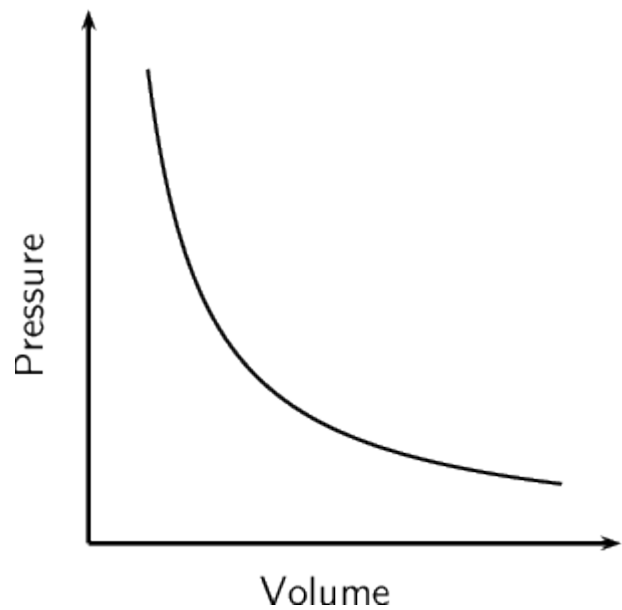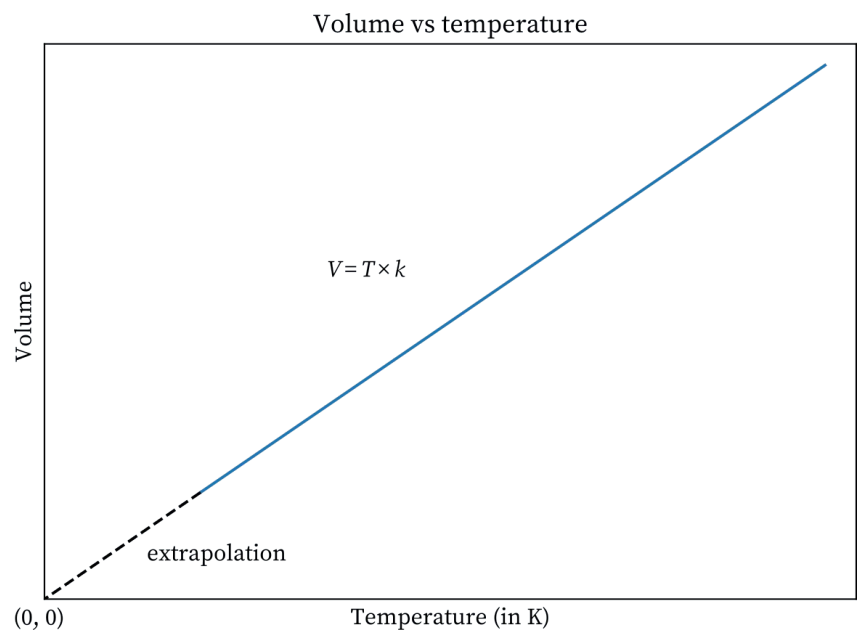# Boyles Law

- As Volume is Doubled pressure is halved
- Liters
- ATM
- Inversly Proportional
- $P_1 V_1 = P_2 V_2$

Charles Law

- As Volume is doubled Temperature is doubled
- Kelvin
- Directly Proportional
- $\dfrac{V_1}{T_1} = \dfrac{V_2}{T_2}$

Volume vs temperature

$V = T \times k$

extrapolation

(0, 0)

Temperature (in K)

ALL TEMPERATURES IN GAS LAWS MUST BE IN KELVIN

KELVIN IS 273.15