

Densities, infrared band strengths, and optical constants of solid methanol

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ABSTRACT

Contact. The increasing capabilities of space missions like the *James Webb* Space Telescope or ground-based observatories like the European Extremely Large Telescope demand high quality laboratory data of species in astrophysical conditions for the interpretation of their findings.

Aims. We provide new physical and spectroscopic data of solid methanol that will help to identify this species in astronomical environments.

Methods. Ices were grown by vapour deposition in high vacuum chambers. Densities were measured via a cryogenic quartz crystal microbalance and laser interferometry. Absorbance infrared spectra of methanol ices of different thickness were recorded to obtain optical constants using an iterative minimization procedure. Infrared band strengths were determined from infrared spectra and ice densities.

Results. Solid methanol densities measured at eight temperatures vary between 0.64 g cm^{-3} at 20 K and 0.84 g cm^{-3} at 130 K . The visible refractive index at 633 nm grows from 1.26 to 1.35 in that temperature range. New infrared optical constants and band strengths are given from 650 to 5000 cm^{-1} (15.4 – $2.0 \mu\text{m}$) at the same eight temperatures. The study was made on ices directly grown at the indicated temperatures, and amorphous and crystalline phases have been recognized. Our optical constants differ from those previously reported in the literature for an ice grown at 10 K and subsequently warmed. The disagreement is due to different ice morphologies. The new infrared band strengths agree with previous literature data when the correct densities are considered.

Key words. solid state: volatile – methods: laboratory: molecular – techniques: spectroscopic – ISM: abundances – infrared: ISM – infrared: planetary systems

1. Introduction

Ice mantles on dust grains in dense clouds in the interstellar medium (ISM) mainly consist of H_2O , CO, and CO_2 . Although it varies from source to source, methanol is in many of them the fourth most abundant species, with an abundance of 6% with respect to water. Slightly lower amounts of NH_3 and CH_4 are also present (Boogert et al. 2015). Methanol is formed on the surface of the grains by hydrogenation of CO to form H_2CO and then CH_3OH (Watanabe & Kouchi 2002; Ioppolo et al. 2011). The distribution of these species in the ice mantle is not homogenous and probably two different layers exist, an inner water-rich layer containing most of the CO_2 and NH_3 , covered by an external water-poor layer formed of CO, the remaining CO_2 , and most of the CH_3OH (Öberg 2016). In particular, the possibility that pure CH_3OH islands are segregated in the water-poor ice layer has been considered by some authors (Bottinelli et al. 2010).

Solid methanol has also been detected on the surface of minor bodies in the solar system, like some trans-Neptunian objects (TNOs) and Centaurus (Merlin et al. 2012). In Enceladus, one of the moons of Saturn, the presence of methanol ice was reported by Hodyss et al. (2009) based on an absorption feature appearing at $3.53 \mu\text{m}$ (233 cm^{-1}) in an infrared spectrum of its surface registered by Cassini. Recently, the detection of a methanol cloud in Enceladus has been reported (Drabek-Maunder et al. 2017), which strengthens the possibility of having this species in solid phase on the surface.

Information about the composition and structure of astrophysical ices, and in particular the presence of solid methanol, is obtained by comparing astrophysical observations in the infrared to laboratory data (Dartois et al. 1999; Pontoppidan et al. 2003; Boogert et al. 2008; Bottinelli et al. 2010). The main infrared absorption features used for solid CH_3OH identification in interstellar ices are the $3.53 \mu\text{m}$ (2833 cm^{-1}) and $9.74 \mu\text{m}$ (1027 cm^{-1}) bands. They correspond to the CH_3 symmetric stretching mode and the CO stretching mode, respectively. The laboratory data most widely used for comparison to observations are those provided by d'Hendecourt & Allamandola (1986), Hudgins et al. (1993), and Kerkhof et al. (1999). The first reference gives IR spectra and band strengths of methanol ice at 10 K in the 2.5 – $20 \mu\text{m}$ range. The second reference presents a thorough infrared study providing a compendium of band strengths and optical constants in the 2.5 – $200 \mu\text{m}$ interval for ices of astrophysical relevance. For pure methanol, optical constants are given for an ice grown at 10 K and subsequently warmed to 50 K , 75 K , 100 K , and 120 K . Band strengths at 10 K are also given. As far as we know, these are the only optical constants available in the literature for solid methanol. The third reference reports band strength variations of methanol absorptions in the mid-IR region with dilution in H_2O ice, or in H_2O and CO_2 ice mixtures. Additional laboratory works provide infrared band strengths of this species at 10 K . Specifically, the data in Palumbo et al. (1999) correspond to the mid-IR (MIR), and the data in Sandford & Allamandola (1993) and Gerakines et al. (2005) to the near-IR (NIR). Finally,

a recent paper by Bouilloud et al. (2015) presents a compilation and new measurements of MIR band strengths of methanol ice at low temperatures. Since methanol ice densities were unknown, the uncertainty on band-strength determination is very large. In particular, a discrepancy among the literature results up to 40% for the strength of the $9.74\mu\text{m}$ band has been pointed out.

Our goal is to provide new data for solid methanol by obtaining, for the first time, actual density values for background deposited ices and by completing the available spectroscopic information on this species in the NIR and MIR spectral ranges. Thus, we have measured densities and the visible refractive index at 633 nm for methanol ices grown at 20 K , 40 K , 60 K , 80 K , 100 K , 110 K , 120 K , and 130 K . We have evaluated IR band strengths and optical constants of solid methanol at those temperatures in the MIR and NIR spectral ranges from 650 cm^{-1} to 5000 cm^{-1} ($15.4\mu\text{m}$ to $2.0\mu\text{m}$).

This paper is organized as follows: the experimental set-up and the procedures to determine densities, visible refractive index, infrared optical constants and band strengths are explained in Sect. 2; results are presented in Sect. 3; Sect. 4 is devoted to a discussion of the results in comparison with previous studies; and finally, our conclusions are summarized in Sect. 5.

2. Experiments and data analysis

Our experiments were performed using two different experimental set-ups, one at the UPV-Alcoy and the other at IEM-CSIC-Madrid. Both have been described in detail in previous publications (Satorre et al. 2017; Molpeceres et al. 2016; Zanchet et al. 2013).

The set-up at Alcoy allows us to determine densities and visible refractive indices. Densities are obtained by dividing the mass deposited in a certain area (g cm^{-2}) by the thickness of the ice formed. The mass is determined by a quartz crystal microbalance (QCMB, 5 MHz AT-cut Q-Sense) that follows the Sauerbrey equation (Sauerbrey 1959),

$$\Delta F = -S \Delta m, \quad (1)$$

where ΔF is the frequency variation (Hz) (right axis in Fig. 1), S is a constant ($\text{g cm}^{-2}\text{ Hz}^{-1}$), and Δm is the mass deposited (g). Thickness was calculated by laser interferometry from the following expression giving the increase in thickness corresponding to one maximum of interference,

$$\Delta x = \frac{\lambda_0}{2n_0 \sqrt{1 - (\frac{\sin \alpha}{n_0})^2}}, \quad (2)$$

where λ_0 is the wavelength of the He–Ne laser (633 nm), n_0 is the refractive index of the ice film at the laser wavelength, and α the incidence angle of the laser from the surface normal. For most ices formed at low temperatures n_0 is unknown. Two interference patterns were obtained at two different angles of incidence (34° and 64° , left axis Fig. 1), giving us a second equation to solve two unknowns. The refractive index can be obtained directly from

$$n^2 = \frac{\sin^2 \beta - \gamma^2 \sin^2 \alpha}{1 - \gamma^2}, \quad (3)$$

where $\gamma^2 = \frac{T_\alpha}{T_\beta}$ and T_α and T_β are the periods of the corresponding angles of incidence measured in seconds.

Films were formed at constant temperature allocating the QCMB at the end of the second stage of a Leybold closed-cycle

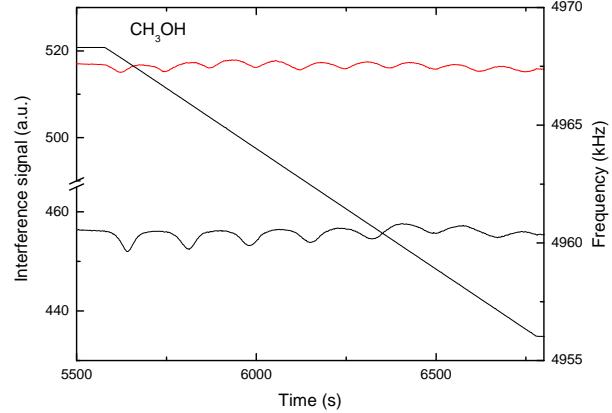


Fig. 1. Laser signals at two angles of incidence (left axis) and straight variation of the frequency of a QCMB (right axis) during methanol deposition at 80 K at constant rate.

He cryostat, where the temperature could be controlled from 12 K to room temperature by an ITC Oxford temperature controller. In these experiments, the temperature was fixed within 1 K . Cold parts were surrounded by a radiation shield inside a high vacuum chamber to prevent contaminants in the QCMB. The pressure in the chamber increases from 5×10^{-8} mbar before accretion to a constant value of 10^{-5} mbar during the deposition. Methanol molecules pass from a prechamber to the deposition chamber through a needle valve. Methanol pressure in the prechamber is maintained constant at 18.7 mbar thanks to both a liquid methanol reservoir at constant temperature and a controlled flux, through another needle valve, from the reservoir to the prechamber. These conditions lead to a deposition rate of $\approx 1\text{ nm s}^{-1}$. The methanol used was from Panreac, UHPLC, 99.9% purity, degassed previous to deposition.

Infrared spectra were obtained in the IEM-CSIC laboratory at Madrid. Ices of methanol were grown in a high vacuum 150 mm diameter cylindrical chamber pumped with a 210 l s^{-1} (for N_2) turbo pump through a DN100 CF flange placed in its base. The chamber is provided with a closed-cycle He cryostat, coupled to a Fourier transform infrared spectrometer and a He–Ne laser to measure ice thickness by interferometry (see Fig. 2). The gas was introduced into the chamber through a needle valve coming from a methanol line (Fisher Scientific, HPLC, 99.99% purity) thermally stabilized at a vapour pressure of $40 \pm 2\text{ mbar}$. This low pressure was achieved and held constant immersing a methanol flask in an ice-water bath. The gas inlet is situated 120 mm below the optical axis transversal plane, and 200 mm above the pump flange. The geometry of the chamber was especially designed to obtain a symmetric deposition on both sides of the IR transparent Si wafer, placed in close contact with the cold head of the cryostat. This was confirmed by monitoring with the He–Ne laser the desorption of the back-side deposit. We found that we have the same thickness in both faces, within 5% error. For the experiments performed in this publication we employed a constant growth rate of 1.2 nm s^{-1} , estimated using Eq. (2), and the time it takes to grow an interference maximum. At this growing rate, the water contamination of the methanol ice, due to the residual water vapour in the chamber, is less than 0.1%. The experiments were performed via the following sequence: first, the desired temperature was set; then gas was introduced in the chamber while He–Ne interference fringes were recorded; after one maximum of interference was grown, the gas inlet valve was closed, and a MIR spectrum was taken.

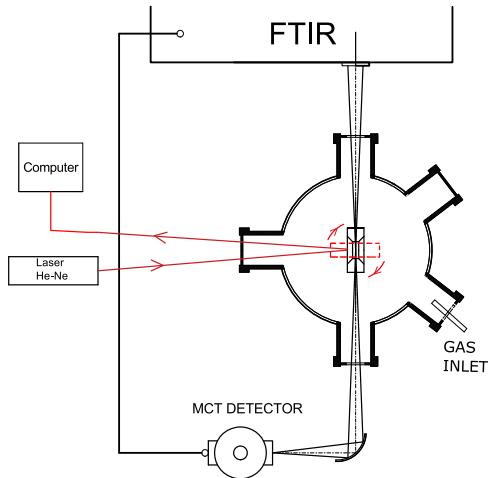


Fig. 2. Scheme of the experimental set-up used to record transmission IR spectra of methanol ices.

This sequence was repeated until four interference maxima were grown and four MIR spectra were obtained. We note that it was necessary to rotate alternatively the Si substrate to face two different windows in the chamber to record He–Ne interferences in quasi normal incidence ($4.0 \pm 0.2^\circ$) or to take FTIR spectra in normal transmission configuration (see Fig. 2). Due to the weakness of the NIR absorptions, no NIR spectra were recorded for those thin layers. Subsequently, at the same temperature, ten more interference maxima were grown and, finally, a near infrared spectrum of the thick ice layer was obtained. Infrared spectra were recorded with a Burker Vertex 70 spectrometer, using a Globar source and a KBr beam splitter for the MIR spectral range, and a CaF beamsplitter and a NIR lamp for the NIR range. The same mercury cadmium telluride detector was used for both regions. We worked at 2 cm^{-1} resolution and an accumulation of 300 scans. In these experiments, ice layer thicknesses were determined from He–Ne laser interference measurements, using Eq. (2) and n_0 from Table 1. To avoid saturation of the MIR absorption bands, only ice layers up to $\approx 1000\text{ nm}$ thick were considered. In the NIR spectral range, ice layers of about 3400 nm were employed.

2.1. Infrared optical constants

Infrared optical constants and refractive index are denoted in the following $n(\text{NIR})+i k(\text{NIR})$ and $n(\text{MIR})+i k(\text{MIR})$ respectively for the NIR and MIR spectral ranges. In this work we set the limits for the NIR region from 5000 ($2.0\text{ }\mu\text{m}$) to 3600 cm^{-1} ($2.7\text{ }\mu\text{m}$), and for the MIR region from 3600 ($2.7\text{ }\mu\text{m}$) to 650 cm^{-1} ($15.4\text{ }\mu\text{m}$). To determine $n(\text{MIR})$ and $k(\text{MIR})$ we have followed a procedure developed in our group and described in detail in Zanchet et al. (2013). This procedure takes advantage of the redundancy of information available when a set of spectra of different thicknesses of the same ice is available. In particular, four mid-IR spectra with thickness between 200 nm and 1000 nm were recorded for each ice. Theoretical transmission spectra were obtained using the Fresnel equations for a multilayer system vacuum/ice/silicon/ice/vacuum, and taking initial input values for $n(\text{MIR})$, $k(\text{MIR})$, n_0 (visible refractive index), and L (ice layer thickness). Then the difference between experimental and theoretical spectra for the four experimental spectra available were minimized (see details in Zanchet et al. 2013). In the minimization procedure the measured n_0 and layer

Table 1. Refractive index at 633 nm and densities of pure CH_3OH ices grown at the indicated temperatures.

T (K)	n_0	Density (g cm^{-3})
20	1.257	0.636
40	1.273	0.664
60	1.271	0.685
80	1.303	0.724
100	1.343	0.798
110	1.346	0.822
120	1.339	0.795
130	1.346	0.838

thickness can vary within experimental error. Additionally, a coherence parameter that varies between 1 and 0 is fitted to consider coherence losses of the IR radiation in its transmission path through the layered system. At the end of the fitting procedure a set of $n(\text{MIR})$ and $k(\text{MIR})$ optical constants is obtained.

For the near infrared spectral range, infrared scattering and coherence losses become important in relation to the intensity of the absorptions, and the model described above, based on Fresnel equations, is not able to reproduce the experimental infrared interference pattern in the 5000 – 3500 cm^{-1} region. For this reason, we decided to record a single NIR spectrum for each ice temperature, corresponding to a thickness of about 3400 nm, and to extract $k(\text{NIR})$ applying Lambert–Beer’s law

$$k(\nu) = \frac{1}{4\pi\nu L} \text{Abs}(\nu), \quad (4)$$

where $\text{Abs}(\nu)$ is the absorbance spectrum (baseline subtracted), L the ice thickness, and ν the frequency. The real component $n(\text{NIR})$ was then obtained by means of the Kramers–Kroning relation between n and k , considering the information we have for n_0 , $n(\text{MIR})$, and $k(\text{MIR})$.

2.2. Band strengths

The integrated band strengths can be obtained from the imaginary part of the refractive index via

$$A = \frac{4\pi}{\rho} \int_{\text{band}} \nu k(\nu) \text{d}\nu, \quad (5)$$

where ρ is the ice density. Ice densities at different temperatures were taken from Table 1. When the imaginary part of the refractive index is not known it is possible to obtain an effective band strength, A' , from the absorbance spectrum and the column density of absorbing species, using

$$A' = \frac{2.303}{\rho L} \int_{\text{band}} \text{Abs}(\nu) \text{d}\nu. \quad (6)$$

The effective band strength can be affected by scattering and IR interferences that disturb the baseline of the absorption spectra and are difficult to disentangle from real absorptions (Hudson et al. 2014). It is interesting to point out that Eqs. (5) and (6) both involve the density of the ice. Ice densities are usually unknown and an A' value is usually estimated assuming an ice density of 1 g cm^{-3} . This was not the case in the present work, where we have measured the densities of the ices in an independent experiment (see Table 1).

3. Results

3.1. Densities and refractive indices.

Figure 3 collects density and refractive indices at eight temperatures covering the range from 20 to 130 K. Each experimental point is the result of two experiments except for 100 K, which corresponds to the mean value of eight different experiments. Experimental errors are calculated from the *t*-student with 95% confidence, from 12 experiments for $T \geq 100$ K, and applied to the complete data set. This procedure gives errors of 0.01 g cm^{-3} for densities and 0.01 for the refractive indices.

The two parameters have the same general behaviour: ρ and n_0 grow with the temperature of deposition from 20 K to 100 K. For higher temperatures both remain almost constant. The lower values correspond to 0.64 g cm^{-3} and 1.26 at 20 K, and the higher values to 0.813 g cm^{-3} and 1.344 for $T \geq 100$ K. In the density plot we have indicated the value used in the bibliography for methanol, 1 g cm^{-3} (Hudgins et al. 1993; Maté et al. 2009; Bouillouard et al. 2015; Torrie et al. 1989).

Refractive indices vary from about 1.26 to 1.34. This last value is almost constant for $T \geq 100$ K, matching the refractive index usually employed in the bibliography, 1.33 (Hudgins et al. 1993; Bouillouard et al. 2015), but disagreeing for lower temperatures of deposition (from 20 to 80 K). Our experimental refractive indices match those obtained by Isokoski et al. (2014; blue squares in Fig. 3). Both sets of experiments show increasing n_0 values from 20 K to 100 K. In both cases the slope seems to vary depending on the deposition temperature. The lowest slope corresponds to temperatures ranging from 40 to 60 K. This seems to be similar to the values reported in Cazaux et al. (2015) for water ice simulations, and would be related to the number of neighbours a molecule is interacting with.

This overall behaviour is similar to that of other molecules studied before: CO₂ (Satorre et al. 2008), NH₃ (Satorre et al. 2013), and C₂H₄ and C₂H₆ (Satorre et al. 2017). An amorphous structure is formed at low temperature that increases its order with increasing deposition temperature until a maximum density (or refractive index) is achieved, corresponding to a crystalline structure. Methanol is atypical because this maximum (at 100 K) does not imply that the crystal is formed, as indicated by the IR spectra. These spectra reveal different crystalline structures at $T \geq 110$ K, but its refractive indices and densities remain almost constant. The small variations in these values in Fig. 3 above 100 K are difficult to assign to structural changes within our error bars.

3.2. Infrared spectra

Figure 4 shows an example of the spectra obtained for methanol ices grown at eight different temperatures, from 20 K to 130 K. An adequate baseline has been subtracted. An amorphous solid is formed in deposits at temperatures equal to or below 100 K. At these temperatures the spectra have a similar shape, denoting the amorphous structure, although some differences can be appreciated. There is an overall tendency for all the bands to become narrower and to increase their maximum peak intensity values with increasing deposition temperature (see e.g. the band at 1027 cm^{-1}). Interestingly, at higher wavenumbers the broad OH stretching band at 3250 cm^{-1} presents a shoulder for ices grown between 20 and 60 K which disappears at 80 K. The band is amplified in Fig. 5a. The presence of that shoulder at 3400 cm^{-1} might be correlated with the density of the ice (see Fig. 3) the spectral signature being larger for lower ice

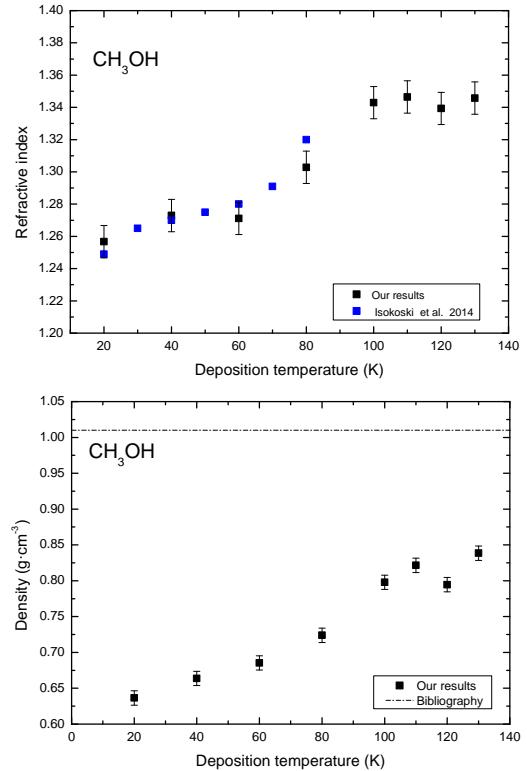


Fig. 3. Variation in density and refractive index of methanol with deposition temperature.

densities. A possible interpretation of these results, in connection with what is observed for water ice, could be the presence of pores in the amorphous methanol ice where methanol molecules are situated with weakly bonded O–H bond or even dangling bonds responsible of the larger wavenumber absorptions in the spectrum.

Between 100 K and 110 K a phase transition from amorphous to polycrystalline solid is revealed by the strong changes that are observed in the IR spectra (see Fig. 5b). However, the crystalline form obtained at 110 K changes when the ice is deposited at 120 K and above. The intermediate phase at 110 K could be a different crystalline form or maybe just a mixture of amorphous and crystalline phases. Nonetheless, taking a closer look to the spectrum of the ice at 110 K, it is observed that there are already very sharp features at 4000 cm^{-1} , 2950 cm^{-1} , and 1450 cm^{-1} , pointing to the presence of a crystalline form at 110 K. It is interesting to note here the result by Isokoski et al. (2014) that follows the interference signal of a He–Ne laser during thermal annealing of methanol ice grown at 20 K. They found an abrupt decay of the signal at about 100 K, which denotes a compaction of the ice before crystallization. This process of compaction but not crystallization agrees with what is observed in our experiments (see Figs. 3–5).

We have also followed the infrared spectral changes of an ice upon annealing at 5 K min^{-1} from 20 K to 130 K. Figure 6 shows two absorption regions chosen to illustrate the differences in the spectra due to the different degrees of crystallization obtained by direct deposition (left panel) or thermal annealing (right panel). At 110 K, 120 K, and 130 K the crystalline ice obtained by annealing does not reach the degree of crystallinity that the ices grown by direct vapour deposition have at these same temperatures.

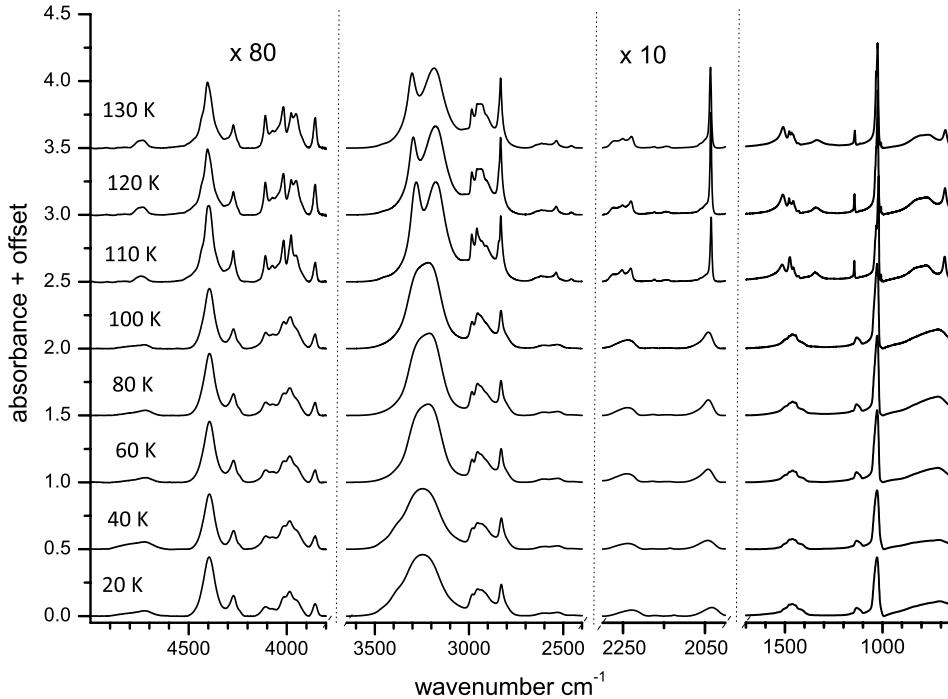


Fig. 4. Infrared spectra of methanol ices grown at the indicated temperatures. Appropriate baselines have been subtracted. The spectra correspond to layers about 990 nm thick. Spectral regions with weak absorptions have been multiplied by a factor indicated in the figure. Vertical offsets have been added to the absorbance spectra for clarity in the representation.

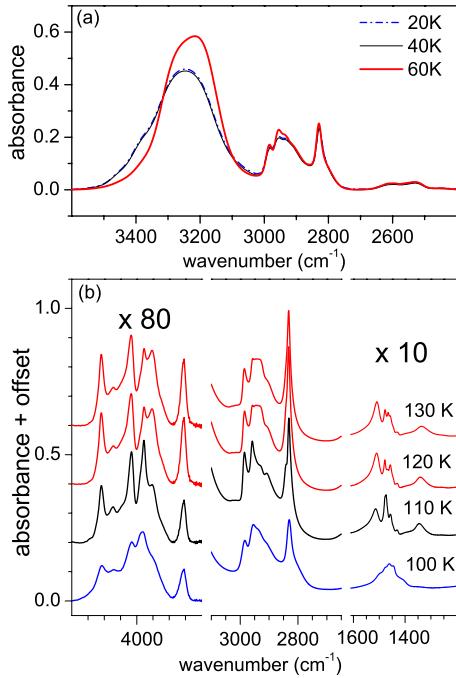


Fig. 5. Enlargement of part of the spectra presented in Fig. 4 to highlight (panel a) the shoulder at higher wavenumbers in the 3250 cm^{-1} band, and (panel b) the differences between compact-amorphous (100 K) and crystalline structures (110 K, 120 K, 130 K).

3.3. Optical constants and band strengths

The NIR and MIR optical constants of solid methanol obtained in this work by the procedure described in Sect. 2.1 are presented in Fig. 7. It is important to remember that for the determination of these magnitudes from the IR spectra it is not necessary to know the ice densities, only the thickness of the ice layers. At the end of the iterative procedure, the simulated spectra reproduce the experimental observations with a root mean square deviation

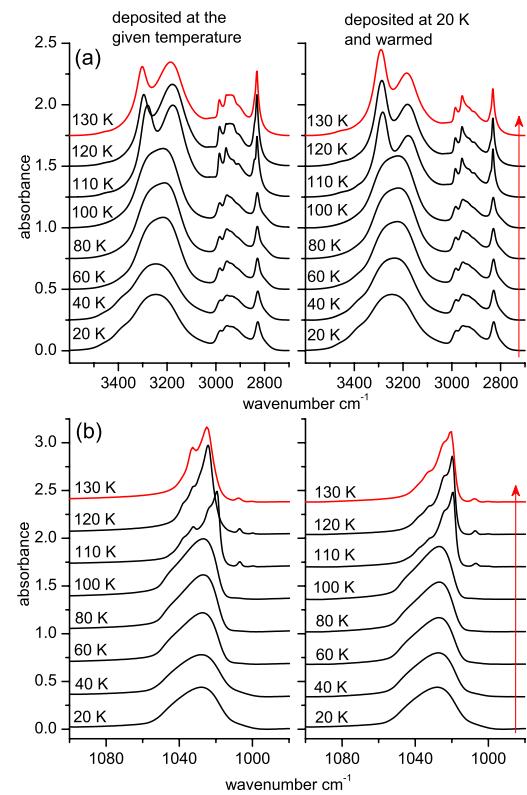


Fig. 6. Infrared spectra of methanol ices grown at the indicated temperatures (left panels) or grown at 20 K and annealed to the indicated temperatures (right panels). Appropriate baselines have been subtracted. The spectra correspond to layers about 990 nm thick. Vertical offsets have been added to the absorbance spectra for clarity in the representation.

of 0.001 in absorbance units. We estimate overall uncertainties of $\approx 1\%$ for the real component of the refractive index n . The uncertainty is also $\approx 1\%$ for the imaginary component, k , except

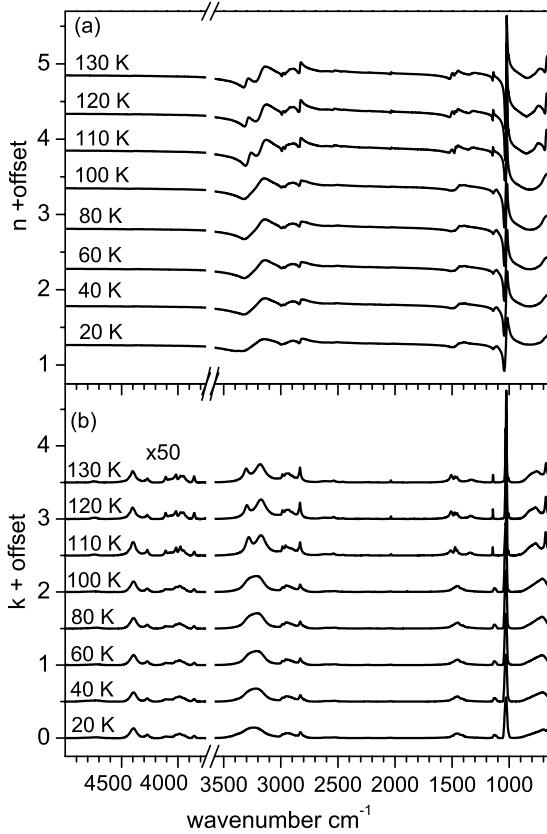


Fig. 7. Infrared real (n) and imaginary (k) part of the refractive index of methanol ices at different deposition temperatures. Vertical offsets have been added to the magnitudes for clarity in the representation. In the lower panel, k multiplied by 50 has been represented in the 3700 and 5000 cm^{-1} spectral range.

when $k < 0.01$, which induces larger errors due to the higher signal-to-noise ratio in the film-substrate transmittance.

In Tables 2 and 3 we present the NIR and MIR band strengths obtained in this work at different temperatures for amorphous and crystalline ices, respectively. Mode assignments are presented in the first two columns. In the third column peak positions are listed, although for most cases several unresolved peaks are present in the integration region, and these positions are somewhat arbitrary. The fourth column indicates the integration intervals employed. We have estimated the absolute band strengths, A , and the effective band strengths, A' , in the MIR region using the expressions in Eqs. (5) and (6), respectively. In the NIR spectral range the two magnitudes coincide since the k values were obtained after subtracting a baseline and applying the Lambert–Beer law. The ice densities needed were taken from Table 1. An average experimental error of 20% was estimated for these magnitudes, being smaller for the stronger features and larger for the weaker ones. This error is mostly due to the criteria used to choose the integration intervals.

In Fig. 8 we have represented the evolution of the band strengths with temperature for two bands. It can be seen that the band strengths do not change appreciably, within experimental error, for temperatures below 100 K. However, some modifications are observed when the ice phase changes from amorphous to crystalline above 100 K. These findings are in good agreement with those of Kerkhof et al. (1999), where no variation was found for the band strength of pure methanol ices grown at 10 K and annealed up to 100 K. In that work, higher temperatures were

not investigated for the pure species. At higher temperatures, the observed changes point to the different crystalline structures formed by deposition at 110 K, 120 K, or 130 K.

4. Discussion

To our knowledge the article Hudgins et al. (1993; hereafter HSAT) presents the most complete set of IR data of pure methanol ice published so far. In particular, they provide the only set of optical constants reported for pure solid methanol in the MIR range. These authors studied a methanol ice grown at 10 K and annealed to higher temperatures. In Fig. 9 we compare the optical constants we obtained for methanol ice at 20 K to those by HSAT at 10 K. It can be seen that the literature values are higher and they approach our data if they are scaled by a factor 0.64. This factor is close to the density we measured for methanol ice at low temperature (see Table 1).

Although the imaginary part of the refractive index, k , as a characteristic of the solid depends on its density, it can be determined without the knowledge of that magnitude. The only ice parameter needed for calculating k is the visible refractive index of the ice, n_0 , which is needed to obtain the ice layer thickness from Eq. (2). The refractive index assumed in HSAT for methanol ice was 1.33, taken from the value for liquid CH_3OH given by Weast (1972). The refractive index we have measured for methanol at 20 K in this work is 1.26. From this data, the ice film thickness considered by HSAT could be underestimated. However, it is likely that the ice grown by HSAT at 10 K will not have the same n_0 or density as the ice grown at 20 K in our work. In their experiments the gas was introduced into the vacuum chamber through a tube directed perpendicular to the cold substrate, while in our work the whole vacuum chamber was filled with a homogenous pressure of methanol gas (background deposition). Random-trajectory background deposition generates more porous ices than directional deposition (Dohnálek et al. 2003; Bossa et al. 2015). Therefore, it is expected that the methanol ice grown by HSAT at 10 K is more compact than the ice grown in the present work at 20 K. A smaller number of absorbers, related to the large porosity of our samples, will be evidenced as a decrease in the imaginary part of the refractive index. On the contrary, a more compact ice will absorb more IR radiation, which means higher k values. These considerations are in agreement with the scaling factor applied to match the two sets of k values represented in Fig. 9. Apart from the overall scaling factor, differences in the shape and relative intensities of some features of the two sets of optical constants can be seen, mostly above 100 K. In Fig. 10 the k values given by HSAT at different temperatures are compared to our data for the bands at $\approx 2833 \text{ cm}^{-1}$ ($3.53 \mu\text{m}$) and at $\approx 1027 \text{ cm}^{-1}$ ($9.74 \mu\text{m}$), two of the peaks most widely employed for methanol identification in astrophysical media. The literature data in Fig. 10 have been scaled for clarity. There is very good agreement in the shape of the two sets of constants for amorphous samples from 10 K up to 100 K (see the four lower lines in both panels). However, the differences in shape are significant between the k values reported in HSAT for crystalline ice at 120 K and the three sets of constants at 110 K, 120 K, and 130 K reported in this work (see the four upper lines in both panels). This situation shows that the structure obtained by direct deposition is more ordered than that obtained by an ice deposited at low temperature and annealed.

Different band strengths for CH_3OH ice at low temperature reported in the literature are compared with our data in Table 4. The A' estimated by HSAT assuming an ice density of 1 g cm^{-3} coincides with our data. As commented above, we believe HSAT

Table 2. Absolute (A) and effective (A') band strengths of CH_3OH amorphous ice grown at the indicated temperatures.

Label	Mode	Peak cm^{-1}	Integration interval cm^{-1}	A	A'	A	A'	$(\times 10^{18} \text{ cm molecule}^{-1})$	A	A'	A	A'	
				100 K	80 K	60 K	40 K		20 K				
ν_1	OH str. CH_3 str.	4760	4908–4649	—	0.039	—	0.053	—	0.05	—	0.092	—	0.028
		4395	4530–4201	—	0.417	—	0.494	—	0.508	—	0.477	—	0.578
		4271	4316–4201	—	0.046	—	0.053	—	0.056	—	0.053	—	0.058
		3983	4201–3886	—	0.339	—	0.31	—	0.305	—	0.337	—	0.289
		3856	3888–3821	—	0.028	—	0.029	—	0.029	—	0.031	—	0.034
ν_3, ν_9	CH ₃ str.	2950	3600–2700	153	145	159	148	155	146	156	150	151	142
ν_3, ν_9		2830	2865–2750	4.51	4.48	4.88	4.74	4.89	4.72	4.06	4.97	5.06	4.86
ν_3, ν_9		2570	2710–2400	2.92	2.75	3.18	2.95	3.11	2.93	3.09	2.99	2.93	2.74
ν_3, ν_9		2240	2300–2204	0.22	0.22	0.24	0.24	0.24	0.24	0.24	0.24	0.22	0.22
ν_3, ν_9	CH ₃ rock	2043	2105–2014	0.31	0.32	0.33	0.33	0.31	0.31	0.31	0.3	0.27	0.24
$\nu_4, \nu_5,$ ν_6, ν_{10}		1460	1655–1302	11	8.91	11.7	9.39	11.3	9.01	11	9.06	11	8.05
ν_7		1126	1178–1103	1.39	1.17	1.57	1.33	1.59	1.32	1.65	1.33	1.61	1.35
ν_8	CO str.	1027	1103–999	15.9	16.9	16.9	17.4	16.4	17	16.4	17.7	16.1	15.6
ν_{12}	Torsion	678	992–639	19	19.4	19.2	23.2	18.3	18.6	19.3	15.1	18.1	22.9

Notes. First two columns: mode assignment (Gálvez et al. 2009); Third column: wavenumber of the strongest peak in the integration interval; Fourth column: integration interval.

Table 3. Absolute (A) and effective (A') band strengths of CH_3OH crystalline ice grown at the indicated temperatures.

Label	Mode	Peak cm^{-1}	Integration interval cm^{-1}	A	A'	$(\times 10^{18} \text{ cm molecule}^{-1})$	A	A'	
				130 K	120 K		110 K		
ν_1	OH str. CH_3 str.	4742	4812–4649	—	0.031	—	0.041	0.028	
		4399	4622–4200	—	0.44	—	0.469	0.527	
		4270	4322–4224	—	0.045	—	0.049	0.058	
		4026	4168–3888	—	0.411	—	0.436	0.402	
		3858	3888–3823	—	0.049	—	0.054	0.033	
ν_1	OH str. CH_3 str.	2950	3600–2700	164	150	179	168	185	178
ν_3, ν_9		2832	2875–2785	7.03	6.51	8.15	7.62	6.24	6.08
ν_3, ν_9	CH ₃ str.	2537	2704–2413	4.21	3.9	4.56	4.25	4.5	4.47
ν_3, ν_9		2251	2302–2206	0.29	0.28	0.33	0.3	0.32	0.33
ν_3, ν_9		2036	2077–2020	0.35	0.38	0.38	0.42	0.26	0.29
ν_4, ν_5	CH ₃ bend CH_3 rock	1460	1742–1420	7.24	8.12	8.11	9.26	6.97	8.62
ν_6, ν_{10}		1336	1421–1235	2.09	1.87	2.03	1.94	1.57	1.81
ν_7	CH ₃ rock	1142	1224–1132	0.68	0.72	0.76	0.85	0.61	0.73
ν_8	CO str.	1027	1133–996	12.8	11.7	14.5	13.1	11.8	11.7
ν_{12}	torsion	720	991–650	18.9	15.1	20.3	22.4	20.7	19.2

Notes. First two columns: mode assignment; third column: wavenumber of the strongest peak in the integration interval; fourth column: integration interval.

methanol ice is denser than that grown in this work, and its density could be close to 1 g cm^{-3} . It could be the case that the actual density of their ice is lower. The agreement found could then be explained by an underestimation of their ice layer thickness values, since this would lead to a similar number of

absorbers (molecules cm^{-2}). It can be seen in the fourth column of Table 4 that correcting HSAT A' values considering the methanol ice density measured in this work, 0.64 g cm^{-3} at 20 K, leads to a large disagreement. In the last column of Table 4 the results of a recent work by Bouilloud et al. (2015) are presented.

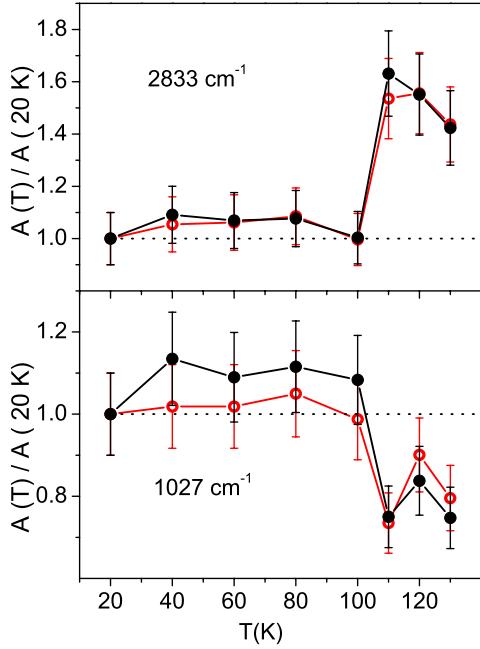


Fig. 8. Temperature dependence of methanol IR band strengths at 2833 cm^{-1} and 1027 cm^{-1} . The two sets of data represented correspond to A (open red circles) and A' (solid black circles).

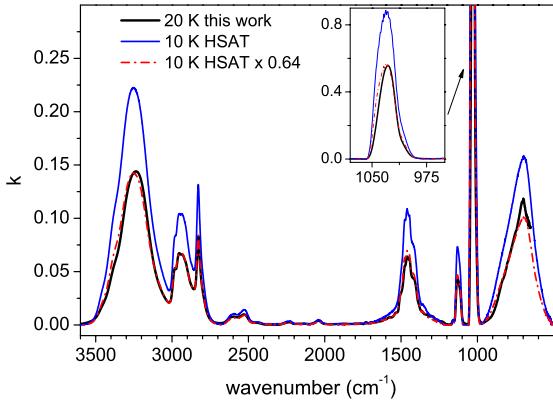


Fig. 9. Comparison of the imaginary part of the refractive index of solid methanol ice at 20 K provided in this work (black) with the values reported by Hudgins et al. (1993; blue). The scaled literature values are represented with a dashed red line.

They correspond to band strengths of methanol ice grown at 25 K by deposition through a capillary that is 10 cm away from the cold surface, with an incidence angle of 45° . This could be considered similar to background deposition and the ice could have low density, 0.64 g cm^{-3} , like those grown in the present work at 20 K. When the values presented in Bouilloud et al. (2015) are corrected assuming an ice density of 0.64 g cm^{-3} , they coincide very well with our data.

5. Conclusions

We have measured new densities, ρ , visible refractive indices at 633 nm , n_0 , IR optical constants, and IR band strengths for methanol ices grown by vapour deposition at 20 K, 40 K, 60 K, 80 K, 100 K, 110 K, 120 K, and 130 K.

Methanol-ice densities and visible refractive indices increase with the temperature of deposition in the interval 20–100 K.

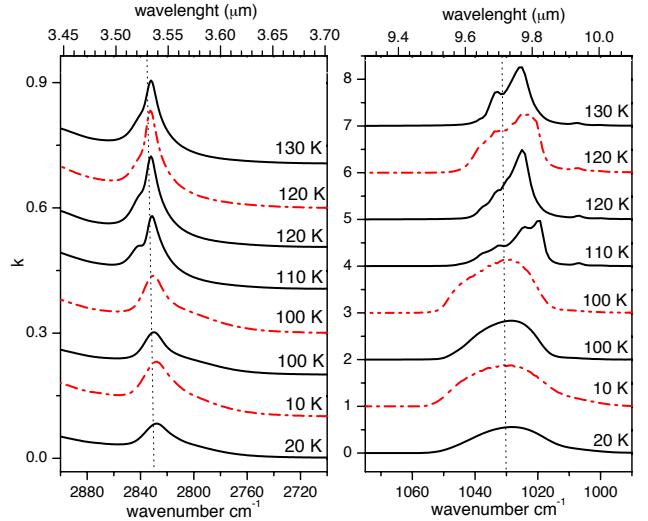


Fig. 10. Comparison between k values from this work (solid black lines) and those from Hudgins et al. (1993; dashed red lines). Literature data have been scaled for clarity in the representation. Vertical offsets have been added to the magnitudes for the same purpose.

They remain almost constant from 100 K to 130 K, which is the highest temperature investigated. The growth is sharper for the interval 60–100 K, in agreement with the n_0 values obtained by Isokoski et al. (2014) in similar accretion conditions. The refractive index does not vary between 40 and 60 K, a behaviour that resembles that obtained by Cazaux et al. (2015) for water ices.

Methanol-ice grown at 100 K presents an anomalous behaviour. It acquires an amorphous structure as dense as the crystalline form. This was revealed by looking at our density and infrared spectra measurements. This particularity has not been observed for any other molecular ice previously studied in our laboratory. For example, CO_2 , C_2H_6 , C_2H_4 , and NH_3 ices have a lower density in amorphous phases than in crystalline phases. This finding coincides with the sudden ice compaction that takes place at 100 K for CH_3OH ice when deposited at 20 K and warmed up as shown in Isokoski et al. (2014). This compaction is not associated with crystallization either, as shown by infrared spectroscopy when annealing CH_3OH to 100 K.

Infrared spectra of crystalline methanol obtained by annealing ices grown at 20 K to 110 K, 120 K, and 130 K show that they are different to the crystalline ices obtained by direct vapour deposition at these temperatures. Optical constants in the NIR and MIR spectral ranges, from 5000 cm^{-1} to 650 cm^{-1} , are provided for amorphous and crystalline methanol ices grown at eight temperatures from 20 K to 130 K. These optical constants differ from the Hudgins et al. (1993) data, due to the different ice-generation conditions which lead to different morphologies and phases.

We also present methanol band strengths in the NIR and MIR spectral ranges at the same eight temperatures mentioned above. Some band strengths vary for crystallized ices up to 30%. This temperature dependence must be considered to properly estimate the amount of methanol in different astrophysical environments.

For low temperature band strengths, we have compared our results with previous literature data collected by Bouilloud et al. (2015). We found that at low temperatures (10 K–25 K), methanol band strengths do not depend appreciably on the ice morphology, but only on the number of absorbers. Therefore, all the literature data agree if the adequate ice densities, which depend on the generation conditions, are taken into account. Our

Table 4. Absolute band strengths ($\times 10^{18}$ cm molecule $^{-1}$).

Peak (cm $^{-1}$)	This work $\rho = 0.64$ (g cm $^{-3}$) 20 K	Hudgins et al. (1993) $\rho = 1$ (g cm $^{-3}$) 10 K	Hudgins et al. (1993) $\rho = 0.64$ (g cm $^{-3}$) 10 K	Bouilloud et al. (2015) $\rho = 0.64$ (g cm $^{-3}$) 25 K
2900	151	160	250	159
2830	5.1	5.3	8.3	—
2558	2.9	2.8	4.4	2.8
1461	11	12	19	10
1126	1.6	1.8	2.8	2.2
1027	16	18	28	17

measurements allowed us to explain discrepancies of about 40% in some band strengths previously reported by Bouilloud et al. (2015).

The spectral shape of the different features in the infrared spectra of methanol ices are distinctive and can help to trace the phase and temperature of methanol ices detected in the ISM or in different solar system objects, and will give information about the formation history of the ices observed.

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