

Stoichiometry Using Copper Lab 31 Answer Key

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Stoichiometry Using Copper Lab 31

In this lab copper was suppose to be recovered to initial values. As the data shows, this did not happen. The final mass of copper (49.5g Cu) ended up significantly more than the original value (1.962 Cu). The final moles of copper (.77 moles Cu) ended up being significantly more than the intital moles of copper (.03 moles Cu).

Stoichiometry Using Copper Lab - AP Chemistry Labs

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Stoichiometry Using Copper Lab 31 Answers Book - dskull

Stoichiometry Using Copper. Purpose: The purpose is to see how the amount of copper (and copper itself) is altered after a series of reactions. ... The main theory used in the lab was the conservation of mass. While copper was subjected to many different types of reactions the mass of copper was projected to stay the same. ... The types of ...

Stoichiometry Using Copper - Alexia's Ap Chemistry Lab ...

Lab 31 Stoichiometry Using Copper.pdf Free Download Here MakingCopper Stoichiometry I - Mt. San Antonio College ... Chemistry 30S/PGL Stoichiometry Lab ... copper ion product in the reaction between copper and silver nitrate ... 0.31 mol/L x 0.025 L ... Stoichiometry - Wikispaces

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Stoichiometry Using Copper Lab. Purpose. To demonstrate the conservation of mass involving chemical reactions using copper and its compounds. ... This should have lowered the final mass of copper because the copper produced was not produced using all the copper compound. Also, when the copper was weighed to determine its final mass, there was a ...

Stoichiometry Using Copper Lab - Weebly

General Chemistry I (FC, 09 - 10) Lab #4: Stoichiometry: The Reaction of Iron with Copper(II) Sulfate Revised 8/19/2009 1 Introduction In this experiment we will use stoichiometric principles to deduce the appropriate equation for the reaction between metallic iron and a solution of copper(II) sulfate. This reaction produces

General Chemistry I (FC, 09 - 10) Lab #4: Stoichiometry ...

there is significant pre-lab preparation that must be done by the instructor. First, I hand out the LAB - Stoichiometry (aluminum to copper). I ask students to skim through the lab and to name the different sections that they will be completing (Procedure, Observations, Data, Analysis, and Conclusion).

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