

L1 What is an electrochemical change?

L2 Why do some materials appear to react more than other materials?

L2 Is it possible to predict whether an electrochemical reaction will occur spontaneously?

L3 How can combinations of species act together as oxidizing or reducing agents?

L3 Can half-reactions be used to predict and explain changes that occur within a chemical system?

### Module 3 Questions

**What properties of metals make them popular choices in the construction and production of materials?**

**How can an understanding of corrosion allow for better selection of materials and development of methods that reduce material damage?**

L3 Can the same substance be the oxidizing agent and the reducing agent in an electrochemical process?

L5 How is the stoichiometric method applied to redox systems?

L5 How can a chemical system be analyzed using redox reactions?

L4 Can corrosion be prevented?

L4 What factors cause corrosion?

L4 What are oxidation numbers and how can they be used to understand redox reactions?