7. $pK_S = 4,75$. $x := [H^+]$

0,1 M Essigsäure $pK_S = 4,75$

$4,75 = -\log(K_S)$

$$K_S = \frac{[H^+][C_2OO^-]}{[C_2OOH]} = 10^{-4,75}$$

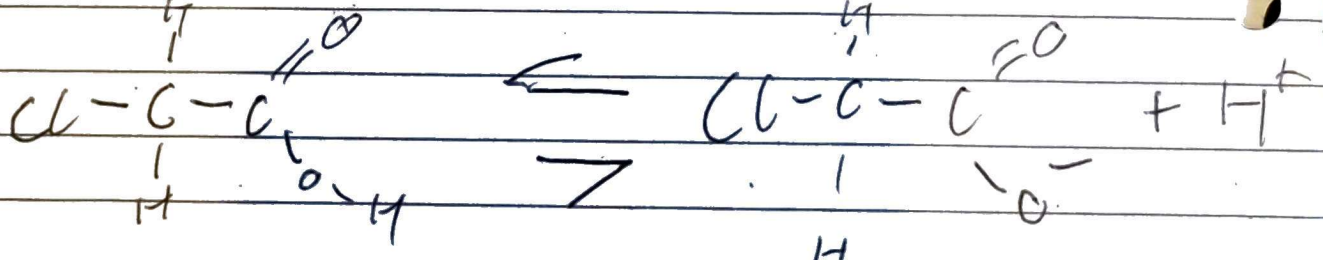
ausgehen: 0,1 M nach Lösung,

$$10x^2 \neq 10^{-4,75}$$

$$x^2 = 10^{-5,75} \quad . \quad x = 0,00133$$

$$-\log(x) = 2,875$$ ✓

8. Chloressigsäure $pK_S = 2,86$



$$10x^2 = 10^{-2,86}$$

$$x^2 = 10^{-3,56}$$

$$x = 0,0117 .$$

$-\log(x) = 1,93$

OK?

1,93 gegen 1,96.