Name _____

Period____

Balancing Equations Practice Number Two

Balance the following equations by placing the appropriate coefficient on the lines.

1)
$$\underline{\hspace{1cm}} Cl_2 + \underline{\hspace{1cm}} Fe \rightarrow \underline{\hspace{1cm}} FeCl_3$$

2)
$$\underline{\hspace{1cm}}$$
 Cl₂+ $\underline{\hspace{1cm}}$ Na \rightarrow $\underline{\hspace{1cm}}$ NaCl

3)
$$\underline{\hspace{1cm}} Cl_2 + \underline{\hspace{1cm}} S_8 \rightarrow \underline{\hspace{1cm}} S_2Cl_2$$

4)
$$__Au_2S_3 + __H_2 \rightarrow __Au + __H_2S$$

5)
$$__Al(C_6H_5)_3 + __HCl \rightarrow __AlCl_3 + __C_6H_6$$

6) ____ Ba(OH)₂ + ____ HNO₃
$$\rightarrow$$
 ____ Ba(NO₃)₂ + ____ H₂O

7)
$$\underline{\hspace{1cm}} C + \underline{\hspace{1cm}} SiO_2 \rightarrow \underline{\hspace{1cm}} CO + \underline{\hspace{1cm}} Si$$

8)
$$\underline{\qquad}$$
 CaCl₂ + $\underline{\qquad}$ AgNO₃ \rightarrow $\underline{\qquad}$ AgCl + $\underline{\qquad}$ Ca(NO₃)₂

9) ____ AgBr + ___ Na₂S₂O₃
$$\rightarrow$$
 ___ Na₃Ag(S₂O₃)₂ + ___ NaBr

10)
$$__$$
 C₆H₆ + $__$ O₂ \rightarrow $__$ H₂O + $__$ CO₂

11) ___
$$C_7H_6O_3 +$$
__ $C_4H_6O_3 \rightarrow$ __ $C_9H_8O_4 +$ __ $H_2O_3 \rightarrow$

12)
$$_$$
 CO + $_$ Fe₂O₃ \rightarrow $_$ Fe + $_$ CO₂

13)
$$_$$
 CO + $_$ NO \rightarrow $_$ CO₂ + $_$ N₂

14) ___
$$Cu(NO_3)_2 \rightarrow$$
 ___ $CuO +$ ___ $NO_2 +$ ___ O_2

15)
$$C_2H_6 + O_2 \rightarrow CO_2 + H_2O$$

16) ___ NaHCO₃
$$\rightarrow$$
 ___ Na₂CO₃ + ___ CO₂ + ___ H₂O

17) ___ NH₃ + ___ NO
$$\rightarrow$$
 ___ N₂ + ___ H₂O

18) ___
$$O_2$$
 + ___ $NH_3 \rightarrow$ ___ NO + ___ H_2O

19)
$$_$$
 $O_2 + _$ $PH_3 \rightarrow _$ $P_4O_{10} + _$ H_2O

20) ___ PbCl₄ + ___ SnCl₂
$$\rightarrow$$
 ___ SnCl₄ + ___ PbCl₂

21)
$$_$$
 BaCO₃ + $_$ HCl \rightarrow $_$ H₂O + $_$ CO₂ + $_$ BaCl₂

22)
$$__BrF_3 + __TiO_2 \rightarrow __TiF_4 + __Br_2 + __O_2$$

23) ___ CaCO₃ + ___ HNO₃
$$\rightarrow$$
 ___ Ca(NO₃)₂ + ___ CO₂ + ___ H₂O

24) ___ CaCO₃ + ___ O₂ + ___ SO₂
$$\rightarrow$$
 ___ CaSO₄ + ___ CO₂

25) ___
$$H_2SO_4 +$$
 ___ $NaCl \rightarrow$ ___ $HCl +$ ___ $NaHSO_4$

26)
$$\underline{\hspace{1cm}}$$
 $H_2SO_4+\underline{\hspace{1cm}}$ $Zn\rightarrow\underline{\hspace{1cm}}$ $ZnSO_4+\underline{\hspace{1cm}}$ H_2

27) ___
$$C_6H_8O_7$$
 + ___ $NaHCO_3$ \rightarrow ___ CO_2 + ___ H_2O + ___ $Na_3C_6H_5O_7$

28) ___
$$C_6H_{10}O_4 +$$
 ___ $NH_3 +$ ___ $H_2 \rightarrow$ ___ $C_6H_{16}N_2 +$ ___ H_2O

29) ___
$$C_6H_8ONC1 +$$
 ___ $NaOH \rightarrow$ ___ $C_6H_7ON +$ ___ $H_2O +$ ___ $NaC1$

30)
$$\underline{\hspace{1cm}} C_2H_3Cl + \underline{\hspace{1cm}} O_2 \rightarrow \underline{\hspace{1cm}} CO_2 + \underline{\hspace{1cm}} H_2O + \underline{\hspace{1cm}} HCl$$