Manufactured Solution for 1D Steady Euler Equations for Hypersonic Flows with Nitrogen Dissociation in Thermal Equilibrium using Maple*

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Abstract

The Method of Manufactured Solutions is a valuable approach for code verification, providing means to verify how accurately the numerical method solves the equations of interest. The method generates a related set of governing equations that has known analytical (manufactured) solution. Then, the modified set of equations may be discretized and solved numerically, and the numerical solution may be compared to the manufactured analytical solution chosen *a priori*.

A choice of analytical solutions for the flow variables of the 1D Steady Euler equations for chemically reacting hypersonic flows in thermal equilibrium and their respective source terms are presented in this document.

1 1D Euler Equations

The conservation of mass, momentum, and total energy for a inviscid compressible fluid composed of a chemically reacting mixture of gases N and N_2 in *thermal equilibrium* may be written as:

$$
\begin{aligned}
\nabla \cdot (\rho_N \boldsymbol{u}) &= \dot{\omega}_N, \\
\nabla \cdot (\rho_{N_2} \boldsymbol{u}) &= \dot{\omega}_{N_2}, \\
\nabla \cdot (\rho \boldsymbol{u}\boldsymbol{u}) &= -\nabla p, \\
\nabla \cdot (\rho \boldsymbol{u}H) &= 0.
\end{aligned}
\tag{1}
$$

where ρ_s is the density of species s (N or N_2), $\rho = \sum_s \rho_s$ is the mixture density and $\boldsymbol{u}$ is the mixture velocity. The total enthalpy, H, may be expressed in terms of the total energy, density, and pressure:

$$
H = E + \frac{p}{\rho},
$$

where the total energy, E, is composed of internal and kinetic components:

$$
E = e^{\text{int}} + \frac{u^2}{2}.
$$

The total internal energy, e^{int}, has contributions from each of the distinct energy *modes*:

$$
\begin{aligned}
e^{\text{int}} &= e^{\text{trans}} + e^{\text{rot}} + e^{\text{vib}} + e^{\text{elec}} + h^0 \\
&= \sum_{s=1}^{ns} c_s e_s^{\text{trans}} + \sum_{s=mol} c_s e_s^{\text{rot}} + \sum_{s=mol} c_s e_s^{\text{vib}} + \sum_{s=1}^{ns} c_s e_s^{\text{elec}} + \sum_{s=1}^{ns} c_s h_s^0,
\end{aligned}
\tag{2}
$$

where $c_s = (\rho_s/\rho)$ is the mass fraction of species s.

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*Work based on Roy, Smith, and Ober (2002); Kirk, Amar, Stogner, Schultz, and Oliver (2009).

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The first four terms on the right of Equation (2) represent the energy due to molecular/atomic translation, molecular rotation, molecular vibration, and electronic excitation. The final term is the heat of formation of the mixture and accounts for the energy stored in chemical bonds (Ait-Ali-Yahia, 1996; Kirk et al., 2009).

Under the approximation that the translational and rotational states of the may be assumed fully populated, translational/rotational energy for each species may be expressed as:

$$e_s^{\text{trans}} + e_s^{\text{rot}} = e_s^{\text{tr}} = C_{v,s}^{\text{tr}} T, \tag{3}$$

where the translational/rotational specific heat, $C_{v,s}^{\text{tr}}$ is given by

$$C_{v,s}^{\text{tr}} = \begin{cases} \frac{5}{2} R_s & \text{for molecules,} \\ \frac{3}{2} R_s & \text{for atoms,} \end{cases}$$

where R_s is the species gas constant, and $R_s = R/M_s$ where R is the universal gas constant and M_s is the species molar mass. The combined term e_s^{tr} in Equation (3) represents the energy due to random thermal translational/rotational motion of a given species.

In contrast to the translational/rotational states, the vibrational energy states are typically not fully populated. One approach for modeling the molecular vibrational energy is through analogy to a harmonic oscillator. In this approach the energy potential between molecular nuclei is modeled as a quadratic function of separation distance. Under this assumption, the vibrational energy for each molecular species can be modeled as:

$$e_s^{\text{vib}} = \begin{cases} \frac{R_s \theta_{vs}}{\exp(\theta_{vs}/T_v) - 1} & \text{for molecules,} \\ 0 & \text{for atoms,} \end{cases}$$

where θ_{vs} is the species characteristic temperature of vibration and T_v is the mixture vibrational temperature. Recall that in the case of thermal equilibrium $T_r = T_t = T_v = T_e \equiv T$.

The energy contained in the excited electronic states for a given species, e_s^{elec}, can be obtained from the assumption that they are in a Boltzmann distribution governed by the electronic excitation temperature T_e (Candler, 1988) as:

$$e_s^{\text{elec}} = R_s \frac{\sum_{i=1}^{\infty} \theta_{is}^{\text{elec}} g_{is} \exp\left(-\theta_{is}^{\text{elec}}/T_e\right)}{g_{0s} + \sum_{i=1}^{\infty} g_{is} \exp\left(-\theta_{is}^{\text{elec}}/T_e\right)}. \tag{4}$$

In practice, Equation (4) can usually be omitted for non-ionized flows such as those considered in this work. Park (1990) observes that electronic transitions in molecules are caused mostly by the impact of free electrons. Since there are no free electrons when there is no ionization, there will be very little electronic excitation. In the present work we choose to *neglect* Equation (4).

Regardless of the thermal state of the mixture, once the translational/rotational temperature T is determined the thermodynamic pressure of the mixture is readily obtained from Dalton's law of partial pressures:

$$p = \sum_{s=1}^{ns} p_s = \sum_{s=1}^{ns} \rho_s R_s T. \tag{5}$$

Therefore, for the Nitrogen dissociation, the expression for internal energy (2) simplifies to:

$$\begin{aligned} e^{\text{int}} &= \sum_{s=1}^{ns} c_s C_{v,s}^{\text{tr}} T + \sum_{s=mol} c_s e_s^{\text{vib}} + \sum_{s=1}^{ns} c_s h_s^0 \\ &= \left(\frac{3}{2} c_{\text{N}} R_{\text{N}} T + \frac{5}{2} c_{\text{N}_2} R_{\text{N}_2} T \right) + \left(\frac{R_{\text{N}_2} \theta_{v\text{N}_2}}{\exp\left(\frac{\theta_{v\text{N}_2}}{T}\right) - 1} \right) + \left(c_{\text{N}} h_{\text{N}}^0 + c_{\text{N}_2} h_{\text{N}_2}^0 \right), \end{aligned} \tag{6}$$

with $c_s = \rho_s/\rho$, $R_s = R/M_s$ where R is the universal gas constant and M_s is the species molar mass for $s = \{\text{N}, \text{N}_2\}$.

1.1 Chemical Kinetics

The rate of production/destruction of the individual species, $\dot{\omega}_s$, is required to close the species continuity equations. For the dissociating Nitrogen flow ($N_2 \rightleftharpoons 2N$) case, let us consider the chemical reactions which occur among the five principal components of dissociating air – N_2, O_2, NO, N, O – but neglecting three species in order to perform this 2–species problem. For this mixture, the single chemical reaction that occur is:

$$N_2 + \mathcal{M} \rightleftharpoons 2N + \mathcal{M}.$$
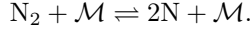
This reaction can occur in either the forward or backward direction, as denoted by the bidirectional arrow. The reaction is presented such that they are endothermic in the forward direction, and $\mathcal{M}$ denotes a generic collision partner, which may be any of the species present in the flow, and is unaltered by the reaction (Kirk et al., 2009). The rate of each reaction is therefore a sum of the forward and backward rates, k_f and k_b:

$$\mathcal{R}_r = \mathcal{R}_{br} - \mathcal{R}_{fr}$$

$$= k_{br} \prod_{s=1}^{ns} \left(\frac{\rho_s}{M_s}\right)^{\beta_{sr}} - k_{fr} \prod_{s=1}^{ns} \left(\frac{\rho_s}{M_s}\right)^{\alpha_{sr}}$$

where α_{sr} and β_{sr} are the stoichiometric coefficients for reactants and products of species s. For the Nitrogen dissociation,

$$\mathcal{R}_1 = \sum_{m \in \mathcal{M}} \left(k_{b_1 m} \frac{\rho_N}{M_N} \frac{\rho_N}{M_N} \frac{\rho_m}{M_m} - k_{f_1 m} \frac{\rho_{N_2}}{M_{N_2}} \frac{\rho_m}{M_m} \right) \tag{7}$$

$$= k_{b_1 N} \frac{\rho_N^3}{M_N^3} - k_{f_1 N} \frac{\rho_{N_2} \rho_N}{2 M_N^2} + k_{b_1 N_2} \frac{\rho_N^2 \rho_{N_2}}{2 M_N^3} - k_{f_1 N_2} \frac{\rho_{N_2}^2}{4 M_N^2}, \tag{8}$$

recalling that $M_{N_2} = 2 M_N$.

The species source terms $\dot{\omega}_s = M_s \sum_{r=1}^{nr} (\alpha_{sr} - \beta_{sr}) (\mathcal{R}_{br} - \mathcal{R}_{fr})$ where nr is the number of reactions can now be expressed in terms of the individual reaction rates as follows:

$$\dot{\omega}_{N_2} = M_{N_2} (\mathcal{R}_1) = 2 M_N \mathcal{R}_1,$$
$$\dot{\omega}_N = M_N (-2\mathcal{R}_1) = -2 M_N \mathcal{R}_1.$$

Note that these source terms sum identically to zero, as required by conservation of mass (Kessler and Awruch, 2004).

The forward rate coefficients k_{fr} can then be expressed in a modified Arrhenius form as

$$k_{fr}(\bar{T}) = C_{fr} \bar{T}^{\eta_r} \exp\left(-E_{ar}/R\bar{T}\right) \tag{9}$$

where C_{fr} is the reaction rate constant, η_r is the so-called pre-exponential factor, E_{ar} is the activation energy. These three constants are determined from curve fits to experimental data (e.g. see Ait-Ali-Yahia (1996)). $\bar{T}$ is a function of the translational/rotational and vibrational temperatures, and in this work, $\bar{T} \equiv T$.

The corresponding backward rate coefficient k_{br} can be found using the principle of detailed balance, which states

$$K_{eq} = \frac{k_{fr}(\bar{T})}{k_{br}(\bar{T})} \tag{10}$$

where K_{eq} is the equilibrium constant and may be obtained either by curve fits or through Gibbs' free energy techniques, such as Park's models. In this work, $K_{eq} = K(T)$.

Therefore:

$$k_{f_1 N} = C_{f_1 N} T^{\eta_{f_1 N}} \exp\left(\frac{-E_{aN}}{RT}\right) \quad \text{and} \quad k_{b_1 N} = \frac{k_{f_1 N}}{K(T)},$$

$$k_{f_1 N_2} = C_{f_1 N_2} T^{\eta_{f_1 N_2}} \exp\left(\frac{-E_{aN_2}}{RT}\right) \quad \text{and} \quad k_{b_1 N_2} = \frac{k_{f_1 N_2}}{K(T)}, \tag{11}$$

are the forward and backward rates for Nitrogen atom and Nitrogen molecule, respectively.

2 Manufactured Solution

Roy et al. (2002) propose the general form of the primitive solution variables to be a function of sines and cosines:

$$\phi(x, y) = \phi_0 + \phi_x f_s \left(\frac{a_{\phi x} \pi x}{L} \right), \tag{12}$$

where $\phi = \rho_{\mathrm{N}}, \rho_{\mathrm{N}_2}, u$ or T, and $f_s(\cdot)$ functions denote either sine or cosine function. Note that in this case, ϕ_x is constant and the subscript does not denote differentiation. Different choices of the constants used in the manufactured solutions for the 2D supersonic and subsonic cases of Euler and Navier-Stokes may be found in Roy et al. (2002).

Therefore, the manufactured analytical solution for each one of the variables in Euler equations are:

$$\begin{aligned}
\rho_{\mathrm{N}}(x) &= \rho_{\mathrm{N}0} + \rho_{\mathrm{N}x} \sin \left(\frac{a_{\rho \mathrm{N}x} \pi x}{L} \right), \\
\rho_{\mathrm{N}_2}(x) &= \rho_{\mathrm{N}_20} + \rho_{\mathrm{N}_2 x} \cos \left(\frac{a_{\rho \mathrm{N}_2 x} \pi x}{L} \right), \\
u(x) &= u_0 + u_x \sin \left(\frac{a_{ux} \pi x}{L} \right), \\
T(x) &= T_0 + T_x \cos \left(\frac{a_{Tx} \pi x}{L} \right).
\end{aligned} \tag{13}$$

Recalling that $\rho = \sum_s \rho_s$, the manufactured analytical solution for the density of the mixture is:

$$\begin{aligned}
\rho(x) &= \rho_{\mathrm{N}} + \rho_{\mathrm{N}_2} \\
&= \rho_{\mathrm{N}0} + \rho_{\mathrm{N}_20} + \rho_{\mathrm{N}x} \sin \left(\frac{a_{\rho \mathrm{N}x} \pi x}{L} \right) + \rho_{\mathrm{N}_2 x} \cos \left(\frac{a_{\rho \mathrm{N}_2 x} \pi x}{L} \right).
\end{aligned} \tag{14}$$

The MMS applied to 1D steady Euler equations for a chemically reacting mixture of N and N_2 in thermal equilibrium consists in modifying Equations (1) by adding a source term to the right-hand side of each equation:

$$\begin{aligned}
\frac{\partial (\rho_{\mathrm{N}} u)}{\partial x} - \dot{\omega}_{\mathrm{N}} &= Q_{\rho_{\mathrm{N}}}, \\
\frac{\partial (\rho_{\mathrm{N}_2} u)}{\partial x} - \dot{\omega}_{\mathrm{N}_2} &= Q_{\rho_{\mathrm{N}_2}}, \\
\frac{\partial (\rho u^2)}{\partial x} + \frac{\partial (p)}{\partial x} &= Q_u, \\
\frac{\partial (\rho u E)}{\partial x} + \frac{\partial (p u)}{\partial x} &= Q_E,
\end{aligned} \tag{15}$$

so the modified set of Equations (15) conveniently has the analytical solutions given in Equations (13) and (14).

Source terms $Q_{\rho_{\mathrm{N}}}$, $Q_{\rho_{\mathrm{N}_2}}$, Q_u and Q_E are obtained by symbolic manipulations of equations above using Maple and are presented in the following sections. The following auxiliary variables have been included in order to improve readability and computational efficiency:

$$\begin{aligned}
\mathtt{Rho}_{\mathrm{N}} &= \rho_{\mathrm{N}0} + \rho_{\mathrm{N}x} \sin \left(\frac{a_{\rho \mathrm{N}x} \pi x}{L} \right), \\
\mathtt{Rho}_{\mathrm{N}_2} &= \rho_{\mathrm{N}_20} + \rho_{\mathrm{N}_2 x} \cos \left(\frac{a_{\rho \mathrm{N}_2 x} \pi x}{L} \right), \\
\mathtt{Rho} &= \mathtt{Rho}_{\mathrm{N}} + \mathtt{Rho}_{\mathrm{N}_2}, \\
\mathtt{U} &= u_0 + u_x \sin \left(\frac{a_{ux} \pi x}{L} \right), \\
\mathtt{T} &= T_0 + T_x \cos \left(\frac{a_{Tx} \pi x}{L} \right).
\end{aligned} \tag{16}$$

The resulting source terms for the 1D steady Euler flow with Nitrogen dissociation in thermal equilibrium described by Equations (1) are obtained through symbolic manipulation using the software Maple and are presented in the following sections.

3 Euler mass conservation equation for Nitrogen atom

The mass conservation equation for Nitrogen atom (N), written as an operator, is:

$$\mathcal{L} = \frac{\partial \rho_N u}{\partial x} - \dot{\omega}_N.$$

Analytically differentiating Equation (13) for ρ_N, and u using operator $\mathcal{L}$ defined above together with suitable substitution for $\dot{\omega}_N$ gives the source term Q_{ρ_N}:

$$
\begin{aligned}
Q_{\rho_N} =\ & \frac{a_{\rho Nx}\pi\rho_{Nx}\, \mathtt{U}}{L}\cos\left(\frac{a_{\rho Nx}\pi x}{L}\right) + \\
& + \frac{a_{ux}\pi u_x\, \mathtt{Rho}_N}{L}\cos\left(\frac{a_{ux}\pi x}{L}\right) + \\
& + \frac{\mathtt{Rho}_N^2}{M_N^2 K(\mathtt{T})}\left(2k_{f_1 N}\,\mathtt{Rho}_N + k_{f_1 N_2}\,\mathtt{Rho}_{N_2}\right) + \\
& - \frac{\mathtt{Rho}_{N_2}}{2 M_N}\left(2k_{f_1 N}\,\mathtt{Rho}_N + k_{f_1 N_2}\,\mathtt{Rho}_{N_2}\right),
\end{aligned}
\tag{17}
$$

where M_N is the molar mass of Nitrogen, and $k_{f_1 N}$ and $k_{f_1 N_2}$ are defined by:

$$
\begin{aligned}
k_{f_1 N} &= C_{f_1 N}\, \mathtt{T}^{\eta_{f_1 N}}\exp\left(\frac{-E_{aN}}{R\,\mathtt{T}}\right), \\
k_{f_1 N_2} &= C_{f_1 N_2}\, \mathtt{T}^{\eta_{f_1 N_2}}\exp\left(\frac{-E_{aN_2}}{R\,\mathtt{T}}\right).
\end{aligned}
\tag{18}
$$

with $\mathtt{Rho}_N$, $\mathtt{Rho}_{N_2}$, $\mathtt{T}$ and $\mathtt{U}$ defined in Equation (16).

4 Euler mass conservation equation for Nitrogen molecule

The mass conservation equation for Nitrogen molecule (N_2), written as an operator, is:

$$\mathcal{L} = \frac{\partial \rho_{N_2} u}{\partial x} - \dot{\omega}_{N_2}.$$

Analytically differentiating Equation (13) for ρ_{N_2}, u and v using operator $\mathcal{L}$ defined above together with suitable substitution for $\dot{\omega}_{N_2}$ gives the source term $Q_{\rho_{N_2}}$:

$$
\begin{aligned}
Q_{\rho_{N_2}} =\ & -\frac{a_{\rho N_2 x}\pi\rho_{N_2 x}\, \mathtt{U}}{L}\sin\left(\frac{a_{\rho N_2 x}\pi x}{L}\right) + \\
& + \frac{a_{ux}\pi u_x\, \mathtt{Rho}_{N_2}}{L}\cos\left(\frac{a_{ux}\pi x}{L}\right) + \\
& - \frac{\mathtt{Rho}_N^2}{M_N^2 K(\mathtt{T})}\left(2k_{f_1 N}\,\mathtt{Rho}_N + k_{f_1 N_2}\,\mathtt{Rho}_{N_2}\right) + \\
& + \frac{\mathtt{Rho}_{N_2}}{2 M_N}\left(2k_{f_1 N}\,\mathtt{Rho}_N + k_{f_1 N_2}\,\mathtt{Rho}_{N_2}\right).
\end{aligned}
\tag{19}
$$

where $k_{f_1 N}$ and $k_{f_1 N_2}$ are given in Equation (18); and $\mathtt{Rho}_N$, $\mathtt{Rho}_{N_2}$ and $\mathtt{U}$ are given in Equation (16).

5 Euler momentum equation

For the generation of the analytical source term Q_u, the x momentum equation in Equation (1) is written as an operator $\mathcal{L}$:

$$\mathcal{L} = \frac{\partial(\rho u^2)}{\partial x},$$

which, when operated in to Equations (13) and (14), provides source term Q_u:

$$
\begin{aligned}
Q_u = {} & \frac{a_{\rho \mathrm{N}x}\pi\rho_{\mathrm{N}x}\,\mathtt{U}^2}{L}\cos\left(\frac{a_{\rho \mathrm{N}x}\pi x}{L}\right) + \\
& -\frac{a_{\rho \mathrm{N}_2 x}\pi\rho_{N_2 x}\,\mathtt{U}^2}{L}\sin\left(\frac{a_{\rho \mathrm{N}_2 x}\pi x}{L}\right) + \\
& -\frac{a_{Tx}\pi T_x R_{\mathrm{N}}\,\mathtt{Rho_N}}{L}\sin\left(\frac{a_{Tx}\pi x}{L}\right) + \\
& -\frac{a_{Tx}\pi T_x R_{\mathrm{N}}\,\mathtt{Rho_{N_2}}}{2L}\sin\left(\frac{a_{Tx}\pi x}{L}\right) + \\
& +\frac{2a_{ux}\pi u_x\,\mathtt{Rho\,U}}{L}\cos\left(\frac{a_{ux}\pi x}{L}\right) + \\
& -\frac{\pi R_{\mathrm{N}}\,\mathtt{T}}{2L}\left(-2a_{\rho \mathrm{N}x}\rho_{\mathrm{N}x}\cos\left(\frac{a_{\rho \mathrm{N}x}\pi x}{L}\right) + a_{\rho \mathrm{N}_2 x}\rho_{N_2 x}\sin\left(\frac{a_{\rho \mathrm{N}_2 x}\pi x}{L}\right)\right),
\end{aligned}
\tag{20}
$$

with $\mathtt{Rho}$, $\mathtt{Rho_N}$, $\mathtt{Rho_{N_2}}$, $\mathtt{T}$ and $\mathtt{U}$ defined in Equation (16).

6 Euler energy equation

The total energy equation is written as an operator:

$$
\mathcal{L} = \frac{\partial \rho u H}{\partial x},
$$

where $H = E + \dfrac{p}{\rho}$ and $E = e^{\mathrm{int}} + \dfrac{u^2}{2}$, with p and e^{int} as defined in Equations (5) and (6), respectively.

Source term Q_E is obtained by operating $\mathcal{L}$ on Equations (13) and (14):

$$
\begin{aligned}
Q_E = {} & -\frac{a_{Tx}\pi T_x\,\alpha\,\theta_{v\mathrm{N}_2}E_{vib,\mathrm{N}_2}\,\mathtt{Rho_{N_2}}\,\mathtt{U}}{L(\alpha-1)\,\mathtt{T}^2}\sin\left(\frac{a_{Tx}\pi x}{L}\right) + \\
& -\frac{5a_{Tx}\pi T_x R_{\mathrm{N}}\,\mathtt{Rho_N}\,\mathtt{U}}{2L}\sin\left(\frac{a_{Tx}\pi x}{L}\right) + \\
& -\frac{7a_{Tx}\pi T_x R_{\mathrm{N}}\,\mathtt{Rho_{N_2}}\,\mathtt{U}}{4L}\sin\left(\frac{a_{Tx}\pi x}{L}\right) + \\
& +\frac{5a_{ux}\pi u_x R_{\mathrm{N}}\,\mathtt{Rho_N}\,\mathtt{T}}{2L}\cos\left(\frac{a_{ux}\pi x}{L}\right) + \\
& +\frac{7a_{ux}\pi u_x R_{\mathrm{N}}\,\mathtt{Rho_{N_2}}\,\mathtt{T}}{4L}\cos\left(\frac{a_{ux}\pi x}{L}\right) + \\
& +\frac{3a_{ux}\pi u_x\,\mathtt{Rho\,U}^2}{2L}\cos\left(\frac{a_{ux}\pi x}{L}\right) + \\
& +\frac{a_{ux}\pi u_x(h^0_{\mathrm{N}}\,\mathtt{Rho_N} + h^0_{\mathrm{N}_2}\,\mathtt{Rho_{N_2}})}{L}\cos\left(\frac{a_{ux}\pi x}{L}\right) + \\
& -\frac{a_{\rho \mathrm{N}_2 x}\pi\rho_{N_2 x}E_{vib,\mathrm{N}_2}\,\mathtt{U}}{L}\sin\left(\frac{a_{\rho \mathrm{N}_2 x}\pi x}{L}\right) + \\
& +\frac{a_{ux}\pi u_x E_{vib,\mathrm{N}_2}\,\mathtt{Rho_{N_2}}}{L}\cos\left(\frac{a_{ux}\pi x}{L}\right) + \\
& +\frac{\pi R_{\mathrm{N}}\,\mathtt{U\,T}}{4L}\left(10a_{\rho \mathrm{N}x}\rho_{\mathrm{N}x}\cos\left(\frac{a_{\rho \mathrm{N}x}\pi x}{L}\right) - 7a_{\rho \mathrm{N}_2 x}\rho_{N_2 x}\sin\left(\frac{a_{\rho \mathrm{N}_2 x}\pi x}{L}\right)\right) + \\
& +\frac{\pi\,\mathtt{U}^3}{2L}\left(a_{\rho \mathrm{N}x}\rho_{\mathrm{N}x}\cos\left(\frac{a_{\rho \mathrm{N}x}\pi x}{L}\right) - a_{\rho \mathrm{N}_2 x}\rho_{N_2 x}\sin\left(\frac{a_{\rho \mathrm{N}_2 x}\pi x}{L}\right)\right) + \\
& +\frac{\pi\,\mathtt{U}}{L}\left(a_{\rho \mathrm{N}x}\rho_{\mathrm{N}x}h^0_{\mathrm{N}}\cos\left(\frac{a_{\rho \mathrm{N}x}\pi x}{L}\right) - a_{\rho \mathrm{N}_2 x}\rho_{N_2 x}h^0_{\mathrm{N}_2}\sin\left(\frac{a_{\rho \mathrm{N}_2 x}\pi x}{L}\right)\right),
\end{aligned}
\tag{21}
$$

where $\mathtt{Rho}$, $\mathtt{Rho_N}$, $\mathtt{Rho_{N_2}}$, $\mathtt{T}$ and $\mathtt{U}$ are given in Equation (16) and:

$$\alpha = \exp\left(\frac{\theta_{vN_2}}{T}\right) \quad \text{and} \quad E_{vib,N_2} = \frac{R_{N_2}\theta_{vN_2}}{(\alpha - 1)}.$$

7   Comments

Source terms Q_{ρ_N}, $Q_{\rho_{N_2}}$, Q_u and Q_E have been generated by replacing the analytical Expressions (13) and (14) into respective Equations (1), followed by the usage of Maple commands for collecting, sorting and factorizing the terms. Files containing C codes for the source terms have also been generated. They are: Euler_1d_steady_chemistry_1T_rho_N_code.C, Euler_1d_steady_chemistry_1T_rho_N2_code.C, Euler_1d_steady_chemistry_1T_u_code.C, and Euler_1d_steady_chemistry_1T_E_code.C.

An example of the automatically generated C file from the source term for the mass conservation equation of N is:

```
#include <math.h>
double SourceQ_rho_N (double x, double M_N, double h0_N, double h0_N2, double Cf1_N, double Cf1_N2,
  double etaf1_N, double etaf1_N2, double Ea_N, double Ea_N2, double Function_to_Calculate_K)
{
  double Q_rho_N;
  double RHO_N;
  double RHO_N2;
  double U;
  double T;
  double kf1_N;
  double kf1_N2;
  double K_eq;

  K_eq = Function_to_Calculate_K;
  RHO_N = rho_N_0 + rho_N_x * sin(a_rho_N_x * PI * x / L);
  RHO_N2 = rho_N2_0 + rho_N2_x * cos(a_rho_N2_x * PI * x / L);
  U = u_0 + u_x * sin(a_ux * PI * x / L);
  T = T_0 + T_x * cos(a_Tx * PI * x / L);
  kf1_N = Cf1_N * pow(T, etaf1_N) * exp(-Ea_N / R / T);
  kf1_N2 = Cf1_N2 * pow(T, etaf1_N2) * exp(-Ea_N2 / R / T);

  Q_rho_N = a_rho_N_x * PI * rho_N_x * U * cos(a_rho_N_x * PI * x / L) / L
    + a_ux * PI * u_x * RHO_N * cos(a_ux * PI * x / L) / L
    + (0.2e1 * kf1_N * RHO_N + kf1_N2 * RHO_N2) * RHO_N * RHO_N * pow(M_N, -0.2e1) / K_eq
    - (0.2e1 * kf1_N * RHO_N + kf1_N2 * RHO_N2) * RHO_N2 / M_N / 0.2e1;
  return(Q_rho_N);
}
```

and the source term for the total energy E of the mixture is:

```
#include <math.h>
double SourceQ_e (double x, double R_N, double R_N2, double h0_N, double h0_N2, double theta_v_N2)
{
  double Q_e;
  double RHO;
  double RHO_N;
  double RHO_N2;
  double U;
  double T;
  double alpha;
  double E_vib_N2;

  alpha = exp(theta_v_N2 / T);
  E_vib_N2 = R_N2 * theta_v_N2 / (alpha - 0.1e1);
```

```
  RHO_N = rho_N_0 + rho_N_x * sin(a_rho_N_x * PI * x / L);
  RHO_N2 = rho_N2_0 + rho_N2_x * cos(a_rho_N2_x * PI * x / L);
  U = u_0 + u_x * sin(a_ux * PI * x / L);
  T = T_0 + T_x * cos(a_Tx * PI * x / L);

 Q_e = -a_Tx * PI * T_x * alpha * theta_v_N2 * E_vib_N2 * RHO_N2 * U
    * sin(a_Tx * PI * x / L) / L / (alpha - 0.1e1) * pow(T, -0.2e1)
  - 0.5e1 / 0.2e1 * a_Tx * PI * T_x * R_N * RHO_N * U * sin(a_Tx * PI * x / L) / L
  - 0.7e1 / 0.4e1 * a_Tx * PI * T_x * R_N * RHO_N2 * U * sin(a_Tx * PI * x / L) / L
  + 0.5e1 / 0.2e1 * a_ux * PI * u_x * R_N * RHO_N * T * cos(a_ux * PI * x / L) / L
  + 0.7e1 / 0.4e1 * a_ux * PI * u_x * R_N * RHO_N2 * T * cos(a_ux * PI * x / L) / L
  + 0.3e1 / 0.2e1 * a_ux * PI * u_x * RHO * U * U * cos(a_ux * PI * x / L) / L
  - a_rho_N2_x * PI * rho_N2_x * E_vib_N2 * U * sin(a_rho_N2_x * PI * x / L) / L
  + a_ux * PI * u_x * E_vib_N2 * RHO_N2 * cos(a_ux * PI * x / L) / L
  + (h0_N * RHO_N + h0_N2 * RHO_N2) * a_ux * PI * u_x * cos(a_ux * PI * x / L) / L
  - (-0.10e2 * a_rho_N_x * rho_N_x * cos(a_rho_N_x * PI * x / L)
    + 0.7e1 * a_rho_N2_x * rho_N2_x * sin(a_rho_N2_x * PI * x / L)) * PI * R_N * U * T / L / 0.4e1
  - (-a_rho_N_x * rho_N_x * cos(a_rho_N_x * PI * x / L)
    + a_rho_N2_x * rho_N2_x * sin(a_rho_N2_x * PI * x / L)) * PI * pow(U, 0.3e1) / L / 0.2e1
  - (-a_rho_N_x * rho_N_x * h0_N * cos(a_rho_N_x * PI * x / L)
    + a_rho_N2_x * rho_N2_x * h0_N2 * sin(a_rho_N2_x * PI * x / L)) * PI * U / L;
  return(Q_e);
}
```

Finally, the gradients of the analytical solutions (13) and (14) have also been computed and their respective C codes are presented in `Euler_1d_chemistry_manuf_solutions_grad_and_code.C`. Therefore,

$$
\begin{aligned}
\nabla \rho_{\mathrm{N}} &= \frac{a_{\rho \mathrm{N}x} \pi \rho_{\mathrm{N}x}}{L} \cos\left(\frac{a_{\rho \mathrm{N}x} \pi x}{L}\right), \quad \nabla \rho_{\mathrm{N}_2} = \frac{a_{\rho \mathrm{N}_2 x} \pi \rho_{\mathrm{N}_2 x}}{L} \sin\left(\frac{a_{\rho \mathrm{N}_2 x} \pi x}{L}\right), \\
\nabla T &= -\frac{a_{Tx} \pi p_x}{L} \sin\left(\frac{a_{Tx} \pi x}{L}\right), \qquad \nabla u = \frac{a_{ux} \pi u_x}{L} \cos\left(\frac{a_{ux} \pi x}{L}\right),
\end{aligned}
\tag{22}
$$

and $\nabla \rho = \nabla \left(\rho_{\mathrm{N}}\right) + \nabla \left(\rho_{\mathrm{N}_2}\right),$ are written in C language as:

```
grad_rho_an_N[0] = rho_N_x * cos(a_rho_N_x * pi * x / L) * a_rho_N_x * pi / L;
grad_rho_an_N2[0] = -rho_N2_x * sin(a_rho_N2_x * pi * x / L) * a_rho_N2_x * pi / L;
grad_rho_an[0] = rho_N_x * cos(a_rho_N_x * pi * x / L) * a_rho_N_x * pi / L
                - rho_N2_x * sin(a_rho_N2_x * pi * x / L) * a_rho_N2_x * pi / L;
grad_u_an[0] = u_x * cos(a_ux * pi * x / L) * a_ux * pi / L;
grad_T_an[0] = -T_x * sin(a_Tx * pi * x / L) * a_Tx * pi / L;
```

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