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Solution Stoichiometry Problems

Stoichiometry Problems Name _____ Chem Worksheet 12-2 Stoichiometry Strategy Amount moles molar mass (g/mol) 22.4 L/mol KNOWN UNKNOWN Mass grams V olume f gas at STP iters Particles atoms, molecules, formula units 6.02 $\times 10^3$ particles/mol Mass grams at STP liters Particles atoms, molecules, formula units Amount B moles molar mass (g/mol)

Stoichiometry Problems Name Chem Worksheet 12-2

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perform stoichiometric calculations. The strategy used for solving these solution stoichiometry problems is to set up the problem so that the units cancel. When the volume of a solution is multiplied by the molarity of a solution the resulting units are moles.

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The units for molarity are USEFUL E U ATIONS written as mol/L or M. This measurement is used to perform stoichiometric calculations. The strategy used for molarity = 3101 solute 1 L = 1000 mL solving these solution stoichiometry problems is to set up L solution the problem so that the units cancel.

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Chemistry: Stoichiometry - Problem Sheet 2 KEY 9) 2 24 2 2 23 2 2 2 4.63 $\times 10$ molecules l 1 mol l 6.02 $\times 10$ moleculesl 1 mol Cl 1mol 71 g Cl Cl \times 546 g Cl 10) 292 g Ag 1 mol Ag 108 g Ag 1 mol Cu 1 mol Ag 63.5 g Cu

Stoichiometry: Problem Sheet 2

Chemistry: Stoichiometry - Problem Sheet 1 Directions: Solve each of the following problems. Show your work, including proper units, to earn full credit. 1. Silver and nitric acid react according to the following balanced equation: $3 \text{Ag(s)} + 4 \text{HNO}_3 \text{(aq)} \rightarrow 3 \text{AgNO}_3 \text{(aq)} + 2 \text{H}_2 \text{O(l)} + \text{NO(g)}$ A.

Stoichiometry: Problem Sheet 1

Use your knowledge of Stoichiometry and the Ideal Gas Law to solve the following problems. The chemical equations given are all balanced. 1. What volume of O_2 is produced when 28.5 g of hydrogen peroxide (H_2O_2) decomposes to form water and oxygen at $150^\circ C$ and 2.0 atm? $2H_2O_2(aq) \rightarrow 2H_2O(l) + O_2(g)$ 2. This reaction uses 18.2 g ...

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problem. For example, when converting 18 inches to centimeters we place the inches on the bottom of the conversion factor so that they cancel with the inches in the 'given'. $254 \text{ in} = 254 \text{ in} \times \frac{2.54 \text{ cm}}{1 \text{ in}}$ - Write the over Insert conversion so that units cancel. Solve. Convert the following measurements using Set up problem using fractions. Show the units that cancel.

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