Name _____   Period _____

Molarity Problems 2

1. What is the molarity of 3.00 moles of KBr diluted to 1.500 liters?

2. If 25.0 grams of KOH are diluted in enough water to make 2.000 liter, what is the molarity of the solution?

3. How many grams of HF are needed to makes 1.00 liters of 3.00 M solution?

4. What volume of H_2SO_4 is needed to have 15.0 moles if the solution is 10.0 M?

5. Concentrated HCl is 12.1 Molar. If I want to prepare 2.00 L of 6.00 M solution how many mL's of concentrate do I need to dissolve?

6. Sulfuric acid comes as an 18.1 Molar concentrate. If I am making 100.0 mL of 1.00 molar how many millilitres of concentrate do I need to dilute?

7. It is possible to buy NaOH as a liquid that is 19.1 M. If you need 1.00 L of 1.50 M how many millilitres of concentrate do you need to dilute?
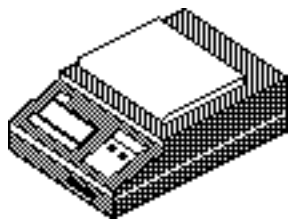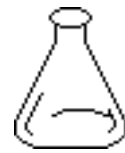
Balance   Beaker   Volumetric   Graduated   Erlenmeyer

Cylinder   Flask   Flask
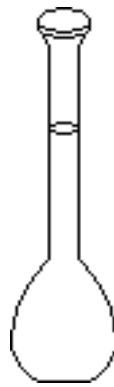
8. You have the equipment pictured above. Describe how you would prepare 1.00 L of 1.00 M HCl from the 12.1 M concentrated solution.

9. You have the equipment pictured above. Describe how you would prepare 1.00 L of 1.00 M NaOH from the solid.

10. Draw a picture of the steps you would follow to make 1.00 L of 1.0 M KOH. Make sure to include all mathematical detail.