

Polymers

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Definition

- Polymer is high molecular weight substance formed by joining a large number of small repeating units called as Monomers.
- In Greek language poly means many & mer means units
- Polymers form very important components in our daily life. The polymers are highly useful in domestic industrial & medical fields.

Examples

- Rubber
- Polyethene, $n\text{CH}_2=\text{CH}_2 \xrightarrow{\text{polymerization}} (\text{CH}_2-\text{CH}_2)_n$
- Polypropylene
- Teflon

The following are the reasons for the extensive use of polymers.

- Most of the polymers are non-toxic & safe to use.
- They have low densities (light in weight) so transportation polymers will be easy.
- They possess good mechanical strength.
- These are resistant to corrosion and will not absorb moisture when exposed to the atmosphere.
- These can function as good thermal & electrical insulators.
- These can be moulded and fabricate easily.
- They possess aesthetic colors.

But the limitations for the use of polymers are

- Some polymers are combustible.
- The properties of polymers are time dependent
- Some of them cannot with stand high temperature.

Classification of polymers

A) Based on source

B) Based on Physical State

Amorphous and semi crystalline

C) Based on Thermal Behavior or Response to Heat.-

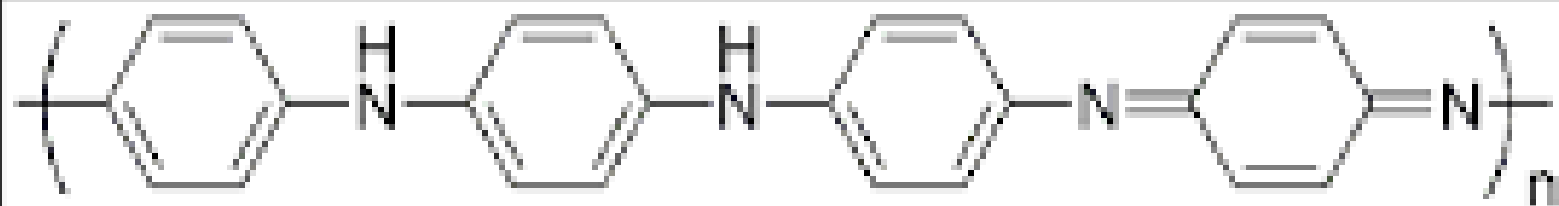
D) Based on end use.-Fibres, Plastics, Elastomers, Films, Paints, Adhesives

E) Based on Origin

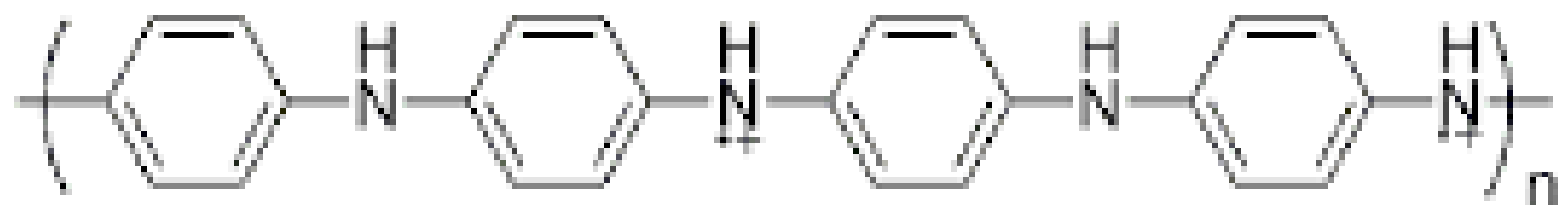
F) Based on Tacticity

Classification of polymers

- G) Based on Conductance
- H) Based on Environment Friendly Nature
- I) Based on the type of monomers
- J) Based on Number of Monomers
- K) Based on Chemical Nature
- L) Based on Growth Mechanism of Polymerization



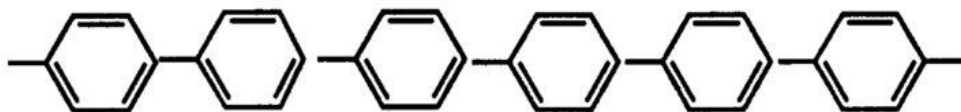
Polyaniline (Emeraldine Base)



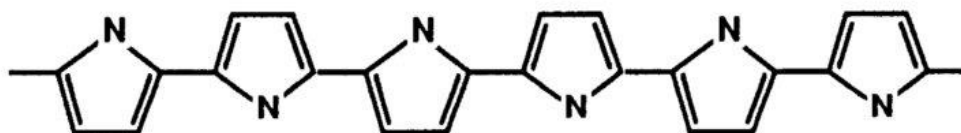
Polyaniline (Emeraldine Salt)



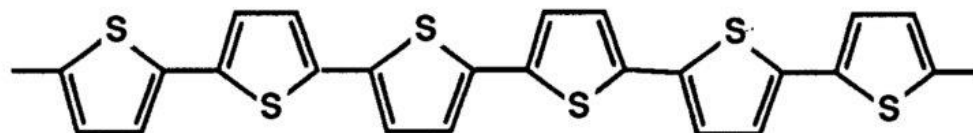
POLYACETYLENE



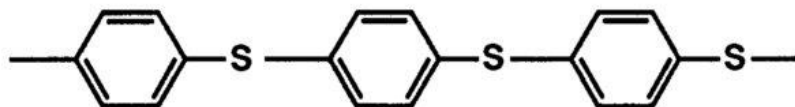
POLYPHENYLENE



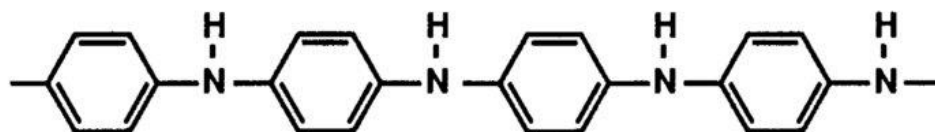
POLYPYRROLE



POLYTHIOPHENE



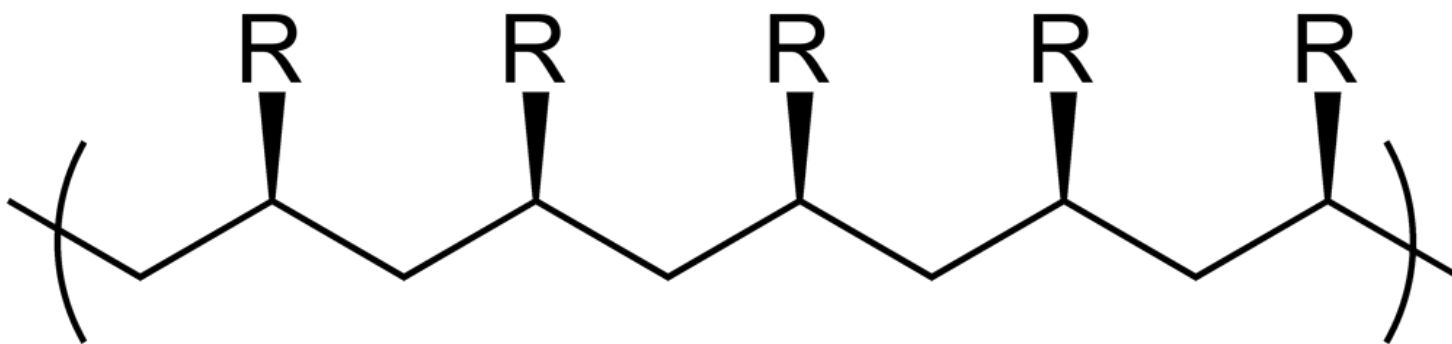
POLY(PHENYLENE-SULFIDE)



POLYANILINE

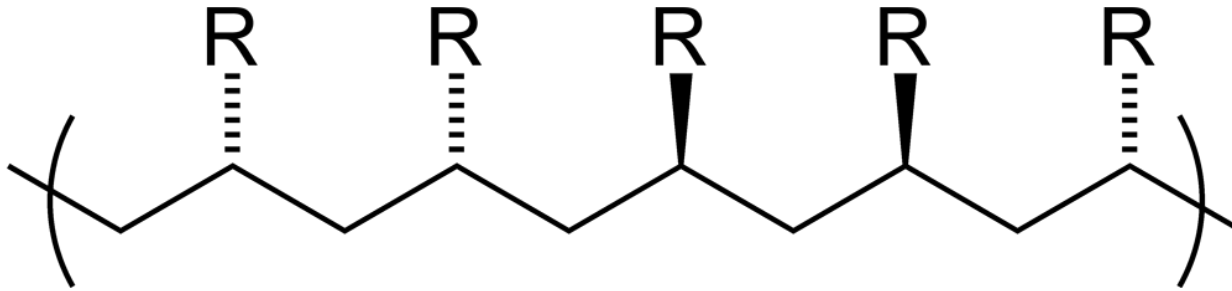
Based on Tacticity

- The arrangement of functional groups on carbon backbone of the polymer is called Tacticity.
 - It is mainly divided into 3 types.
- 1) **Isotactic polymers:** Those polymers in which the functional groups are arranged on the same side are called Isotactic polymers. E.g.:- PVC

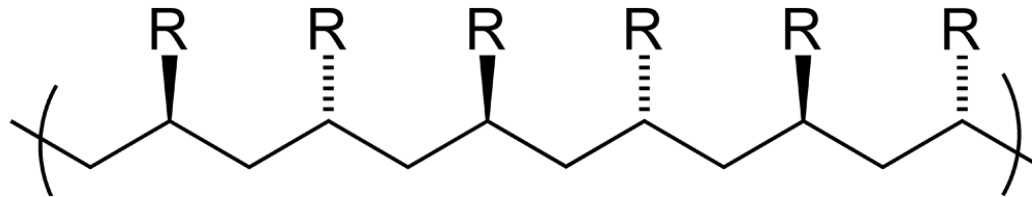


2) **Atactic polymers:** When there is no regular arrangement of functional groups on the back bone of the polymer chain these polymers are called atactic polymers.

E.g.: PVC (Poly Vinyl chloride)



- **Syndiotactic Polymers:** The polymers with alternate arrangement of functional groups are called syndiotactic polymers for e.g.:- PVC

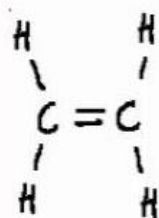


Difference between condensation and addition polymerisation:-

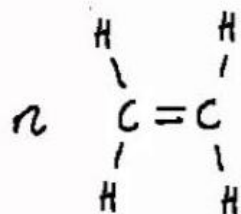
	Condensation polymerisation		Addition polymerisation
(1)	It is also known as step growth polymerisation	(1)	It is also known as chain growth polymerization
(2)	It takes place in monomers having reactive functional groups	(2)	It takes place only in monomers having multiple bonds.
(3)	It takes place with elimination of simple molecule like H_2O , NH_3 , HCl etc.,	(3)	It takes place without elimination of simple molecule.
(4)	Repeat units of monomers are different	(4)	Repeat units & monomers are same.
(5)	The polymer is formed in gradual steps	(5)	Reaction is fast and polymer is formed at once.
(6)	The molecular mass of polymer increases throughout the reaction	(6)	There is very little change in the molecular mass throughout the reaction
(7)	Product obtained may be thermosetting/thermoplastic	(7)	Product obtained are thermoplastic
(8)	E.g.:- Bakelite, polyester ,polyamides etc.,	(8)	E.g.:-Polyethylene, PVC, poly styrene.

Addition polymers

monomer



or



high pressure

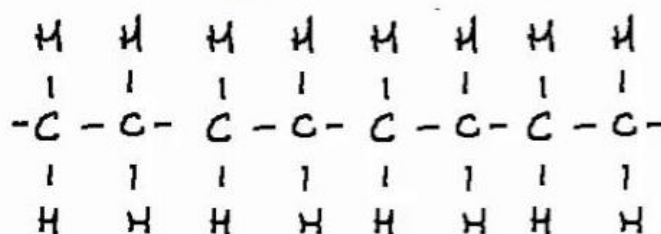


catalyst

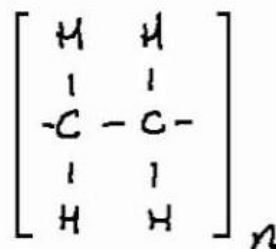
1. conc. H_2SO_4
2. peroxide free radicals



Polymer

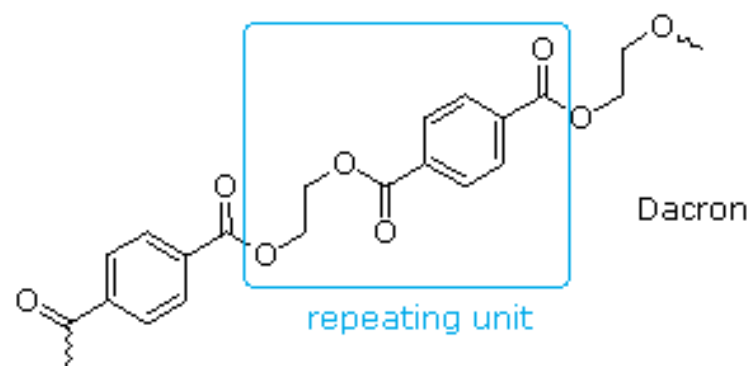
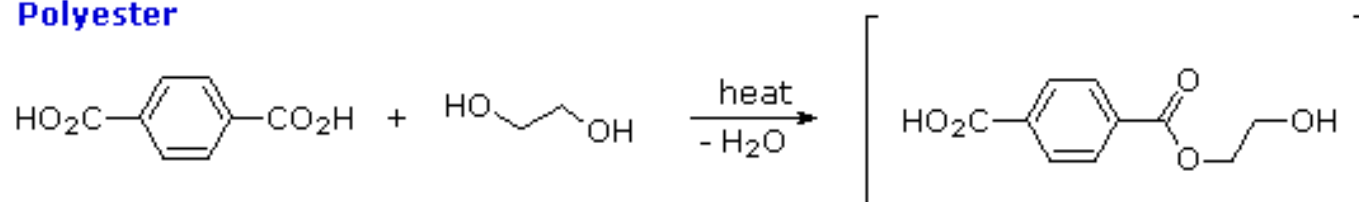


polymer units



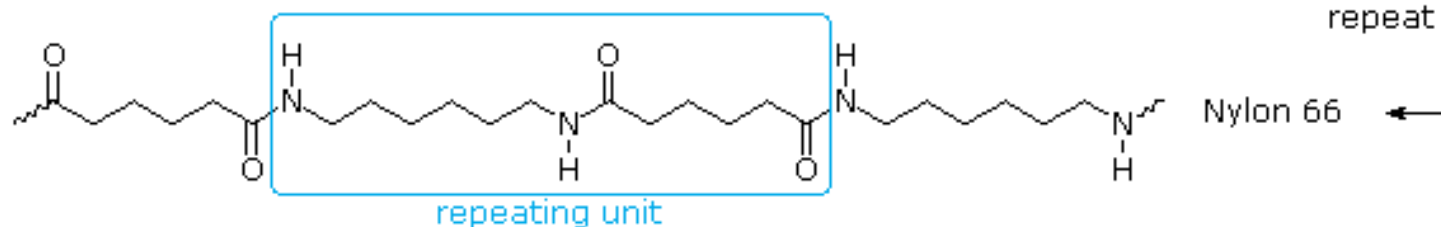
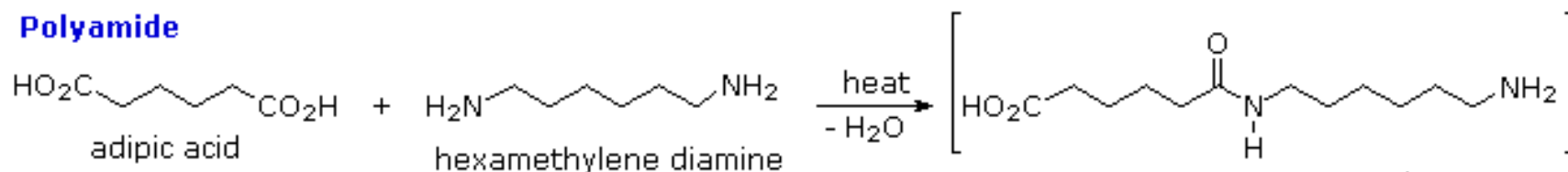
Examples of Condensation Polymers

Polyester



repeat

Polyamide



repeat

Difference between thermoplastic & thermosetting resins:-

Thermoplastic resins (or) Polymers		Thermosetting resins	
(1)	These are produced by addition polymerization	(1)	These are produced by condensation polymerization.
(2)	The resins are made of long chains attached by weak Vander Waal's force of attraction	(2)	The resins have three dimensional network structure connected bonds.
(3)	On heating they soften and on cooling become stiff chemical nature won't change.	(3)	On heating they become stiff & hard. No change on cooling. Chemical nature changes.
(4)	They can be remoulded	(4)	They cannot be remoulded because once set means they are permanently set
(5)	Scrap (waste product) can be used	(5)	Scrap cannot be used
(6)	The resins are soft, weak and less brittle	(6)	The resins are usually hard, strong tough & more brittle
(7)	These are easily soluble in some organic substances E.g.:- PVC, polyethylene etc.,	(7)	Resins are not soluble in organic Solvents E.g.:- Nylon, Bakelite etc.,
(8)	Contain long chain polymer with no cross linkage.	(8)	They have 3D network structure.

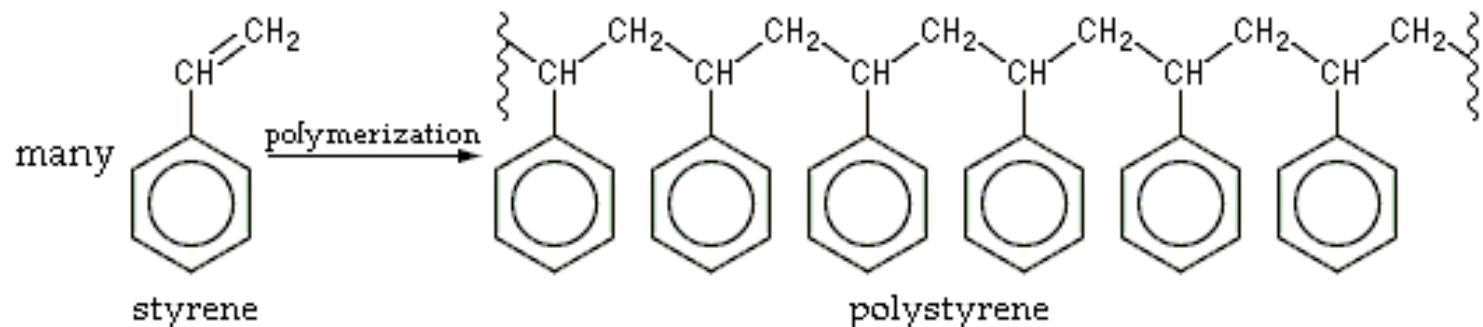
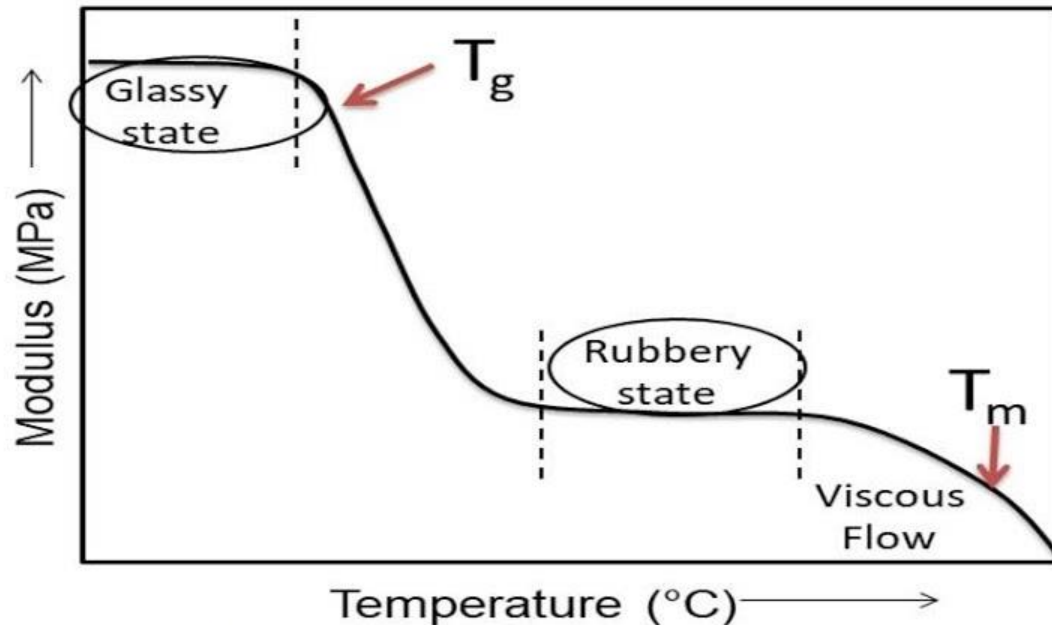
Glass Transition Temperature, T_g

- The glass transition temperature, often called T_g (or “ $T_{sub\ g}$ ”), is an important property when considering polymers for a particular end-use.
- The glass transition temperature is the temperature below which the physical properties of plastics change in a manner similar to those of a glassy or crystalline state, and above which they behave like rubbery materials.
- A plastic's T_g , is the temperature below which molecules have little relative mobility.
- T_g is usually applicable to wholly or partially amorphous plastics.
- A plastic's properties can be dramatically different above and below its T_g .

The temperature at which the polymers undergo the transition from glassy to rubbery state

Above T_g : Polymers are soft and flexible like rubber due to some mobility

Above T_g : Physical and mechanical properties of polymer change



Five Regions of Viscoelastic behavior

Viscoelasticity = (Viscous + Elastic) behavior

Phase I : Glassy region

- ✓ Relatively high modulus
- ✓ Very hard
- ✓ High resistance to flow

Phase II : Leathery/Glass-transition region

- ✓ Sharp decrease in the elastic modulus
- ✓ Deformation not totally recoverable

Phase III : Rubbery region

- ✓ both elastic and viscous components present
- ✓ Modulus falling rate stabilizes
- ✓ Elastic – high strain rate
- ✓ Viscous – low strain rate

Phase IV : Rubbery flow region

- ✓ Viscosity starts dominating
- ✓ Modulus starts falling

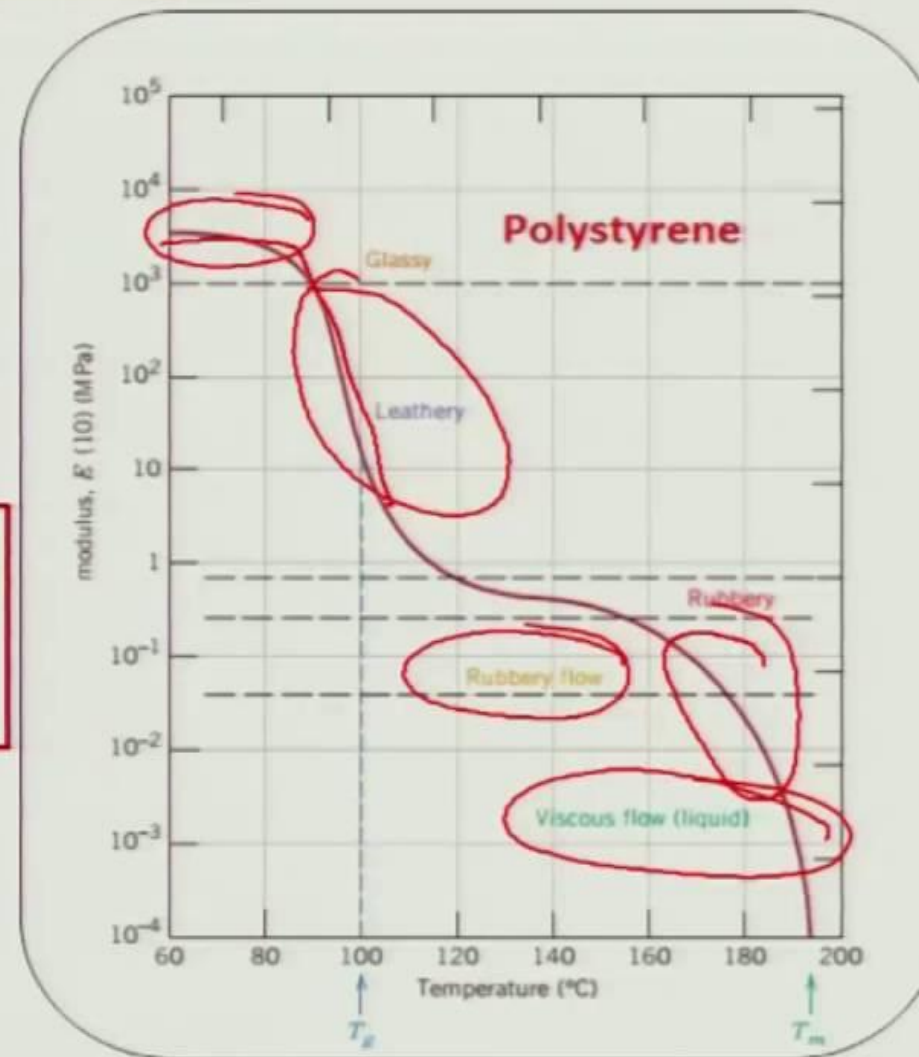
Phase V : Viscous flow region

- ✓ Modulus drops steeply

Finally - decomposition!



Smart Materials Structures and Systems
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Glass Transition Video

- <https://www.youtube.com/watch?v=fVQ9PbIbSlo>
- <https://www.youtube.com/watch?v=3AtwffJS7Lw&t=20s>
- <https://www.youtube.com/watch?v=7UUGaswIXDw>
- <https://www.youtube.com/watch?v=Q1VtAXeMn74>

Molecular Mass of Polymer

- The molecular mass of polymer is an important property of polymer because many important properties are influenced by molecular mass. Polymers with higher molecular mass are tougher and more resistant. Their viscosities and softening temperature are also higher. Thus polymers with molecular mass are often required for particular purposes.
- Molecular mass of polymer is not a fixed or constant value like organic compound. Their molecular mass is controlled by polymerization reaction, which in turn depend upon availability of functional group, charge carrier, life time of charge carriers. Because of random nature of growth process, the product of polymerization process is mixture of chains of different length.
- Hence polymers are poly disperse mixture of various molecular mass polymers. Therefore molecular mass of polymers is average molecular mass.

Average molecular mass of polymers can be expressed in following ways

-

1. Number Average Molecular Mass (\bar{M}_n)
2. Weight Average Molecular Mass (\bar{M}_w)
3. Z-Average Molecular Mass (\bar{M}_z)
4. Viscosity average Molecular Mass (\bar{M}_v)

Methods for determining the molecular weights of polymer

A) Primary or Absolute methods, which are capable of determining molecular weights from first principles like colligative property measurements (Which gives M_n) , light scattering measurements (which gives M_w)

B) Secondary or relative methods, Which require calibration with samples of known molar mass. Examples Viscosity measurements which gives M_v .

Number average Molecular Mass (M_n)

In this number of molecules are involved in the molecular weight measurements. Hence the molecular weight is obtained by colligative property measurements and known as number average molecular weight.

Let us consider N_1, N_2, N_3, \dots be the number of molecules of polymer in a sample solution. And M_1, M_2, M_3, \dots be the their molecular weight then Number average molecular weight is given by

$$\overline{M}_n = \frac{\sum_{i=1}^N N_i M_i}{\sum_{i=1}^N N_i} = \frac{\sum_{i=1}^N w_i}{\sum_{i=1}^N w_i / M_i}$$

Polymer solution of polyethene	Number of molecules	Their molecular weight
Sample has	100 N_1	150 M_1
	200 N_2	200 M_2
	300 N_3	300 M_3

Weight Average molecular Weight

- Let us consider N_1, N_2, N_3, \dots be the number of molecules of polymer in a sample solution.
- And M_1, M_2, M_3, \dots be the their molecular weight
- And W_1, W_2, W_3, \dots be the their weight(mass) then Weight average molecular weight is given by

$$\overline{M}_w = \frac{\sum_{i=1}^N N_i M_i^2}{\sum_{i=1}^N N_i M_i} = \frac{\sum_{i=1}^N w_i M_i}{\sum_{i=1}^N w_i} :$$

Polymer solution of polyethene	Molecular weight	Their weight
Sample has	100	15g
	200	20g
	300	30g

Z-average molecular Weight

Let us consider N_1, N_2, N_3, \dots be the number of molecules of polymer in a sample solution.

And M_1, M_2, M_3, \dots be the their molecular weight, then Z-average molecular weight is given by

$$M_z = \frac{\sum_i N_i M_i^3}{\sum_i N_i M_i^2}$$

Viscosity-average molecular weight (M_v)

- viscosity, like light scattering, is greater for the larger-sized polymer molecules than the smaller ones, and is much closer to M_w than M_n

$$\overline{M}_v = \left[\sum_{i=1}^{\infty} w_i M_i^a \right]^{1/a} = \left[\frac{\sum_{i=1}^{\infty} N_i M_i^{a+1}}{\sum_{i=1}^{\infty} N_i M_i} \right]^{1/a}$$

w = weight fraction
 N = number of moles
 M = molecular weight
 a = A constant

- When $a = 1$, $M_v = M_w$, usually $a \sim 0.5-0.9$
- a is a measure of the the hydrodynamic volume of the polymer
- varies with polymer, solvent and temperature

Mark-Houwink equation

- Mark-Houwink equation is valid for **broad** molecular weight distribution:

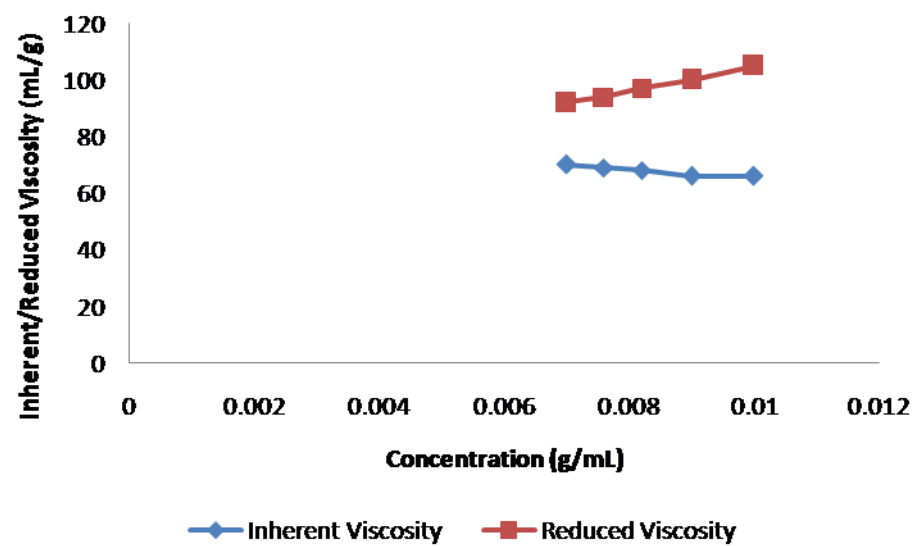
$$[\eta] = K_m \overline{M}_v^a$$

$[\eta]$ = intrinsic viscosity

\overline{M}_v = viscosity average molecular weight

K_m, a = constants for particular polymer-solvent combination

For flexible polymer chains a is 0.5 - 0.8, for stiff and rod like chains a is 2.0.



Degree of Polymerization

- It represents the average number of monomer unit in a polymer chain.

- $\overline{DP_n} = \overline{M_n} / M_0$

- $\overline{DP_w} = \overline{M_w} / M_0$

Polydispersity Index

- It is used as a measure of molecular weight distribution and is defined as
- $$PDI = \overline{M}_w / \overline{M}_n$$
- In case of monodisperse systems (Natural polymers and synthetic polymers made by anionic polymerization) , PDI =1 Unity since $\overline{M}_n = \overline{M}_w$
- In most other cases, $PDI > 1$ or $\overline{M}_w > \overline{M}_n$

Numerical 1

- In a polymer, there are 100 molecules of molecular weight 100, 200 molecules of molecular weight 1000 and 300 molecules of molecular weight 10000. Find M_n , M_w and PDI.
- $$M_n = \frac{N_1 M_1}{N_1} + \frac{N_2 M_2}{N_2} + \frac{N_3 M_3}{N_3}$$
-

Answer 1

$$\text{Solution. } \bar{M}_n = \frac{\sum N_i M_i}{\sum N_i} = \frac{N_1 M_1 + N_2 M_2 + N_3 M_3}{N_1 + N_2 + N_3}$$
$$= \frac{100 \times 100 + 200 \times 1000 + 300 \times 10,000}{100 + 200 + 300}$$

$$= \frac{3.21 \times 10^6}{600} = 5.35 \times 10^3$$

$$\bar{M}_w = \frac{\sum N_i M_i^2}{\sum N_i M_i} = \frac{100 \times (100)^2 + 200 \times (1000)^2 + 300 \times (10,000)^2}{100 \times 100 + 200 \times 1000 + 300 \times 10,000}$$
$$= 9.4 \times 10^3$$

$$\text{PDI} = \frac{\bar{M}_w}{\bar{M}_n} = 1.757$$

Numerical 2

A polydisperse sample of polystyrene is prepared by mixing three monodisperse samples in the following proportions:

1 g, 10,000 molecular weight

2 g, 50,000 molecular weight

2 g, 100,000 molecular weight

Using this information, calculate the number-average molecular weight, weight-average molecular weight, and PDI of the mixture.

Answer

$$\bar{M}_n = \frac{\sum_{i=1}^3 N_i M_i}{\sum_{i=1}^3 N_i} = \frac{\sum_{i=1}^3 W_i}{\sum_{i=1}^3 (W_i / M_i)} = \frac{1 + 2 + 2}{\frac{1}{10,000} + \frac{2}{50,000} + \frac{2}{100,000}} = 31,250$$

$$\bar{M}_w = \frac{\sum_{i=1}^3 N_i M_i^2}{\sum_{i=1}^3 N_i M_i} = \frac{\sum_{i=1}^3 W_i M_i}{\sum_{i=1}^3 W_i} = \frac{10,000 + 2(50,000) + 2(100,000)}{5} = 62,000$$

$$\text{PDI} = \frac{\bar{M}_w}{\bar{M}_n} = \frac{62,000}{31,250} = 1.98$$

Numerical 3

- An equimolecular (i.e. equimolar) mixture of samples of molecular mass 50 000 and 100 000, what will be their M_n and M_w .

Answer

$$\begin{aligned}\overline{M}_n &= \frac{(0.5 \times 50\,000) + (0.5 \times 100\,000)}{1.0} \\ &= 75\,000\end{aligned}$$

$$\begin{aligned}\overline{M}_w &= \frac{(0.5 \times 50\,000^2) + (0.5 \times 100\,000^2)}{75\,000} \\ &= 83\,000\end{aligned}$$

Numerical 4

- Consider a polymer sample comprising of 5 moles of polymer molecules having molecular weight of 40,000 g/mol and 15 moles of polymer molecules having molecular weight of 30,000 g/mol.

Calculate M_n and M_w

- Ans: $M_n = 32500$ and $M_w = 33076.9$

Numerical 5

Example 4. 42 gm of propene was polymerized by radical polymerization process and \overline{DP} was found to be 1000. Calculate the number of molecules of PP produced.

Solution. As \overline{DP} of PP = $\frac{\text{Numbers of Propene molecules}}{\text{Number of PP molecules formed}}$

Hence, number of PP molecules formed

$$= \frac{\text{Number of propene molecules}}{\overline{DP} \text{ of PP}}$$

$$= \frac{42 \text{ gm} \times (6.023 \times 10^{23} \text{ molecules/42 gm})}{1000}$$

$$= 6.023 \times 10^{20} \text{ molecules}$$

Numerical 6

Example 1. A polymer sample consists of 10% by weight of macromolecules of molecular weight 10,000 and 90% by weight of macromolecules with molecular weight 100,000. Calculate \overline{M}_n and \overline{M}_w .

Solution. As
$$\overline{M}_n = \frac{\sum N_i M_i}{\sum N_i} = \frac{\sum W_i}{\sum N_i}$$

where W_i = weight of i^{th} constituent,

N_i = number of moles of i^{th} constituent,

and M_i = molecular weight of i^{th} constituent

Given $W_1 = 10$ gms. and $W_2 = 90$ gm.

$$\therefore \sum W_i = W_1 + W_2 = 10 + 90 = 100 \text{ gm.}$$

Numerical 7

- Find M_w of polypropylene given its degree of polymerization as 10000.
- Propylene $\text{CH}_3\text{CH}=\text{CH}_2$ = 42 mol wt

- **Numerical 8**

- Calculate the M_n M_w and PDI if the polymer mixture contains 25% polymer with Mol wt 16800, 35% polymer with Mol. Wt. 33600 and 40% polymer with mol. Wt 25200.

- **Numerical 10**
- The M_n of polystyrene of 10^5 g/mol. Find its DP_n
- Styrene $C_6H_5CH=CH_2$

Numerical 9

Calculate M_n and M_w if a mixture of polymer has composition 10 % of polymer with mol wt 10000, 50 % of polymer with mol wt 15000 and remaining of polymer with mol wt 20000.

$$N_1 = 10$$

Numerical 11

- Calculate the M_n and M_w based on following data
- A) $[-CH_2-CH(CH_3)-]_{400}$ is 25%
- B) $[-CH_2-CH(CH_3)-]_{800}$ is 40%
- C) $[-CH_2-CH(CH_3)-]_{600}$ is 35%
- $M_1 = 400 \times 42 = 16800$ $N_1 = 25$
- $M_2 = 800 \times 42 = 33600$ $N_2 = 40$
- $M_3 = 600 \times 42 = 25200$ $N_3 = 35$

Since $N_i = \frac{W_i}{M_i}$

Hence $N_1 = \frac{10}{10,000}$ and $N_2 = \frac{90}{100,000}$

Now $\bar{M}_n = \frac{\sum W_i}{\sum N_i} = \frac{W_1 + W_2}{N_1 + N_2} = \frac{10 + 90}{\frac{10}{10,000} + \frac{90}{100,000}} = 5.26 \times 10^4$

Similarly, $\bar{M}_w = \frac{\sum N_i M_i^2}{\sum N_i M_i} = \frac{N_1 M_1^2 + N_2 M_2^2}{N_1 M_1 + N_2 M_2} = \frac{N_1 M_1^2 + N_2 M_2^2}{W_1 + W_2}$

$$= \frac{\frac{10}{10,000} \times (10,000)^2 + \frac{90}{100,000} (100,000)^2}{10 + 90}$$

$$= 9.1 \times 10^4$$

Compounding of Plastics

- In order to impart certain definite properties to the finished products, resins are compounded with certain other substances. The process is called as compounding of plastics.

1. Binders or Resins

- The product of polymerization is called resins and this forms the major portion of the body of plastics.
- It is the binder, which holds the different constituents together.
- Thermosetting resins are usually, supplied as linear – polymers of comparatively low molecular weight, because at this stage they are fusible and hence, mouldable.
- The conversion of this fusible form into cross-linked infusible form takes place, during moulding itself, in presence of catalysts etc.
- A binder may compose of 30-100%.
- The binders used may be natural or synthetic or cellulose derivatives.

2. Fillers

- Fillers are generally added to thermosetting plastics to increase elasticity and crack resistance.
- Fillers improve thermal stability, strength, non combustibility, water resistance, electrical insulation properties & external appearance.
- E.g.: - Mica(Hardness), cotton(Shock resistance) , carbon(electrical conductivity) black, graphite, BaSO₄ etc. Asbestos (Heat resistance)

3. Plasticizers

- Plasticizers are substances added to enhance the plasticity of the material and to reduce the cracking on the surface.
- Plasticizers are added to the plastics to increase the flexibility & toughness. Plasticizers also increase the flow property of the plastics.
- Commonly used plasticizers are vegetable oils, camphor, esters of steric, oleic acid, tributyl phosphate, triphenyl phosphate etc.

4. Dyes and Pigments

- These are added to impart the desired colour to the plastics and give decorative effect.
- e.g.: -Inorganic Pigments: Lead chromate (yellow), ferro cyanide (blue)
- Organic Pigments: Alizarin Red, Indigo Dye, Azo dyes.

5. Lubricants

- Includes Oils, waxes and soaps.
- They help in easy moulding and glossy finish to final product.
- The lubricant prevent the plastic materials from sticking to the mould.

6. Catalysts

- They are added to plastics to accelerate polymerizations reaction.
- ZnCl_2 , CaO , ammonia, benzoyl peroxides

7. Stabilizers

- They are added to improve thermal stability during processing. Eg. Polymers of vinyl chloride show a tendency to undergo decomposition at moulding temperatures. So during their moulding heat stabilizers are added.
- i) Opaque moulding stabilizers: Salts of lead, like white lead, lead chromate, red lead, lead silicate etc.
- li) Transparent Moulding stabilizers: Stearates of Lead, chromium and barium.

- In addition small quantities of antioxidants, antiseptics, fungicides, flame retardants etc. are added as ingredient of plastics.

FABRICATION OF PLASTICS

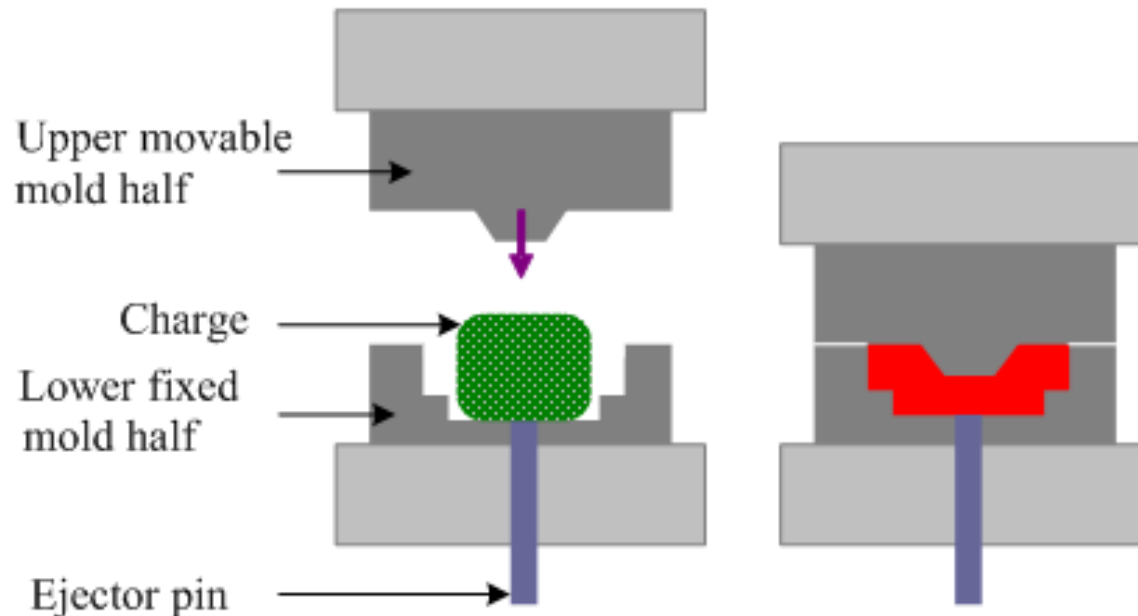
- Giving any desired shape to the plastics (granules or powders) by using mould under the application of heat and pressure.
 - A proper method is to be selected depending on the shape and type of resin used.
 - Methods involves partial melting of resinous mass by heating.
 - In case of thermo-plasts molten resin is introduced in die/mould and desired shape could be achieved by compression and further cooling.
 - In case of thermo-sets partially polymerized mass or raw materials are introduced in the die/mould, which further cured at high temperature in the mould itself to achieve desired shape.
1. In case of thermoplasts, curing is done at room temperature (low temperature)
 2. In case of thermosets, curing is done at high temperature to obtain desired cross-linking

TYPES OF FABRICATION

- [I} Compression Molding (Suitable for Thermosets / Thermoplasts)
- [II] Injection Molding (Suitable for Thermoplasts)
- [III] Transfer Molding (Suitable for Thermosets)
- [IV] Extrusion Molding (Suitable for Thermoplasts)

COMPRESSION MOULDING

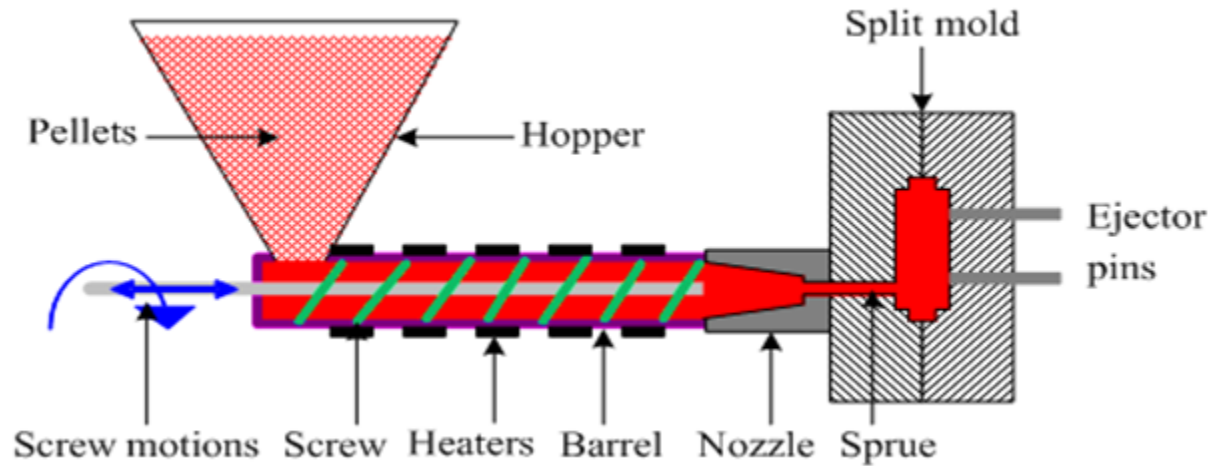
1. Common and oldest method for molding thermosetting / thermoplastic materials
2. Compression of raw materials or soften resinous mass is done in the mould /die under heat and pressure



COMPRESSION MOULDING

1. Predetermined quantity of raw materials is introduced carefully in the mould, further compressed by hydraulic pressure (2000 to 10000 psi)
2. Molten or soften resinous mass gets filled in the cavity of mould.
3. Curing is done by heating (Thermosetting) or by cooling (Thermoplastics)
4. Finally moulded article is separated from the mould by opening the mould apart.
5. Applications : Electric switch boxes, Ash trays, cabinets for radio, television, computers etc.

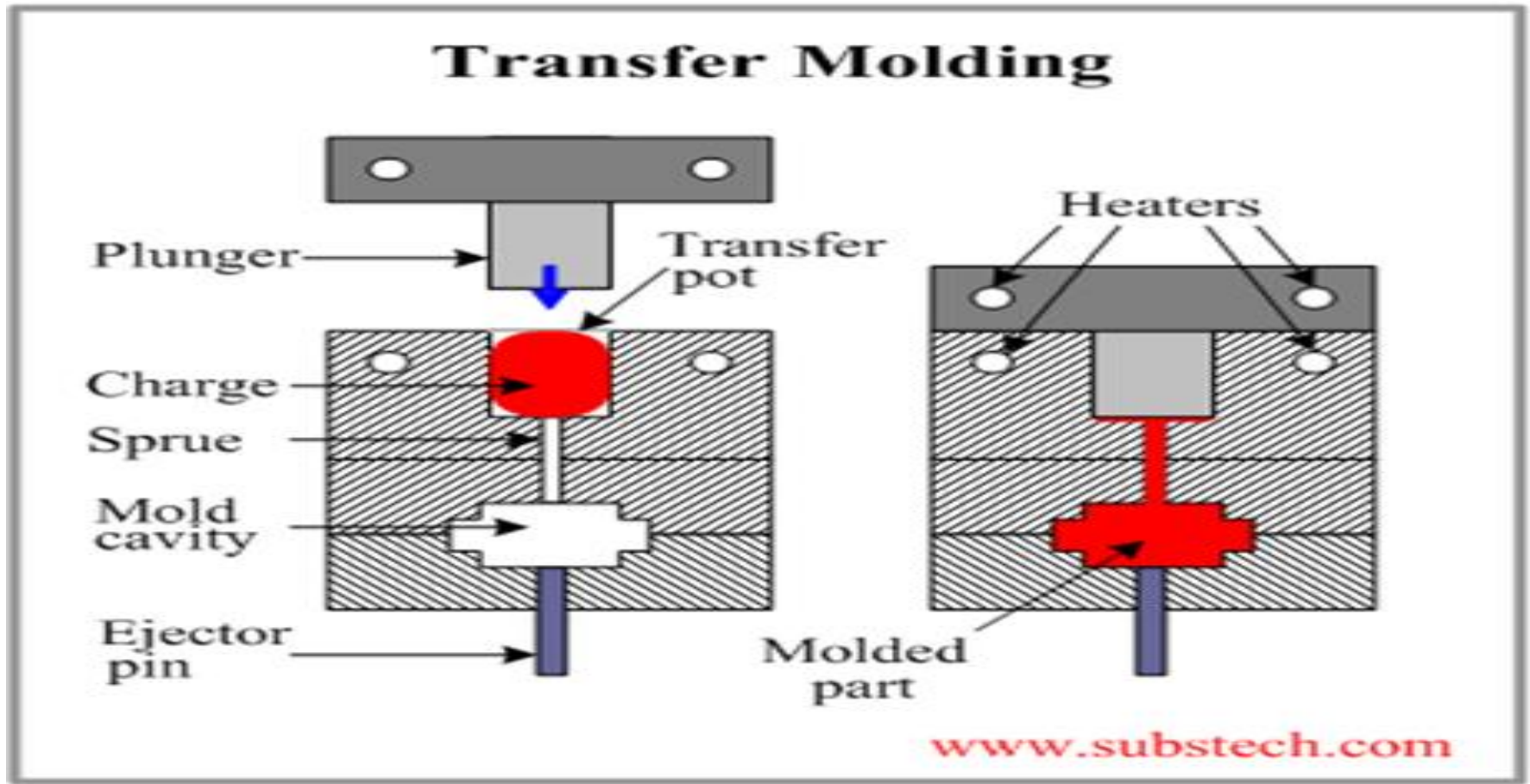
INJECTION MOULDING



INJECTION MOULDING

- This method is especially used for thermoplastic materials
- Powder or granular resin is heated in a cylinder and injected at a controlled rate in a mould with the help of piston plunger or screw.
- Piston plunger or screw is used to force the material in mould.
- Pressure up to 2000 kg/cm² (100 psi) is used
- Once the article is formed mould is cooled and half mould is opened to remove the finished article.
- Disadvantage of the method is formation of air bubbles or cavities in the articles
- Applications: Smaller but large volume articles such as, pen caps, bottle caps, cups, containers, mechanical parts

TRANSFER MOULDING



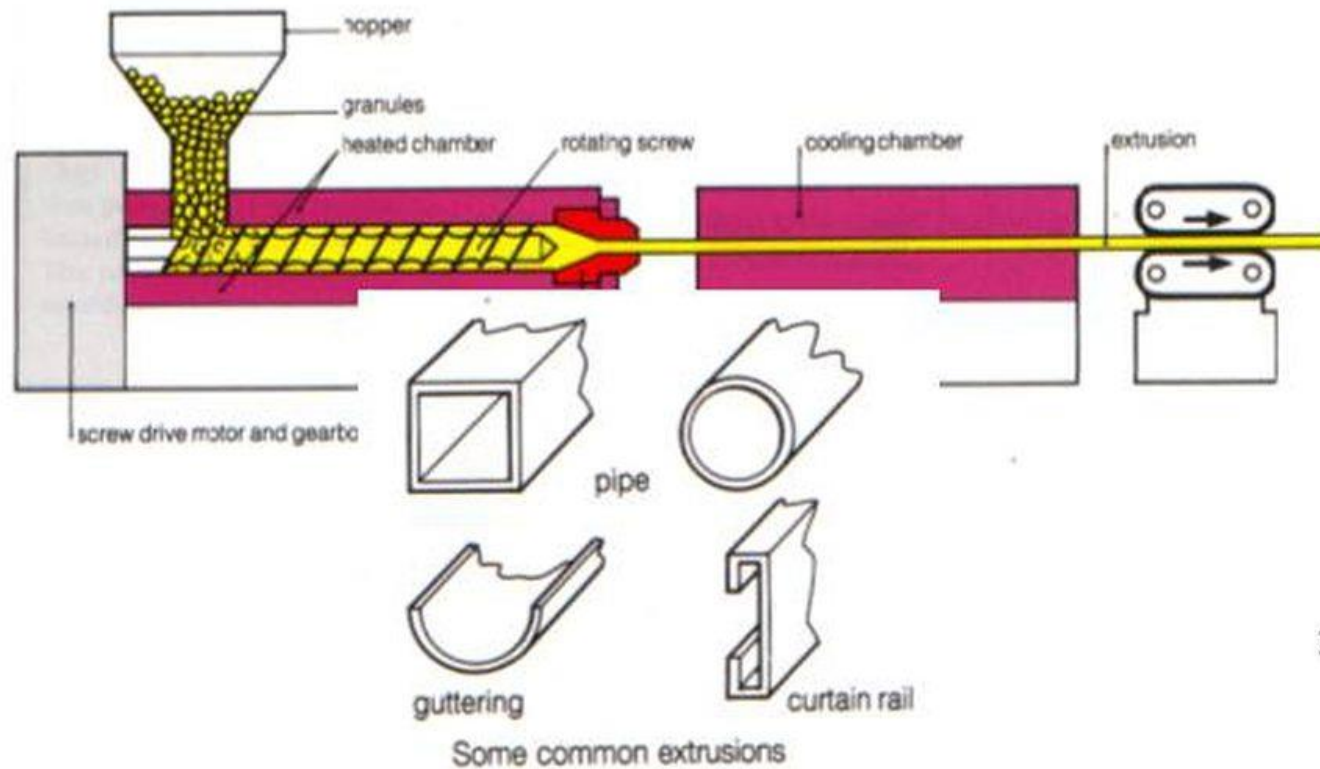
TRANSFER MOULDING

- The method combines features of both Compression Molding (hydraulic pressing of molding materials - thermosets) and Injection Molding (ram-plunger and filling the mold through a sprue).
- The method is used especially for molding thermosetting resins (thermosets)
- Products with relatively intricate designs could be fabricated with this method
- Powdered raw materials are heated at certain low temperature to soften and then introduced through an orifice or sprue in the mould
- Then it is cured in the mould at high temperature for certain time
- Finally the moulded article is removed by separation of mould

Advantages:

- Articles with intricate shapes could be designed
- Aerospace and automobile parts, car body, helmets
- The articles produced are blister free
- Fine wires and glass fibers can be inserted in the mould
- Even thick pieces can be cured completely and uniformly

Extrusion Molding



Extrusion Moulding

- Whenever continuous molding of material like wires, cables, and sheets is required extrusion molding is used. The thermoplastic materials are molded by this method. They undergo continuous molding to form articles of uniform cross section. In this method, the thermoplastic material are heated to plastic condition and pushed by means of screw conveyor in to a mould cavity having required outer shape of articles to be manufactured. Here the plastic mass gets cooled due to atmospheric exposure. A long conveyor carries away the cooled products continuously.

Preparation, Properties and uses of IMPORTANT POLYMERS:

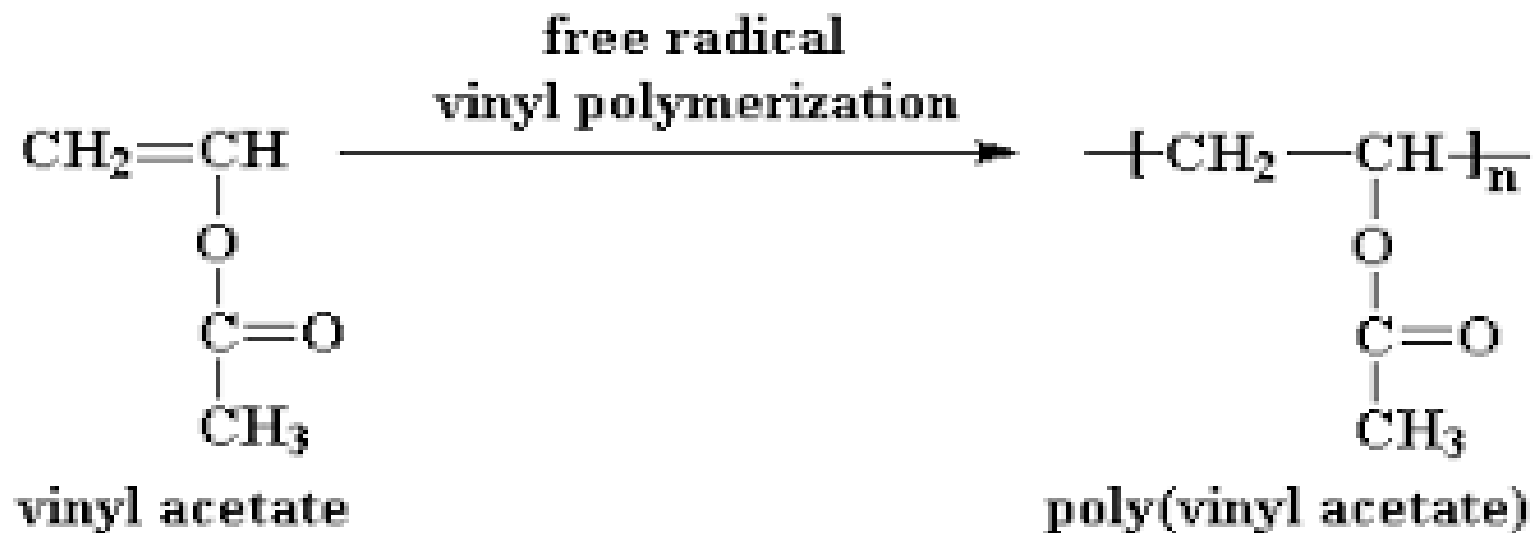
- POLY VINYL ACETATE (PVAc):

Polyvinyl acetate is a substance produced through the combining of many units of monomeric vinyl acetate ($\text{CH}_3\text{COOCH}=\text{CH}_2$). The number of units so combined is typically between 100 and 5,000. This translates to an average molecular weight of between 850 and 40,000. Polyvinyl acetate may be used as is or modified through chemical reactions to produce other important polymeric substances

- Preparation: Fe^{2+} (Ferrous), Fe^{3+} (Ferric), Hg^{+1} (Mercurous), Hg^{2+} (Mercuric)

Monomeric vinyl acetate was once prepared by reacting acetylene (C_2H_2) with anhydrous acetic acid in the presence of a mercurous sulfate (Hg_2SO_4) catalyst; poly vinyl acetate is prepared by free radical vinyl polymerization as follows

- $\text{C}_2\text{H}_2 + \text{CH}_3\text{COOH} \rightarrow \text{CH}_3\text{COOCH}=\text{CH}_2$



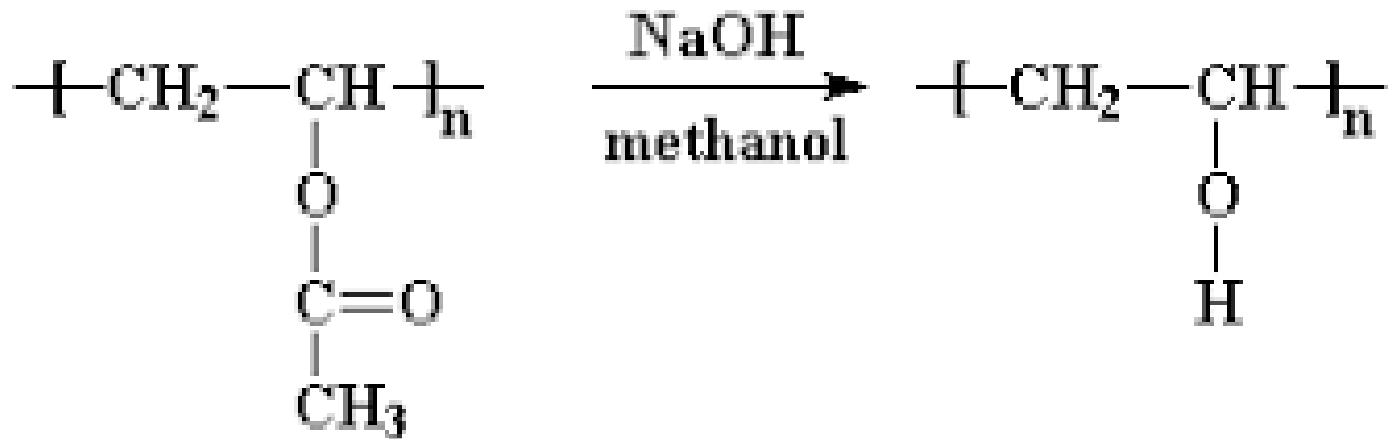
- **Properties:**

- Polyvinyl acetate is an amorphous polymer, not a crystalline one.
- It is Clear colourless transparent material.
- The hardest of the polyvinyl esters, polyvinyl acetate offers good adhesion to most surfaces.
- Unlike some other thermoplastics, it will not turn yellow.
- Polyvinyl acetate does not cross-link, thus becoming insoluble in water, and it can be dissolved in many solvents other than water.
- One slow-drying formulation combines 5 to 15 percent polyvinyl acetate with ethyl alcohol (ethanol).
- A fast-drying counterpart combines the same amount of polyvinyl acetate with acetone (dimethyl ketone).
- Lower molecular weight polymers become gum like and hence used for making chewing gums. Since it harms less when taken orally.

- **USES:**
- Emulsified polyvinyl acetate is used in water-based adhesives, including pastes and glues.
- One of the uses for emulsified polyvinyl acetate is in bookbinding.
- Depending upon the necessary lifetime of the book, the polyvinyl acetate chosen will either be copolymeric or homopolymeric.
- Polyvinyl acetate offers acceptable gap-filling capability. It may be used as a resinous component of latex paints, offering compatibility with a wide-range of other paint chemicals.
- Polyvinyl acetate may be used in the lamination of metal foils. Non-emulsified, or waterless, polyvinyl acetate is useful as a thermosetting adhesive.

Poly Vinyl Alcohol (PVA):

- Polyvinyl alcohol, also known as PVOH, PVA, or PVAL, is a synthetic polymer that is soluble in water. It is effective in film forming, emulsifying, and has an adhesive quality. It has no odor and is not toxic, and is resistant to grease, oils, and solvents. It is ductile but strong, flexible, and functions as a high oxygen and aroma barrier.
- **Preparation:**
 - Unlike most vinyl polymer, PVA is not prepared by polymerization of the corresponding monomer
 - Polyvinyl alcohol was first prepared in 1924 by hydrolyzing polyvinyl acetate in ethanol with potassium hydroxide.
 - Polyvinyl alcohol is produced commercially from polyvinyl acetate, usually by a continuous process. The acetate groups are hydrolyzed by ester interchange with methanol in the presence of anhydrous sodium methylate or aqueous sodium hydroxide. The physical characteristics and its specific functional uses depend on the degree of polymerization and the degree of hydrolysis. Polyvinyl alcohol is classified into two classes namely: partially hydrolyzed and fully hydrolyzed. Partially hydrolyzed PVA is used in the foods.
 - $\text{CH}_2=\text{CH}-\text{OH}$ CH_3CHO
 - Vinyl alcohol Acetaldehyde (Tautomerism)



poly(vinyl acetate)

poly(vinyl alcohol)

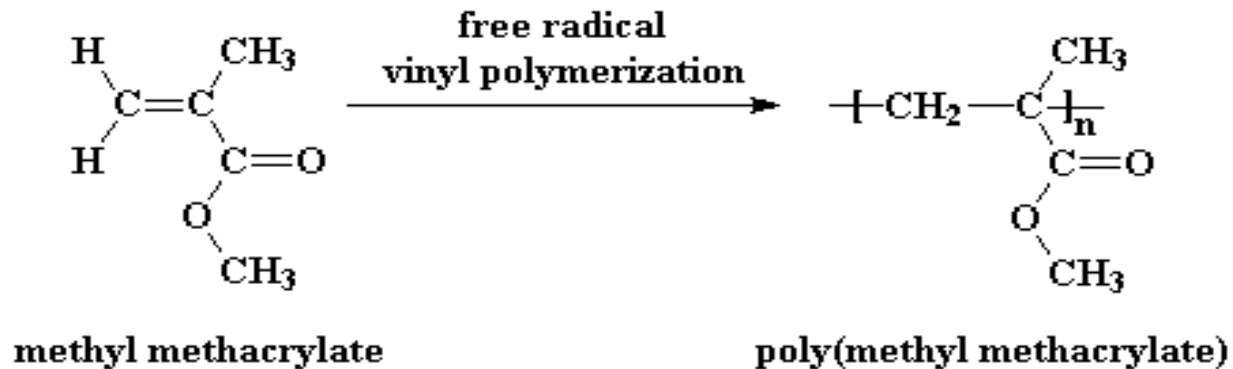
- PVA is an atactic material that exhibits crystallinity
- Polyvinyl alcohol has excellent film forming, emulsifying and adhesive properties.
- It is also resistant to oil, grease and solvents.
- It has high tensile strength and flexibility, as well as high oxygen and aroma barrier properties.
- However, these properties are dependent on humidity: water absorbed at higher humidity levels acts as a plasticiser, which reduces the polymer's tensile strength, but increases its elongation and tear strength.

Uses.

- PVA is used in a variety of medical applications because of its biocompatibility, low tendency for protein adhesion, and low toxicity. Specific uses include cartilage replacements, contact lenses, and eye drops.
- Polyvinyl alcohol is used as an aid in suspension polymerizations
- Nowadays PVA-based polymers are being used widely in additive manufacturing. For example, 3D printed oral dosage forms demonstrate great potential in the pharmaceutical industry.
- Medically, it may also be used as the embolic agent in a Uterine Fibroid Embolization (UFE).

POLYMETHYL METHACRYLATE (PMMA)

- Poly(methyl methacrylate) (PMMA), also known as acrylic or acrylic glass as well as by the trade names Crylux, Plexiglas, Acrylite, Lucite, and Perspex among several others (see below), is a transparent thermoplastic often used in sheet form as a lightweight or shatter-resistant alternative to glass. The same material can be used as a casting resin, in inks and coatings, and has many other uses.
- Preparation:
- PMMA is routinely produced by emulsion polymerization, solution polymerization, and bulk polymerization. Generally, radical initiation is used.



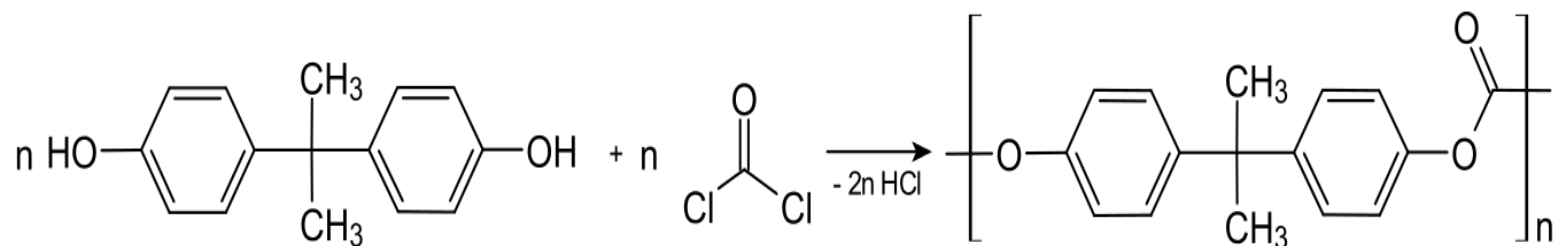
- **Properties:**
- **CH₂=CH-COOH Acrylic Acid.**
- **CH₂=CH-COOCH₃ Methyl Acrylate**
- **CH₂=C(CH₃)-COOCH₃ Methyl methacrylate**
- PMMA is a strong, tough, and lightweight material. It has a density of 1.17–1.20 g/cm³, which is less than half that of glass.
- It also has good impact strength, higher than both glass and polystyrene; however, PMMA's impact strength is still significantly lower than polycarbonate
- PMMA transmits up to 92% of visible light (3 mm thickness), and gives a reflection of about 4% from each of its surfaces due to its refractive index (1.4905 at 589.3 nm).
- It filters ultraviolet (UV) light at wavelengths below about 300 nm
- PMMA swells and dissolves in many organic solvents; it also has poor resistance to many other chemicals due to its easily hydrolyzed ester

Uses

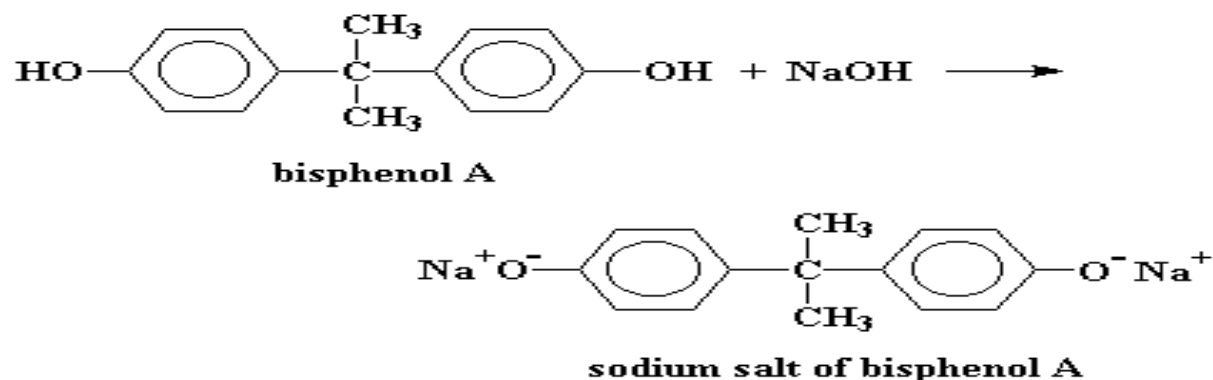
- Being transparent and durable, PMMA is a versatile material and has been used in a wide range of fields and applications such as rear-lights and instrument clusters for vehicles, appliances, and lenses for glasses.
- PMMA in the form of sheets affords to shatter resistant panels for building windows, skylights, bulletproof security barriers, signs & displays, sanitary ware (bathtubs), LCD screens, furniture and many other applications.
- It is also used for coating polymers based on PMMA provides outstanding stability against environmental conditions with reduced emission of VOC.
- Methacrylate polymers are used extensively in medical and dental applications where purity and stability are critical to performance.

POLY CARBONATES (PC)

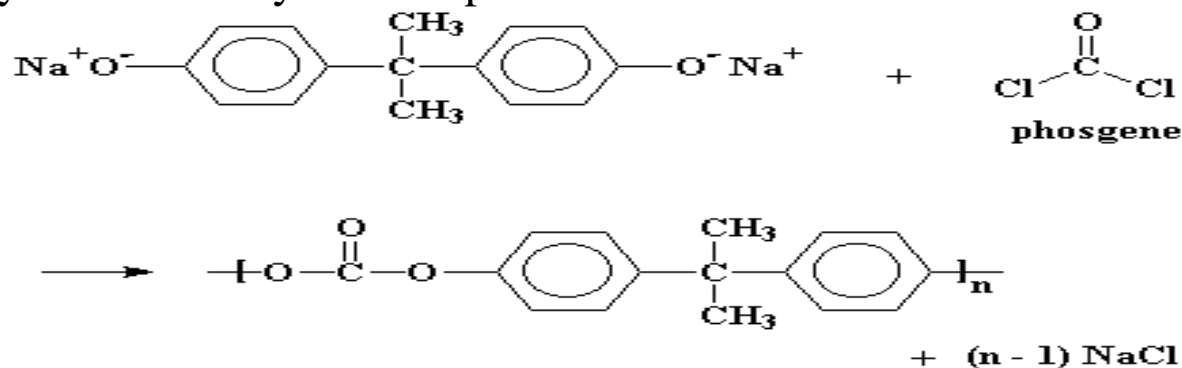
- Polycarbonates (PC) are a group of thermoplastic polymers containing carbonate groups in their chemical structures. Polycarbonates used in engineering are strong, tough materials, and some grades are optically transparent. They are easily worked, molded, and thermoformed.
- CO₃
- Preparation:
 1. The main polycarbonate material is produced by the reaction of bisphenol A (BPA) and phosgene COCl₂
 2. The overall reaction can be written as follows:



- The first step of the synthesis involves treatment of bisphenol A with sodium hydroxide, which deprotonates the hydroxyl groups of the bisphenol A



The diphenoxide ($\text{Na}_2(\text{OC}_6\text{H}_4)_2\text{CMe}_2$) reacts with phosgene to give a chloroformate, which subsequently is attacked by another phenoxide. The net reaction from the diphenoxide is:



- **Properties:**

- Polycarbonate is a durable material. Although it has high impact-resistance, it has low scratch-resistance.
- The characteristics of polycarbonate compare to those of polymethyl methacrylate (PMMA, acrylic), but polycarbonate is stronger and will hold up longer to extreme temperature.
- Polycarbonate has a glass transition temperature of about 147 °C (297 °F; 420 K)
- Unlike most thermoplastics, polycarbonate can undergo large plastic deformations without cracking or breaking.

- **Uses:**

- **ELECTRONICS:**

- Polycarbonate is mainly used for electronic applications that capitalize on its collective safety features. Being a good electrical insulator and having heat-resistant and flame-retardant properties, it is used in various products associated with electrical and telecommunications hardware. It can also serve as a dielectric in high-stability capacitors

- **CONSTRUCTION:**

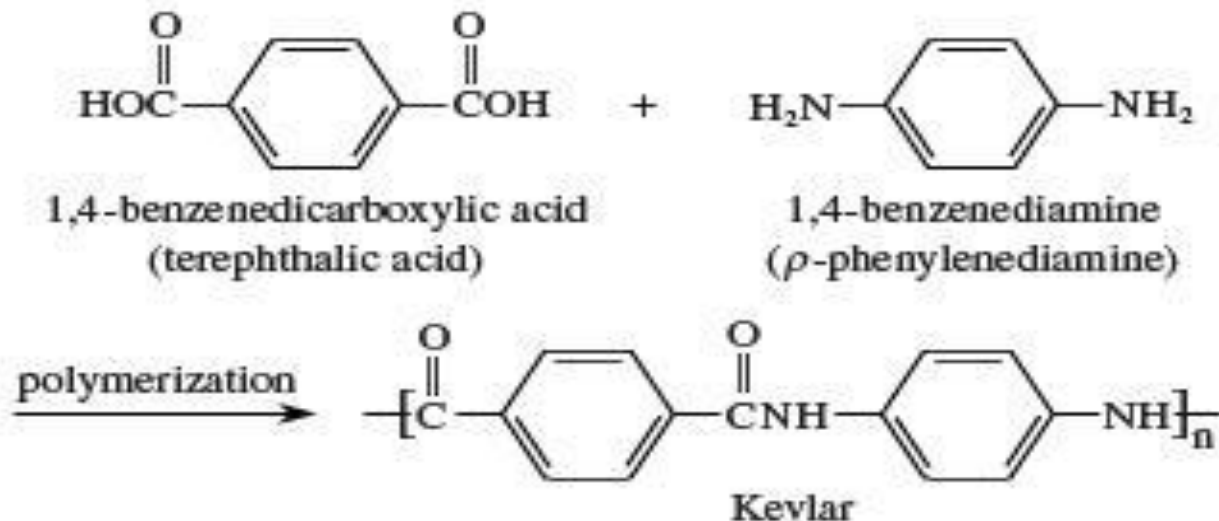
- The second largest consumer of polycarbonates is the construction industry, e.g. for domelights, flat or curved glazing, and sound walls, which all use extruded flat solid or multiwall sheet, or corrugated sheet.

- **DATA STORAGE:**

- A major application of polycarbonate is the production of Compact Discs, DVDs. These discs are produced by injection molding polycarbonate into a mold cavity that has on one side a metal stamper containing a negative image of the disc data, while the other mold side is a mirrored surface.
- Due to its low weight and high impact resistance, polycarbonate is the dominant material for making automotive headlamp lenses.

Poly-Paraphenylene Terephthalamide (KEVLAR):

- Poly-paraphenylene terephthalamide – branded Kevlar is a heat-resistant and strong synthetic fiber,
- Preparation:
- Kevlar is synthesized in solution from the monomers 1,4-phenylene-diamine (para-phenylenediamine) and terephthaloyl chloride in a condensation reaction yielding hydrochloric acid as a byproduct.

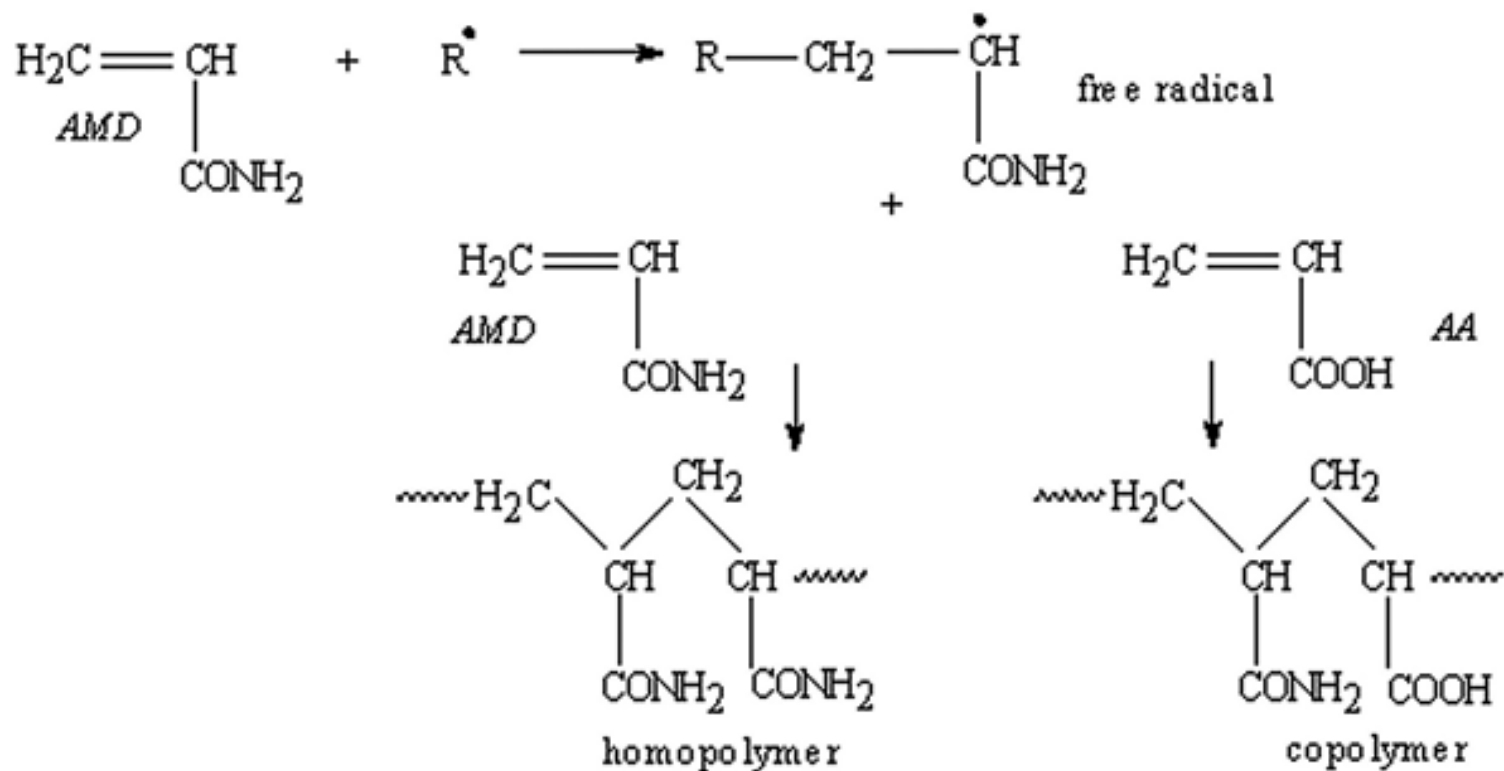


- Properties:
- When Kevlar is spun, the resulting fiber has a tensile strength of about 3,620 MPa,[14] and a relative density of 1.44.
- Kevlar maintains its strength and resilience down to cryogenic temperatures (−196 °C); in fact, it is slightly stronger at low temperatures.
- Kevlar's structure consists of relatively rigid molecules which makes them exceptional strong
- They are high tensile strength-to-weight ratio; by this measure it is 5 times stronger than steel.

Uses

- Kevlar has many applications, ranging from bicycle tires and racing sails to bulletproof vests,
- It is also used to make modern marching drumheads that withstand high impact.
- Kevlar is a well-known component of personal armor such as combat helmets, ballistic face masks, and ballistic vests.
- It is used for motorcycle safety clothing, especially in the areas featuring padding such as shoulders and elbows.
- Kevlar is often used in the field of cryogenics for its low thermal conductivity and high strength relative to other materials for suspension purposes.

Polyacrylamide



- **Properties:**

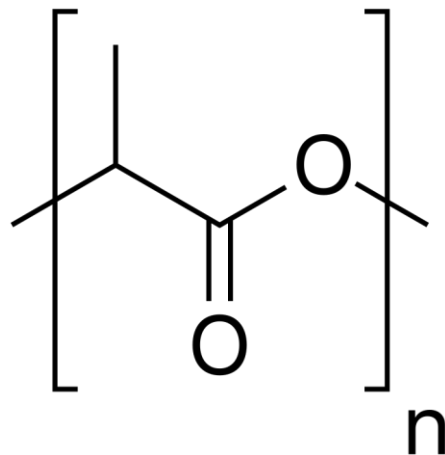
- Linear polyacrylamide is a water-soluble polymer.
- It is typically non-ionic polymer but due to hydrolysis of some amide groups they could convert into carboxylic groups giving polyacrylamide some weak an ionic properties.

- **Uses**

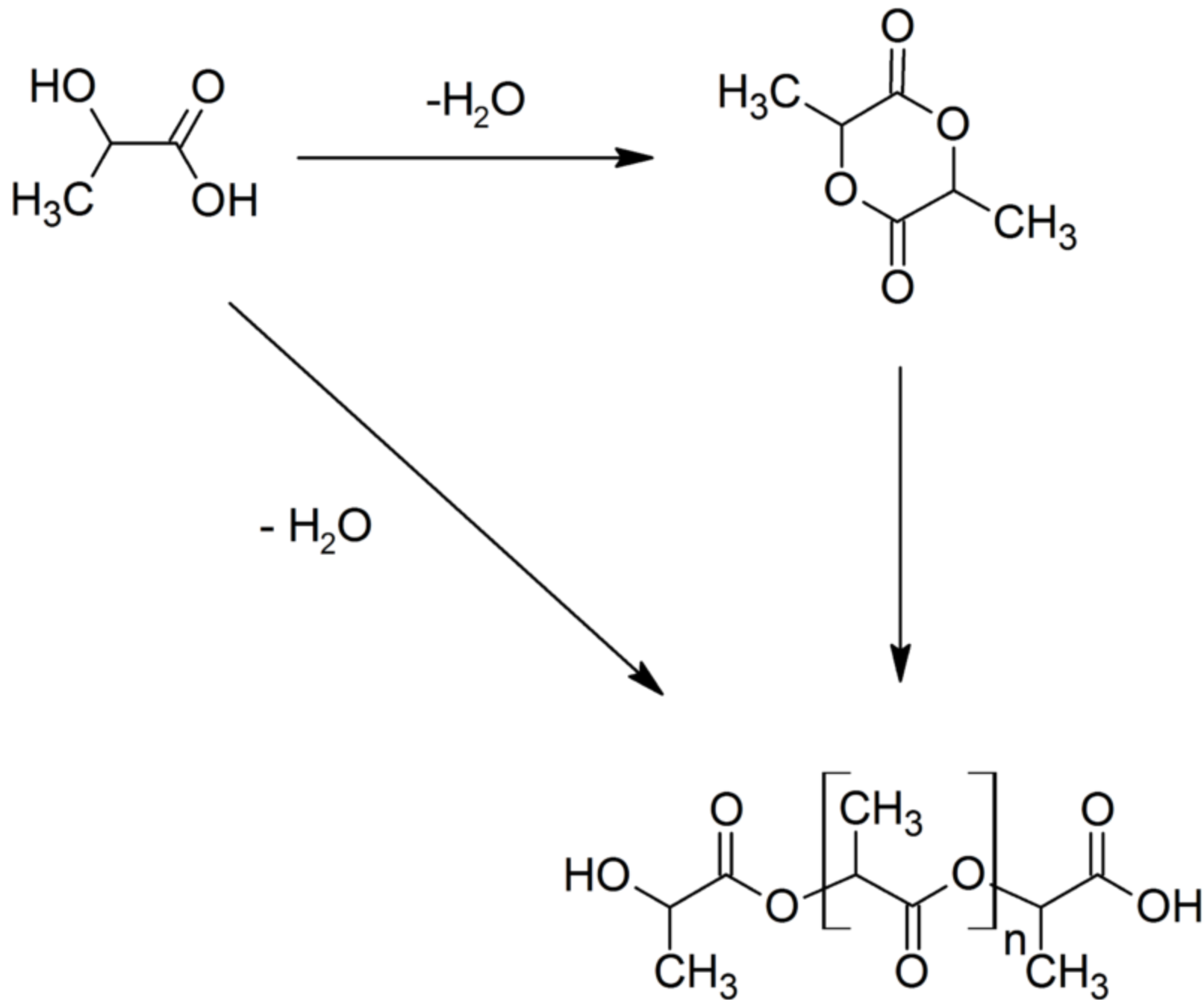
- One of the largest uses for polyacrylamide is to flocculate solids in a liquid. This process applies to water treatment, and processes like paper making and screen printing.
- Polyacrylamide can be supplied in a powder or liquid form, with the liquid form being subcategorized as solution and emulsion polymer. Even though these products are often called 'polyacrylamide', many are actually copolymers of acrylamide
- Another common use of polyacrylamide and its derivatives is in subsurface applications such as Enhanced Oil Recovery.

Polylactic acid

- Polylactic acid, also known as poly(lactic acid) or polylactide (abbreviation PLA) is a thermoplastic polyester with backbone formula $(C_3H_4O_2)_n$ or $[-C(CH_3)HC(=O)O-]_n$, formally obtained by condensation of lactic acid $C(CH_3)(OH)HCOOH$ with loss of water (hence its name). It can also be prepared by ring-opening polymerization of lactide $[-C(CH_3)HC(=O)O-]_2$, the cyclic dimer of the basic repeating unit.



- Synthesis
- The monomer is typically made from fermented plant starch such as from corn, cassava, sugarcane or sugar beet pulp.
- Several industrial routes afford usable (i.e. high molecular weight) PLA. Two main monomers are used: lactic acid, and the cyclic di-ester, lactide. The most common route to PLA is the ring-opening polymerization of lactide with various metal catalysts (typically tin octoate) in solution or as a suspension. The metal-catalyzed reaction tends to cause racemization of the PLA, reducing its stereoregularity compared to the starting material (usually corn starch).[7]
- The direct condensation of lactic acid monomers can also be used to produce PLA. This process needs to be carried out at less than 200 °C; above that temperature, the entropically favored lactide monomer is generated. This reaction generates one equivalent of water for every condensation (esterification) step. The condensation reaction is reversible and subject to equilibrium, so removal of water is required to generate high molecular weight species. Water removal by application of a vacuum or by azeotropic distillation is required to drive the reaction toward polycondensation. Molecular weights of 130 kDa can be obtained this way. Even higher molecular weights can be attained by carefully crystallizing the crude polymer from the melt. Carboxylic acid and alcohol end groups are thus concentrated in the amorphous region of the solid polymer, and so they can react. Molecular weights of 128–152 kDa are obtainable thus.



Properties

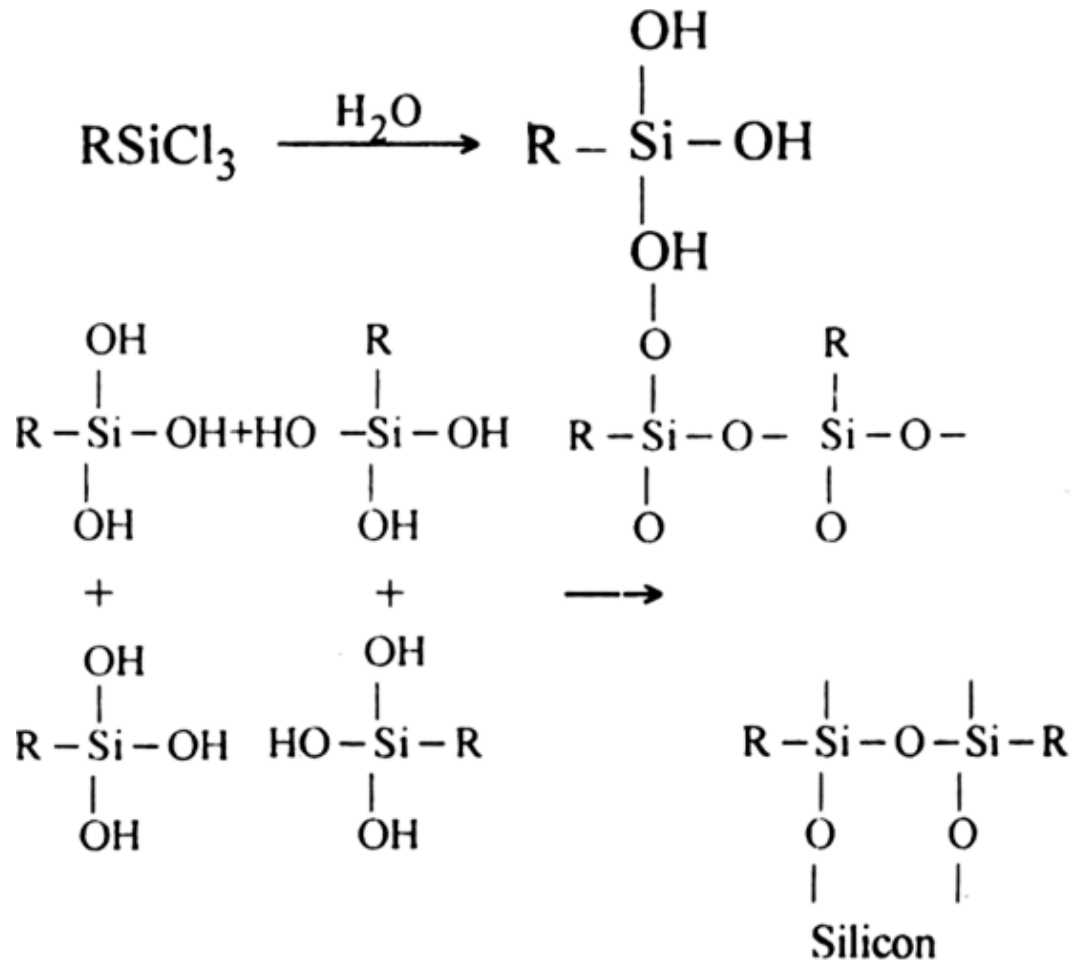
- PLA polymers range from amorphous glassy polymer to semi-crystalline and highly crystalline polymer with a glass transition 60–65 °C, a melting temperature 130-180 °C, and a Young's modulus 2.7–16 GPa.[13][14][15] Heat-resistant PLA can withstand temperatures of 110 °C.[16] The basic mechanical properties of PLA are between those of polystyrene and PET.[13] The melting temperature of PLLA can be increased by 40–50 °C and its heat deflection temperature can be increased from approximately 60 °C to up to 190 °C by physically blending the polymer with PDLA (poly-D-lactide). PDLA and PLLA form a highly regular stereocomplex with increased crystallinity. The temperature stability is maximised when a 1:1 blend is used, but even at lower concentrations of 3–10% of PDLA, there is still a substantial improvement. In the latter case, PDLA acts as a nucleating agent, thereby increasing the crystallization rate[citation needed]. Biodegradation of PDLA is slower than for PLA due to the higher crystallinity of PDLA[citation needed]. The flexural modulus of PLA is higher than polystyrene and PLA has good heat sealability.
- Racemic PLA and pure PLLA have low glass transition temperatures, making them undesirable because of low strength and melting point. A stereocomplex of PDLA and PLLA has a higher glass transition temperature, lending it more mechanical strength.

Uses

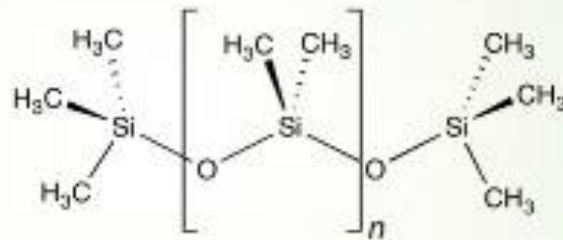
- Consumer goods
- PLA is used in a large variety of consumer products such as disposable tableware, cutlery, housings for kitchen appliances and electronics such as laptops and handheld devices, and microwavable trays. (However, PLA is not suitable for microwavable containers because of its low glass transition temperature.) It is used for compost bags, food packaging and loose-fill packaging material that is cast, injection molded, or spun.[35] In the form of a film, it shrinks upon heating, allowing it to be used in shrink tunnels. In the form of fibers, it is used for monofilament fishing line and netting. In the form of nonwoven fabrics, it is used for upholstery, disposable garments, awnings, feminine hygiene products, and diapers.
- PLA has applications in engineering plastics, where the stereocomplex is blended with a rubber-like polymer such as ABS. Such blends have good form stability and visual transparency, making them useful in low-end packaging applications.
- PLA is used for automotive parts such as floor mats, panels, and covers. Its heat resistance and durability are inferior to the widely used polypropylene (PP), but its properties are improved by means such as capping of the end groups to reduce hydrolysis.[35]

- Agricultural
- In the form of fibers, PLA is used for monofilament fishing line and netting for vegetation and weed prevention. It is used for sandbags, planting pots, binding tape and ropes .[35]
- Medical
- PLA can degrade into innocuous lactic acid, so it is used as medical implants in the form of anchors, screws, plates, pins, rods, and mesh.[35] Depending on the exact type used, it breaks down inside the body within 6 months to 2 years. This gradual degradation is desirable for a support structure, because it gradually transfers the load to the body (e.g. to the bone) as that area heals. The strength characteristics of PLA and PLLA implants are well documented.

Silicone polymer



- By varying the -Si-O- **chain lengths, side groups, and crosslinking**, silicones can be synthesized with a wide variety of properties and compositions.
- They can vary in consistency from liquid to gel to rubber to hard plastic.
- The most common siloxane is **linear polydimethylsiloxane (PDMS)**, a silicone oil.



- The second largest group of silicone materials is based on silicone resins.
- Si-O bonds are much stronger than a typical C-O single bond. (809 kJ/mol compared to 538 kJ/mol)

Properties

- Silicones exhibit many useful characteristics, including:[1]
- Low thermal conductivity
- Low chemical reactivity
- Low toxicity
- Thermal stability (constancy of properties over a wide temperature range of –100 to 250 °C)
- The ability to repel water and form watertight seals.
- Does not stick to many substrates, but adheres very well to others, e.g. glass
- Does not support microbiological growth
- Resistance to oxygen, ozone, and ultraviolet (UV) light. This property has led to the widespread use of silicones in the construction industry (e.g. coatings, fire protection, glazing seals) and the automotive industry (external gaskets, external trim).
- Electrical insulation properties. Because silicone can be formulated to be electrically insulative or conductive, it is suitable for a wide range of electrical applications.

Uses

- Applications of silicone polymers Health care applications
 - Moulded parts - drainage accessories, external feeding, laboratory tubing, respiratory masks, drug delivery devices etc.
 - Dental impression materials, Extrusion and tubing - blood handling equipments, urological applications etc.
 - Breast implants
- Cosmetics & Toiletries
 - Hair care products mainly conditioning and hair styling products, but siloxanes may also be used in shampoos and colors.
 - Skin care products such as body care, facial care and sun screen products.
 - Shaving products such as preshave and after-shave lotion and shaving foam.
 - Liquid soap and shower gels.
- Sealants used for construction
 - One component RTV silicones: mainly used in the construction industry for sealants around windows and doors, in bathrooms, expansion joints between dissimilar materials.