

Simple proof of the second law of thermodynamics that uses the Kullback-Leibler divergence.

Brendan Lucas

Consider an isolated thermodynamic system a random discrete energy variable E . The first law of thermodynamics states that the internal energy $U = \langle E \rangle = \sum_i p_i E_i$ is constant in an isolated system. The system may start in a nonequilibrium state, but it approaches an equilibrium state $p_i = \frac{e^{-\beta E_i}}{\sum_j e^{-\beta E_j}}$. We consider the set of discrete energy levels $\mathcal{E} \ni E_i$ not to change in time, and the temperature of the system to be a constant $\beta = 1/k_B T$.

p_{now} is our model of the system's energy now, and p_{eq} is the unique marginal of \mathcal{E} 's equilibrium; p_{eq} is "the heat death of the universe." $\beta = 1/k_B T$ is inverse temperature, or coldness. p_{now} is compared against p_{eq} with the Kullback-Leibler divergence (KLD). p_{now} can be any marginal over \mathcal{E} , 1. Define KLD:

$$D[p_{\text{now}}||p_{\text{eq}}] = \sum_{i=1}^N p_{\text{now}}(E_i) \ln \left(\frac{p_{\text{now}}(E_i)}{p_{\text{eq}}(E_i)} \right) = \sum_{i=1}^N p_{\text{now}}(E_i) \ln p_{\text{now}}(E_i) - \sum_{i=1}^N p_{\text{now}}(E_i) \ln p_{\text{eq}}(E_i)$$

2. Expand and add a special zero

$$\begin{aligned} D[p_{\text{now}}||p_{\text{eq}}] &= \sum_{i=1}^N p_{\text{now}}(E_i) \ln p_{\text{now}}(E_i) - \sum_{i=1}^N p_{\text{now}}(E_i) \ln p_{\text{eq}}(E_i) \\ &\quad + \left(\sum_{i=1}^N p_{\text{eq}}(E_i) \ln p_{\text{eq}}(E_i) - \sum_{i=1}^N p_{\text{eq}}(E_i) \ln p_{\text{eq}}(E_i) \right) \end{aligned}$$

3. Convert to macroscopic variables with $\frac{1}{k_B} S = -\sum_{i=1}^N p_i \ln p_i$

$$\begin{aligned} D[p_{\text{now}}||p_{\text{eq}}] &= \frac{1}{k_B} \Delta S - \sum_{i=1}^N p_{\text{now}}(E_i) \ln p_{\text{eq}}(E_i) \\ &\quad + \sum_{i=1}^N p_{\text{eq}}(E_i) \ln p_{\text{eq}}(E_i) \\ &= \frac{1}{k_B} \Delta S + \sum_{i=1}^N [p_{\text{eq}}(E_i) - p_{\text{now}}(E_i)] \ln p_{\text{eq}}(E_i) \end{aligned}$$

4. Plug in the “heat death” partition function: $p_{\text{eq}}(E_i) = \frac{e^{-\beta E_i}}{\sum_{j=1}^N e^{-\beta E_j}} = \frac{e^{-\beta E_i}}{Z}$,

$$\begin{aligned} D[p_{\text{now}}||p_{\text{eq}}] &= \frac{1}{k_B} \Delta S + \sum_{i=1}^N [p_{\text{eq}}(E_i) - p_{\text{now}}(E_i)] \ln \left(\frac{e^{-\beta E_i}}{\sum_{j=1}^N e^{-\beta E_j}} \right) \\ &= \frac{1}{k_B} \Delta S + \sum_{i=1}^N [p_{\text{now}}(E_i) - p_{\text{eq}}(E_i)] \cdot [\beta E_i] + \sum_{i=1}^N [p_{\text{now}}(E_i) - p_{\text{eq}}(E_i)] \ln Z \end{aligned}$$

5. Notice that

$$\sum_{i=1}^N [p_{\text{now}}(E_i) - p_{\text{eq}}(E_i)] \ln Z = \ln Z \left[\sum_{i=1}^N p_{\text{now}}(E_i) - \sum_{i=1}^N p_{\text{eq}}(E_i) \right] = \ln Z - \ln Z = 0$$

So,

$$D[p_{\text{now}}||p_{\text{eq}}] = \frac{1}{k_B} \Delta S + \sum_{i=1}^N [p_{\text{now}}(E_i) - p_{\text{eq}}(E_i)] \cdot [\beta E_i]$$

6. The following step gives a difference in the expectation values of energy if and only if every E_i is the same in p_{eq} and p_{now} ; the energy levels did not change, but only their probabilities changed.

$$\begin{aligned} D[p_{\text{now}}||p_{\text{eq}}] &= \frac{1}{k_B} \Delta S + \sum_{i=1}^N p_{\text{now}}(E_i) \beta E_i - \sum_{i=1}^N p_{\text{eq}}(E_i) \beta E_i \\ &= \frac{1}{k_B} \Delta S - \beta \Delta \langle E \rangle \end{aligned}$$

7. Internal energy $U = \langle E \rangle$ is a constant,

$$k_B D[p_{\text{now}}||p_{\text{eq}}] = \Delta S,$$

8. KLD is non-negative, so

$$\Delta S \geq 0$$

QED, that is the second law of thermodynamics. **The entropy of an isolated thermodynamic system can only increase or stay the same, but never decrease!**