

## ***Saturated Unsaturated Supersaturated Graph Answers***

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### **Saturated Unsaturated Supersaturated Graph Answers**

Unsaturated, Saturated, and Supersaturated Examples. A supersaturated solution is one that has more solute than it can hold at a certain temperature. Typically when the temperature of a solution is increased, more particles can be dissolved, thus increasing the amount of solute. A supersaturated solution goes through all...

### **Types of Solutions: Saturated, Supersaturated, or Unsaturated**

A graph can illustrate what solution is saturated and unsaturated. If the point is on the line, then the solution is saturated, while if it is below the line, the solution is unsaturated.

### **How does a graph show a saturated solution ... - answers.com**

Best Answer: How are you using the term "supersaturated"? If that is used properly, it is a true solution containing no solid but containing a concentration of solute greater than a saturated solution. If you are using it to mean a saturated solution in contact with excess solid, then two of the solutions are supersaturated.

### **Chemistry question: unsaturated, saturated, or ...**

A graph can illustrate what solution is saturated and unsaturated. If the point is on the line, then the solution is saturated, while if it is below the line, the solution is unsaturated.

### **When is a solution saturated - answers.com**

In a nutshell, unsaturated solutions have points below the line on the graph, saturated is on the line and supersaturated is above the line on the graph.

### **Solubility Curves - Saturated, Unsaturated, Supersaturated Solutions**

Saturated. Unsaturated. Miscible. Immiscible . supersaturated. The graph below (on the next page) is known as a solubility curve and can be used to determine if a particular solution is saturated at a given set of conditions and how much of the solute is dissolved under those conditions.

### **Solubility Curve Practice Problems Worksheet 1**

A supersaturated solution contains more solute at a given temperature than is needed to form a saturated solution.. Increased temperature usually increases the solubility of solids in liquids. For example, the solubility of glucose at 25 °C is 91 g/100 mL of water.

### **Saturated and Supersaturated Solutions - Chemistry | Socratic**

Chemistry saturated, unsaturated or supersaturated question.? So im stuck on a question At 90 degrees celsius, 10g of potassium chlorate dissolved in 100g of water is the solution saturated, unsaturated or supersaturated and how do you know so i know what the 3 terms means, i just dont know a way o figure it out with the above info.

### **Chemistry saturated, unsaturated or supersaturated ...**

On the line=saturated (full can not hold anymore solute) Below the line=unsaturated (can hold more solute) Above the line= supersaturated (holding more solute than it should, unstable condition) ... Answer--> The graph show the decrease in solubility as temperature increases: 1/04.

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