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# List of Abbreviations

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<b>BDC</b>	Benzene-1,4-dicarboxylate
<b>BTC</b>	Benzene-1,3,5-tricarboxylic acid
<b>BO</b>	Born–Oppenheimer
<b>CMM</b>	Center for Molecular Modeling
<b>CN</b>	Coordination number
<b>CV</b>	Collective variable
<b>DFT</b>	Density functional theory
<b>FES</b>	Free energy surface
<b>FTIR</b>	Fourier-transform infrared
<b>GCMC</b>	Grand canonical Monte Carlo
<b>HF</b>	Hartree–Fock
<b>HPC</b>	High–performance computing
<b>IR</b>	Infrared
<b>IUPAC</b>	International Union of Pure and Applied Chemistry
<b>LFEP</b>	Lowest free energy path
<b>M–L</b>	Metal–linker coordination
<b>MD</b>	Molecular dynamics
<b>MOF</b>	Metal–organic framework
<b>MPV</b>	Meerwein–Ponndorf–Verley
<b>MTD</b>	Metadynamics
<b>MTK</b>	Martyna–Tobias–Klein
<b>MTO</b>	Methanol–to–olefins
<b>NEB</b>	Nudged elastic band
<b>PCP</b>	Porous coordination polymer
<b>PHVA</b>	Partial Hessian Vibrational Analysis
<b>PES</b>	Potential energy surface
<b>PXRD</b>	Powder X–Ray diffraction
<b>QM</b>	Quantum mechanics
<b>SBU</b>	Secondary building unit
<b>SCF</b>	Self–consistent field
<b>SXRD</b>	Single–crystal X–ray diffraction
<b>TPS</b>	Transition path sampling
<b>TS</b>	Transition state
<b>TST</b>	Transition state theory
<b>UiO</b>	Universitetet i Oslo (University of Oslo)
<b>WHAM</b>	Weighted histogram analysis method
<b>ZPE</b>	Zero point energy



# List of Symbols

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$\ddagger$	Transition state
$\Delta E^\ddagger$	Electronic energy barrier
$\Delta H^\ddagger$	Enthalpy barrier
$\Delta G^\ddagger$	Free energy barrier
$B$	Bulk modulus
$E$	Electronic energy
$E_{ZPE}$	Zero-point vibrational energy
$F$	Helmholtz free energy
$G$	Gibbs free energy
$H$	Enthalpy
$h$	Planck constant
$k$	Rate coefficient unimolecular reaction
$k_B$	Boltzmann constant
$N$	Number of particles
$N_A$	Avogadro constant
$N_{dof}$	Number of degrees of freedom
$p$	Pressure
$P$	Product
$q_X$	Partition function of species $X$
$q_{vib}$	Vibrational partition function
$R$	Reactant
$S$	Entropy
$T$	Temperature
$U$	Internal energy
$Q$	Partition function
$\nu_i$	Vibrational frequency

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# Samenvatting

Summary in Dutch [here](#)



# Summary

Summary in English here



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## **Part I**

**Towards modeling catalytic  
processes in UiO-66  
metal-organic framework at  
operating conditions**



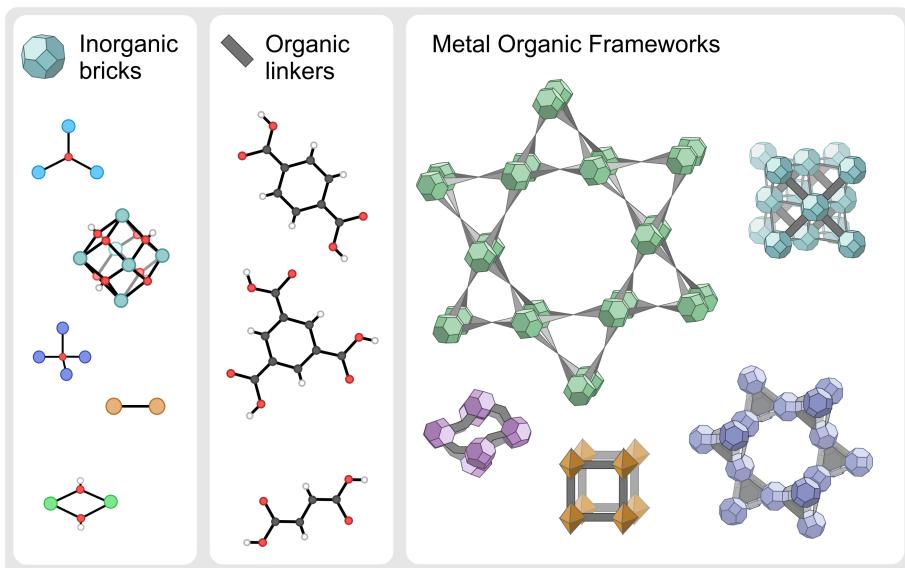
# 1

## Introduction

In order for the wheel to turn, for life to be lived, impurities are needed, and the impurities of impurities in the soil, too, as is known, if it is to be fertile. – Primo Levi –

Catalysis, from the greek for “down” and “loosen”, is at the heart of almost every industrially relevant chemical process. A catalyst intervenes in the reaction mechanism, allowing chemical species to meet each other and react with a specific mechanism, without being consumed. Not only can it increase the rate of a given reaction, allowing processes that would not happen spontaneously, but it can also steer the reaction to specific end products, allowing selectivity. Catalysts can be divided in homogeneous and heterogeneous, according to the phase where they are located with respect to the reactants. Homogeneous catalysts share the same phase with the reactants, and can be for instance molecules dissolved in a solvent, such as with organometallic compounds. These type of catalysts are used in many industrial processes, but have the drawback of being difficult to separate from the products. From a modeling and experimental point of view, they are the easiest to study, and it is easier to hypothesize reaction mechanisms from experimental data. Heterogeneous catalysts, on the other hand, are located in a different phase with respect to the reactants, and for this reason they have the advantage of being easily separable from the reaction products. Even if the term does not refer to a specific phase, often heterogeneous catalysts is synonymous for solid catalysts. In order to exert their function, these materials must possess specific active sites

that are easy to reach for the reactants, where these can adsorb, react, desorb, and ultimately diffuse back in the bulk to leave space for a new cycle. For these processes to occur, the number of active sites and the area of contact between the two phases (or surface area) must be sufficiently high. In this sense, nanoporous materials, which possess pores with diameter of  $\geq 100$  nm, have drawn a lot of attention for their enormous catalytic potential. Their pore structure provides them with an exceptionally high surface area, facilitating diffusion of reactants inside the material, and allowing shape selectivity to give specific products. In general, the study of solid heterogeneous catalysts is more difficult than their homogeneous counterparts due to their higher complexity. It is not always clear, in fact, where exactly the active sites may be located in the material, and how they interact with molecules. Moreover, these materials are far from perfect, and it is often the imperfections they contain that give them their catalytic properties. In this sense, computational design offers a perfect platform not only to understand the behavior of heterogeneous catalysts, but also to understand structure-activity relationship and design structures to target specific applications. Example of industrially relevant nanoporous heterogeneous catalysts are zeolites, which have been the workhorses of petrochemistry for the last decades. In this work of thesis, we focus on a new class of nanoporous materials which combine the crystalline nanoporous structure of zeolites with the versatility in synthesis of metal complexes and for this reason has shown great potential for catalysis.



**Figure 1.1:** Schematic representation of the building block design in Metal-Organic Frameworks

## 1.1 Metal Organic Frameworks

Metal organic frameworks (MOFs) are one of the most intriguing class of materials of current science. These promising materials, first called 'porous coordination polymers' (PCPs) were discovered in the late 50's, but only at the end of the century with the works of Robson [1,2], Kitagawa [3,4], Yaghi [5] and Ferey [6], the scientific community started to understand their full potential. If at first these materials were seen as a laboratory curiosity, with the focus on the discovery of new structures, in the last few decades, the field has seen an incredible explosion in scientific and industrial interest, with new applications being continuously explored [7]. MOFs are hybrid nanoporous materials that are composed by metal or metal-oxo clusters connected by multtopic organic linkers, to form multidimensional pore structures. Compared to the already established zeolites, MOFs can be constructed without templating agents, with a far greater number of metals and with an exceptional structural diversity. In fact, their particular building block design (Fig. 1.1), that makes use of secondary building units (SBUs), allows the creation of an almost infinite number of crystalline structures with different topology and chemical composition. In principle, the nature of the SBUs and their association can be finely tuned [8], allowing control on properties such as pore shape and size, functionalization, surface area, or response to chemical and physical stimuli [9, 10]. Moreover, multiple physical or chemical functionalities can be integrated in the crystals at the same time [11]. This tunability, along

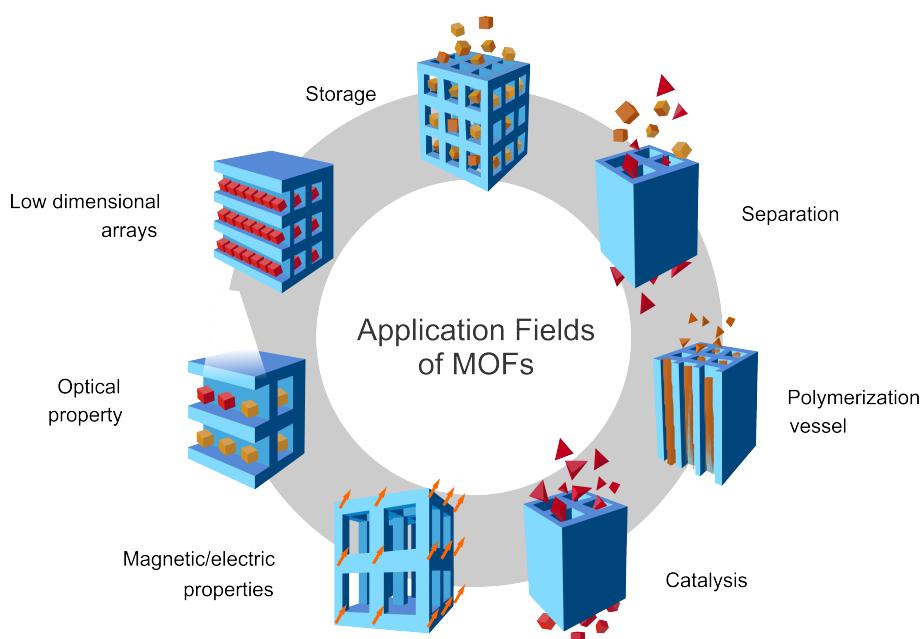


Figure 1.2: Some of the applications of MOFs

with their high crystallinity, metal content and porosity, allows their application in different industrially relevant fields, such as catalysis, gas storage and separation, drug delivery or sensing. More specific applications are being further explored, such as warfare agents decomposition, magnetic applications, or membrane separation. After the discovery of MOFs, with their tunability and ease in functionalization, the study of their application in catalysis followed naturally, and was one of the earliest documented applications [12]. Their high porosity, in particular, allows chemical species to diffuse in the pores and access the active sites, and the high metal content offers the possibility to have many guest interactions sites. Exploiting this property, a plethora of MOFs has been synthesized possessing unsaturated metal sites within the pores which offer different Lewis acidity. In this sense, provided the structures are stable at reaction conditions (i.e. no leaking is observed), MOFs can be considered true heterogeneous catalysts, where the active sites are inherent part of the framework. We can differentiate between three generations of porous systems, following the nomenclature proposed by Kitagawa [13]. First generation MOFs are defined as having a guest–molecules supported pore system that collapses when these are evacuated, and for this reason found very limited use for practical applications. Those of second generation are more robust and have permanent porosity that is retained even in absence of guest molecules. These materials show high potential for catalysis and other applications and are the main object of this dissertation. Finally, third generation MOFs are characterized by flexible pores that can reversibly change shape with the presence of guest molecules, or upon certain stimuli, such as temperature or pressure. Matsuka et al. <sup>1</sup> identified five types of response mechanisms in MOFs, denoted as shrinking, expanding, reshaping, swelling and gate opening or closing. A perspective on the types of stimuli that can induce responsive in MOFs has been reported by Coudert <sup>2</sup>.

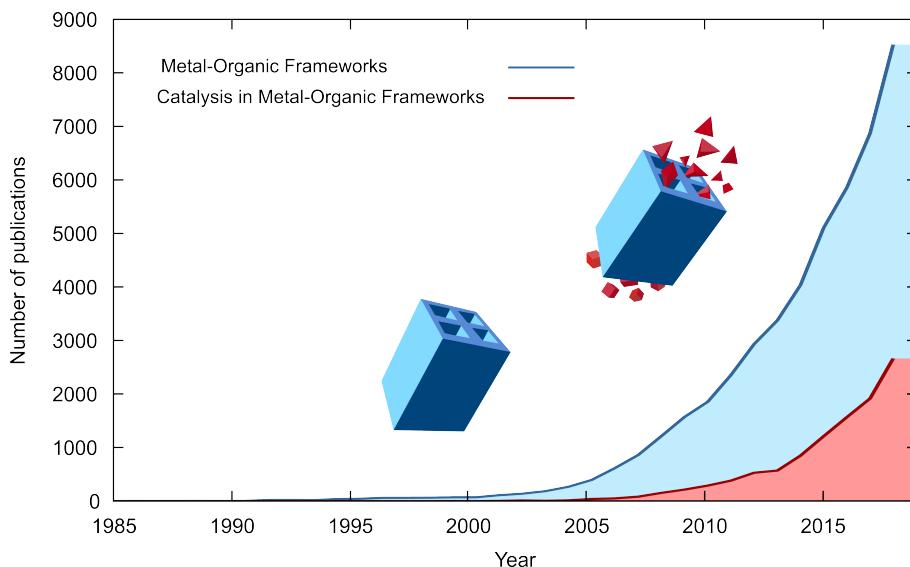
### 1.1.1 Post-synthetic modification

One of the most intriguing concepts is the isoreticular synthesis, by which inorganic or organic SBUs can be replaced by topologically identical (or similar) building blocks, giving rise to whole MOF families which derive from a specific precursor and span a range of pore size and functionality. For instance, the pore size can be significantly increased up to the mesoporous range by using longer isoreticular linkers, such as in the IRMOF series [14]. When it is not possible to introduce functionalities with direct synthesis, post synthetic modification [15], which makes use of the building block design, has become a well established procedure that allows the preparation of MOF materials with specific functionality. Via this technique, it is possible to modify the crystal after the synthesis, allowing to finely tune the properties of the material. Post synthetic modifications include encapsulation of guest molecules or nanoparticles in the pores, modifications of the linkers without breaking the metal–ligand bond, or post synthetic exchange

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<sup>1</sup>Matsuka jacs 2004

<sup>2</sup>chem mater 2015 <https://pubs.acs.org/doi/full/10.1021/acs.chemmater.5b00046>

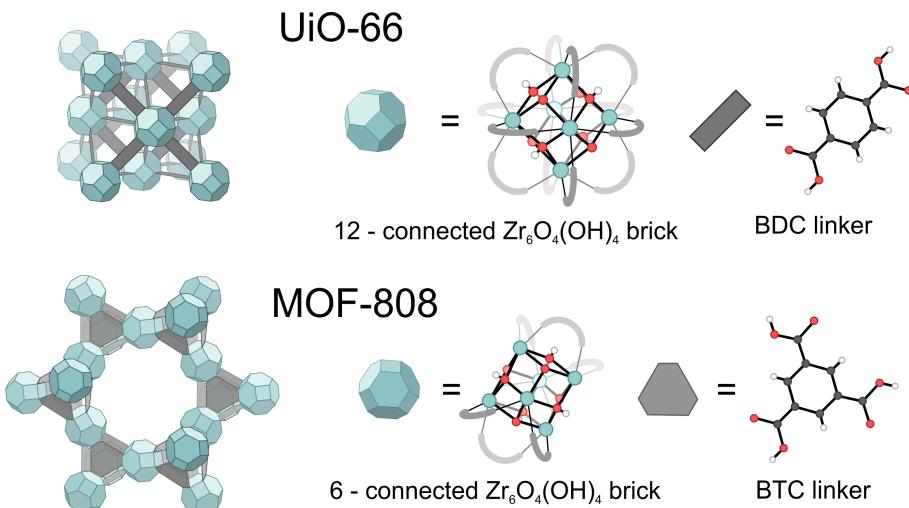


**Figure 1.3:** Number of publications on MOFs and catalysis in MOFs in the past 25 years showing the importance of catalysis in MOFs.(source: WebofScience)

(PSE) of linkers and metals, where building blocks are dynamically exchanged. PSE can also involve terminal ligands, or ligands adsorbed on the bricks that do not function as linkers, such as modulators. This way, building blocks can be exchanged, but also eliminated to create vacancies if the stability of the material allows it. Defect-containing MOFs have become an active field in MOF research, as defect sites can play a key role in the performance of the material. Post synthetic modification has been used as a strategy to efficiently introduce defective sites in MOFs.

### 1.1.2 Zr-based MOFs

In general, to function at operating conditions, materials need to retain their thermal, mechanical and chemical stability at those conditions. For instance, mechanical stability is needed when compressing MOF in pellets or other shapes for industrial processes [16], while chemical stability is crucial for any application, such as drug delivery, molecular separation, or catalysis [17]. Catalysis often also requires LISthermal stability, as the materials must be able to resist harsher conditions for certain processes such as in petrochemistry. However, the metal-ligand coordination bond that makes MOFs so tunable is also regarded as one of their main drawbacks [18–20], as it is responsible for the lower structural stability when compared to already established nanoporous catalysts such as zeolites. For example, the first synthesized MOFs such as  $\text{Cu}_2^+$  trimesate HKUST-1, or MOF-5, composed by  $\text{Zn}_2^+$  clusters and BDC linkers were degraded by water even at



**Figure 1.4:** Structure of UiO-66 (top) and MOF-808 (bottom)

mild conditions [21–24]. Recently, a class of robust MOFs have been synthesized showing an unprecedented stability [25]. Zr-based MOFs [26] which exploit the robustness of the Zr–O bond, show an outstanding stability and are at present time one of the most studied classes of MOFs. Zirconium is an ubiquitous metal that is present in biological systems and has low toxicity, as well as limited cost. This makes these materials particularly promising for applications in catalysis, gas sorption, and drug delivery. The vast majority of these materials is characterized by Zr(IV) atoms in a high coordination state, which interact strongly with the oxygens of carboxylate linkers of various topology. These MOFs are characterized by a Zr<sub>6</sub>O<sub>4</sub>(OH)<sub>4</sub> cluster in which each of the 6 zirconium atoms is connected to 4 oxygen atoms (two  $\mu_3$ -OH and two  $\mu_3$ -O), each of which is connected to three zirconium atoms, forming a polyhedron (Fig. 1.4). Each zirconium atom can in turn form 4 other bonds with ligands, accommodating up to 24 metal–ligand bonds per cluster. Each zirconium atom can therefore form a total of 8 coordination bonds in a square–antiprismatic geometry, yielding a rich range of possible structures that can be synthesized. The dual Lewis acid/base nature of the Zr–carboxylate bonds, along with the high metal oxidation state, gives rise to strong interactions between the SBUs, thus allowing processes such as PSM without compromising the stability of the structures. Different structures can be constructed with this SBU and different linker topology and connectivity ranging from 12, like in UiO-66, to 6 as in MOF-808 (see Fig. 1.4). These materials can further be functionalized by post synthetic modification. The stability of Zr-based MOFs is related to the connectivity between the inorganic and organic SBUs, as well as the number of zirconium–ligand bonds. However, open metal sites in these materials are present only when the connectivity of the linkers is less than 12. Zirconium atoms that remain undercoordinated in these materials are Lewis acid

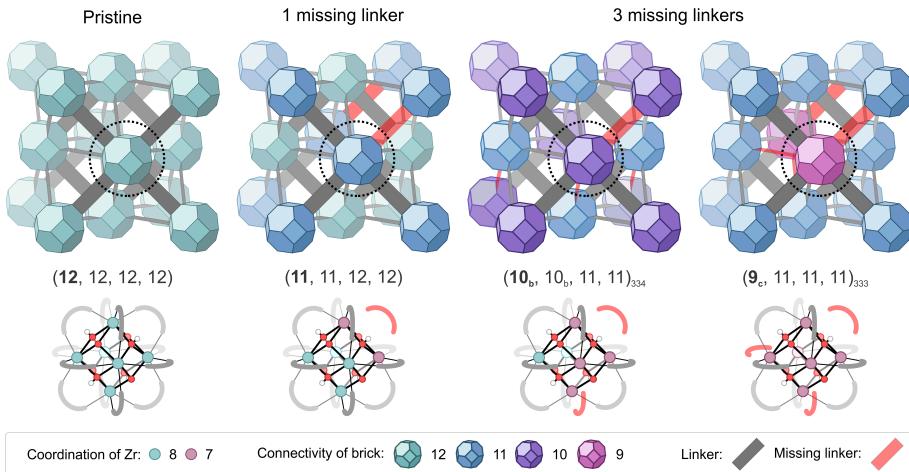
sites where reactants can adsorb and that can function as catalytic centers.

### 1.1.3 UiO-66

The precursor of the whole class of Zr-based MOFs is the Zr–therephthalate based UiO-66, which was first synthesized at the Universiteit i Oslo by Lillerud and coworkers [27]. This material is characterized by an extremely high connectivity that gives rise to an exceptional structural stability, which makes UiO-66 one of the most widely investigated MOFs up to date. In this material, each  $Zr_6$  SBU is connected to 12 terephthalate (or benzenedicarboxylate (BDC)) linkers forming a cubic close packed structure with a space group  $F\bar{m}3m$ , No. 225. In this structure there are two different cavities of octahedral and tetrahedral shape, with window sizes of 10 Å and 25 Å, respectively. Each octahedral cage shares triangular windows with eight tetrahedral cages. This results in an extremely robust material which is stable up to 375°C and in most solvents and pH conditions. Moreover, the  $Zr_6O_4(OH)_4$  bricks be reversibly dehydrated upon thermal treatment at temperatures between 523 and 573 K. Up to two water molecules can be formed this way, yielding a  $Zr_6O_6$  brick, where the zirconium atoms have a coordination of 7 [28]. A whole family of isoreticular MOFs can be derived from UiO-66 by using linkers of different size, spanning from fumaric acid [29] up to terphenyldicarboxylic acid [30]. Interpenetrated MOFs with UiO-66 topology have been also reported if longer linkers are used [31]. Moreover, different functional groups can be appended to the phenyl rings, such as bromo, amino, nitro, or naphthalene. Garibay and Cohen showed that UiO-66-NH<sub>2</sub> can be further modified to yield new functionalized frameworks [32]. Also the inorganic SBUs can be modified, for instance introducing titanium or hafnium [33]. The exceptional thermal and chemical stability of UiO-66 allows all these modifications of the structure, and for this reason, this material is often considered a perfect MOF archetype.

## Defects

In theory, the perfect crystalline UiO-66 structure, where every zirconium atom is 8-fold coordinated, does not possess undercoordinated Lewis acid sites available for catalysis. However, it became clear that the synthesized UiO-66 showed a behavior that deviated from the theory and pointed towards the presence of defective sites, such as symmetry-forbidden reflections in the PXRD pattern, metal-linker ration obtained by thermogravimetric analysis (TGA), higher than expected surface area, appearance of O-H stretching bands in the FTIR spectrum etc. [28, 34]. It has been generally accepted that the material contains defects in the form of missing linkers or clusters, and that these are not only naturally occurring during synthesis, but their number can easily be tuned by adapting the synthesis conditions, such as temperature and presence of modulator [35, 36]. Defects can influence the material properties by a large extent. The beneficial role of defects in UiO-66 has been explored in many applications such as gas storage and separation [35, 37],



**Figure 1.5:** Defective UiO–66 unit cells with different number of missing linkers

sensing [38], drug delivery [39] and catalysis [40, 41]. Following the classification proposed by Sholl, these defective sites in MOFs can be either point vacancies (such as missing linkers or clusters) or extended ones such as in the case of a surface [42] or a region in the material. At low defect concentration, a random distribution of isolated point defects can be found, whereas at higher concentrations clustering could occur, if the presence of a vacancy is influenced by other vacancies in proximity. This complexity makes the study of defective sites and disorder a current challenge in MOF research. In particular, the control of their effect on heterogeneous catalysis is heavily limited if their location, type and dispersion are unknown. This randomness can be partially limited by judicious defect engineering during and post-synthesis which is in turn dependent on their understanding at the molecular level. In order to identify and characterize defects in MOFs, it is particularly important to obtain atomic scale resolution on the defect sites and understand how they impact certain properties. The physical properties of defective UiO–66 differ according to the number of defects and their location, as has been extensively studied both theoretically and experimentally. A decrease in the connectivity in the structure will naturally lead to a decrease in stability of the material, but the extremely high connectivity of UiO–66 allows the presence of numerous missing linkers or cluster without loss of crystallinity. In this context, Rogge et al. report the influence of all possible configurations of one to two linker vacancies on the stability of UiO–66, showing that the equilibrium volume is not affected by such vacancies [43]. However properties such as bulk modulus and loss-of-crystallinity pressure, which in turn influence the stability, are affected by the relative missing linker configuration. De Vos et al. [44] reports the electronic properties for all possible configuration of defective UiO–66 with up to three missing linkers, showing that some configurations are energetically more stable than others, and that the number and position of missing linkers affect the band gap of the

material. From these studies performed small unit cells it is already clear that the number of possible defective configurations can be extremely high, and it is still unclear to what extent such point defects are disordered in the material. A regular distribution of point defects that involve missing clusters can lead to different phases in the material, from fcu (non-defective), to bcu (missing linkers with open channels), reo (missing clusters) to scu (missing linkers and missing clusters). Recently, Cliffe et al. [45] showed in a combined theoretical-experimental work that missing clusters on UiO-66(Hf) were correlated and formed nanodomains in the material, characterized by a reo phase. Missing clusters could in principle increase the catalytic activity of the material more than missing linkers due to the larger pore size.

The number of defects can be measured by ..

The role of modulator is critical

### Active sites for catalysis on UiO-66

One of the breakthroughs of MOF research was the discovery that UiO-66 could contain a high number of open metal sites. As reported earlier, active sites for catalysis in UiO-66 and other Zr-MOFs are present when the zirconium connectivity is decreased from its equilibrium value of 8. A first way to obtain coordinatively unsaturated zirconium sites is by creation of defects, in the form of missing linkers or clusters. The inherent vacancies in UiO-66 bring unsaturated zirconium Lewis acid sites [34, 35, 40, 46, 47] and at the same time enable the reactants accessibility to the sites, increasing the pore size. The synthesis of defective UiO-66 can be performed via modulators such as formic acid or trifluoroacetic acid (TFA), that are competing with BDC linkers in binding to the inorganic SBU. Vermoortele et al. first showed in a dual computational–experimental study that the Lewis catalyzed cyclization of citronellal can only be done on UiO-66 in case of missing linkers [48]. For Meerwein reduction, another Lewis catalyzed reaction, the catalytic activity of UiO-66 could be significantly increased by making use of TFA, that introduced a large number of linker vacancies. Additionally, the non-modulated material that contained only a small number of defects showed nearly no catalytic activity [40]. Such relationship between catalytic activity and amount of defects has been reported for different Lewis catalyzed reactions giving evidence that the catalytic centers are located on the defective sites. The molecular characterization of the defective sites on UiO-66 has been an ongoing research interest in MOF literature. Lillerud and coworkers [49] reported that when linker vacancies were present, XRD characterization of the material showed two type of Zr-bonded oxygen atoms. The first type was BDC oxygen, and the other was associated to water species coordinated to the zirconium on the defect site. Still from XRD measurements, the group of Yaghi <sup>3</sup> proposed that the two Zr-bonded oxygens belonged to physisorbed water molecules and that the charge compensating species was a hydroxyl group. Insight from molecular

<sup>3</sup>cite Trickett

simulations helped to shed light into the characterization of the active sites at the molecular level. Ling and Slater<sup>4</sup> performed MD simulations starting from such structure and showed a progressive stabilization towards a structure where the two adjacent zirconium atoms were coordinated to a physisorbed water molecule and a hydroxyl group, as shown in Fig. 1.6. These sites has been also confirmed by a comprehensive computational study of different adsorption possibilities of up to three water molecules on defective UiO–66 performed by Vandichel et al.<sup>5</sup> and by simulations performed in this work of thesis. Such species are strongly interacting with the zirconium atoms and are sources of Bronsted sites that can play an active role in catalytic processes, as will be shown in **PAPER I** and can in principle be considered as part of the catalyst. In general, such sites play an active role in processes involving proton transfers. Serre and coworkers report that a hydrogen bonded network that spans the octahedral and tetrahedral cages is responsible for the high proton conductivity showed by UiO-66 at high temperatures<sup>6</sup>. Kitagawa et al. showed the positive role of defects in proton conductivity in the UiO-66 material, as defective sites provide both larger pores and zirconium-bonded water species that are proton donors<sup>7</sup>. Farha and coworkers make use of the acidic protons belonging to physisorbed water molecule on the defective site for the quantification of defects, by means of potentiometric titration<sup>8</sup>. The chemistry associated with the hydroxyl group face topology changes on the inorganic SBU has been investigated by Yang et al.<sup>9</sup>. The interaction between defect-coordinating water molecules and zirconium atoms is strong, and UiO–66 can be partially hydrolyzed [50] while retaining its stability, as will be further discussed in this dissertation. The defect-coordinating species can be removed by thermal activation at T>470 K<sup>10</sup>. The more loosely bonded physisorbed water is the first to be removed, followed by the chemisorbed water. The dehydrated active site obtained this way is missing a  $\mu_3$ -OH proton and has two 7-fold coordinated zirconium open metal sites as shown in Fig. 1.6. The second activation process by which open metal sites can be created on UiO–66 is the reversible dehydration of the Zr<sub>6</sub>O<sub>4</sub>(OH)<sub>4</sub> cornerstone to obtain Zr<sub>6</sub>O<sub>6</sub> by removal of up to two water molecules performed at a temperature range between 523 and 573 K. The 7-fold coordinated zirconium atoms are Lewis acid sites, but the large pores created by linker vacancies are missing, and the active sites are less accessible than in the defective material.

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<sup>4</sup>cite ling and slater

<sup>5</sup>cite vandichel 2016

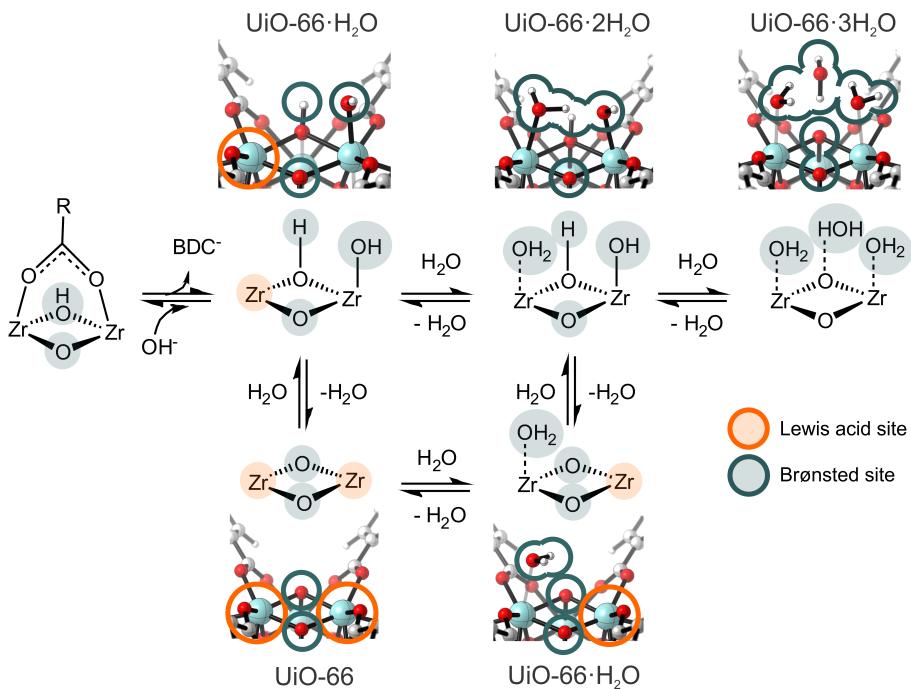
<sup>6</sup><https://onlinelibrary.wiley.com/doi/full/10.1002/anie.201510855>

<sup>7</sup><https://pubs.acs.org/doi/abs/10.1021/acs.chemmater.5b00665>

<sup>8</sup>cite Farha paper potentiometric titration

<sup>9</sup>(Yang, Bernales et al. 2016)

<sup>10</sup>check



**Figure 1.6:** Lewis and Brønsted site when a linker is removed from the  $\text{UiO-66}$  zirconium brick

## Functionalization

As anticipated earlier, the nature of the active site can also be influenced by linker functionalization<sup>11</sup>. The presence of electron-donating substituents showed an increase in catalytic activity for jasminaldehyde condensation<sup>12</sup>. The effect of other functional groups was subsequently studied by the same authors for Lewis catalyzed citronellal cyclization [48], who report a positive effect of electron-withdrawing groups. The beneficial role of amino groups was reported as well by Timofeeva et al.<sup>13</sup> for acetylation of benzaldehyde, as well as by Cirujano et al.<sup>14</sup> for esterification. The electron donating effect of amino groups would in principle lead to a decrease in the Lewis acidity of the defective zirconium atoms, lowering the reaction rate. Therefore, a change in mechanism in the presence of BDC-NH<sub>2</sub> was hypothesized. However, Hajek et al. further studied aldol condensation in a computational work, and reported a beneficial, but passive role of amino groups, which is further confirmed in **PAPER I** for Fischer esterification.

### 1.1.4 Active sites on MOF-808

MOF-808, formed by the inorganic Zr<sub>6</sub>O<sub>4</sub>(OH)<sub>4</sub> cornerstone and BTC linkers, represent the least connected MOF in the Zr-MOF family<sup>15</sup>, and its brick can be regarded as an extreme case of defective UiO-66. The rigid BTC tritopic linkers stabilize the structure and yield exceptionally wide channels which are required for the diffusion of substrates. The octahedral crystals contain 6 BTC linkers per inorganic node and are characterized by a *spn* topology. In the material, each zirconium atom is bonded to four oxygens from the inorganic brick and two oxygens from the BTC linkers. The other two bonds that are needed to reach the total coordination of 8 are provided by species present in solution that do not contribute to the framework connectivity. In the as-synthesized material, the zirconium atoms are capped with the relative mobile formate modulator. The material at this stage does not contain catalytically active sites and must be activated by post-synthetic treatment. By hot filtration, the formate groups can be eliminated and replaced by water and hydroxyl groups. In the catalytically active material, each brick is connected to six BTC linkers, six water molecules and six hydroxyl groups, and possesses a high number of potential Brønsted sites. Klet et al. report the presence of four<sup>16</sup> different types of protons in the material<sup>17</sup>. Lewis acid sites can be created by thermal activation, similarly as in UiO-66. In this case, hydroxyl groups remain connected to the zirconium atoms and the material possesses both Lewis and Brønsted sites. The hydroxyl groups could also be used as anchors to

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<sup>11</sup>citeKandiah2010a, Kandiah2010b. Kim citeKim2012

<sup>12</sup>cite vermoortele 2011

<sup>13</sup>cite Timofeeva2014

<sup>14</sup>cite Cirujano2015

<sup>15</sup>cite Furukawa 2014

<sup>16</sup>check

<sup>17</sup>cite Klet 2016

incorporate new features, for instance by molecular docking of a single-site catalyst to immobilize it on a MOF-scaffold. MOF-808 is a promising catalyst that is still less studied than UiO-66, but has a lot of potential for possible applications. For instance, it shows a higher catalytic activity for certain Lewis catalyzed reactions, such as Meerwein-Ponndorf-Verley (MPV) reduction<sup>18</sup>. Moreover, MOF-808 represents the first evidence of a superacid in MOFs, showing superacidity after post-synthetic treatment with sulfuric acid<sup>19</sup>.

## 1.2 Outline and goal of the thesis

In this chapter, the characteristics of MOFs that make them appealing as catalysts have been shown, as well as the current challenges that MOF research is facing. The properties of catalytically active sites on defective UiO-66 and MOF-808 were studied by means of molecular modeling techniques. Molecular modeling can give fundamental insight into the properties of MOFs, and at present time a plethora of different techniques are available for this purpose. To model processes in MOFs, knowledge of both power and limitations of these methods and experimental insight into the scientific question are required to find the appropriate combination of techniques.

This PhD thesis is organized as follows:

- In Chapter 2, a theoretical overview of the state of the art modeling techniques is given. Particular attention is drawn on how these techniques can be applied in the case of MOF materials to obtain insight into structural and catalytic properties at operating conditions.
- In Chapter 3, the main results of this PhD thesis are summarized. The links between theory and experiment are highlighted throughout the chapter. All results are the result of fruitful collaborations and have been published in international peer-reviewed journals.
- In Chapter 4, the main conclusions of this thesis work, as well as perspectives on future research are given.

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<sup>18</sup>cite Plessers Catalysts, 2016, Mautschke Catalysis Science., 2018

<sup>19</sup>cite <https://pubs.acs.org/doi/10.1021/ja507119n>



# 2

## Modeling metal organic frameworks

The understanding of catalytic processes in MOFs is a very challenging task. MOFs are materials of complex nature, and reactive processes in these materials are elusive and difficult to track on a purely experimental basis. Molecular simulations offer an alternative approach that starts from the construction of models that can explain, complement, and predict experiments. With growing computational power, computational models can aim at giving a more and more accurate description of materials at operating conditions, narrowing the gap between theory and experiment. Often, the structural properties and chemical transformations that take place on the active sites need to be investigated using a combination of multiple computational techniques, that allow to tackle the problem from different points of view. In this chapter, an overview of the different computational methods that can be used to study reactive processes in MOFs will be given.

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### 2.1 Framework topology

A crucial decision when performing simulations lies in the choice of the model system, and what should be included in it. When choosing a model to represent the system under study, there is always a fine balance between accuracy and computational cost. On the one hand, it is crucial to use a computational model that captures all the relevant properties of the material and mimics the experimental

structure as close as possible. On the other, it is often convenient to approximate and neglect certain properties in favor of a larger scale description of the processes. The focus in this work are the active sites that can be used for catalysis, therefore an accurate electronic description of this region of the material is imperative. Nevertheless, the activity of these sites for chemical reactions can also be influenced by other factors, such as the pore size or functionalization. Therefore, to describe active sites in MOFs and other nanoporous materials, which can have rather large unit cells and non-periodic structural defects, the first question that needs to be asked is how to account for periodicity. Two conceptually different approaches, which are described below, can be used.

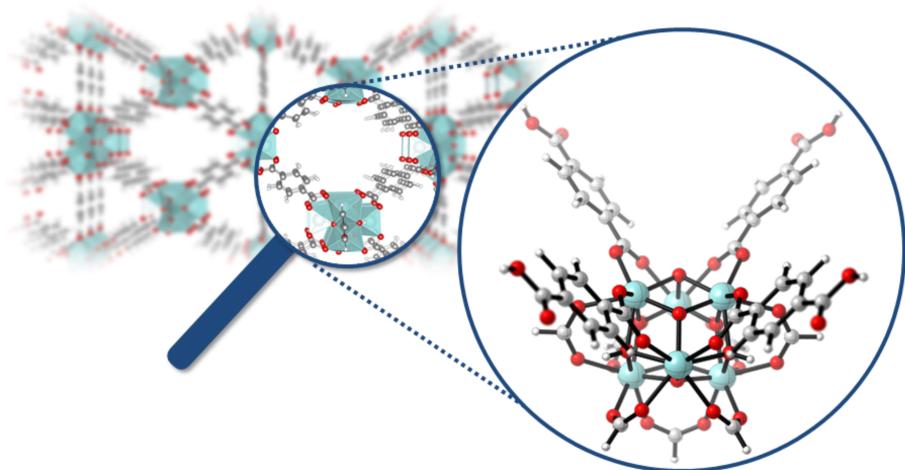
## Extended cluster model

A very computationally efficient approach consists in neglecting periodicity and extracting a finite cluster of atoms from the periodic structure. This cluster model, displayed in Figure 2.1, contains the active sites and their surroundings but consists in a limited number of atoms, which decrease the computational cost. This allows both a more accurate treatment of the electronic structure, and a screening of different possible geometrical configurations of adsorbates, which is useful in the search for transition states. Moreover, very efficient transitions state searching algorithms have been developed for such systems in Gaussian, the most widely used code for cluster calculations. When cutting a cluster, particular attention has to be drawn to the termination of bonds and the charge compensation, that have to be done in the most realistic way. The rest of the crystal structure does not surround the external cluster atoms. Some of these atoms need to be fixed in order to mimic the periodic environment and prevent nonphysical deformations that would affect an estimation of the entropy [51].

Cluster calculations are an excellent way to benchmark and do a first qualitative screening of reactions and possible configurations and have been for long the standard computational tool when studying reaction in nanoporous materials. However, they are not adequate to correctly describe complex reactive processes, where confinement effect of the pores and structural rearrangements can play an important role. The role of solvent in the pores can also be crucial for the outcome of a reaction and cannot be explicitly studied by cluster models. Periodic calculations resolve this shortcoming, and as computational power grows, the heterogeneous catalysis community is shifting towards these more expensive, however more accurate models.

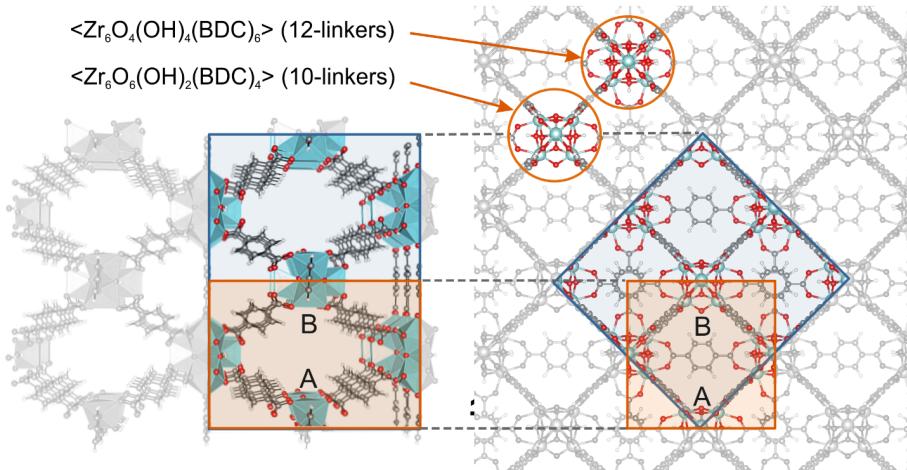
## Periodic model

Periodic models enable to describe the whole topology of the framework. These calculations make use of periodic boundary conditions (PBC), which allow to simulate bulk phases with a limited number of atoms. In this model, the unit cell is replicated infinite times in each direction. When one atom disappears from one side



**Figure 2.1:** Extended cluster model cut from the periodic structure of UiO-66. The cluster contains the active site, the brick and the linkers in the closest proximity to the active site.

of the unit cell it will reappear on the opposite side and each atom interacts with its neighbors in the same unit cell but also in the adjacent ones. Spurious interactions between the atoms can be avoided by applying a *minimum image convention*, for which each atom interacts with its nearest neighbor or periodic image. In the case of long range interactions such as the electrostatic other techniques need to be used, such as Ewald summation [52], where the potential is divided into a short range contribution, calculated in real space, and a long range contribution, calculated in reciprocal space using a Fourier transform. In the case of UiO-66, the conventional unit cell [27] contains 4 Zr bricks (Figure 2.2, blue). In the calculations of this thesis, BDC linkers have been removed in the unit cell to introduce defects which are active sites in catalysis. Different amounts of missing linkers with different topologies have been considered. An interesting topology is the one denoted as type 6 in the work of Rogge et al. [41] that is characterized by a channel which offers good perspectives for the diffusion of guest molecules. This unit cell (displayed in blue in Figure 2.2) can be reduced by symmetry to a 2-brick unit cell (in orange, Figure 2.2) which offers the best compromise between accuracy and computational cost. This reduced unit cell is used in most of the calculations performed in this thesis.



**Figure 2.2:** Representation of the unit cells containing the defect. In blue, the conventional 4-brick unit cell, in orange, the 2-brick unit cell used for the calculations. The two different bricks are highlighted in orange. The 10-fold coordinated brick has two terephthalate linkers missing, one at site A and one at the opposite site B.

## 2.2 Theoretical methods

### Electronic energy methods

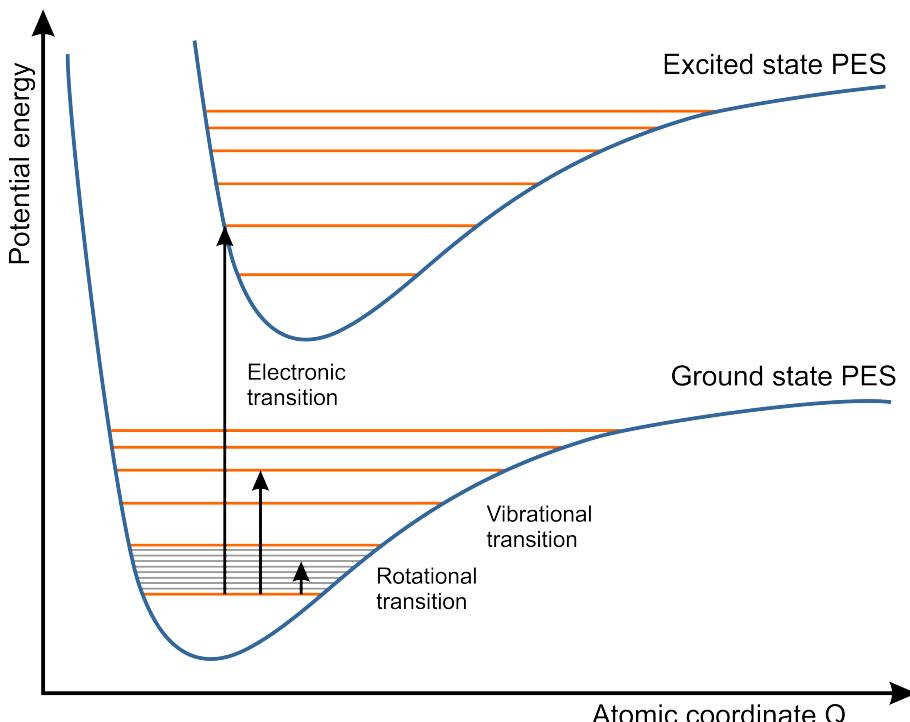
A basic quantity to study any chemical or physical transformation is the potential energy surface (PES), which is a function of the coordinates of all the atoms of the system. The PES is always the reference quantity in our simulations and every atomic configuration can be represented as a point in this hypersurface, with a given value of potential energy (Figure 2.3). Ideally, by calculating the value of the PES for each atomic configuration we can obtain all information on the system and on the transformations that can occur. However, the complexity of this surface escalates quickly with the number of atoms, and the sampling of its relevant regions represents the main challenge of molecular simulations. The information gained by exploring the PES is tightly connected to the experimental observables. Statistical physics acts like a bridge between the microscopic insight that is gained through molecular simulations and the macroscopic properties which are measured experimentally. In principle, all macroscopic properties of a system can be derived from its wavefunction. To calculate it, the stationary Schrödinger equation is solved:

$$\hat{H} |\psi\rangle = E |\psi\rangle$$

where  $\psi$  is the wavefunction,  $\hat{H}$  is the Hamiltonian of the system, and  $E$  is the total energy. The resolution of this equation is at the heart of computational

chemistry and will in principle provide the exact description of matter, but it is nevertheless extremely difficult to solve for most of the electron systems. The presence of electron–electron interactions makes it a highly coordinated problem, and for this reason, different approximations need to be applied, to remove the interactions that have a minimal contribution to the energy.

### Born–Oppenheimer approximation

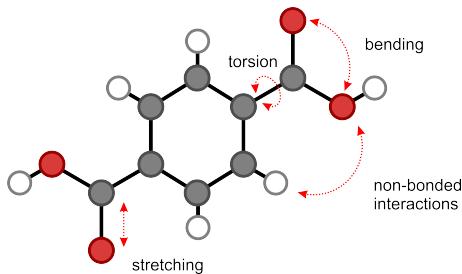


**Figure 2.3:** The two lowest PES in the BO approximation for a diatomic molecule. In blue, the electronic PES for the ground state and first electronic excited state (UV-Vis transition). In orange, the vibrational energy levels (IR transitions), in grey the rotational levels (microwave transitions).

For all calculations performed in this work, we rely on the so-called Born–Oppenheimer (BO) approximation [53]. In this treatment, nuclei are considered as classical points which move in the potential energy surface generated by the electrons (Figure 2.3). This way, to each nuclear configuration a corresponding electronic energy can be assigned, and nuclear coordinates enter in the Schrödinger equation only as parameters, allowing to construct a BO surface, or PES. This approximation holds since nuclei are much slower than electrons, therefore the motion of electron is instantaneous from the nuclei point of view. This approximation is not always possible, especially when dealing with light nuclei such as hydrogen.

In these cases, nuclear quantum effects can have an impact on the measured properties [54]. In most cases, the electronic ground state is also not interacting with the higher electronic states because of the high energy difference. In the BO approximation, the electronic energy levels are also considered fully separated and do not interact with each other. For this reason, the approximation is also called adiabatic approximation. Additional interactions have to be considered when two surfaces lie close to each other, for instance in the neighborhood of conical intersections.

## Force Fields



**Figure 2.4:** Representation of some of the molecular modes taken into account in a generic force field model.

The simplest way to describe interactions between atoms which determine the PES is the so-called “balls and springs” model. In this treatment, all interactions are represented by interatomic classical potentials which are parametrized to reproduce the results of more accurate quantum mechanical calculations. In this work, generic force field calculations have been used in some cases to give preliminary input structures for more costly *ab initio* calculations, through which the description of chemical transformations is possible. Force fields are often constituted by harmonic potentials which do not allow bonds being broken and formed (Figure 2.4). Reactive force fields, such as ReaxFF [55] are currently being developed, but their application in complex heterogeneous reactions is still an ongoing chemical challenge and is out of the scope of this work. For this reason, the description of reactive processes needs a more advanced treatment, where electronic distributions are explicitly taken into account.

## Density Functional Theory

Density functional theory (DFT) has become the method of choice for the study of chemical systems, due to its good trade-off between computational cost and accuracy of the obtained results. DFT began in the 1920’s with the work of Thomas and Fermi [56, 57], but it was only in the '60s that it became a complete and accurate theory, shown in the work of Kohn, Hohenberg and Sham [58]. The

fundamental property that DFT describes is electron density as opposed to many body electron wavefunctions, which allows to reduce enormously the number of variables in the case of complex systems. Two fundamental theorems by Hohenberg and Kohn state that there is a unique relation between electronic density and total wavefunction, therefore the ground state density allows us to determine all properties of the system. Moreover, the ground state density can be obtained from a minimization of the total energy functional with a variational method by solving the so-called Kohn-Sham equations. The global minimum value of the functional determines the exact ground state of the system. This way it is possible to obtain the total energy of the system and the forces which act on the atoms, two quantities which are needed in all the simulations performed in this work.

In principle, DFT is an exact method, but the minimization of energy is far from trivial. Kohn and Sham [59] introduced a method which replaces the many-body problem with an auxiliary system of non-interacting particles, allowing a fast solution of the eigenvalue problem. What needs to be added in this treatment is an additional functional which describes exchange and correlation. Nowadays one of the greatest challenges in DFT consists in the search for an accurate expression for the exchange-correlation functional. The simplest is known as Local Density Approximation (LDA) initially proposed by Kohn and Sham [59] and can also be adapted to include spin in the Local Spin Density Approximation (LSDA) [60]. A more refined method is the Generalized Gradient Approximation (GGA) which involves the calculation of the gradient of electron density and includes functionals such as B88 [61], LYP [62] and PBE [63, 64], used in this thesis. More recent functionals are the so-called hybrid functionals, which include the Hartree-Fock (HF) exchange, such as B3LYP [61, 62, 65], which is a combination of B88, LYP and LDA with HF, and PBE0 [66], which mixes PBE with HF. These functionals can give a more accurate electronic description of the system but are computationally very expensive. As compromise between accuracy and computational cost, what is often performed in the simulation of this thesis is a geometry optimization with PBE, and a single point calculation to refine the energies with B3LYP.

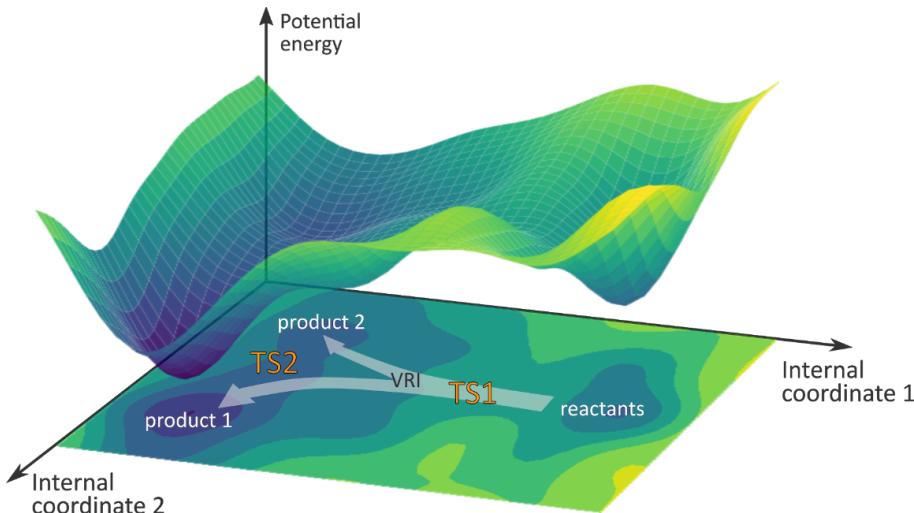
### Dispersion interactions

In this thesis we often encounter non-covalent interactions which need to be treated with high accuracy, such as the adsorption of guest molecules on the Zr Lewis acid sites or interaction between solvent molecules. One of the challenges of DFT methods is the description of long range dispersive interactions such as London forces, which are commonly referred to as van der Waals interactions. These interactions are due to many particle electron correlation effects which are present also in absence of charges and can have a significant impact on the non-covalent interaction energy. To tackle this problem, various dispersion schemes have been proposed. One of the most used is currently the Grimme–D3 method [67], where a damped  $-C_6R^{-6}$  function is added to the DFT functional. Recently, more advanced dispersion schemes have been developed, such as the many body

dispersion scheme [68], or the one of Tkatchenko and collaborators [69], although for the systems we are studying not many benchmarks of these new methods have been performed so far [70].

## Geometry optimization

In order to obtain molecular structures that have physical significance and their relative energies, the arrangement of the atoms needs to be optimized. There are generally two types of molecular structures that we need to find in our simulations, the equilibrium geometries, which correspond to minima of the PES, and the transition state geometries, which correspond to first order saddle points, as displayed in Figure 2.5. These points are characterized by null first derivatives of the energy (the total forces acting on each atom are sufficiently close to zero), all positive second derivatives for local minima and one negative second derivative for first order saddle points, which correspond to transition states.



**Figure 2.5:** Schematic representation of the potential energy surface and the stationary points.

The geometrical optimization of reactants and products consists in a minimization of the energy along the nuclear coordinates. Often the starting point is the experimental structure which can be obtained from diffraction data. In the most used codes several minimization methods are implemented, each characterized by a different computational cost and robustness, such as steepest descent, conjugated gradient or simulated annealing. The algorithms will find local minima, and do not guarantee that the system will be in a global minimum, therefore the minimizations must start from a sufficiently good guess. The search for transition states is far from trivial and often requires an iterative procedure involving different methods and requiring a good knowledge about the system and chemical process under study.

In the calculations performed in this thesis, we often start from an equilibrium structure and as a first guess, we adapt the bond lengths and angles to be close to the transition state with a molecular editor such as Zeobuilder [71]. These bond lengths are then fixed, and the rest of the structure is reoptimized. The Hessian of this partly optimized structure needs to be then computed, and the vibrational modes analyzed, to check which (if any) negative frequency corresponds to the transition state. The system can be optimized again without the constraint using an improved dimer method along a selected eigenvector [72] which is followed by the optimization with a quasi-Newton method [73]. In some difficult cases, the TS search can be initiated on simpler cluster models and optimized in a code in which methods are usually implemented to directly find the TS structure. This way, a first guess of the TS geometry might be obtained and further transferred to the periodic model. In both transition states and minima, if there are superfluous negative frequencies, these need to be removed. Often, it is sufficient to minimize the energy along that vibrational mode. Single point energy calculations can be performed for different values of the displacement along that mode, and the lowest point in energy can be used as starting point for a subsequent minimization of all the coordinates.

### Cell optimization

In the case of periodic systems, not only the structure, but also the unit cell needs to be optimized. This is not trivial, as when using a finite plane wave basis set the number of plane waves depends on the volume of the unit cell. If the volume changes during the optimization, artificial forces which go under the name of Pulay stress can arise. This would require many iterations to optimize the volume. In this thesis, another approach was used [74] which relies on an equation of state fit. For a given volume, for instance taken from experimental data, the unit cell is optimized. Then a set of equally spaced different volumes is defined and for each of these points the geometry and unit cell parameters are optimized. This way it is possible to construct an energy-volume curve, which for a rigid system can be fitted with a Birch-Murnaghan equation of state, allowing to extract the volume  $V_0$  which corresponds to the minimum electronic energy  $E_0$ .

$$E(V) = E_0 + \frac{9V_0B_0}{16} \left\{ \left[ \left( \frac{V_0}{V} \right)^{\frac{2}{3}} - 1 \right]^3 B'_0 + \left[ \left( \frac{V_0}{V} \right)^{\frac{2}{3}} - 1 \right]^2 \left[ 6 - 4 \left( \frac{V_0}{V} \right)^{\frac{2}{3}} \right] \right\}$$

Where  $B_0$  and  $B'_0$  are the bulk modulus and its derivative. A new structure is then generated at this given volume and coordinates and unit cell parameters are optimized again.

### Molecular vibrations

As seen in the previous paragraph, for many purposes in this thesis we need to calculate the second order derivatives (Hessian matrix) of the PES, which are

associated to molecular vibrations. First of all, the Hessian gives us information about the curvature of the surface and the nature of the stationary points encountered during the minimization. The second order derivatives are obtained by displacing the atoms in the three directions and calculating the energies, then the Hessian is diagonalized to determine the eigenvectors that correspond to the vibrational motions. From the Hessian we can calculate the vibrational frequencies, which open the door to a lot more information on the system than a single point calculation. Single point calculations are performed at 0 K, but even at this temperature nuclei vibrate around their equilibrium positions, and this movements are responsible for the vibrational entropy. We can approximate these motions with those of harmonic oscillators, by using the vibrational frequencies constructed from the Hessian. These frequencies can then be used to estimate the value of the vibrational entropy at finite temperatures, as will be explained later. In the calculations performed in this thesis, due to computational limits, a partial Hessian approach was used when dealing with reactions, as implemented in the TAMkin toolkit [75]. The quantity that needs to be derived from these calculations is the change in free energy, which mainly depends on the parts of the system that change during the reaction, in the case of a heterogeneous catalyst the active site and the adsorbed reactants. Therefore, restricting the entropy calculations only to this part of the system is a good approximation that allows to decrease enormously the computational cost [76]. This approach has been used in the calculation of the free energy barriers for the Fischer esterification on UiO-66 (**PAPER I**), where the atoms taken into account were the adsorbed reactants and four atoms of the active sites in their immediate proximity, as displayed in Figure 2.6.

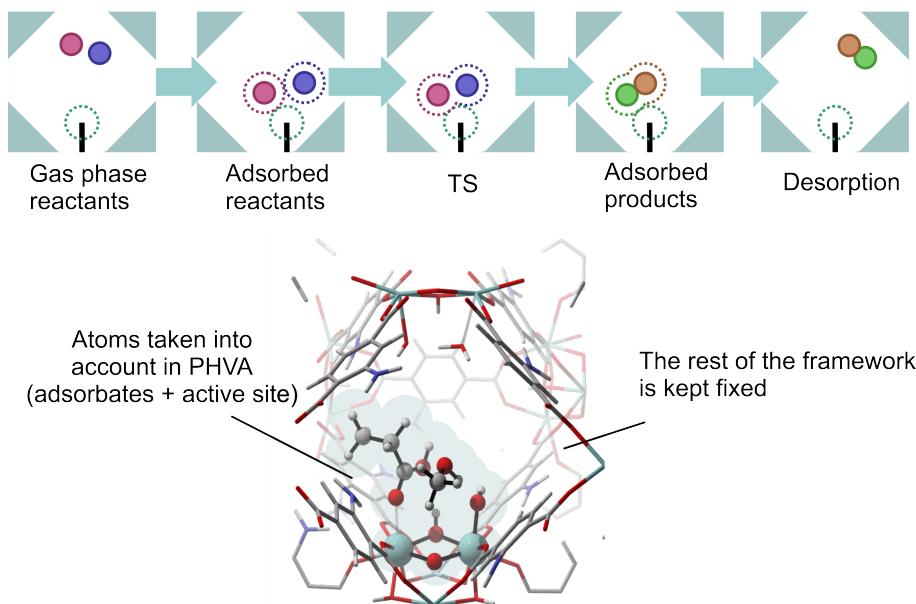
## 2.3 Free energy

The central thermodynamic quantity that determines the outcome of a reaction is the free energy change associated to the process. In general, a chemical system will undergo changes in a direction that minimizes its free energy, until an equilibrium is reached. Knowing the difference in free energy between reactants and products allows us to know the equilibrium constant for a given reaction. The Gibbs Free energy can be decomposed in an enthalpic and an entropic contribution, that can be evaluated from the simulations knowing the molecular partition functions. Initially, the total internal energy  $U$  has to be obtained from the electronic energy  $\varepsilon_0$ , the zero-point vibrational energy  $E_{ZPE}$  and the molecular partition function  $Q$  at constant number of particles  $n$  and volume  $V$ :

$$U = U_0 + RT^2 \left( \frac{\partial \ln Q}{\partial T} \right)_{n,V}$$

$$U_0 = \varepsilon_0 + E_{ZPE}$$

Where  $R$  is the gas constant, equal to the product  $N_A \cdot k_B$  between Avogadro's number and Boltzmann constant. The molecular partition function  $Q$  can be split



**Figure 2.6:** representation of the atoms taken into account in the PHVA approach. Top: a schematic representation of a reactive process in nanoporous material, bottom: a snapshot from the static calculations where the atoms of the active site and the adsorbates are highlighted

in its translational, rotational, and vibrational components:

$$Q = Q_{trans} Q_{rot,ext} Q_{vib}$$

The enthalpy  $H$  corresponds to the total energy plus the work associated to the change in volume.

$$H = U + p_0 V$$

The entropy  $S$  can be directly obtained from the partition function:

$$S = R \ln Q + RT \left( \frac{\partial \ln Q}{\partial T} \right)_{n,V}$$

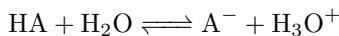
Finally, the Gibbs free energy  $G$  will be:

$$G = H - TS = U_0 + p_0 V - RT \ln Q$$

Where in the case of non-interacting particles  $p_0 V = RT$  following the ideal gas law.

## Equilibrium

As explained above, there is a tight connection between free energy and equilibrium concentrations in chemical reactions. As an example, an equilibrium which is of outmost importance in chemistry is the acidic dissociation of species in aqueous solution:



The acidic dissociation constant ( $K_a$ ) is an important equilibrium constant in chemistry and is equal to the ratio between the concentration of products and reactants when the reaction reaches the equilibrium. It is often reported with its negative decimal logarithm as  $pK_a$ .

$$K_a = \frac{[\text{A}^-][\text{H}^+]}{[\text{HA}]}$$

The equilibrium constant is equal to the Gibbs free energy change from reactants to products.

$$\Delta G = RT \ln K_a$$

$$pK_a = \frac{\ln 10}{RT} \Delta G$$

## Transition state theory

A chemical reaction is a process that through rearrangement of the atoms transforms one stable state into another. Every elementary reaction can be represented as a minimum energy path connecting two minima along the potential energy surface. Furthermore, along this reaction path the existence of a saddle can be

postulated, which is the highest point in energy that needs to be crossed to go to the product state. The saddle point is typically called transition state or activated complex and it is the basis for the transition state theory developed by Eyring in the 1930's, one of the most successful chemical theories which allows to explain reaction rates of elementary chemical reactions. The assumption of the theory is that there is a quasi-equilibrium between reactants and activated complex, and the rate constant can be obtained by the size of the energy barrier and by the frequency at which the system can cross the barrier. This is possible because the barrier acts as a bottleneck in the reaction, its crossing is a rare event and all the kinetics depends only on it. For a unimolecular reaction, the rate constant can be derived from the partition functions of reactant, transition state and their energy difference:

$$k(T) = \frac{k_B T}{h} \frac{q_{TS,\ddagger}}{q_R} e^{-\frac{\Delta E^\ddagger}{k_B T}}$$

Where  $k_B$  is the Boltzmann constant,  $h$  is the Planck constant,  $q_R$  and  $q_{TS,\ddagger}$  are the molecular partition functions of reactants and activated complex for all coordinates except the reaction coordinate, evaluated from the zero-point vibrational level.

$$q_{vib,i} = \prod_{i=1}^{N_{dof}} \frac{1}{1 - e^{-\frac{h\nu_i}{k_B T}}}$$

Where  $N_{dof}$  are the number of vibrational degrees of freedom of the system. The energy difference  $\Delta E^\ddagger$  includes electronic energy and zero-point vibrational energy difference at 0 K:

$$\begin{aligned} \Delta E^\ddagger &= E_0^{TS} - E_0^R + \Delta E_{0,vib} \\ \Delta E_{0,vib} &= \sum_{i=0}^{N_{dof}-1} \frac{h\nu_i^{TS,\ddagger}}{2} - \sum_{i=0}^{N_{dof}} \frac{h\nu_i^R}{2} \end{aligned}$$

This theory has some limitations, and may fail in the case of labile intermediates, when nuclei deviate from a classical behavior, or at high temperatures. For a given reaction, in fact, there will be many paths characterized by different barriers, and at low temperature only the lowest one will be likely to be crossed. When the kinetic energy is high enough, many other paths will be activated. The transition state will occupy a larger region of the PES, and it will not be possible to derive entropy from the vibrational partition functions.

## 2.4 Exploring the free energy surface

Static calculations, where molecular vibrations are approximated using harmonic oscillators, can fail to give an accurate representation of the entropy when there is a high configurational freedom. When the PES is flat with respect to  $k_B T$ , the system at equilibrium can evolve in a larger region of the PES and move along more than one minimum. In this case, vibrational frequencies are anharmonic and it is

not possible to represent the system by approximating around one single minimum. Therefore, static calculations are not always sufficient in describing the system at operating conditions. In this view, molecular dynamics (MD) techniques, which follow the time evolution of the system, can resolve this shortcoming.

## Ab initio Molecular Dynamics

From MD simulations, thermodynamic properties such as free energy can be obtained taking into account a whole region of the PES instead of a single point. This is based on the ergodic theorem, that states that the time average of equilibrium properties is equal to the ensemble average, in the limit of a sufficient long simulation. MD simulations are based on solving Newton's equations of motion:

$$M_i \ddot{\mathbf{R}}_i = \mathbf{F}_i = -\nabla_i V$$

where  $M_i$  and  $\mathbf{R}_i$  are the mass of a given nucleus and its coordinates,  $\mathbf{F}_i$  the forces that act on it, which correspond to the gradient  $\nabla_i V$  of the PES. There are many ways to calculate these quantities and to integrate the equations of motion, and at present time, chemists and physicists can choose between a plethora of MD techniques which span a whole range of complexity, accuracy and computational cost. In the calculations performed in this thesis, potential energy and forces on the PES are calculated from first principles by means of DFT to account for the full dynamic behavior of the material by ab initio molecular dynamics (AIMD). The calculation of electronic properties which define the PES is decoupled from the propagation of nuclear motions, in a method called Born-Oppenheimer Molecular Dynamics (BOMD). Other famous AIMD methods, which differ by how the calculations of electronic potential and the equation of motion are combined, are the Car-Parrinello MD (CPMD) [77], where a fictitious electronic kinetic energy is added to the lagrangian, or the Ehrenfest MD, based on the namesake theorem [78, 79]. The first MD calculations were performed in the microcanonical (NVE) ensemble, where total energy, number of particles and volume are fixed. However, in experiments it is often the temperature that is fixed, not the energy. In general, the choice of the ensemble depends on the thermodynamic quantities that need to be determined. Nowadays there are many thermodynamic ensembles in which the simulation can be performed. The most convenient for a comparison with experiments are the canonical (NVT), with fixed number of molecules, volume and temperature, or the isothermal-isobaric (NpT), with fixed number of molecules and temperature, but where the volume can fluctuate. In order to have a fixed average temperature, some control of the kinetic energy of the atoms is needed. Various thermostats, which differ by speed and robustness, are implemented in every MD code. In this thesis, Nose'-Hoover thermostat was used, where the system is connected to a heat bath. The pressure is also controlled in simulations by means of a barostat. The most commonly used is the one developed by Martyna, Tobias, and Klein (MTK) [80].

from here add references

### Towards modeling at operating conditions

With the growth in computational time, the new challenge is constituted by modeling the system at operating conditions. Many chemical reactions, especially when performed at mild conditions, involve the presence of a solvent. In order to move closer to modeling the system at operating conditions, the solvent in the pores can also be taken into account. This adds a lot of degrees of freedom to the system, and for this reason often an implicit description of the solvent is done, such as in the Periodic Continuum Model (PCM). In the case of the work in this thesis, however, it is necessary to fully model the solvent molecules, as they are actively involved in proton transfers. To do so, the number of solvent molecules that can fit in the unit cell needs to be estimated. Monte Carlo method (MC) is an alternative approach to MD to explore the PES for complex systems. It was initially developed for the calculation of multidimensional integrals and is nowadays largely used in chemistry, especially when dealing with adsorption. In the framework of this thesis, it has been applied in the Gran Canonical ensemble (fixed chemical potential, volume and temperature) to determine the number of solvent molecules that could fill the pores of the material at standard conditions.

### Enhanced MD methods

MD simulations can offer valuable insights into the behaviour of a chemical system at equilibrium conditions. From these simulations, many properties can be extracted, such as equilibrium geometries, vibrational spectra, diffusion coefficients, structural parameters etc. Configurations associated to higher (or lower) values of potential energies will be sampled for shorter (or longer) times, and in principle, if a certain process is sufficiently sampled, based on the ergodic theorem we can know its equilibrium constant, and in turn the free energy barrier associated to it. However, chemical reactions, where bonds are broken and formed, are generally rare events that will not be sampled with a regular exploration of the PES. If the barrier is high compared to the  $k_B T$ , the probability that such event would spontaneously occur during the simulation time is practically none. For this reason, different enhanced sampling techniques have been developed to enhance the sampling of certain regions of the PES.

### Choice of collective variable

The PES is a highly dimensional surface, defined by the positions of all atoms in the system. However, often a reactive process can be described by few important coordinates called “collective variables” (CV), which are projections of the high dimensional space. In fact, what can be considered a chemical configuration is an ensemble of microstates that are different in terms of absolute coordinates of each

atom, but all contribute to the same macrostate. In the simplest case, a single CV can represent the reaction coordinate for the process.

The choice of the right collective variable is crucial to describe the correct process. In the case of the simulations of this work, we made use of distances, angles and coordination numbers. Coordination numbers represent a smart choice compared to distances, because for each pair of atoms, the coordination is zero for distances higher than a certain threshold. This allows to consider the possibility that different atoms could decoordinate and coordinate again to a certain atom

rewrite this

what are the main choices, distances, CN, angles

## Metadynamics

### Umbrella sampling

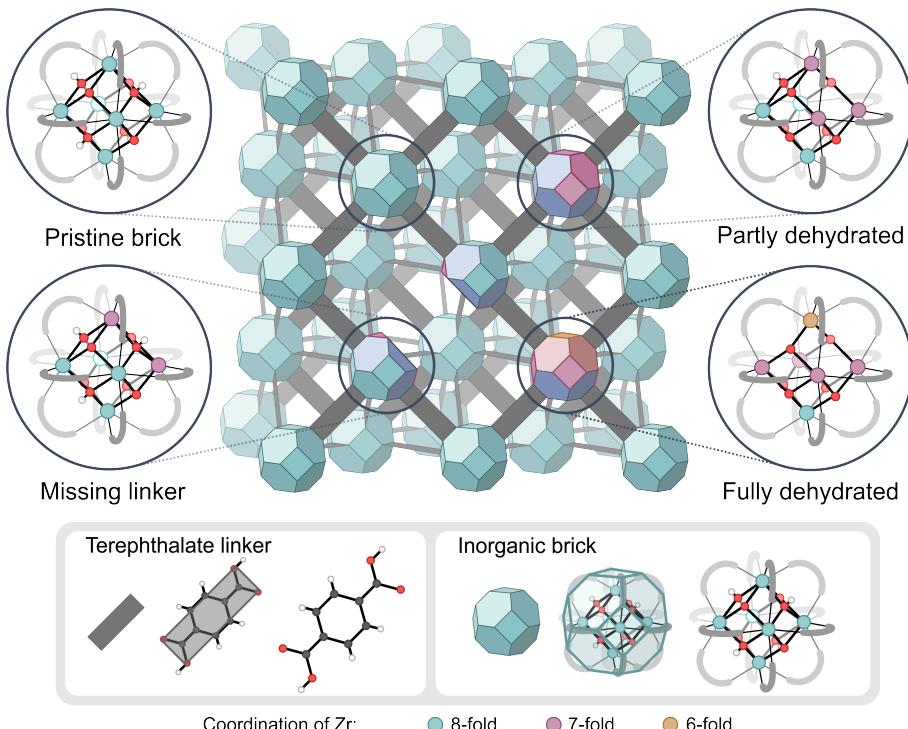
# 3

## Major research results

This Chapter illustrates the main research results obtained in the framework of this thesis. The main goal of this work was the study of the nature of active sites on UiO–66 and MOF–808 upon activation processes, and how the solvent and the functionalization influenced their behavior. The role of active sites on defective UiO–66 was studied for Fischer esterification of free fatty acids (FFA), at it became clear that solvent played an unexpected active and beneficial role in the reaction mechanism. Different molecular modeling techniques based on static and dynamic methods have been applied to gain insight into the interaction between active sites and reactants in the material, as has been introduced in Chapter 2. Contrary to reactions in zeolites, processes in MOFs are often performed ad mild conditions in the presence of a solvent, which adds complexity to the model. So far, solvent in MOFs had been studied only with classic approaches, but with the increase in computational power it was possible to include a full *ab initio* treatment of water and methanol solvent in the UiO–66 pores and to go towards an operando description of activation processes in MOFs. The most important scientific results will be highlighted in this Chapter. More details are to be found in the original articles, enclosed in Part II.

### 3.1 Activation processes in zirconium–MOFs

One of the main challenges in MOF research is the understanding of how active sites are created and how they impact the properties of the material. For this purpose, molecular modeling offers a platform that allows to study such activation processes and nature of active sites at the molecular level. In this sense, UiO–66, characterized by an exceptional stability, is the perfect MOF archetype where different activation and PSM processes can take place without disrupting the stability of the structure. This MOF represents a showcase example, as the findings can in principle be extended to other MOFs that are more difficult to study.

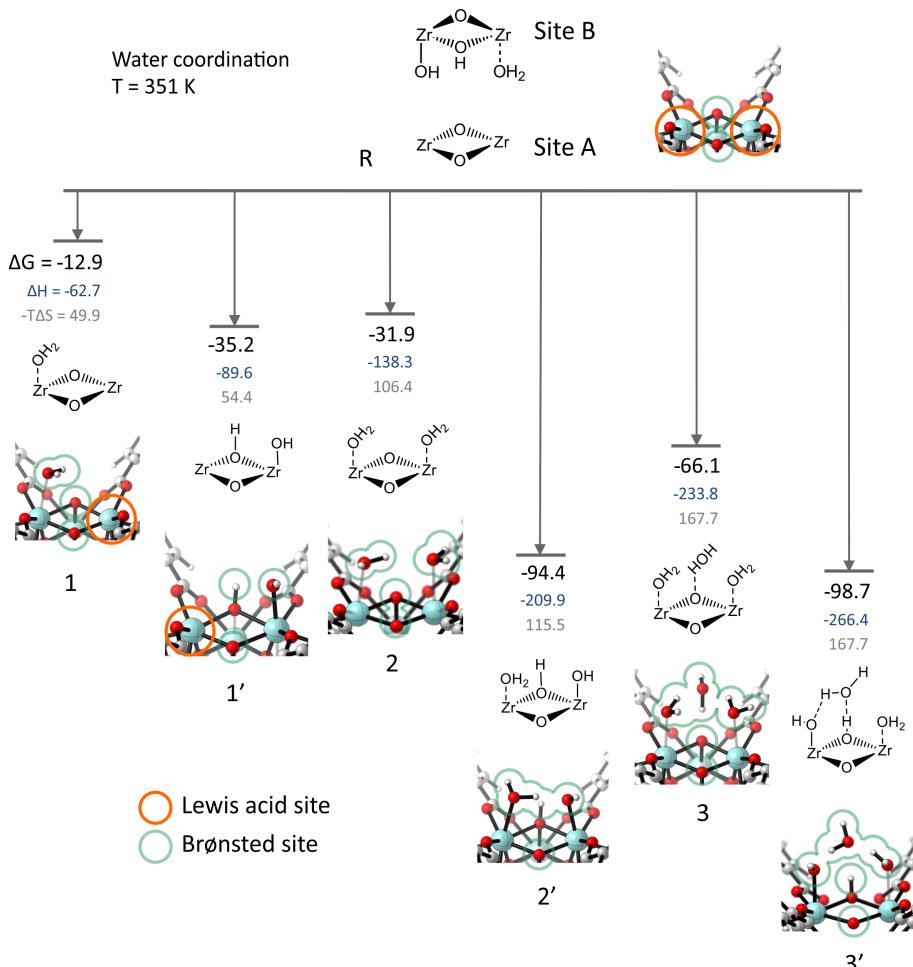


**Figure 3.1:** Schematic representation of the UiO-66 structure with possible configurations of the bricks that give rise to coordinatively unsaturated Zr atoms. The colors indicate the coordination of the Zr atoms.

#### 3.1.1 Missing linker defects

In **PAPER I** we investigated the local defect topology upon removal of a linker on UiO–66 as starting point to investigate the catalytic role of defect coordinating species in the mechanism of Fischer esterification. When studying the catalytic behaviour of UiO–66, an essential ingredient to understand the reactivity of the

inorganic SBU is the knowledge of the molecular structure of the active sites upon removal of a linker. On these sites, defect coordinating species can be adsorbed and have an impact on the catalytic properties of the MOF by introducing additional sites that can play an active role in reactions. As a thought experiment, we can



**Figure 3.2:** Coordination free energies at reaction temperature of 351 K of one, two and three water molecules at coordinatively unsaturated Zr-bricks in defective UiO-66 with respect to a water coordination free site (site R). The structure of the opposite site B corresponds with configuration 2' with two water molecules and consistently used in all periodic calculations considered in the figure. Free energies (in black) are given in kJ/mol, and their decomposition into enthalpic  $\Delta H$  (blue) and entropic  $-T\Delta S$  (grey) contributions. Energies are resulting from periodic calculations with PBE-D3 level of theory. In each configuration Lewis acid and Brønsted sites are indicated.

imagine that upon removal of one of the twelve negatively charged BDC linkers from the inorganic  $\text{Zr}_6\text{O}_4(\text{OH})_4$  SBU, the positive charge left on the brick has to be compensated. Charge neutralization can be accomplished by either coordinating a negative ion such as hydroxyl group to one of the zirconium atoms, or by removing a proton from the brick. This latter case is characterized by two zirconium open metal sites, and is the type of active site that is obtained upon thermal treatment of the brick at  $T > 150$  C. It was previously accepted that these Lewis sites were responsible for the catalytic activity of UiO—66, but recent studies point towards the active role of Brønsted sites in the neighbourhood of defective zirconium atoms <sup>1</sup>. The presence of multiple sites makes it difficult to establish a simple structure–activity relation. In particular, it is important to understand from a mechanistic point of view how the presence of defect coordinating species may affect the catalytic activity on Zr-MOFs. This research was done in collaboration with the group of Dr. Francesc X. Llabrés i Xamena as experimental partners. The coordination of water species near the active sites can occur with different configuration schematically displayed in Fig. 3.2. In all structures the presence of possible Lewis and Brønsted sites which may play a role in reactions is highlighted. The simplest case taken as reference configuration is the unsaturated site obtained upon thermal activation, containing two  $\mu_3$ -oxygens bridging the Zr-atoms which may act as Brønsted sites. From this structure, the coordination of one physisorbed water molecule to one of the zirconium atoms is energetically favorable. A decrease in energy is observed when the molecule is deprotonated to the oxo atom as in configuration 1'. The chemisorption of this water molecule shields the Lewis character of the zirconium atom, but introduces an additional Brønsted site in proximity to the open metal site. The most stable configurations are observed in presence of 2 or 3 adsorbed water molecules on the adjacent zirconium atoms. Two physisorbed water molecules of configuration 2' can be adsorbed to the two zirconium atoms, followed by an immediate dissociation of one of the two molecules into a hydroxyl group and a proton on the adjacent  $\mu_3$  oxygen. This configuration is characterized by a free energy difference of -94.4 kJ/mol at reaction temperature of 351 K compared to the dehydrated defective site. The presence of two zirconium coordinating species has been reported by Lillerud <sup>2</sup> by means of Single-crystal X-Ray diffraction (SXRD), showing that the material is most stable when all zirconium atoms are fully coordinated. When starting from the thermally activated material that contains open Lewis sites, at standard condition, water present in the atmosphere will immediately coordinate to restore the 8-fold coordination of the zirconium atoms to give structure 2. This structure is consistent with the one proposed by the group of Farha <sup>3</sup> who identified three types of protons from potentiometric titration:  $\mu_3$ –OH, Zr–OH and Zr–OH<sub>2</sub>. A third bridging hydroxyl species as charge neutralizing

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<sup>1</sup>(Canivet, Fateeva et al. 2014, Ghosh, Colon et al. 2014, Øien, Wragg et al. 2014, Canivet, Vandichel et al. 2016, Ling and Slater 2016, Liu, Klet et al. 2016, Vandichel, Hajek et al. 2016)

<sup>2</sup>cite Øien

<sup>3</sup>cite Klet Liu 2016

species was proposed by Yaghi<sup>4</sup> from XRD data. They propose two physisorbed water molecules bridged by an OH<sup>-</sup> counterion that is stabilized by a hydrogen–bond interaction with the neighboring  $\mu_3$ -OH of the brick. However, a recent study Ling and Slater<sup>5</sup> did not succeed in finding a corresponding minimum on the PES. In this work, we observe that the  $\mu_3$ -OH atom immediately deprotonates in proximity of such OH<sup>-</sup> anion (configuration 3). Similarly as in the previous case, this structure can further stabilized by a deprotonation of one of the two physisorbed water molecules, to yield configuration 3' of Fig. 3.2, in agreement with a previous report by Vandichel et al.<sup>6</sup>. Adsorption of reactants involved on the Fischer esterification reaction was also taken into account. Similar considerations on the deprotonation processes can be drawn when considering methanol instead of water as reported more in detail in **PAPER I** and **PAPER II**. The energies obtained clearly demonstrate that water molecules preferentially adsorb on the zirconium atoms and the only limit to the adsorption of more water molecules lies in the entropic penalty.

### 3.1.2 Nature of active sites for Fischer esterification

Among the reactions that can be catalyzed by this MOF, Fischer esterification is an important process in the production of biodiesel, a biofuel that is obtained from renewable sources, such as oils and animal fats. UiO-66 has been shown to be a stable and reusable Lewis catalyst with high conversion rate for the reaction<sup>7</sup>. Experimental findings performed on both hydrated and dehydrated UiO-66 show that water has a beneficial role in the reaction, but a theoretical rationalization of the underlying causes was missing. Moreover, amino functionalization was shown to increase the reaction rate. To address these questions, the role of active sites on UiO-66 during the Fischer esterification reaction was studied in **PAPER I**. Two possible lowest activated reaction pathways were identified for the hydrated and dehydrated active site.

The proposed reaction mechanism is shown in Fig. 3.3. From the reactant configuration, the physisorbed water molecule is displaced by the carboxylic acid that coordinates on the zirconium atom. The resulting configuration was identified after a series of static calculations probing possible geometries. In this configuration, the acid carbonyl group is bonded to the zirconium, and at the same time methanol is hydrogen bonded to the hydroxyl group coordinated to zirconium. The adsorption of the acid on the Lewis acid site gives to the carboxylic carbon a more electrophilic character, making it more prone to interact with the alcohol. The oxygen of methanol is at the same time made more nucleophilic due to the hydrogen–bonding interaction with the Brønsted basic site situated in close

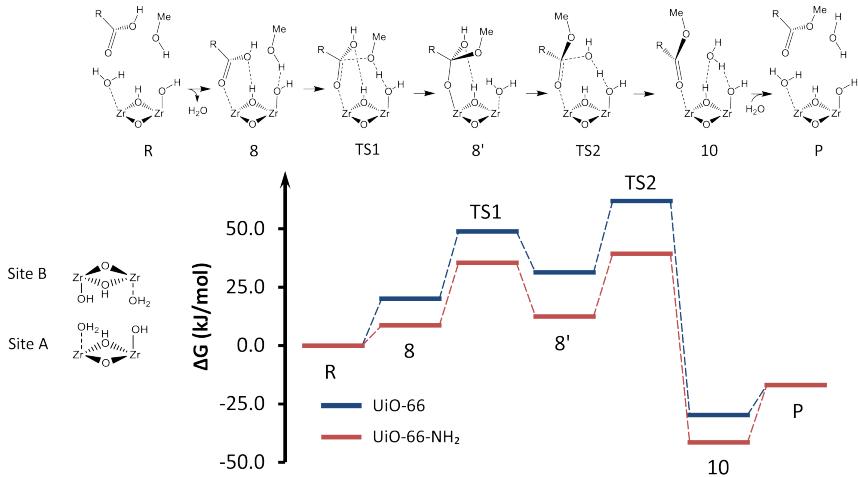
<sup>4</sup>cite Trickett

<sup>5</sup>cite Ling Slater

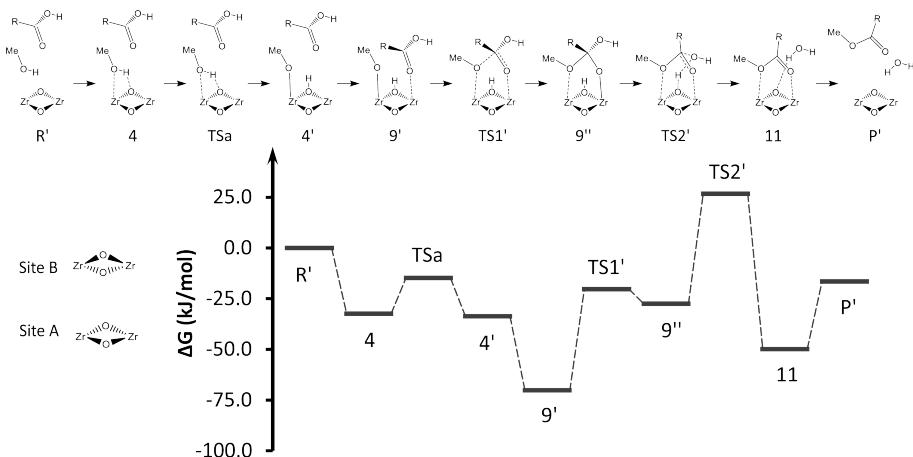
<sup>6</sup>cite Vandichel 2016

<sup>7</sup>cite Cirujano2015, Cirujano2016

### Hydrated material



### Dehydrated material



**Figure 3.3:** Mechanism and free energy profile for the esterification of propionic acid with methanol on a hydrated and defective UiO-66 material (blue), a hydrated defective UiO-66 material with amino functionalization of the BDC linkers (red), and on a dehydrated defective UiO-66 (black). Periodic calculations at B3LYP-D3//PBE-D3 level of theory, T=351 K. R corresponds with an empty frame with one linker defect and a pool with all reactants to guarantee mass balance. In P the defective Zr-brick is coordinated with two water molecules (configuration 2'). P' corresponds to the empty frame with the ester as final product and remaining water molecules in gas phase.

proximity. This favors the condensation between activated carboxylic carbon of the acid and methanol. In both transition states, the hydroxyl group plays a role, first as proton acceptor, then as proton donator, while the acid maintains the bond with the carbonyl oxygen. The low energy barriers associated to the two transition states are 28.9 and 30.6 kJ/mol at 351 K. The reaction can proceed both ways, until an equilibrium is reached. This mechanism is characterized by a dual participation of Brønsted and Lewis sites, and overlays with the experimental findings.

Upon thermal treatment at 150 C, UiO–66 loses the adsorbed solvent molecules without compromising the structure of the inorganic SBU<sup>8</sup>, contrary to the dehydroxylation with release of two water molecule that takes part at T > 300 C<sup>9</sup>. In principle, these open metal sites should be more catalytically active, but a decrease in catalytic activity is experimentally observed. In the proposed mechanism, two Lewis acid sites, the zirconium atoms, and one Brønsted basic site, the  $\mu_3$  oxygen, play an active role in the reaction. This mechanism is characterized by three transition states, in which: 1) methanol is deprotonated to the oxo atom, 2) an adduct is formed between the electrophilic carbon and the oxygen of methanol 3) the  $\mu_3$ -OH group deprotonates to form water. This reaction is characterized by higher activation barriers (a total barrier of 90 kJ/mol at 351 K), which explain the lower catalytic activity of the dehydrated material.

Other mechanisms that do not make use of either Lewis or Brønsted sites were investigated without success, as the energy barriers were too high to be likely to occur. Both proposed mechanisms are characterized by a dual participation of Lewis and Brønsted sites that work complementary to each other. The presence of acid and basic centers within molecular distances has been shown to be essential in the performance of the catalytic reaction as they cooperate in a concerted way during the chemical transformation. Most of the previous mechanistic studies on the UiO-66 material merely focused on the Lewis acidity of the under-coordinated sites, but it has become more and more clear that the bifunctional nature of the UiO–66 catalyst will play an important role in its future applications.

### 3.1.3 Role of defect topology

The adsorption of species present in the reaction environment of PSLE process was studied in the framework of **PAPER IV**. In this process, we wanted to see how defect-free and defective UiO–66 material would interact with water and methanol species. To obtain this information, we made use of a larger unit cell, to be able to represent different amount of missing linker defects. This increase in complexity reflects in the higher computational cost of modeling such structure. The defective structures taken into account in the modeling are obtained by removal of linkers

<sup>8</sup>cite Valenzano 2016

<sup>9</sup>cite Valenzano 2011

from the conventional unit cell containing four inorganic  $\text{Zr}_6(\text{O})_4(\text{OH})_4$  bricks<sup>10</sup>. Different numbers of missing linkers were modeled, to represent the different amount of defects in the experimental samples. A case with a low amount of defects is modeled by a unit cell with one missing linker. In this structure, two bricks are 12-fold coordinated, and two bricks are 11-fold coordinated, with an average of 11.5 linkers per brick. A second type of unit cell was taken with three missing linkers, corresponding to a higher amount of defects, with an average of 10.5 linkers per brick. De Vos et al. and Rogge et al.<sup>11</sup> shown in their comprehensive studies that there are multiple topologically diverse possibilities to remove linkers from a 4-brick unit cell. In this work, we chose two possible, discrete cases to represent different coordination of the inorganic brick with this amount of defects. In a first unit cell with three missing linkers, two bricks are 10-fold coordinated and two are 11-fold coordinated. The other represents a more extreme case, with one 9-fold coordinated brick and three 11-fold coordinated bricks. The structures taken into account are denoted as  $(10_b, 10_b, 11, 11)_{334}$  and  $(9_c, 11, 11, 11)_{333}$  in the work of De Vos et al. The zirconium atoms on the defect sites have been capped by all species which can be present in the framework during the activation process, studying different combinations by means of water, hydroxo species, methanol, methoxide and formate. In our models, all Zr atoms have been capped and are fully (8-fold) coordinated, being the most stable state at experimental conditions. On each defect site there is always a negatively charged species, to compensate for the removal of a carboxylate. Given the symmetry of the unit cell, when possible, multiple choices for the positioning of these species were taken into account and modeled. The removal of a negative charge can be also compensated by removal of a proton, as obtained in the dehydrated material, with subsequent physisorption of two neutral species, but these configurations are higher in energy, as seen in **PAPER I**. These results show a clear enthalpic preference for the methanol/methoxide pair in detriment of formate and water. The defective material synthesized with formate modulator is promptly attacked by methanol species present in solution. The obtained results are in line with a previous report<sup>12</sup> and highlight the preference of zirconium active sites for MeOH/MeO- substitution as opposed to H2O/OH-, alternatively assumed as preferential charge balancing element for missing linker defects<sup>13</sup>.

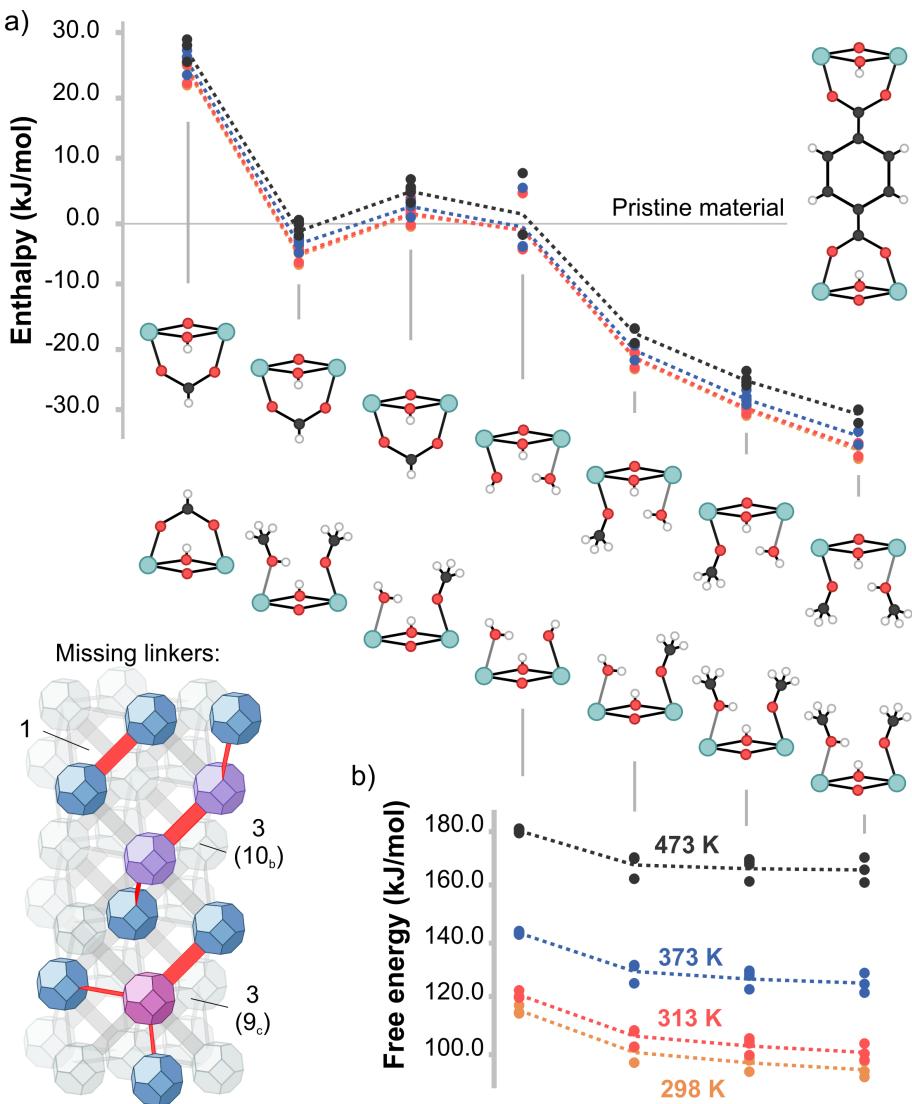
No substantial difference in enthalpy is observed between the unit cells with different amount of defects (energy in each defective structure represented by a dot in Fig. 3.4). This is an indication that the thermodynamics that governs the process does not depend on the starting number of missing linkers, in agreement with the experimental observations of **PAPER IV**. The effect of temperature on the free energy was also investigated on the relative stability of defective configurations

<sup>10</sup>cite cavka 2008

<sup>11</sup>cite Rogge, De Vos

<sup>12</sup>cite Yang et al 2016

<sup>13</sup>cite Trickett 2015, Ling/Slater 2016



**Figure 3.4:** Energy diagrams for defective UiO-66 unit cells. Each dot represents a possible distribution of missing-linker defects (1 or 3 in total) within the unit cell; the connecting dotted line represents a weighted average. Values are normalized by the number of missing linkers in the unit cell. (a) Enthalpy difference between defect sites capped in different ways versus the non-defective material at  $T = 298, 313, 373, 473\text{ K}$ . (b) Temperature-dependence of the free energy difference of the defective structures indicated above (capped with  $\text{H}_2\text{O}/\text{OH}^-$ ,  $\text{H}_2\text{O}/\text{MeO}^-$  and  $\text{MeOH}/\text{MeO}^-$ ) versus the non-defective structure. A representation of the clusters with different missing linker connectivities is also provided.

with respect to the pristine unit cell, where the linker is not missing. The first observation is that configurations involving water and methanol are more stable at lower temperature, thus the formation of missing linker defect through ligand exchange is enthalpically driven and favored by lower temperatures. Secondly, at 473 K the energy of the non-defective framework is much lower than the defective cases, which explains why the synthesis of the non-defective UiO–66 is favorable at such high temperature<sup>14</sup>.

### 3.1.4 Activation by dehydration

The other process that can lead to activation of the UiO–66 material is the reversible dehydration of the brick performed at  $T > 523$  K. The dehydration mechanism may have a decisive effect on certain catalytic reactions, where next to the Lewis acid site also the neighboring Brønsted base or acid site may take a cooperative role in the reaction mechanism. At elevated temperatures and low pressures, the Zr core gets dehydrated and rearranged, and two water molecules are subsequently removed from the brick 3.1. The fully dehydrated  $Zr_6O_6$  brick contains coordinated vacancies<sup>15</sup>, with coordination of the zirconium atoms ranging from 8 to 6. The structure of UiO–66 is however preserved and the brick can easily be rehydrated with a reversible mechanism, giving evidence that the inorganic SBU can undergo dynamic processes. During these structural rearrangements, the presence of a band related to hydroxy groups was observed by Nishida by means of ultrafast IR<sup>16</sup>. The mechanism of dehydration was first studied by Vandichel et al. by means of nudged elastic band calculations<sup>17</sup>, showing the presence of loose hydroxyl groups and partially decoordination linkers. These findings point towards an intrinsic mobility of the framework structure, that can accommodate rearrangements without disrupting its stability. However, the static study of this process possesses serious limitations. In **PAPER III**, we follow on the fly the fast dynamic of the UiO–66 material at dehydration temperature of 573 K by means of umbrella sampling simulations. The results show an intrinsic dynamic behavior of the material, with open metal sites being created by continuous changes in the network connectivity due to labile M–L bonds. We identify two types of motions of the linkers, namely translation along the axis connecting the two adjacent zirconium atoms, and rotation along the linker axis connecting the two inorganic SBUs, as shown in Fig. 3.5. At the same time, these motions are accompanied by a high mobility of hydroxy groups created by decoordination of the  $\mu_3$ -OH groups. These results show that linkers are more mobile than originally anticipated. The high connectivity between the two SBUs of UiO–66 allows all these reversible rearrangements with activation barriers that can be easily accessible at experimental conditions.

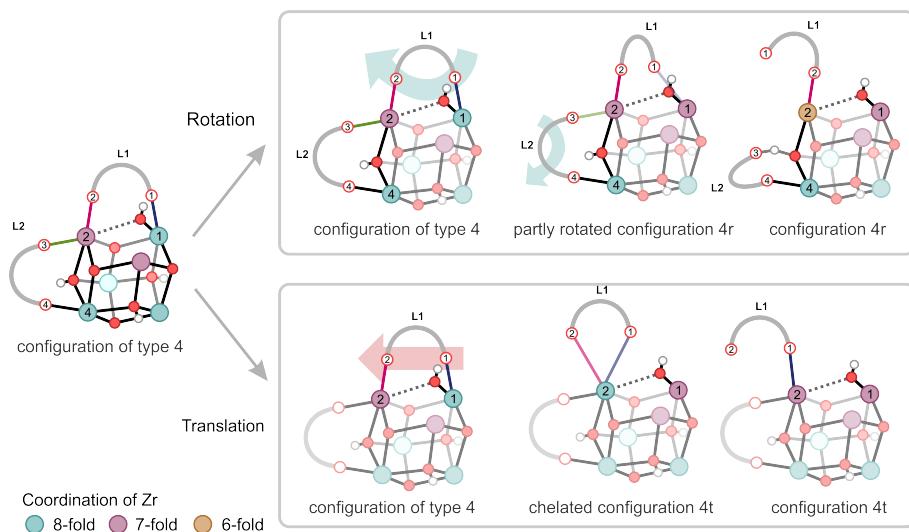
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<sup>14</sup>cite the one on synthesis

<sup>15</sup>cite Valenzano 2011, Shearer 2013, DeCoste 2013, Vandichel 2015

<sup>16</sup>cite Nishida 2014

<sup>17</sup>cite Vandichel 2016



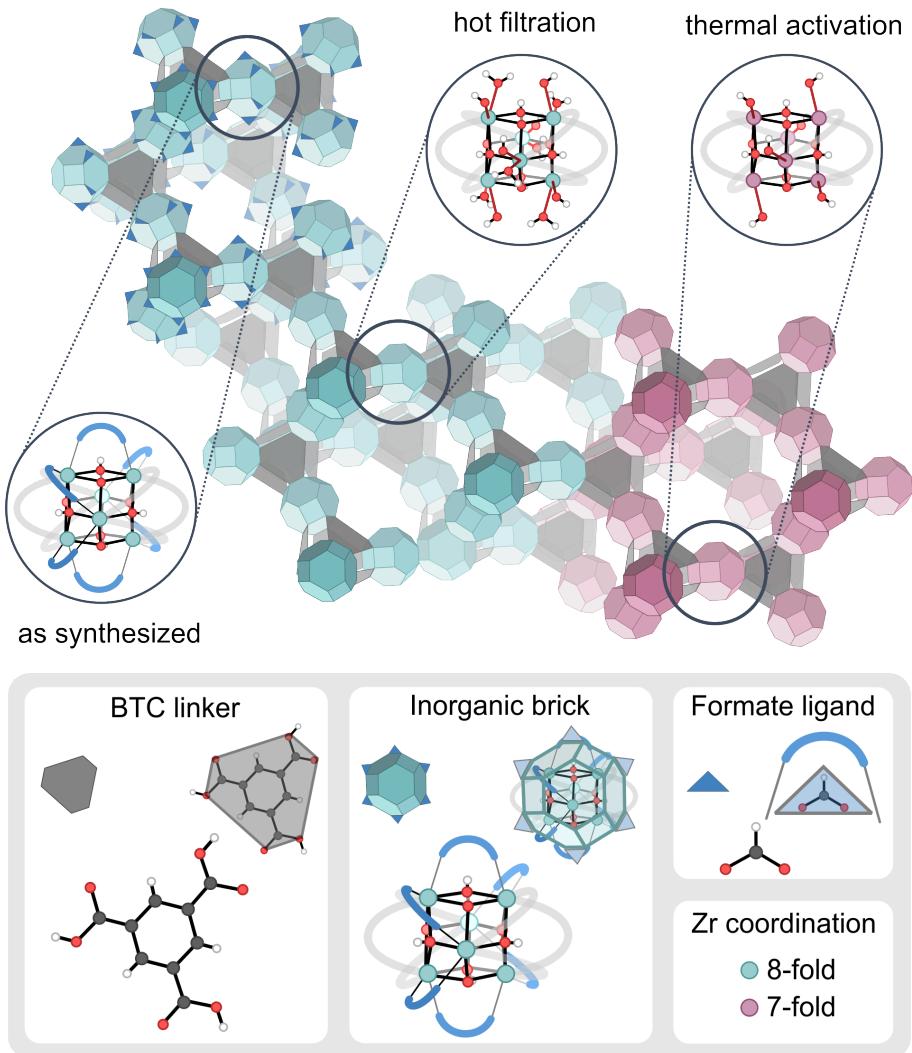
**Figure 3.5:** Umbrella sampling in two windows of  $CV = 1.45$  and  $1.51$ , showing two distinct motions of the linkers. On the left column, a translation of the linker L1 generates a chelated structure and a subsequent shift in the carboxylic oxygen connected to Zr2 (configuration 4t). On the right column, a rotation of the linker L1 and a partial decoordination of linker L2 forming a hydrogen bond with an  $\mu_3$ -OH hydroxyl group is shown (configuration 4r). A proton transfer between the carboxylic oxygen O3 and the bridging  $\mu_3$ -O is also observed and is an indication of the occurrence of an intrinsic dynamic acidity. Colors indicate the coordination number of Zr atoms.

### 3.1.5 Thermal activation of MOF–808

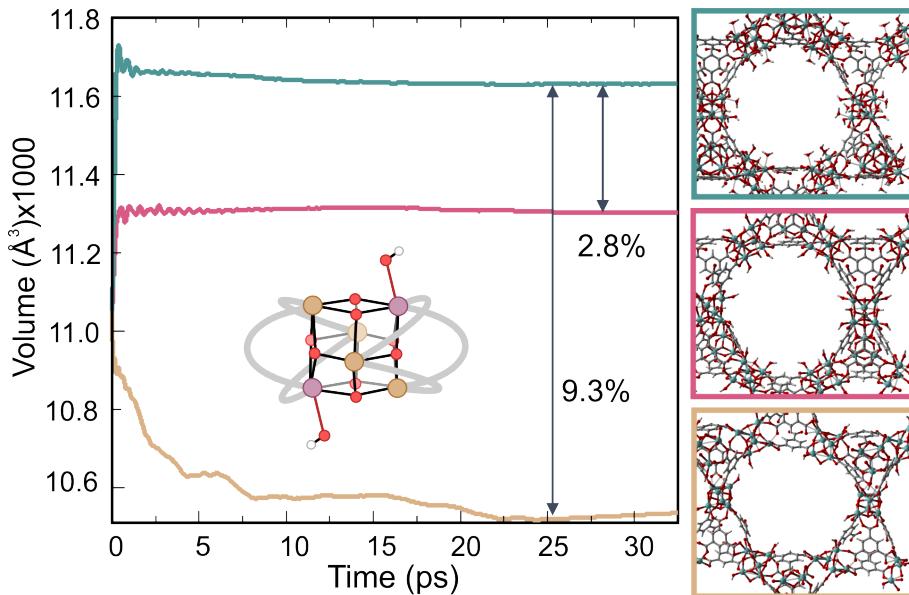
The active sites and coordination changes upon thermal activation have been studied in MOF–808 in **PAPER VI**. MOF–808 shares the  $\text{zr}_6\text{o}_4\text{oh}_4$  brick with UiO–66, but the brick is connected to only six tritopic BTC linkers, making it the least connected MOF in the Zr–MOF family. The material possesses a high catalytic potential due to the intrinsic presence of defective sites of complex nature. In the as synthesized material (Fig. 3.6), the zirconium atoms that are not connected to BTC linkers are capped by formate groups. The material is activated by hot filtration, whereby each formate is replaced by water and hydroxyl group, for a total of six water molecules and six hydroxyl groups per inorganic SBU. Each pair of adjacent zirconium atoms has a similar configuration as the stable configuration 2 of Fig. 3.2, but the presence of multiple Brønsted sites in close proximity gives rise to a more complex nature of the active sites. By thermal activation, similarly to UiO–66, water can be removed from the active sites to open Lewis acid sites for catalysis. In a second stage, up to four of the hydroxyl groups could be also decoordinated by extracting a  $\mu_3\text{--OH}$  proton from the brick to form water, giving rise to mixed–coordination bricks with 6 and 7-fold coordinated zirconium atoms. In **PAPER VI**, the behaviour and stability of the material upon activation processes was investigated by means of a series of independent MD simulations at 300 K at variable unit cell parameters. We show that the dehydration of the inorganic brick substantially affects the stability of the structure. In the hydrated form, the material possesses a high number of Brønsted sites that show dynamic acidity in the form of proton transfer between water and hydroxyl groups that are located in close proximity. The physisorbed water is removed leading to a homogeneous distribution of 7-fold coordinated zirconium atoms in the brick. Upon this dehydration, only a slight decrease in the unit cell volume is observed (Fig. 3.7), and the structure remains stable, even though it possesses a high amount of undercoordinated sites. Moreover, these Lewis sites are located in proximity to Brønsted sites arising from hydroxyl groups, making dehydrated MOF–808 a dual heterogeneous catalyst. Further removal of the hydroxyl groups together with the  $\mu_3\text{--OH}$  protons causes a collapse of the structure. Due to the overall low structure connectivity the adsorbed hydroxyl groups have to be considered as inherent part of the framework composition and their separation results in the collapse of the material. In contrast to UiO–66, that can undergo reversible dehydration processes where the coordination of the zirconium atoms can be lowered to 6 without disrupting the structure, MOF–808 cannot sustain such decrease in coordination.

### 3.1.6 Linker functionalization on UiO–66

The following results concern the linker functionalization on UiO–66 and its effects on the catalytic activity of the material. The easiness of functionalization of UiO–66 can be exploited to tune its catalytic properties.



**Figure 3.6:** Schematic representation of as synthesized and upon activation MOF-808 structures



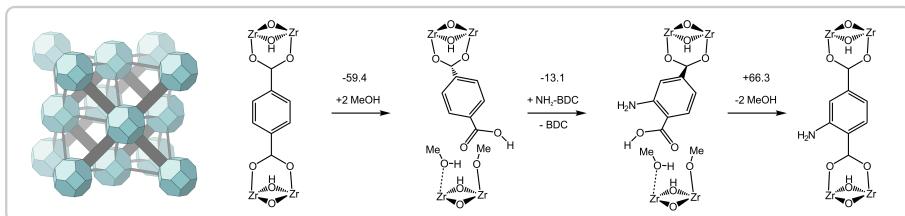
**Figure 3.7:** Change of volume in time of the three investigated structures with different Zr coordination

### Post synthetic linker exchange

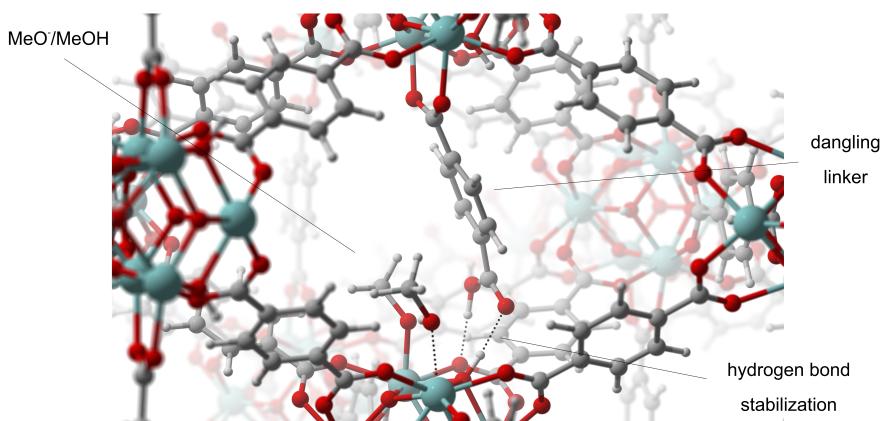
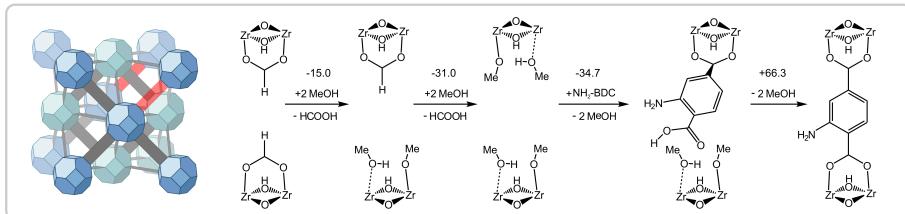
As reported in Chapter 1, functionalization in UiO–66 can be induced by PSLE, exploiting the robustness of the framework, which can easily undergo coordination changes. In **PAPER IV**, the mechanism was investigated in a dual experimental–computational study performed in collaboration with the group of Prof. Rob Ameloot. The PSLE in UiO–66 that leads to functionalization of the BDC linkers with amino groups was performed in methanol at mild conditions (40°C). Results show that the initial amount of missing linkers did not have an effect on the final composition of the material, pointing towards low energy barriers for the exchange. Moreover, the process was accompanied by an initial lowering of the BET surface area. In light of these experimental observations and the computational findings reported in Fig. 3.4, a hypothetical exchange mechanism was modeled starting from both pristine and defective material (Fig. 3.8). In the proposed mechanism, linker exchange is initiated by coordination of methanol on the zirconium sites, causing a partial hydrolysis of one of the BDC linkers. This metastable state is characterized by a dangling linker which is stabilized by hydrogen bonds with other carboxylic oxygen atoms and the  $\mu_3\text{--OH}$  hydrogen from the brick, in a similar fashion as what observed in **PAPER III**. This structure causes a hindrance of the pore in agreement with the reduction in BET surface area. From this configuration, exchange of mono-coordinated BDC linkers with  $\text{NH}_2\text{-BDC}$  quickly ensues, with partial coordination of the new linker. In a following step, methanol is desorbed from the active site and the linker can connect to the zirconium atoms,

reestablishing the binding between the two bricks. This preferential adsorption of BDC-NH<sub>2</sub> is explained by the lowering of the enthalpy of about 7 kJ/mol per linker. When starting from a defective material, a similar mechanism can be proposed. In this second scenario, the synthesized material contains formate on both defective sides. Formate will be substituted by methanol present in solution, and the defect healing can proceed in a similar way as in the previous case. All these rearrangements are supposedly characterized by low activation barriers, as the final composition of the material does not depend on the initial percentage of functionalized or missing linkers. During the process, the zirconium atoms rapidly change their coordination number, however, giving preference to 8-fold coordinated active sites that enhance the stability of the structure.

Non-defective material:



Defective material:



**Figure 3.8:** Proposed PSE mechanism in non-defective and defective UiO-66. MeOH facilitates ligand exchange through the creation and stabilization of defects. Enthalpy differences are given in kJ/mol at 313 K (PSE temperature)

### Effect of linker functionalization on Fischer esterification in UiO–66

In order to understand the effect of the functionalization of the BDC linkers on the catalytic activity of defective UiO–66, the esterification reaction (Fig. 3.3) was investigated also on the amino functionalized UiO–66-NH<sub>2</sub>. Amino functionalization on UiO–66 brings an electrodonating group that would in principle lower the acidity of the Lewis acid site. However, experimental insights into the Fischer esterification reaction showed that the amino functionalized UiO–66 was more catalytically active than its non-functionalized counterpart. For this reason, it was speculated that amino groups located in proximity to the defect sites would play an active role during the reaction. Our computational results show no possible reaction pathway that involves proton transfers with the amino groups. We therefore studied the reaction following the same pathway that we propose for the hydrated brick (Fig. 3.3) in the non-functionalized material. The decrease in Lewis acidity upon amino functionalization is indeed confirmed by the increase of the Zr–O distances of the adsorbates <sup>18</sup>. However, for UiO–66-NH<sub>2</sub>, stronger stabilization of the adsorbates is observed, as well as a slight decrease in the overall energy barrier of about 11 kJ/mol, which confirms the experimental findings <sup>19</sup>. The amino groups, although not playing an active role in the reaction mechanism, indirectly modulate the properties of the Lewis and Brønsted sites. A stronger adsorption of the reactants is caused by the formation of a network of hydrogen bonds with the amino groups that cannot be seen in the pristine material. Moreover, amino groups provide additional sites where solvent can form hydrogen bonds, as displayed in Fig. 3.9. This indirect positive effect of amino groups on the catalytic properties was observed as well by Hajek et al. for aldol condensation <sup>20</sup>. The effect of other functional groups was further analyzed (see Supplementary Material of PAPER I). Electron-withdrawing substituents such as –NO<sub>2</sub> did not decrease the energy barriers, although increasing the Lewis acidity of the metal. Once again, this confirms the dual Lewis/Brønsted character of the UiO–66 catalyst, that can be enhanced by the presence of additional sites within molecular distance.

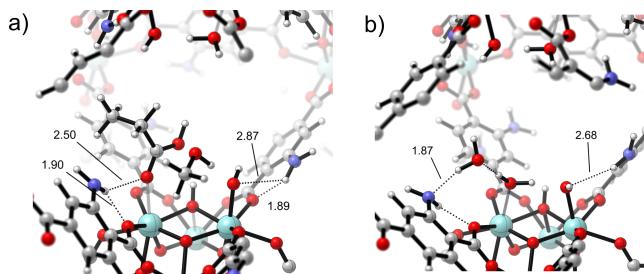
## 3.2 Role of solvent: towards modeling at operating conditions

So far, a simple model to represent the active sites was used, where only the species immediately coordinated to the zirconium atoms were considered. A more complex model can be constructed by including microsolvation, with additional solvent molecules. However, the previous findings point towards a complex nature of the active sites in the material, where solvent may play an active role in reactive

<sup>18</sup>cite Vermoortele, Vandichel et al. 2012

<sup>19</sup>cite Cirujano 2015, Cirujano 2015

<sup>20</sup>cite Vandichel, Hajek et al. 2015

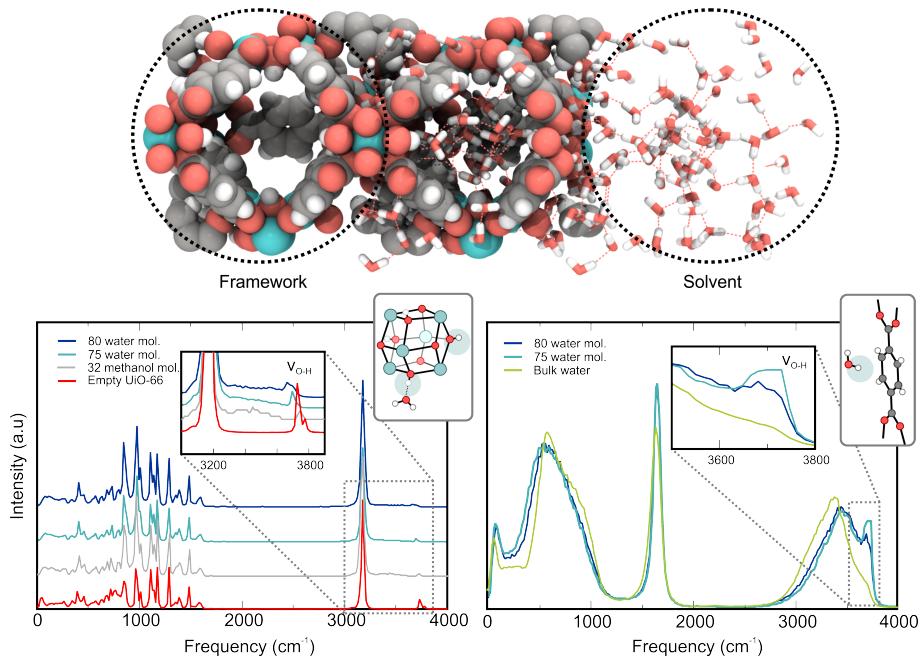


**Figure 3.9:** Network of hydrogen bonds between acid, methanol, hydroxyl group and amino groups. a) reactive complex 8, b) an additional water molecule present in solution.

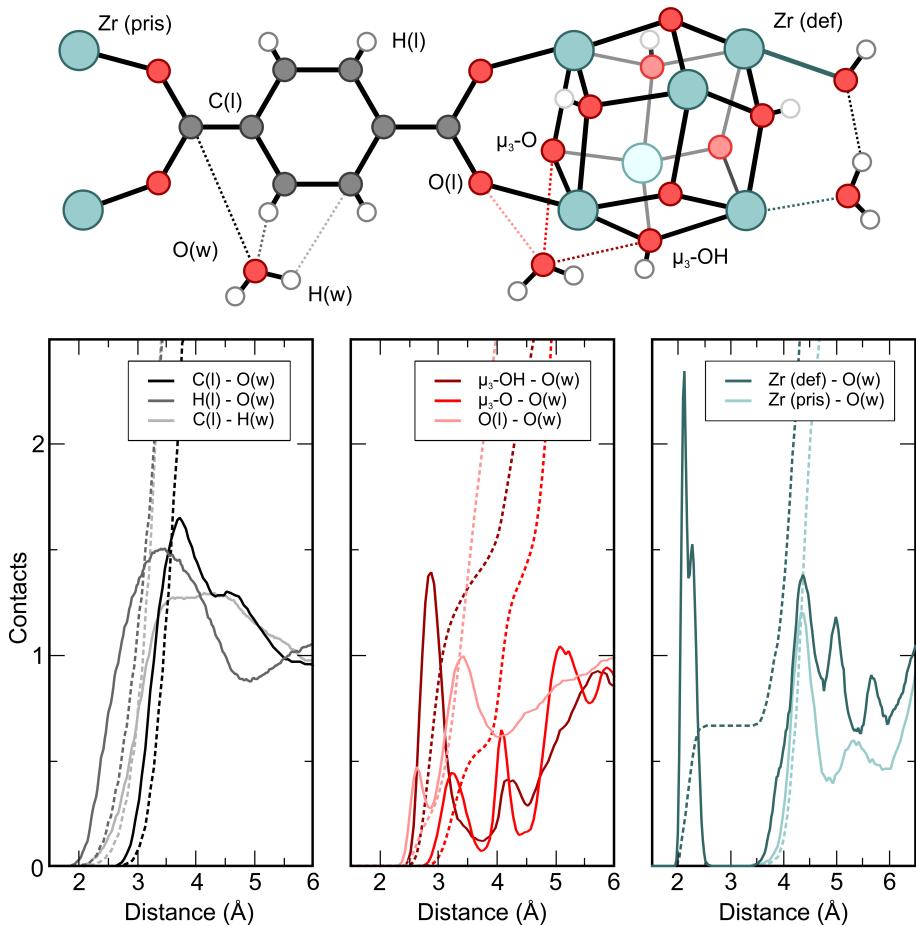
processes and structural modifications. For this reason, in **PAPER II** and **PAPER V** we went a step forward in complexity by including an explicit treatment of the solvent in the pores of the UiO–66 material. This increase in complexity of the model allows to take into account the processes that can occur at operating conditions, and to take into account the full solvent environment within the pores at realistic temperatures and pressures.

### 3.2.1 Interaction between UiO–66 and solvent

### 3.2.2 Interaction between methanol molecules on the UiO–66 defective site



**Figure 3.10:** Top: schematic representation of the empty pore, pore with the solvent, and confined solvent without the material. Bottom: vibrational density of states obtained from the velocity autocorrelation function power spectra of selected atoms of the simulation. Bottom left: solvated material compared to the empty material. Bottom right: water in the pores compared to bulk water.



**Figure 3.11:** Radial distribution functions or pair correlation functions  $g(r)$  between oxygen and hydrogen of water ( $O(w)$ ,  $H(w)$ ), and different atoms of the material obtained from the simulation with 80 water molecules in the unit cell. Full lines indicate the  $g(r)$ , dashed lines indicate its integral. Left panel: RDFs between water and linker carbons and hydrogens ( $C(I)$ ,  $H(I)$ ); middle panel: RDFs between water and oxygen atoms of the linkers and bricks ( $O(I)$ ,  $\mu_3\text{-OH}$ ,  $\mu_3\text{-O}$ ); right panel: RDFs between water and zirconium atoms of defective and pristine bricks ( $Zr(\text{def})$ ,  $Zr(\text{pris})$ ).



## **Part II**

# **Published papers**





## Publication List

Updated Apr 2019

### Publications in international peer-reviewed journals

1. C. Caratelli, J. Hajek, F. G. Cirujano, M. Waroquier, F. X. Llabrés i Xamena, V. Van Speybroeck, *Nature of active sites on UiO-66 and beneficial influence of water in the catalysis of Fischer esterification*, Journal of Catalysis, **352**, 401–414 (2017)  
IF: 6.844
2. C. Caratelli, J. Hajek, S.M.J. Rogge, S. Vandenbrande, E.J. Meijer, M. Waroquier, V. Van Speybroeck, *Influence of a confined methanol solvent on the reactivity of active sites in UiO-66*, ChemPhysChem, **19**, 420–4290 (2018)  
IF:3.075
3. J. Hajek, C. Caratelli, R. Demuynck, K. De Wispelaere, L. Vanduyfhuys, M. Waroquier, V. Van Speybroeck, *On the intrinsic dynamic nature of the rigid UiO-66 metal-organic framework*, Chemical Science, **8**, 2723–2732 (2018)  
IF:8.688

4. J. Marreiros, C. Caratelli, J. Hajek, A. Krajnc, G. Fleury, B. Bueken, D. De Vos, G. Mali, M. Roeffaers, V. Van Speybroeck, R. Ameloot, *Active Role of Methanol in Post-Synthetic Linker Exchange in the Metal-Organic Framework UiO-66*, Chemistry of Materials, **31** (4), 1359–1369 (2019)  
IF: 9.890
5. C. Caratelli, J. Hajek, E.J. Meijer, B. Ensing, M. Waroquier, V. Van Speybroeck, *Dynamic interplay between defective UiO-66 and protic solvents in activated processes*, in preparation
6. J. Hajek, C. Caratelli, ... M. Waroquier, V. Van Speybroeck, *Investigating zirconium coordination changes in MOF-808 upon activation processes*, in preparation
7. M. Filez, C. Caratelli, M. Rivera-Torrente, F. Muniz-Miranda, M. Hoek, M. F.M. Altelaar, A. J.R. Heck, A.k Dutta Chowdhury, V. Van Speybroeck, B. M. Weckhuysen, *Molecular Nucleation Mechanism of Zeolitic Imidazolate Frameworks*, in preparation

## Conference contributions

### Oral presentations

1. Dynamic interplay between defective UiO-66 and confined solvent: insights into a reaction environment at operating conditions  
C. Caratelli, J. Hajek, S. M.J. Rogge, A. Lamine, M. Waroquier, V. Van Speybroeck  
XXth Netherlands' Catalysis and Chemistry Conference (NCCC XX), Noordwijkerhout, The Netherlands, Mar 4–6 2019
2. Modeling nanoporous materials at the nanoscale: the role of high performance computing in materials science  
C. Caratelli  
HPC-UGent User Meeting, Ghent, Belgium, Jan 28 2019
3.  

**TALK OF JOAO AT MOF 2018**

  
J. Marreiros, C. Caratelli, J. Hajek, V. Van Speybroeck  
IVth International Conference on Metal-Organic Frameworks and Open Framework Compounds (MOF 2018), Auckland, New Zealand, 9–13 December 2018

4. Investigating solvent effect in the acidity of UiO-66 metal organic framework for catalysis  
C. Caratelli, J. Hajek, A. Tiwari, S. M.J. Rogge, B. Ensing, M. Waroquier, E.J. Meijer, V. Van Speybroeck  
IVth International Conference on Metal-Organic Frameworks and Open Framework Compounds (MOF 2018), Auckland, New Zealand, 9–13 December 2018
5. Towards a molecular level understanding of chemical and physical phenomena in metal-organic frameworks  
J. Wieme, C. Caratelli, R. Demuynck, A. De Vos, J. Hajek, A. E. J. Hoffman, A. Lamaire, K. Lejaeghere, S. M.J. Rogge, S. Vandenbrande, L. Vanduyfhuys, M. Waroquier, V. Van Speybroeck  
Congrès français des MOFs, Paris, France, 16–18 May 2018
6. Influence of structural topology on the catalytic properties of Zr based MOFs: the case of UiO-66 and MOF-808  
J. Hajek, C. Caratelli, M. Waroquier, V. Van Speybroeck  
XIXth Netherlands' Catalysis and Chemistry Conference (NCCC XIX), Noordwijkerhout, The Netherlands, 5–7 March 2018
7. Influence of a confined methanol solvent on the reactivity of active sites on UiO-66  
C. Caratelli, J. Hajek, M. Waroquier, E.J. Meijer, V. Van Speybroeck  
2nd DEFNET School, Bochum, Germany, 18–21 September 2017
8. First principle study of active sites on UiO-66 for Fischer esterification  
C. Caratelli, J. Hajek, F. G. Cirujano, M. Waroquier, F.X. Llabrés i Xamena, V. Van Speybroeck  
Europacat 2017, Florence, Italy, 27–31 August 2017
9. First principle characterization of active sites on UiO-66 and their role in the catalysis of Fischer esterification  
C. Caratelli, J. Hajek, F. G. Cirujano, M. Waroquier, F. X. Llabrés i Xamena, V. Van Speybroeck  
XVIIIth Netherlands' Catalysis and Chemistry Conference (NCCC XVIII), Noordwijkerhout, The Netherlands, 6–8 March 2017
10. Nature of active sites on UiO-66 and UiO-66-NH<sub>2</sub> in the catalysis of Fischer esterification  
C. Caratelli, J. Hajek, G. Cirujano, A. Corma, M. Waroquier, F.X. Llabrés i Xamena, V. Van Speybroeck  
Chemical Research in Flanders Symposium (CRF-1), Blankenberge, Belgium, 24–26 October 2016
11. Mechanistic study of Fischer esterification on UiO-66 and UiO-66-NH<sub>2</sub>  
C. Caratelli, J. Hajek, F.G. Cirujano, A. Corma, M. Waroquier, F.X. Llabrés

i Xamena, V. Van Speybroeck  
1st DEFNET School, Valencia, Spain, 21–24 June 2016

## Poster presentations

1. How the connectivity of stable Zr-based MOFs affects the metal coordination and the nature of the active sites  
J. Hajek, C. Caratelli, M. Waroquier, V. Van Speybroeck  
MOFSIM 2019, Ghent, Belgium, Apr 10–12 2019
2. Dynamic creation of active sites on UiO-66 by interaction with protic solvents  
C. Caratelli, J. Hajek, S. M.J. Rogge, M. Waroquier, B. Ensing, E.J. Meijer, V. Van Speybroeck  
MOFSIM 2019, Ghent, Belgium, Apr 10–12 2019
3. Dynamic creation of active sites on UiO-66 by interaction with protic solvents  
C. Caratelli, J. Hajek, S. M.J. Rogge, M. Waroquier, B. Ensing, E.J. Meijer, V. Van Speybroeck  
1st KNCV-CTC symposium, Amsterdam, The Netherlands, Mar 26 2019
4. How the connectivity of stable Zr-based MOFs affects the metal coordination and the nature of the active sites  
J. Hajek, C. Caratelli, M. Waroquier, V. Van Speybroeck  
XXth Netherlands' Catalysis and Chemistry Conference (NCCC XX), Noordwijkerhout, The Netherlands, Mar 4–6 2019
5. Exploring the intrinsic dynamics of rigid Zr-based MOFs  
J. Hajek, C. Caratelli, R. Demuynck, K. De Wispelaere, L. Vanduyfhuys, M. Waroquier, V. Van Speybroeck  
IVth International Conference on Metal-Organic Frameworks and Open Framework Compounds (MOF 2018), Auckland, New Zealand, 9–13 December 2018
6. Modeling reactive processes in nanoporous materials: a look into the complexity  
C. Caratelli, K. De Wispelaere, J. Hajek, P. Cnudde, S. M.J. Rogge, S. Vandenbrande, R. Demuynck, L. Vanduyfhuys, M. Waroquier, V. Van Speybroeck  
Ghent, Belgium, Jun 1 2018
7. Investigating the outstanding dynamic behavior of protons on UiO-66 defective sites  
C. Caratelli, J. Hajek, A. Tiwari, M. Waroquier, B. Ensing, E. Jan Meijer, V. Van Speybroeck  
EuroMOF 2017, Delft, The Netherlands, Oct 29 -- Nov 1 2017

8. Influence of a confined methanol solvent on the reactivity of active sites in UiO-66  
C. Caratelli, J. Hajek, S. M.J. Rogge, S. Vandenbrande, E.J. Meijer, M. Waroquier, V. Van Speybroeck  
XIXth Netherlands' Catalysis and Chemistry Conference (NCCC XIX), Noordwijkerhout, The Netherlands, Mar 5–7 2018
9. Catalytic sites on UiO-66 for Fischer esterification  
C. Caratelli, J. Hajek, G. Cirujano, A. Corma, M. Waroquier, F.X. Llabrés i Xamena, V. Van Speybroeck  
MOLSIM 2017: Understanding Molecular Simulations, Amsterdam, The Netherlands, Jan 9–20 2017
10. Catalytic sites on UiO-66 for Fischer esterification  
C. Caratelli, J. Hajek, G. Cirujano, A. Corma, M. Waroquier, F.X. Llabrés i Xamena, V. Van Speybroeck  
Annual IAP Meeting IAP-PAI P7/05, Liège, Belgium, 12 October 2016
11. Catalytic role of UiO-66 and UiO-66-NH<sub>2</sub> in Fischer esterification: a mechanistic study  
C. Caratelli, J. Hajek, G. Cirujano, A. Corma, M. Waroquier, F.X. Llabrés i Xamena, V. Van Speybroeck  
3rd DEFNET workshop, Ghent, Belgium, 22–24 March 2016
12. Catalytic role of UiO-66 and UiO-66-NH<sub>2</sub> in Fischer esterification: a mechanistic study  
C. Caratelli, J. Hajek, G. Cirujano, A. Corma, M. Waroquier, F.X. Llabrés i Xamena, V. Van Speybroeck  
XVIIth Netherlands' Catalysis and Chemistry Conference (NCCC XVII), Noordwijkerhout, The Netherlands, 7–9 March 2016



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