Let's look at adiabatic compression. We know that when you compress a container adiabatically, it heats up because you’re putting energy into the system through work. We can relate a molecule’s temperature to its speed with the equation (1/2)mv^2=(3/2)kT. This tell us that being compressed makes molecules speed up, but why?\n To figure this out, we need to look how molecules behave when they collide with objects. When an ideal gas molecule hits another object, it undergoes an elastic collision. When that object is a massive stationary wall, the gas molecule is reflected by the wall and its kinetic energy is unchanged. When the wall is moving, this is not true.\nSo to understand why compression heats a system, we need to describe molecules’ collisions with moving walls. Try to do that using the simulation.