

# Thermodynamics & Statistical Physics

## Chapter 1. Basic Laws in Thermodynamics

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## 1 Chpt 1. Basic Laws in Thermodynamics

### • 1.1 Thermal equilibrium state

- 1.2 Law of thermal equilibrium and the temperature
- 1.3 Equation of state
- 1.4 Work
- 1.5 First law of thermodynamics
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- Thermal SP (additional): temperature ( $T$ ).

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- Thermometric scale  $T$  (K) (Kelvin).
- Ideal gas scaling:  $T = 273.16\text{K} \times \lim_{p_t \rightarrow 0} (p/p_t)$ .
- Daily use, Celsius scale  $(T - 273.15)^\circ\text{C}$ .

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- Using  $\left( \frac{\partial T}{\partial V} \right)_p \left( \frac{\partial p}{\partial T} \right)_V \left( \frac{\partial V}{\partial p} \right)_T = -1$ ,  
 $\therefore \alpha = \kappa_T \beta p$ .

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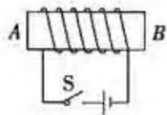
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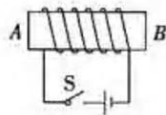
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- Summary:  $\delta W = \sum_i Y_i dy_i$ ,  
where  $Y_i$  is the generalized force, and  $y_i$  is the generalized displacement.

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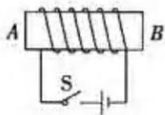
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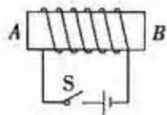


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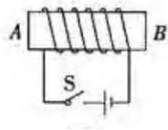
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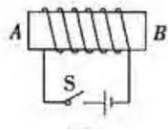
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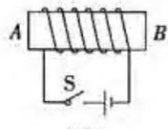


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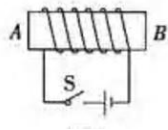
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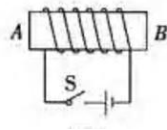
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where  $\frac{\mu_0 H^2}{2}$  is the magnetic energy density (work by increasing the magnetic field);  $M$  is the magnetization (work done by magnetize the material).

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- “Almost” obeys in all physical processes.

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 $C_p = \lim_{\Delta T \rightarrow 0} \left( \frac{\Delta U + p \Delta V}{\Delta T} \right)_p = \left( \frac{\partial U}{\partial T} \right)_p + p \left( \frac{\partial V}{\partial T} \right)_p$ .
- Enthalpy:  $H \equiv U + pV$ .  
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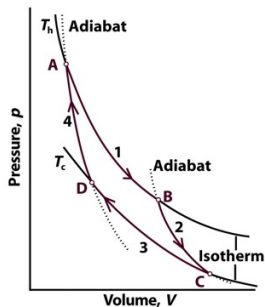


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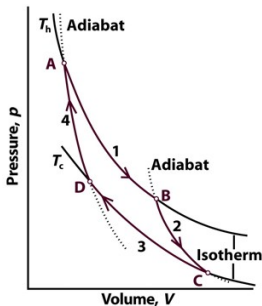


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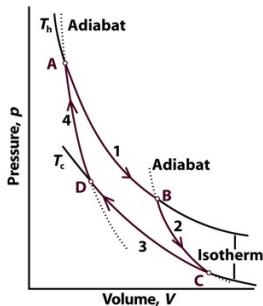


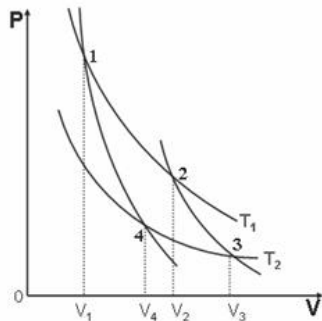
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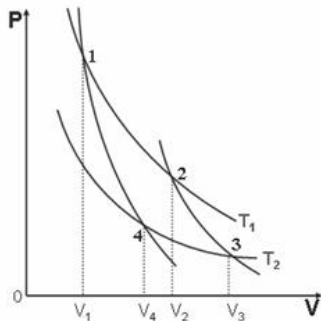
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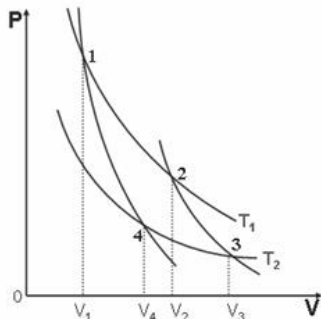
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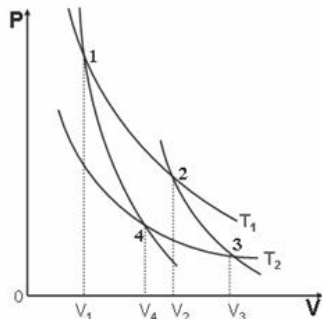
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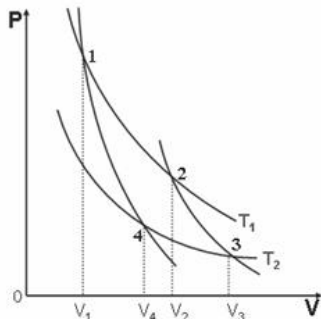
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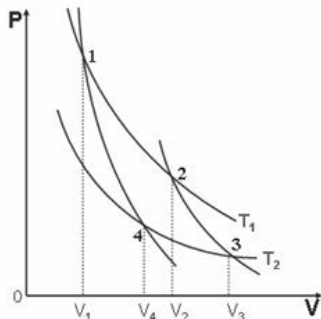
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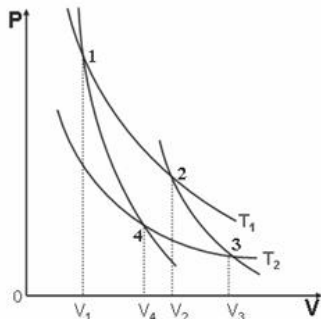
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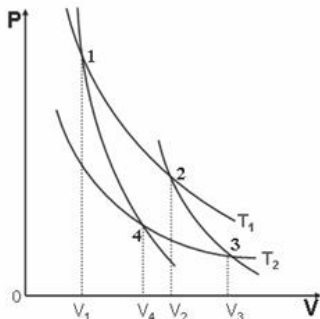
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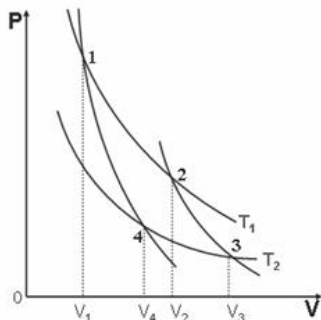
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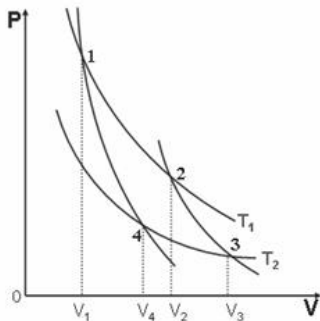
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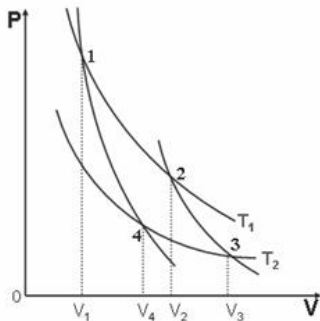
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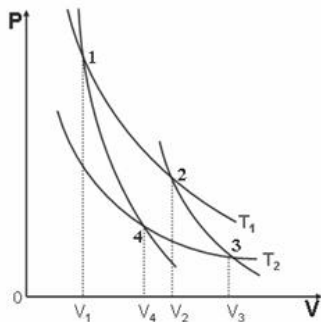


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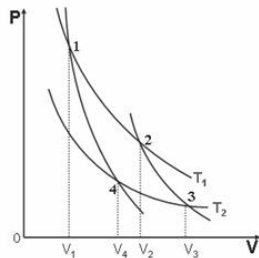
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- 1 Isothermal expansion: absorb heat  $Q_1 = RT_1 \ln \frac{V_2}{V_1}$ ,
  - 2 Adiabatic expansion:  $Q = 0$ ,
  - 3 Isothermal compression: release heat  $Q_2 = RT_2 \ln \frac{V_3}{V_4}$ ,
  - 4 Adiabatic compression:  $Q = 0$ .
- Total work:  $W = Q_1 - Q_2 = RT_1 \ln \frac{V_2}{V_1} - RT_2 \ln \frac{V_3}{V_4}$ .

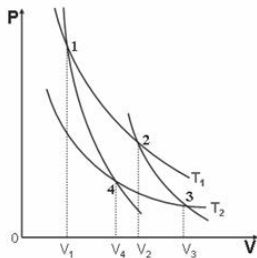
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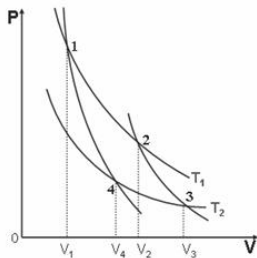
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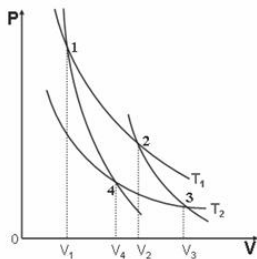


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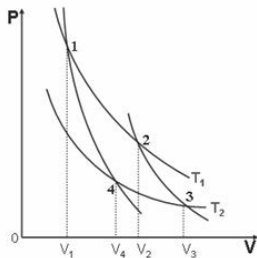
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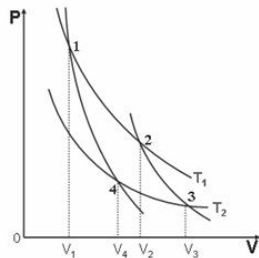
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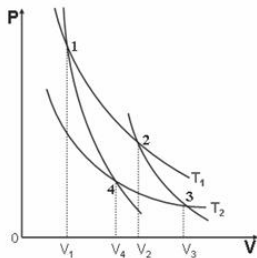
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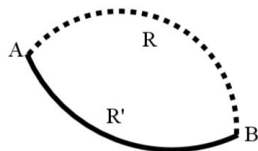
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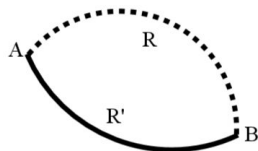
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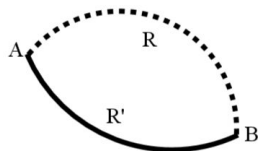


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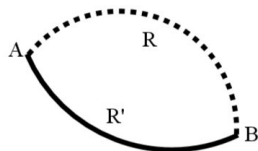
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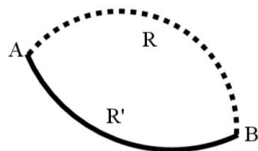
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- 1st law:  $dU = \delta Q + \delta W = \delta Q - pdV$   
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 (Basic equation of thermodynamics).



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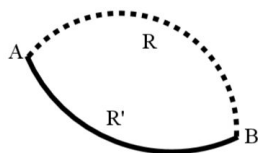
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$$S_B - S_A \equiv \int_A^B \frac{\delta Q_R}{T} \text{ or } dS \equiv \frac{\delta Q}{T}.$$

- 1st law:  $dU = \delta Q + \delta W = \delta Q - pdV$   
 $\Rightarrow dS = \frac{dU + pdV}{T}, \text{ or } dU = TdS - pdV$   
 (Basic equation of thermodynamics).
- $S = \sum S_i$  for extensible system.



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## 1 Chpt 1. Basic Laws in Thermodynamics

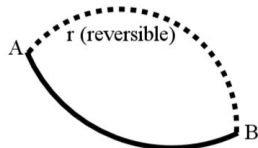
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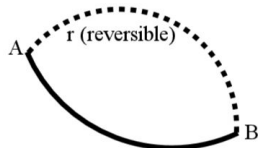
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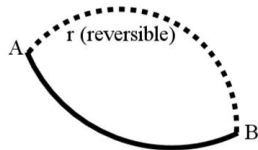


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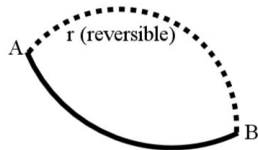
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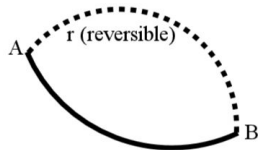
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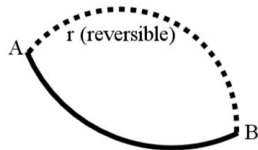
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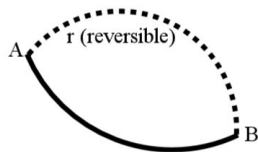
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- As  $S$  a state function, which doesn't depend on the process, the directly contacting A and B also derives  $\Delta S = Q(\frac{1}{T_2} - \frac{1}{T_1}) > 0$ .

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# Table of contents I

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## Table of contents II

- 1.12 Thermodynamic temperature scale
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