

# Manufactured Solution for 1D Steady Euler Equations for Hypersonic Flows with Nitrogen Dissociation in Thermal Nonequilibrium\*

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## Contents

<b>1</b>	<b>Mathematical Model</b>	<b>1</b>
1.1	Chemical Kinetics . . . . .	3
1.2	Vibrational/Electronic Energy Production & Vibrational Relaxation . . . . .	4
<b>2</b>	<b>Manufactured Solutions</b>	<b>5</b>
2.1	1D Steady Euler Equations in Thermochemical Nonequilibrium . . . . .	6
2.2	Mass Conservation of Nitrogen Atom . . . . .	6
2.3	Mass Conservation of Nitrogen Molecule . . . . .	7
2.4	Momentum Conservation . . . . .	7
2.5	Total Energy Conservation . . . . .	8
2.6	Vibrational Energy Conservation . . . . .	8
<b>3</b>	<b>Comments</b>	<b>9</b>
<b>A</b>	<b>List of model / manufactured solution parameters</b>	<b>12</b>

## Abstract

The Method of Manufactured Solutions is a valuable approach for code verification, providing means to verify how accurately the numerical method solves the equations of interest. The method generates a related set of governing equations that has known analytical (manufactured) solution. Then, the modified set of equations may be discretized and solved numerically, and the numerical solution may be compared to the manufactured analytical solution chosen *a priori*.

A choice of analytical solutions for the flow variables of the 1D steady Euler equations for chemically reacting hypersonic flows in thermal nonequilibrium and their respective source terms are presented in this document.

## 1 Mathematical Model

Using the two-temperature model, the conservation of mass, momentum, and energy for a inviscid compressible fluid composed of a mixture of gases N and N<sub>2</sub> in thermochemical nonequilibrium may be written

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\*Work based on Kirk, Amar, Stogner, Schultz, and Oliver (2009).

as:

$$\nabla \cdot (\rho_N \mathbf{u}) = \dot{\omega}_N, \quad (1)$$

$$\nabla \cdot (\rho_{N_2} \mathbf{u}) = \dot{\omega}_{N_2}, \quad (2)$$

$$\nabla \cdot (\rho \mathbf{u} \mathbf{u}) = -\nabla p, \quad (3)$$

$$\nabla \cdot (\rho \mathbf{u} H) = 0, \quad (4)$$

$$\nabla \cdot (\rho e_V \mathbf{u}) = \dot{\omega}_V, \quad (5)$$

where  $\rho_s$  is the density of species  $s$  (N or  $N_2$ ),  $\rho = \sum_s \rho_s$  is the mixture density and  $\mathbf{u}$  is the mixture velocity and  $e_V$  is the vibrational/electronic energy. The total enthalpy,  $H$ , may be expressed in terms of the total energy, density, and pressure:

$$H = E + \frac{p}{\rho},$$

where the total energy,  $E$ , is composed of internal and kinetic components:

$$E = e^{\text{int}} + \frac{u^2}{2}.$$

The total internal energy,  $e^{\text{int}}$ , has contributions from each of the distinct energy *modes*:

$$\begin{aligned} e^{\text{int}} &= e^{\text{trans}} + e^{\text{rot}} + e^{\text{vib}} + e^{\text{elec}} + h^0 \\ &= \sum_{s=1}^{ns} c_s e_s^{\text{trans}} + \sum_{s=\text{mol}} c_s e_s^{\text{rot}} + \sum_{s=\text{mol}} c_s e_s^{\text{vib}} + \sum_{s=1}^{ns} c_s e_s^{\text{elec}} + \sum_{s=1}^{ns} c_s h_s^0, \end{aligned} \quad (6)$$

where  $c_s = (\rho_s/\rho)$  is the mass fraction of species  $s$ .

The first four terms on the right of Equation (6) represent the energy due to molecular/atomic translation, molecular rotation, molecular vibration, and electronic excitation. The final term is the heat of formation of the mixture and accounts for the energy stored in chemical bonds (Ait-Ali-Yahia, 1996; Kirk et al., 2009). In the two-temperature model, the translational and rotational modes may modeled with a single translational/rotational temperature  $T \equiv T_t = T_r$ , while the vibrational and electronic energy are governed by a separate temperature  $T_V \equiv T_v = T_e$ .

Under the approximation that the translational and rotational states of the may be assumed fully populated, translational/rotational energy for each species may be expressed as:

$$e_s^{\text{trans}} + e_s^{\text{rot}} = e_s^{\text{tr}} = C_{v,s}^{\text{tr}} T, \quad (7)$$

where the translational/rotational specific heat,  $C_{v,s}^{\text{tr}}$  is given by

$$C_{v,s}^{\text{tr}} = \begin{cases} \frac{5}{2} R_s & \text{for molecules,} \\ \frac{3}{2} R_s & \text{for atoms,} \end{cases}$$

where  $R_s$  is the species gas constant, and  $R_s = R/M_s$  where  $R$  is the universal gas constant and  $M_s$  is the species molar mass. The combined term  $e_s^{\text{tr}}$  in Equation (7) represents the energy due to random thermal translational/rotational motion of a given species.

In contrast to the translational/rotational states, the vibrational energy states are typically not fully populated. One approach for modeling the molecular vibrational energy is through analogy to a harmonic oscillator. In this approach the energy potential between molecular nuclei is modeled as a quadratic function of separation distance. Under this assumption, the vibrational energy for each molecular species can be modeled as:

$$e_s^{\text{vib}} = \begin{cases} \frac{R_s \theta_s^{\text{vib}}}{\exp(\theta_s^{\text{vib}}/T_v) - 1} & \text{for molecules,} \\ 0 & \text{for atoms,} \end{cases} \quad (8)$$

where  $\theta_s^{\text{vib}}$  is the species characteristic temperature of vibration and  $T_v$  is the mixture vibrational temperature.

The energy contained in the excited electronic states for a given species,  $e_s^{\text{elec}}$ , can be obtained from the assumption that they are in a Boltzmann distribution governed by the electronic excitation temperature  $T_e$  (Candler, 1988) as:

$$e_s^{\text{elec}} = R_s \frac{\sum_{i=1}^{\infty} \theta_{s,i}^{\text{elec}} g_{s,i} \exp(-\theta_{s,i}^{\text{elec}}/T_e)}{\sum_{i=1}^{\infty} g_{s,i} \exp(-\theta_{s,i}^{\text{elec}}/T_e)}. \quad (9)$$

where  $g_{s,i}$  is the degeneracy of the electronic level  $i$  of species  $s$ , and  $\theta_{s,i}^{\text{elec}}$  is its characteristic temperature. In practice, a simple cut off criteria is applied to select the number of electronic energy levels in the calculation of the series. The number of electronic energy levels for N is 3 and number of electronic energy levels for  $N_2$  is 15 ?.

Recalling that in the case of thermal equilibrium  $T_e = T_v = T_V$ , Equations (8) and (9) are simplified to:

$$e_s^{\text{vib}} = \begin{cases} \frac{R_s \theta_s^{\text{vib}}}{\exp(\theta_s^{\text{vib}}/T_V) - 1} & \text{for molecules,} \\ 0 & \text{for atoms,} \end{cases} \quad e_s^{\text{elec}} = R_s \frac{\sum_{i=1}^{nel_s} \theta_{s,i}^{\text{elec}} g_{s,i} \exp(-\theta_{s,i}^{\text{elec}}/T_V)}{\sum_{i=1}^{nel_s} g_{s,i} \exp(-\theta_{s,i}^{\text{elec}}/T_V)}.$$

where  $nel_s$  is the number of electronic energy levels of species  $s$ .

Therefore,

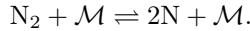
$$\rho e_V(T_V) = \sum_{s=\text{mol}} \rho_s e_s^{\text{vib}}(T_V) + \sum_{s=1}^{ns} \rho_s e_s^{\text{elec}}(T_V). \quad (10)$$

Regardless of the thermal state of the mixture, once the translational/rotational temperature  $T$  is determined the thermodynamic pressure of the mixture is readily obtained from Dalton's law of partial pressures:

$$p = \sum_{s=1}^{ns} p_s = \sum_{s=1}^{ns} \rho_s R_s T. \quad (11)$$

## 1.1 Chemical Kinetics

The rate of production/destruction of the individual species  $s$ ,  $\dot{\omega}_s$ , is required to close the species continuity equations. For the dissociating Nitrogen flow ( $N_2 \rightleftharpoons 2N$ ) case, let us consider the chemical reactions which occur among the five principal components of dissociating air –  $N_2$ ,  $O_2$ ,  $NO$ ,  $N$ ,  $O$  – but neglecting three species in order to perform this 2-species problem. For this mixture, the single chemical reaction that occur is:



This reaction can occur in either the forward or backward direction, as denoted by the bidirectional arrow. The reaction is presented such that they are endothermic in the forward direction, and  $\mathcal{M}$  denotes a generic collision partner, which may be any of the species present in the flow, and is unaltered by the reaction (Kirk et al., 2009). The rate of each reaction is therefore a sum of the forward and backward rates,  $k_f$  and  $k_b$ :

$$\begin{aligned} \mathcal{R}_r &= \mathcal{R}_{br} - \mathcal{R}_{fr} \\ &= k_{br} \prod_{s=1}^{ns} \left( \frac{\rho_s}{M_s} \right)^{\beta_{sr}} - k_{fr} \prod_{s=1}^{ns} \left( \frac{\rho_s}{M_s} \right)^{\alpha_{sr}} \end{aligned}$$

where  $\alpha_{sr}$  and  $\beta_{sr}$  are the stoichiometric coefficients for reactants and products of species  $s$ . For the Nitrogen dissociation,

$$\begin{aligned} \mathcal{R}_1 &= \sum_{m \in \mathcal{M}} \left( k_{b_1 m} \frac{\rho_N}{M_N} \frac{\rho_N}{M_N} \frac{\rho_m}{M_m} - k_{f_1 m} \frac{\rho_{N_2}}{M_{N_2}} \frac{\rho_m}{M_m} \right) \\ &= k_{b_1 N} \frac{\rho_N^3}{M_N^3} - k_{f_1 N} \frac{\rho_{N_2} \rho_N}{2M_N^2} + k_{b_1 N_2} \frac{\rho_N^2 \rho_{N_2}}{2M_N^3} - k_{f_1 N_2} \frac{\rho_{N_2}^2}{4M_N^2}, \end{aligned} \quad (12)$$

recalling that  $M_{N_2} = 2M_N$ .

The species source terms  $\dot{\omega}_s = M_s \sum_{r=1}^{nr} (\alpha_{sr} - \beta_{sr}) (\mathcal{R}_{br} - \mathcal{R}_{fr})$  where  $nr$  is the number of reactions can now be expressed in terms of the individual reaction rates as follows:

$$\begin{aligned}\dot{\omega}_{N_2} &= M_{N_2} (\mathcal{R}_1) = 2M_N \mathcal{R}_1, \\ \dot{\omega}_N &= M_N (-2\mathcal{R}_1) = -2M_N \mathcal{R}_1.\end{aligned}$$

Note that these source terms sum identically to zero, as required by conservation of mass (Kessler and Awruch, 2004).

The forward rate coefficients  $k_{fr}$  can then be expressed in a modified Arrhenius form as

$$k_{fr}(\bar{T}) = C_{fr} \bar{T}^{\eta_r} \exp(-E_{ar}/R\bar{T}) \quad (13)$$

where  $C_{fr}$  is the reaction rate constant,  $\eta_r$  is the so-called pre-exponential factor,  $E_{ar}$  is the activation energy. These three constants are determined from curve fits to experimental data (e.g. see Ait-Ali-Yahia (1996)). The effective temperature,  $\bar{T}$ , is a function of the translational/rotational and vibrational temperatures, and in this work it is taken as:

$$\bar{T} = T^q T_V^{1-q}, \quad 0 \leq q \leq 1. \quad (14)$$

The corresponding backward rate coefficient  $k_{br}$  can be found using the principle of detailed balance, which states

$$K_{eq} = \frac{k_{fr}(\bar{T})}{k_{br}(\bar{T})} \quad (15)$$

where  $K_{eq}$  is the equilibrium constant and may be obtained either by curve fits or through Gibbs' free energy techniques<sup>1</sup>. In this work,  $K_{eq} = K(T)$ .

Therefore:

$$\begin{aligned}k_{f_1N} &= C_{f_1N} \bar{T}^{\eta_{f_1N}} \exp\left(\frac{-E_{aN}}{R\bar{T}}\right) \quad \text{and} \quad k_{b_1N} = \frac{k_{f_1N}}{K(T)}, \\ k_{f_1N_2} &= C_{f_1N_2} \bar{T}^{\eta_{f_1N_2}} \exp\left(\frac{-E_{aN_2}}{R\bar{T}}\right) \quad \text{and} \quad k_{b_1N_2} = \frac{k_{f_1N_2}}{K(T)},\end{aligned} \quad (16)$$

are the forward and backward rates for Nitrogen atom and Nitrogen molecule, respectively.

## 1.2 Vibrational/Electronic Energy Production & Vibrational Relaxation

For the case of thermal nonequilibrium it remains to define the vibrational/electronic energy source term,  $\dot{\omega}_V$ , which appears in Equation (5). This term represents the production/destruction of vibrational/electronic energy in the gas, and is due to both the creation of molecules with some vibrational/electronic energy and the transfer of energy between the various modes in the gas:

$$\dot{\omega}_v = \dot{Q}_v + \dot{Q}_{\text{transfer}}. \quad (17)$$

When molecular species are created in the gas at rate  $\dot{\omega}_s$ , they contribute vibrational/electronic energy at the rate:

$$\dot{Q}_{vs} = \dot{\omega}_s (e_s^{\text{vib}} + e_s^{\text{elec}}),$$

so the net vibrational energy production rate is then simply:

$$\dot{Q}_v = \sum_{s=1}^{ns} \dot{\omega}_s (e_s^{\text{vib}} + e_s^{\text{elec}}). \quad (18)$$

There is also energy transfer among the various energy modes in the gas. Strictly speaking, one such energy transfer is vibration-vibration coupling between the various molecules in the gas. However, implicit in the use of a single vibrational energy equation is the assumption that the molecular vibrational energies equilibrate

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<sup>1</sup>Examples of curve fits for calculating the equilibrium constant may be found in Park (1990) whereas Gibbs' free energy techniques are employed in McBride et al. (2002).

very rapidly and thus are adequately characterized with a single vibrational temperature  $T_V$ . There is also energy transfer between translational and vibrational modes as well as rotational and vibrational modes (Kirk et al., 2009). These latter two exchanges are grouped together and represented as a single vibrational energy transfer rate  $\dot{Q}_s^{\text{tr-vib}}$ . In this work we adopt the Landau-Teller model for the vibrational energy transfer for a given species:

$$\dot{Q}_s^{\text{tr-vib}} = \rho_s \frac{\hat{e}_s^{\text{vib}} - e_s^{\text{vib}}}{\tau_s^{\text{vib}}}, \quad (19)$$

where  $\hat{e}_s^{\text{vib}}$  is the species equilibrium vibrational energy (Equation (8) evaluated at temperature  $T$ ) and the vibrational relaxation time  $\tau_s^{\text{vib}}$  is given by:

$$\tau_s^{\text{vib}} = \frac{\sum_{r=1}^{ns} \chi_r}{\sum_{r=1}^{ns} \chi_r / \tau_{sr}^{\text{vib}}} \quad (20)$$

where  $\chi_r$  is given by

$$\chi_r = c_r \frac{M}{M_r}, \quad \text{with} \quad M = \left( \sum_{s=1}^{ns} \frac{c_s}{M_s} \right)^{-1} \quad (21)$$

and

$$\tau_{sr}^{\text{vib}} = \frac{1}{P} \exp \left[ A_{sr} \left( T^{-1/3} - 0.015 \mu_{sr}^{1/4} \right) - 18.42 \right], \quad (22)$$

$$A_{sr} = 1.16 \times 10^{-3} \mu_{sr}^{1/2} (\theta_s^{\text{vib}})^{4/3}, \quad (23)$$

$$\mu_{sr} = \frac{M_s M_r}{M_s + M_r}. \quad (24)$$

Combining (19) and (18) yields the desired net vibrational energy source term

$$\dot{\omega}_V = \sum_{s=1}^{ns} \dot{Q}_s^{\text{tr-vib}} + \sum_{s=1}^{ns} \dot{\omega}_s (e_s^{\text{vib}} + e_s^{\text{elec}}). \quad (25)$$

## 2 Manufactured Solutions

Roy et al. (2002) propose the general form of the primitive solution variables to be a function of sines and cosines:

$$\phi(x, y) = \phi_0 + \phi_x f_s \left( \frac{a_{\phi x} \pi x}{L} \right), \quad (26)$$

where  $\phi = \rho_N, \rho_{N_2}, u, T$  or  $T_V$ , and  $f_s(\cdot)$  functions denote either sine or cosine function. Note that in this case,  $\phi_x$  is constant and the subscript does not denote differentiation. Different choices of the constants used in the manufactured solutions for the 2D supersonic and subsonic cases of Euler and Navier-Stokes may be found in Roy et al. (2002).

Therefore, the manufactured analytical solution for each one of the variables in Euler equations are:

$$\begin{aligned} \rho_N(x) &= \rho_{N0} + \rho_{Nx} \sin \left( \frac{a_{\rho Nx} \pi x}{L} \right), \\ \rho_{N_2}(x) &= \rho_{N_20} + \rho_{N_2x} \cos \left( \frac{a_{\rho N_2x} \pi x}{L} \right), \\ u(x) &= u_0 + u_x \sin \left( \frac{a_{ux} \pi x}{L} \right), \\ T(x) &= T_0 + T_x \cos \left( \frac{a_{Tx} \pi x}{L} \right), \\ T_V(x) &= T_{V0} + T_{Vx} \cos \left( \frac{a_{T_Vx} \pi x}{L} \right). \end{aligned} \quad (27)$$

Recalling that  $\rho = \sum_s \rho_s$ , the manufactured analytical solution for the density of the mixture is:

$$\begin{aligned} \rho(x) &= \rho_N + \rho_{N_2} \\ &= \rho_{N0} + \rho_{N_20} + \rho_{Nx} \sin \left( \frac{a_{\rho Nx} \pi x}{L} \right) + \rho_{N_2x} \cos \left( \frac{a_{\rho N_2x} \pi x}{L} \right). \end{aligned} \quad (28)$$

## 2.1 1D Steady Euler Equations in Thermochemical Nonequilibrium

The MMS applied to 1D steady Euler equations for a chemically reacting mixture of N and N<sub>2</sub> in thermal nonequilibrium consists in modifying Equations (1)–(5) by adding a source term to the right-hand side of each equation:

$$\begin{aligned}
\frac{\partial(\rho_N u)}{\partial x} - \dot{\omega}_N &= Q_{\rho_N}, \\
\frac{\partial(\rho_{N_2} u)}{\partial x} - \dot{\omega}_{N_2} &= Q_{\rho_{N_2}}, \\
\frac{\partial(\rho u^2)}{\partial x} + \frac{\partial(p)}{\partial x} &= Q_u, \\
\frac{\partial(\rho u E)}{\partial x} + \frac{\partial(pu)}{\partial x} &= Q_E, \\
\frac{\partial(\rho e_V)}{\partial t} + \frac{\partial(\rho e_V u)}{\partial x} - \dot{\omega}_V &= Q_{e_V},
\end{aligned} \tag{29}$$

so the modified set of Equations (29) conveniently has the analytical solutions given in Equations (27) and (28).

Source terms  $Q_{\rho_N}$ ,  $Q_{\rho_{N_2}}$ ,  $Q_u$ ,  $Q_E$  and  $Q_{e_V}$  are obtained by symbolic manipulations of equations above using Maple and are presented in the following sections. The following auxiliary variables have been included in order to improve readability and computational efficiency:

$$\begin{aligned}
\text{Rho}_N &= \rho_{N0} + \rho_{Nx} \sin\left(\frac{a_{\rho Nx} \pi x}{L}\right), \\
\text{Rho}_{N_2} &= \rho_{N_2 0} + \rho_{N_2 x} \cos\left(\frac{a_{\rho N_2 x} \pi x}{L}\right), \\
\text{Rho} &= \text{Rho}_N + \text{Rho}_{N_2}, \\
U &= u_0 + u_x \sin\left(\frac{a_{ux} \pi x}{L}\right), \\
T &= T_0 + T_x \cos\left(\frac{a_{Tx} \pi x}{L}\right), \\
T_V &= T_{V0} + T_{Vx} \cos\left(\frac{a_{TVx} \pi x}{L}\right),
\end{aligned} \tag{30}$$

which simply are the manufactured solutions, and the derivatives:

$$\begin{aligned}
\frac{\partial e_N^{\text{elec}}}{\partial x} &= \frac{(e_N^{\text{elec}})^2 M_N a_{TVx} \pi T_{Vx}}{R T_V^2 L} \sin\left(\frac{a_{TVx} \pi x}{L}\right) - \frac{R a_{TVx} \pi T_{Vx}}{(M_N T_V^2 L)} \sin\left(\frac{a_{TVx} \pi x}{L}\right) \text{AUX1}, \\
\frac{\partial e_{N_2}^{\text{elec}}}{\partial x} &= \frac{2(e_{N_2}^{\text{elec}})^2 M_N a_{TVx} \pi T_{Vx}}{R T_V^2 L} \sin\left(\frac{a_{TVx} \pi x}{L}\right) - \frac{R a_{TVx} \pi T_{Vx}}{2(M_N T_V^2 L)} \sin\left(\frac{a_{TVx} \pi x}{L}\right) \text{AUX2}, \\
\text{AUX1} &= \frac{\sum_{i=1}^{n_{elN}} (\theta_{N,i}^{\text{elec}})^2 g_{N,i} \exp(-\theta_{N,i}^{\text{elec}} / T_V)}{\sum_{i=1}^{n_{elN}} g_{N,i} \exp(-\theta_{N,i}^{\text{elec}} / T_V)}, \\
\text{AUX2} &= \frac{\sum_{i=1}^{n_{elN_2}} (\theta_{N_2,i}^{\text{elec}})^2 g_{N_2,i} \exp(-\theta_{N_2,i}^{\text{elec}} / T_V)}{\sum_{i=1}^{n_{elN_2}} g_{N_2,i} \exp(-\theta_{N_2,i}^{\text{elec}} / T_V)}.
\end{aligned} \tag{31}$$

## 2.2 Mass Conservation of Nitrogen Atom

The mass conservation equation for Nitrogen atom (N), written as an operator, is:

$$\mathcal{L} = \frac{\partial \rho_N u}{\partial x} - \dot{\omega}_N.$$

Analytically differentiating Equation (27) for  $\rho_N$  and  $u$  using operator  $\mathcal{L}$  defined above together with suitable substitution for  $\dot{\omega}_N$  gives the source term  $Q_{\rho_N}$ :

$$Q_{\rho_N} = \frac{a_{\rho N x} \pi \rho_{N x} \mathbf{U}}{L} \cos\left(\frac{a_{\rho N x} \pi x}{L}\right) + \frac{a_{u x} \pi u_x \mathbf{Rho}_N}{L} \cos\left(\frac{a_{u x} \pi x}{L}\right) - \dot{\omega}_N \quad (32)$$

with  $\mathbf{Rho}$  and  $\mathbf{U}$  defined in Equation (30). The source term  $\dot{\omega}_N$  is given by:

$$\dot{\omega}_N = -2M_N \mathcal{R}_1$$

where  $M_N$  is the molar mass of Nitrogen and  $\mathcal{R}_1$  is the reaction given according to Eq. (12):

$$\mathcal{R}_1 = \frac{k_{f_1 N}}{K(\mathbf{T})} \frac{\mathbf{Rho}_N^3}{M_N^3} - k_{f_1 N} \frac{\mathbf{Rho}_{N_2} \mathbf{Rho}_N}{2M_N^2} + \frac{k_{f_1 N_2}}{K(\mathbf{T})} \frac{\mathbf{Rho}_N^2 \mathbf{Rho}_{N_2}}{2M_N^3} - k_{f_1 N_2} \frac{\mathbf{Rho}_{N_2}^2}{4M_N^2}, \quad (33)$$

$K(\mathbf{T})$  is a to-be-defined function for the equilibrium constant (13), and  $k_{f_1 N}$  and  $k_{f_1 N_2}$  are defined by:

$$\begin{aligned} k_{f_1 N} &= C_{f_1 N} \bar{T}^{\eta_{f_1 N}} \exp\left(\frac{-E_{aN}}{R\bar{T}}\right), \\ k_{f_1 N_2} &= C_{f_1 N_2} \bar{T}^{\eta_{f_1 N_2}} \exp\left(\frac{-E_{aN_2}}{R\bar{T}}\right), \\ \bar{T} &= \mathbf{T}^q \mathbf{T}_V^{1-q}. \end{aligned} \quad (34)$$

### 2.3 Mass Conservation of Nitrogen Molecule

The mass conservation equation for Nitrogen molecule ( $N_2$ ), written as an operator, is:

$$\mathcal{L} = \frac{\partial \rho_{N_2} u}{\partial x} - \dot{\omega}_{N_2}.$$

Analytically differentiating Equation (27) for  $\rho_{N_2}$  and  $u$  using operator  $\mathcal{L}$  defined above together with suitable substitution for  $\dot{\omega}_{N_2}$  gives the source term  $Q_{\rho_{N_2}}$ :

$$Q_{\rho_{N_2}} = -\frac{a_{\rho N_2 x} \pi \rho_{N_2 x} \mathbf{U}}{L} \sin\left(\frac{a_{\rho N_2 x} \pi x}{L}\right) + \frac{a_{u x} \pi u_x \mathbf{Rho}_{N_2}}{L} \cos\left(\frac{a_{u x} \pi x}{L}\right) - \dot{\omega}_{N_2} \quad (35)$$

where  $\mathbf{Rho}$  and  $\mathbf{U}$  defined in Equation (30) and:

$$\dot{\omega}_{N_2} = M_{N_2} \mathcal{R}_1 = 2M_N \mathcal{R}_1$$

with  $\mathcal{R}_1$  given in Equation (33).

### 2.4 Momentum Conservation

For the generation of the analytical source term  $Q_u$ ,  $x$ -momentum equation (3) is written as an operator  $\mathcal{L}$ :

$$\mathcal{L} = \frac{\partial(\rho u^2)}{\partial x} + \frac{\partial p}{\partial x},$$

which, when operated in to Equations (27) and (28), provides source term  $Q_u$ :

$$\begin{aligned} Q_u &= \frac{a_{\rho N x} \pi \rho_{N x} \mathbf{U}^2}{L} \cos\left(\frac{a_{\rho N x} \pi x}{L}\right) + \\ &\quad - \frac{a_{\rho N_2 x} \pi \rho_{N_2 x} \mathbf{U}^2}{L} \sin\left(\frac{a_{\rho N_2 x} \pi x}{L}\right) + \\ &\quad + \frac{2a_{u x} \pi u_x \mathbf{Rho} \mathbf{U}}{L} \cos\left(\frac{a_{u x} \pi x}{L}\right) + \\ &\quad - \frac{a_{T x} \pi T_x R(2\mathbf{Rho}_N + \mathbf{Rho}_{N_2}) \sin\left(\frac{a_{T x} \pi x}{L}\right)}{2LM_N} + \\ &\quad - \frac{\pi R \mathbf{T}}{2LM_N} \left[ -2a_{\rho N x} \rho_{N x} \cos\left(\frac{a_{\rho N x} \pi x}{L}\right) + a_{\rho N_2 x} \rho_{N_2 x} \sin\left(\frac{a_{\rho N_2 x} \pi x}{L}\right) \right] \end{aligned} \quad (36)$$

with  $\mathbf{Rho}$ ,  $\mathbf{Rho}_N$ ,  $\mathbf{Rho}_{N_2}$ ,  $\mathbf{T}$  and  $\mathbf{U}$  defined in Equation (30).

## 2.5 Total Energy Conservation

The total energy equation is written as an operator:

$$\mathcal{L} = \frac{\partial \rho u H}{\partial x},$$

where  $H = E + \frac{p}{\rho}$  and  $E = e^{\text{int}} + \frac{u^2}{2}$ , with  $p$  and  $e^{\text{int}}$  as defined in Equations (11) and (6), respectively.

Source term  $\dot{Q}_E$  is obtained by operating  $\mathcal{L}$  on Equations (27) and (28):

$$\begin{aligned} \dot{Q}_E = & \frac{3a_{ux}\pi u_x \text{Rho } U^2}{2L} \cos\left(\frac{a_{ux}\pi x}{L}\right) + \\ & + \frac{(h_N^0 + e_N^{\text{elec}})a_{\rho N x} \pi \rho_{N x} U}{L} \cos\left(\frac{a_{\rho N x} \pi x}{L}\right) + \\ & - \frac{(h_{N_2}^0 + e_{N_2}^{\text{elec}} + e_{N_2}^{\text{vib}})a_{\rho N_2 x} \pi \rho_{N_2 x} U}{L} \sin\left(\frac{a_{\rho N_2 x} \pi x}{L}\right) + \\ & + \frac{(10 \text{Rho}_N + 7 \text{Rho}_{N_2})a_{ux} \pi u_x R T}{4LM_N} \cos\left(\frac{a_{ux}\pi x}{L}\right) + \\ & + \frac{(h_N^0 + e_N^{\text{elec}})a_{ux} \pi u_x \text{Rho}_N}{L} \cos\left(\frac{a_{ux}\pi x}{L}\right) + \\ & + \frac{(h_{N_2}^0 + e_{N_2}^{\text{elec}} + e_{N_2}^{\text{vib}})a_{ux} \pi u_x \text{Rho}_{N_2}}{L} \cos\left(\frac{a_{ux}\pi x}{L}\right) + \\ & - \frac{(10 \text{Rho}_N + 7 \text{Rho}_{N_2})a_{Tx} \pi T_x R U}{4LM_N} \sin\left(\frac{a_{Tx}\pi x}{L}\right) + \\ & - \frac{e_{N_2}^{\text{vib}} \theta_{N_2}^{\text{vib}} a_{TV x} \pi T_{V x} \text{Rho}_{N_2} U}{L T_V^2} \sin\left(\frac{a_{TV x} \pi x}{L}\right) + \\ & - \frac{2(e_{N_2}^{\text{vib}})^2 a_{TV x} \pi T_{V x} M_N \text{Rho}_{N_2} U}{LR T_V^2} \sin\left(\frac{a_{TV x} \pi x}{L}\right) + \\ & - \frac{\pi U^3}{2L} \left[ -a_{\rho N x} \rho_{N x} \cos\left(\frac{a_{\rho N x} \pi x}{L}\right) + a_{\rho N_2 x} \rho_{N_2 x} \sin\left(\frac{a_{\rho N_2 x} \pi x}{L}\right) \right] + \\ & - \frac{\pi R U T}{4LM_N} \left[ -10a_{\rho N x} \rho_{N x} \cos\left(\frac{a_{\rho N x} \pi x}{L}\right) + 7a_{\rho N_2 x} \rho_{N_2 x} \sin\left(\frac{a_{\rho N_2 x} \pi x}{L}\right) \right] + \\ & + \frac{\partial e_N^{\text{elec}}}{\partial x} \text{Rho}_N U + \frac{\partial e_{N_2}^{\text{elec}}}{\partial x} \text{Rho}_{N_2} U \end{aligned} \quad (37)$$

where  $\text{Rho}$ ,  $\text{Rho}_N$ ,  $\text{Rho}_{N_2}$ ,  $T$  and  $U$  are given in Equation (30), the derivatives  $\frac{\partial e_N^{\text{elec}}}{\partial x}$  and  $\frac{\partial e_{N_2}^{\text{elec}}}{\partial x}$  are given in Equation (31), and:

$$\begin{aligned} e_N^{\text{elec}} &= \frac{R}{M_N} \frac{\sum_{i=1}^{nel_N} \theta_{N,i}^{\text{elec}} g_{N,i} \exp(-\theta_{N,i}^{\text{elec}} / T_V)}{\sum_{i=1}^{nel_N} g_{N,i} \exp(-\theta_{N,i}^{\text{elec}} / T_V)}, \\ e_{N_2}^{\text{elec}} &= \frac{R}{2M_N} \frac{\sum_{i=1}^{nel_{N_2}} \theta_{N_2,i}^{\text{elec}} g_{N_2,i} \exp(-\theta_{N_2,i}^{\text{elec}} / T_V)}{\sum_{i=1}^{nel_{N_2}} g_{N_2,i} \exp(-\theta_{N_2,i}^{\text{elec}} / T_V)} \\ e_{N_2}^{\text{vib}} &= \frac{R}{2M_N} \frac{\theta_{N_2}^{\text{vib}}}{\exp(\theta_{N_2}^{\text{vib}} / T_V) - 1}. \end{aligned} \quad (38)$$

## 2.6 Vibrational Energy Conservation

The vibrational energy equation is written as an operator:

$$\mathcal{L} = \frac{\partial e_V \rho u}{\partial x},$$



where  $e_V = e^{\text{vib}} + e^{\text{elec}}$ . For the N-N<sub>2</sub> mixture,  $e_V$  is given by Equation (10):

$$e_V = \frac{\rho_{N_2}}{\rho} e_{N_2}^{\text{vib}} + \frac{\rho_N}{\rho} e_N^{\text{elec}} + \frac{\rho_{N_2}}{\rho} e_{N_2}^{\text{elec}}.$$

Therefore, source term  $Q_{e_V}$  is obtained by operating  $\mathcal{L}$  on Equations (27) and (28):

$$\begin{aligned} Q_{e_V} = & \frac{e_N^{\text{elec}} a_{\rho N x} \pi \rho_{N x} \mathbf{U}}{L} \cos\left(\frac{a_{\rho N x} \pi x}{L}\right) + \\ & - \frac{(e_{N_2}^{\text{elec}} + e_{N_2}^{\text{vib}}) a_{\rho N_2 x} \pi \rho_{N_2 x} \mathbf{U}}{L} \sin\left(\frac{a_{\rho N_2 x} \pi x}{L}\right) + \\ & + \frac{e_N^{\text{elec}} a_{u x} \pi u_x \mathbf{Rho}_N}{L} \cos\left(\frac{a_{u x} \pi x}{L}\right) + \\ & + \frac{(e_{N_2}^{\text{elec}} + e_{N_2}^{\text{vib}}) a_{u x} \pi u_x \mathbf{Rho}_{N_2}}{L} \cos\left(\frac{a_{u x} \pi x}{L}\right) + \\ & - \frac{e_{N_2}^{\text{vib}} \theta_{N_2}^{\text{vib}} a_{T_V x} \pi T_{V x} \mathbf{Rho}_{N_2} \mathbf{U}}{L T_V^2} \sin\left(\frac{a_{T_V x} \pi x}{L}\right) + \\ & - \frac{2(e_{N_2}^{\text{vib}})^2 a_{T_V x} \pi T_{V x} M_N \mathbf{Rho}_{N_2} \mathbf{U}}{L R T_V^2} \sin\left(\frac{a_{T_V x} \pi x}{L}\right) + \\ & + \frac{\partial e_N^{\text{elec}}}{\partial x} \mathbf{Rho}_N \mathbf{U} + \frac{\partial e_{N_2}^{\text{elec}}}{\partial x} \mathbf{Rho}_{N_2} \mathbf{U} + \\ & - \dot{\omega}_V \end{aligned} \quad (39)$$

where  $\mathbf{Rho}$ ,  $\mathbf{Rho}_N$ ,  $\mathbf{Rho}_{N_2}$ ,  $\mathbf{T}$  and  $\mathbf{U}$  are given in Equation (30),  $e_N^{\text{elec}}$ ,  $e_{N_2}^{\text{elec}}$  and  $e_{N_2}^{\text{vib}}$  are given in Equation (38). Source term  $\dot{\omega}_V$  is given according to Section 1.2 as:

$$\dot{\omega}_V = \dot{\omega}_N e_N^{\text{elec}} + \dot{\omega}_{N_2} (e_{N_2}^{\text{vib}} + e_{N_2}^{\text{elec}}) + \mathbf{Rho}_{N_2} (e_{N_2,eq}^{\text{vib}} - e_{N_2}^{\text{vib}}) \left[ \frac{2 \mathbf{Rho}_N}{(2 \mathbf{Rho}_N + \mathbf{Rho}_{N_2}) \tau_{N_2 N}^{\text{vib}}} + \frac{\mathbf{Rho}_{N_2}}{(2 \mathbf{Rho}_N + \mathbf{Rho}_{N_2}) \tau_{N_2 N_2}^{\text{vib}}} \right],$$

with

$$\begin{aligned} e_{N_2,eq}^{\text{vib}} &= \frac{R}{2 M_N \exp(\theta_{N_2}^{\text{vib}} / \mathbf{T}) - 1}, \\ \tau_{N_2 N_2}^{\text{vib}} &= \frac{\exp\left(\frac{29}{25000} M_N^{1/2} (\theta_{N_2}^{\text{vib}})^{4/3} \left[\mathbf{T}^{-1/3} - \frac{3}{2000} M_N^{1/4}\right] - 18.42\right)}{P}, \\ \tau_{N_2 N}^{\text{vib}} &= \frac{\exp\left(\frac{29}{75000} (6 M_N)^{1/2} (\theta_{N_2}^{\text{vib}})^{4/3} \left[\mathbf{T}^{-1/3} - \frac{1}{200} (54 M_N)^{1/4}\right] - 18.42\right)}{P}. \end{aligned} \quad (40)$$

and  $P$  is given by the Dalton's law (11):

$$P = \frac{\mathbf{Rho}_N \mathbf{T} R}{M_N} + \frac{\mathbf{Rho}_{N_2} \mathbf{T} R}{2 M_N}.$$

### 3 Comments

Source terms  $Q_{\rho_N}$ ,  $Q_{\rho_{N_2}}$ ,  $Q_u$ ,  $Q_E$  and  $Q_{e_V}$  have been generated by replacing the analytical Expressions (27) and (28) into respective Equations (1)–(5), followed by the usage of Maple commands for collecting, sorting and factorizing the terms. Files containing C codes for the source terms have also been generated. They are: `Euler_1d_steady_chemistry_2T_rho_N.code.C`, `Euler_1d_steady_chemistry_2T_rho_N2.code.C`, `Euler_1d_steady_chemistry_2T_u.code.C`, `Euler_1d_steady_chemistry_2T_E.code.C` and `Euler_1d_steady_chemistry_2T_eV.code.C`.

An example of the automatically generated C file from the source term for the mass conservation equation of N<sub>2</sub> is:

```

#include <math.h>

double SourceQ_rho_N2 (double x)
{
    double Q_rho_N2;
    double RHO_N;
    double RHO_N2;
    double U;
    double T;
    double TV;
    double kf1_N;
    double kf1_N2;
    double K;
    double R1;
    double w_dot_N2;
    double T_bar;
    RHO_N = rho_N_0 + rho_N_x * sin(a_rho_N_x * PI * x / L);
    RHO_N2 = rho_N2_0 + rho_N2_x * cos(a_rho_N2_x * PI * x / L);
    U = u_0 + u_x * sin(a_ux * PI * x / L);
    T = T_0 + T_x * cos(a_Tx * PI * x / L);
    TV = Tv_0 + Tv_x * cos(a_Tvx * PI * x / L);
    T_bar = pow(T, q) * pow(TV, 0.1e1 - q);
    K = calculate_equilibrium_constant_K(T);
    kf1_N2 = Cf1_N2 * pow(T_bar, etaf1_N2) * exp(-Ea_N2 / R / T_bar);
    kf1_N = Cf1_N * pow(T_bar, etaf1_N) * exp(-Ea_N / R / T_bar);
    R1 = kf1_N * pow(RHO_N, 0.3e1) / K * pow(M_N, -0.3e1) - kf1_N * RHO_N2 * RHO_N * pow(M_N, -0.2e1) / 0.2e1
        + kf1_N2 * RHO_N * RHO_N * RHO_N2 / K * pow(M_N, -0.3e1) / 0.2e1
        - kf1_N2 * RHO_N2 * RHO_N2 * pow(M_N, -0.2e1) / 0.4e1;
    w_dot_N2 = 0.2e1 * M_N * R1;
    Q_rho_N2 = -a_rho_N2_x * PI * rho_N2_x * U * sin(a_rho_N2_x * PI * x / L) / L
        + a_ux * PI * u_x * RHO_N2 * cos(a_ux * PI * x / L) / L - w_dot_N2;
    return(Q_rho_N2);
}

```

and the source term for the vibrational energy  $e_V$  of the mixture is:

```

#include <math.h>

double SourceQ_eV (double x, int energy_level_N, int energy_level_N2, double *g_N, double *g_N2,
    double *theta_e_N, double *theta_e_N2)
{
    double Q_eV;
    double e_elec_N;
    double e_elec_N_num;
    double e_elec_N_den;
    double e_elec_N2;
    double e_elec_N2_num;
    double e_elec_N2_den;
    double e_vib_N2;
    double e_vib_eq_N2;
    double RHO;
    double RHO_N;
    double RHO_N2;
    double U;
    double T;
    double TV;
    double P;
    double tau_vib_N2_N;
    double tau_vib_N2_N2;
    double De_elecN2_Dx;
    double De_elecN_Dx;
    double AUX1;
    double AUX2;
    double w_dot_N;
    double w_dot_N2;
    double kf1_N;
    double kf1_N2;
    double R1;

```

```

double K;
double T_bar;
int i;
double w_dot_V;
RHO_N = rho_N_0 + rho_N_x * sin(a_rho_N_x * PI * x / L);
RHO_N2 = rho_N2_0 + rho_N2_x * cos(a_rho_N2_x * PI * x / L);
RHO = RHO_N + RHO_N2;
U = u_0 + u_x * sin(a_ux * PI * x / L);
T = T_0 + T_x * cos(a_Tx * PI * x / L);
TV = Tv_0 + Tv_x * cos(a_Tvx * PI * x / L);
P = RHO_N * R * T / M_N + RHO_N2 * R * T / M_N / 0.2e1;
T_bar = pow(T, q) * pow(TV, 0.1e1 - q);
K = calculate_equilibrium_constant_K(T);
AUX1 = 0.0e0;
AUX2 = 0.0e0;
e_elec_N_num = 0.0e0;
e_elec_N2_num = 0.0e0;
e_elec_N_den = 0.0e0;
e_elec_N2_den = 0.0e0;
for (i = 1; i <= energy_level_N; i++)
{
    e_elec_N_den = e_elec_N_den + g_N[i - 1] * exp(-theta_e_N[i - 1] / TV);
    e_elec_N_num = e_elec_N_num + theta_e_N[i - 1] * g_N[i - 1] * exp(-theta_e_N[i - 1] / TV);
    AUX1 = AUX1 + pow(theta_e_N[i - 1], 0.2e1) * g_N[i - 1] * exp(-theta_e_N[i - 1] / TV);
}
for (i = 1; i <= energy_level_N2; i++)
{
    e_elec_N2_den = e_elec_N2_den + g_N2[i - 1] * exp(-theta_e_N2[i - 1] / TV);
    e_elec_N2_num = e_elec_N2_num + theta_e_N2[i - 1] * g_N2[i - 1] * exp(-theta_e_N2[i - 1] / TV);
    AUX2 = AUX2 + pow(theta_e_N2[i - 1], 0.2e1) * g_N2[i - 1] * exp(-theta_e_N2[i - 1] / TV);
}
e_elec_N = R * e_elec_N_num / M_N / e_elec_N_den;
e_elec_N2 = R * e_elec_N2_num / M_N / e_elec_N2_den / 0.2e1;
e_vib_N2 = R * theta_v_N2 / M_N / (exp(theta_v_N2 / TV) - 0.1e1) / 0.2e1;
e_vib_eq_N2 = R * theta_v_N2 / M_N / (exp(theta_v_N2 / T) - 0.1e1) / 0.2e1;
De_elecN_Dx = e_elec_N * e_elec_N * M_N * a_Tvx * PI * Tv_x * sin(a_Tvx * PI * x / L) / R
    * pow(TV, -0.2e1) / L - R * AUX1 * a_Tvx * PI * Tv_x * sin(a_Tvx * PI * x / L) / M_N / e_elec_N_den
    * pow(TV, -0.2e1) / L;
De_elecN2_Dx = 0.2e1 * e_elec_N2 * e_elec_N2 * M_N * a_Tvx * PI * Tv_x * sin(a_Tvx * PI * x / L) / R
    * pow(TV, -0.2e1) / L - R * AUX2 * a_Tvx * PI * Tv_x * sin(a_Tvx * PI * x / L) / M_N / e_elec_N2_den
    * pow(TV, -0.2e1) / L / 0.2e1;
tau_vib_N2_N = exp(0.29e2 / 0.75000e5 * sqrt(0.6e1) * sqrt(M_N) * pow(theta_v_N2, 0.4e1 / 0.3e1)
    * (pow(T, -0.1e1 / 0.3e1) - pow(0.54e2, 0.1e1 / 0.4e1) * pow(M_N, 0.1e1 / 0.4e1) / 0.200e3)
    - 0.921e3 / 0.50e2) / P;
tau_vib_N2_N2 = exp(0.29e2 / 0.25000e5 * sqrt(M_N) * pow(theta_v_N2, 0.4e1 / 0.3e1)
    * (pow(T, -0.1e1 / 0.3e1) - 0.3e1 / 0.200e3 * pow(M_N, 0.1e1 / 0.4e1))
    - 0.921e3 / 0.50e2) / P;
kf1_N2 = Cf1_N2 * pow(T_bar, etaf1_N2) * exp(-Ea_N2 / R / T_bar);
kf1_N = Cf1_N * pow(T_bar, etaf1_N) * exp(-Ea_N / R / T_bar);
R1 = -RHO_N * RHO_N2 * kf1_N * pow(M_N, -0.2e1) / 0.2e1
    - RHO_N2 * RHO_N2 * kf1_N2 * pow(M_N, -0.2e1) / 0.4e1
    + RHO_N * RHO_N * RHO_N2 * kf1_N2 / K * pow(M_N, -0.3e1) / 0.2e1
    + pow(RHO_N, 0.3e1) * kf1_N / K * pow(M_N, -0.3e1);
w_dot_N = -0.2e1 * M_N * R1;
w_dot_N2 = -w_dot_N;
w_dot_V = e_elec_N * w_dot_N + (e_vib_N2 + e_elec_N2) * w_dot_N2
    + (e_vib_eq_N2 - e_vib_N2) * RHO_N2 * (0.2e1 * RHO_N / (0.2e1 * RHO_N + RHO_N2) / tau_vib_N2_N
    + RHO_N2 / (0.2e1 * RHO_N + RHO_N2) / tau_vib_N2_N2);
Q_eV = e_elec_N * a_rho_N_x * PI * rho_N_x * U * cos(a_rho_N_x * PI * x / L) / L
    - (e_vib_N2 + e_elec_N2) * a_rho_N2_x * PI * rho_N2_x * U * sin(a_rho_N2_x * PI * x / L) / L
    + e_elec_N * a_ux * PI * u_x * RHO_N * cos(a_ux * PI * x / L) / L
    + (e_vib_N2 + e_elec_N2) * a_ux * PI * u_x * RHO_N2 * cos(a_ux * PI * x / L) / L
    - e_vib_N2 * a_Tvx * PI * Tv_x * theta_v_N2 * RHO_N2 * U * sin(a_Tvx * PI * x / L) / L * pow(TV, -0.2e1)
    - 0.2e1 * e_vib_N2 * e_vib_N2 * M_N * a_Tvx * PI * Tv_x * RHO_N2 * U * sin(a_Tvx * PI * x / L) / R
    * pow(TV, -0.2e1) / L
    + De_elecN_Dx * RHO_N * U + De_elecN2_Dx * RHO_N2 * U - w_dot_V;
return(Q_eV);
}

```

Finally, the gradients of the analytical solutions (27) and (28) have also been computed and their respective C codes are presented in `Euler_1d_chemistry_manuf_solutions_grad_and_code.C`. Therefore,

$$\begin{aligned}
\nabla \rho_N &= \frac{a_{\rho N x} \pi \rho_{N x}}{L} \cos\left(\frac{a_{\rho N x} \pi x}{L}\right), & \nabla \rho_{N_2} &= \frac{a_{\rho N_2 x} \pi \rho_{N_2 x}}{L} \sin\left(\frac{a_{\rho N_2 x} \pi x}{L}\right), \\
\nabla T &= -\frac{a_{T x} \pi T_x}{L} \sin\left(\frac{a_{T x} \pi x}{L}\right), & \nabla u &= \frac{a_{u x} \pi u_x}{L} \cos\left(\frac{a_{u x} \pi x}{L}\right), \\
\nabla T_V &= -\frac{a_{T_V x} \pi T_{V x}}{L} \sin\left(\frac{a_{T_V x} \pi x}{L}\right), & \nabla \rho &= \nabla(\rho_N) + \nabla(\rho_{N_2}),
\end{aligned} \tag{41}$$

are written in C language as:

```

grad_rho_an_N[0] = rho_N_x * cos(a_rho_N_x * pi * x / L) * a_rho_N_x * pi / L;
grad_rho_an_N2[0] = -rho_N2_x * sin(a_rho_N2_x * pi * x / L) * a_rho_N2_x * pi / L;
grad_rho_an[0] = rho_N_x * cos(a_rho_N_x * pi * x / L) * a_rho_N_x * pi / L
    - rho_N2_x * sin(a_rho_N2_x * pi * x / L) * a_rho_N2_x * pi / L;
grad_u_an[0] = u_x * cos(a_u_x * pi * x / L) * a_u_x * pi / L;
grad_T_an[0] = -T_x * sin(a_T_x * pi * x / L) * a_T_x * pi / L;
grad_Tv_an[0] = -Tv_x * sin(a_Tv_x * pi * x / L) * a_Tv_x * pi / L;

```

## References

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## A List of model / manufactured solution parameters

There are a variety of parameters present in the thermochemical nonequilibrium flow of a inviscid fluid, due to both fluid properties and the constants arising from the chosen manufactured solutions.

**Fluid properties:**  $R$ ,  $M_N$ ,  $h_N^0$ ,  $h_{N_2}^0$ ,  $\theta_{N_2}^{\text{vib}}$ ,  $\theta_N^{\text{elec}}$ ,  $\theta_{N_2}^{\text{elec}}$ ,  $g_N$ ,  $g_{N_2}$ ,  $nel_N$ ,  $nel_{N_2}$ ,  $C_{f_1 N}$ ,  $\eta_{f_1 N}$ ,  $E_{aN}$ ,  $C_{f_1 N_2}$ ,  $\eta_{f_1 N_2}$ ,  $E_{aN_2}$  and  $q$ , where  $\theta_{N_2}^{\text{elec}}$ ,  $\theta_N^{\text{elec}}$ ,  $g_N$ ,  $g_{N_2}$  are lists of values and their size is either  $n_{elN}$  or  $n_{elN_2}$ , depending on the species  $N$  or  $N_2$ .

In the C codes, these values are represented, respectively, by:

```

R, M_N, h0_N, h0_N2, theta_v_N2, theta_e_N, theta_e_N2, g_N, g_N2, energy_level_N,
energy_level_N2, Cf1_N, etaf1_N, Ea_N, Cf1_N2, etaf1_N2, Ea_N2, q.

```

**Manufactured solutions:**  $\rho_{N0}, \rho_{Nx}, a_{\rho Nx}, \rho_{N20}, \rho_{N2x}, a_{\rho N2x}, u_0, u_x, a_{ux}, T_0, T_x, a_{Tx}, T_{V0}, T_{Vx}, a_{TVx}, L$ .

In the C codes, these values are represented, respectively, by:

`rho_N_0, rho_N_x, a_rho_N_x, rho_N2_0, rho_N2_x, a_rho_N2_x, u_0, u_x, a_ux, T_0, T_x, a_Tx, Tv_0, Tv_x, a_Tvx, L`.

**Examples of data values:** The data for each of the species are given in Table 1:  $M_s$  is the species molar mass,  $R_s$  is the species gas constant,  $h_s^0$  is the activation energy, and  $\theta_s^{\text{vib}}$  is the species characteristic temperature of vibration for the 5-species model (Kessler and Awruch, 2004).

Table 1: Parameters for Nitrogen atom and molecule.

	N <sub>2</sub>	N
$M_s$ [kg/kg-mol]	28.02	14.01
$R_s$ [J/ kg K]	296.7	593.6
$h_s^0$ [K]	0	$33.59 \times 10^6$
$\theta_s^{\text{vib}}$ [K]	3393	-

Moreover, the characteristic electronic temperature  $\theta_s^{\text{elec}}$ , degeneracy of the electronic mode  $g_s$  and total number of electronic levels  $nel_s$  of Nitrogen atom and molecule are presented in Table 2 (Kirk et al., 2009).

Table 2: Electronic data of N and N<sub>2</sub>

Species $s$	$nel_s$	$\theta_s^{\text{elec}} [K]$	$g_s$
N	3	0	4
		$2.76647 \times 10^4$	10
		$4.14931 \times 10^4$	6
N <sub>2</sub>	15	0	1
		$7.22316 \times 10^4$	3
		$8.57786 \times 10^4$	6
		$8.60503 \times 10^4$	6
		$9.53512 \times 10^4$	3
		$9.80564 \times 10^4$	1
		$9.96827 \times 10^4$	2
		$1.04898 \times 10^5$	2
		$1.11649 \times 10^5$	5
		$1.22584 \times 10^5$	1
		$1.24886 \times 10^5$	6
		$1.28248 \times 10^5$	6
		$1.33806 \times 10^5$	10
		$1.40430 \times 10^5$	6
		$1.50496 \times 10^5$	6