

## *Solubility Ksp Worksheet 2 Answers*

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### **Solubility Ksp Worksheet 2 Answers**

sp and Molar Solubility Problems Worksheet 1. Use the chemical AgCl to describe solubility, molar solubility and solubility product 2. Write balanced equations and solubility product expressions for the following compounds a. CuBr b. ZnC<sub>2</sub>O<sub>4</sub> c. Ag<sub>2</sub>CrO<sub>4</sub> d. Hg<sub>2</sub>Cl<sub>2</sub> e. AuCl<sub>3</sub> f. Mn<sub>3</sub>(PO<sub>4</sub>)<sub>3</sub> 3.

### **Ksp and Molar Solubility Problems Worksheet**

SCH4U1 Name: EA02 Date: Solubility Product, K<sub>sp</sub> - Worksheet 2 - ANSWERS 1. Calculate the solubility (in mol/L) for each of the following compounds: a. FeS K<sub>sp</sub> =  $3.7 \times 10^{-19}$  solubility =  $6.1 \times 10^{-10}$  mol/L b.

### **Solubility Product, K<sub>sp</sub> - Worksheet 2 - ANSWERS**

Solubility Product Worksheet - Answers 1) What is the concentration of a saturated silver (I) acetate solution? K<sub>sp</sub>(AgC<sub>2</sub>H<sub>3</sub>O<sub>2</sub>) =  $1.94 \times 10^{-3}$ . Since K<sub>sp</sub> = [Ag<sup>+</sup>][C<sub>2</sub>H<sub>3</sub>O<sub>2</sub><sup>-</sup>], and the concentration of silver ions is the same as the concentration of acetate ions, we can set up the following equation:  $1.94 \times 10^{-3} = x^2$  x = 0.0440 M

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### **SOLUBILITY KSP WORKSHEET 4 ANSWERS**

Solubility Unit Problems Worksheet: 1. Calculate the K<sub>sp</sub> for each of the salts whose solubility is listed below.: a) CaSO<sub>4</sub> =  $5.0 \times 10^{-3}$  mol/L : b) MgF<sub>2</sub> =  $2.7 \times 10^{-3}$  mol/L : c) AgC<sub>2</sub>H<sub>3</sub>O<sub>2</sub> = 1.02 g/100 mL : d) SrF<sub>2</sub> = 12.2 mg/100 mL : 2. Calculate : a) the solubility in moles/L of each of three salts and

### **Solubility & Ksp Problems Worksheet**

Solubility, Ksp Worksheet 1 1. How many milliliters of 0.20 M AlCl<sub>3</sub> ... is equal to the solubility of CaF<sub>2</sub> in pure water. (B) is less than the solubility of CaF<sub>2</sub> ... Solubility, Ksp Answer Key WORKSHEET 1 1) A 2) A 3) C 4) A 5) D 6) B 7) C 8) C 9) C 10) A 11) a) drives rxn to the left ...

### **Solubility, Ksp Worksheet 1 - Ms. Kissinger**

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### **Solubility product constant (Ksp) Worksheet.doc - Google Docs**

Chemistry 12 Unit 3 – Solubility Equilibrium WS3.1 Solubility and Ksp.docx 1 Worksheet 3.1 Solubility Product Calculations 1. Write the K<sub>sp</sub> expression for these compounds. a) FeS b) Bi<sub>2</sub>S<sub>3</sub> c) Ag<sub>2</sub>CO<sub>3</sub> d) Al(OH)<sub>3</sub> e) PbI<sub>2</sub> f) Cu(OH)<sub>2</sub> g) PbCrO<sub>4</sub> h) Sb<sub>2</sub>S<sub>3</sub> 2. Why are solids left out of the K<sub>sp</sub> expression? 3. Barium sulphate, BaSO<sub>4</sub>

### **Worksheet 3.1 Solubility Product Calculations**

Worksheet 7—More Solubility Problems Answer Key 1. A solution is made with NaI and NaCl such that it is 0.01 M in both I<sup>-</sup> and Cl<sup>-</sup>. To 1 L of this solution 0.01 moles Ag(NO<sub>3</sub>) are added (you can ignore any volume change). The

### **Worksheet 7—More Solubility Problems Answer Key**

Test your knowledge of solubility and how to use a solubility constant ( $K_{sp}$ ) in calculations with an interactive quiz and printable worksheet....

### **Quiz & Worksheet - Use a Solubility Constant ( $K_{sp}$ ) in ...**

Solubility Worksheet #2 Level 1 1 Name \_\_\_\_\_ Date \_\_\_\_\_ We have considered some conditions that allow solutes to dissolve in solvents, but we have not yet considered the solubility of a solute. Solubility is defined as the maximum mass of solute that can dissolve in a certain quantity of solvent at a

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