat input from a combustion reaction. Then, through a series of steps, as the system settles into its final equilibrium state, work is extracted.

Equilibrium thermodynamics - Wikipedia

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BIBLIOGRAPHY [1] C. J. Adkins, Equilibrium Thermodynamics, 3rd edition. Cambridge University Press, Cambridge, 1983. [2] Robert A. Alberty, "Use of Legendre Transforms in Chemical Thermodynamics (IUPAC

Thermodynamics and Chemistry

the study of the transformations of energy from one form into another. First Law: Heat and Work are both forms of Energy. in any process, Energy can be changed from one form to another (including heat and work), but it is never created or destroyed: Conservation of Energy.

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