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Welcome back to Week 3 of your HSC Chemistry Syllabus Notes! In the prior weeks, we have explored what equilibrium reactions are, the factors that affects an equilibrium's position, how we can use Le Chatelier's Principle to predict which direction an equilibrium will shift and many other smaller things under those concepts.

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A solution of sucrose (molar mass = 342 g mol⁻¹) has been prepared by dissolving 68.5 g of sucrose in 100 g of water. The freezing point of the solution obtained will be (K_f for water = 1.86 K kg mol⁻¹) [CBSE AIPMT 2010]

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