

Photochemistry and Photophysics

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Chapter 1

Welcome

The notes have been prepared in a package called BookDown for RStudio so that the equations are accessible to screen readers. However, by providing the notes as a .html webpage I can also embed short videos to further describe some of the topics. I am using this format as it is a legally accessible format. Further you can download the document in a format that suits you (either pdf or epub) to view offline, or change the way this document appears for ease of reading (top left of this browser window).

This course is structured such that in our ‘in person time’ we will be workshoping problems. I believe I have already set enough questions in this set of notes so that they are all ready ahead of schedule, but I will bring some ‘reserve’ questions if they are required. The reason I have chosen to workshop in our live time is because it is the best opportunity for interaction, this course has no tutorials associated with it and so I wanted to be able to help you as much as I could.

Please ensure that you have read through the accompanying notes for each workshop ahead of our inperson time.

1.1 Workshops for Photochemistry & Photophysics

The answers for questions will only be shared in our workshops which will (hopefully) be recorded on ReView and shared *via* Moodle. I will make available ‘handwritten’ or typed answers if required, please contact me.

Please download the app [UniDoodle]{<http://www.unidoodle.com>} prior to our workshops, this is a mechanism where you can ‘draw or write’ at me.

The schedule for this class is:

Week 1 - Intro to this document, summary of expected prior knowledge

Prep for week 2 - Chapter 3

Week 2 - Workshop on Absorbance 4

Prep for week 3 - Chapter 5

Week 3 - Workshop on Emission 6

Prep for week 4 - Chapter 7

Week 4 - Workshop on Non-emissive Decay of the Excited State 8

Prep for week 5 - Chapter 9

Week 5 - Workshop on Other Excited State Structures 10

Week 6 - Workshop covering all material from course

Week 7 - Workshop on past exam questions

I am astmatic, therefore I will be abiding by the university policy of, in lecturers I expect you to either:

- wear a mask
- visibly wear a sunflower lanyard
- or leave

1.2 Additional Resources

Whilst this resource is designed to stand alone, you may find it beneficial to read around the subject or else listen to the concept bites. Please also note that you can (and should) ask questions in the workshops - this is a good use of everyone's time because if you are struggling with something you won't be the only one.

1.3 Report errors

If you spot any errors or areas of lack of clarity, please message me in Teams or preferably report on the error log below:

Loading...

Version history

I may update this work for typos or clarity if people contact me about this (see above), but there will be no substantive changes to the document in this academic year.

reaction coordinate, fluor workshop version 221021

pdf works version 131021

No longer COVID version 300921

COVID the first version finished 291020

The initial commit of this book is dated 25th September 2020.

Chapter 2

Introduction to Photochemistry & Photophysics

Some context on this course.

Photochemistry is by definition from IUPAC ‘the branch of chemistry concerned with the chemical effects of light’¹. However, in this course we are going to explore beyond just chemical reactions to also examine some of the fundamental photo physics of systems. Photophysics is closely related to what you have previously studied in your first and second years, in that it is a study of the processes of absorption and emission of light and the kinetics of these processes.

Light is the very reason why there is life on earth, the outstanding beauty of processes such as photosynthesis or vision speak to the power of this branch of chemistry. Photochemistry offers a way for us to ‘cheat’ traditional chemical reactions, in that we no longer need thermal energy to get over an activation barrier, instead we use the energy of a photon. It is only by understanding how we can best utilise the bounteous resource of the sun’s light that we as a species can hope to have a long and fruitful future on this planet.

It is just over 100 years since quantum theory revolutionised our understanding of matter and light, not quite 125 years since the discovery of the electron, but in that time we have achieved wonderful things, things that have revolutionised the way we live. From lasers in cd and blue ray players to their use in scanning bar codes, glow sticks - fun at festivals, essential in emergencies, or more prosaic uses such as sun creams, self cleaning windows or display technology. In this course you will learn about a range of photochemical and photophysical processes that

¹IUPAC Goldbook, <http://goldbook.iupac.org/terms/view/P04588> (accessed July 2020).

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underpin our modern lives and also learn about some of the current challenges and interest in the field.

Spectroscopy

- techniques

Photochemistry

- reactions
- molecular structure

Photophysics

- kinetics
- thermodynamics
- quantum mechanics

Photophysical processes

transitions which convert between excited states or between an excited state and a ground state of a molecule

Photochemical processes

reactions or rearrangements which occur as a consequence of excitation from the ground state

2.1 Assumed background knowledge

The course relies heavily on first and second year kinetics concepts as well as building on the quantum mechanics from second year and some of the spectroscopy you did in first year. Whilst I have done my best to ensure that there is a minimum of non-physical chemistry content required to get the most out of this course if I want to talk about molecules some very basic knowledge from both organic and inorganic chemistry, much of which will at least have been mentioned at ‘A’-level or IB.

I will talk a lot about and bonding, but this has been covered in quantum mechanics, a double bond has one and one bond. I will also talk about shape and structure, a carbon with only bonds is 3 dimensional and tetrahedral, a carbon with 3 and 1 bond is trigonal planar, and one with 2 and 2 bonds is linear.

However we need to think about the fact that a molecular orbital exists over a hole molecule. If we reduce bond order overall we change the shape of the molecule. This will be discussed in the absorption (Chapter 3) section of this course.

If you need more help on structure, bonding or metal complexes Chemistry3 is an excellent book to start with, or please ask for help.

I also recommend:

- Principles & Applications of Photochemistry, Brian Wardle,Wiley, ISBN 0470014938 (available as an ebook from the library).
- Photochemistry, Carol E.Wayne & Richard P.Wayne, OUP (primer), ISBN 0198558864 (a really nice primer well worth the money if you are interested in photochemistry)
- Principles of Molecular Photochemistry, Nicholas J.Turro,V. Ramamurthy, J.C. Scaiano, University Science Books, ISBN 1891389573
- The book Principles of Fluorescence Spectroscopy (3rd Edition) by Joseph R. Lakowicz (Springer, ISBN 0387312781) may also be useful for the sections on fluorescence.

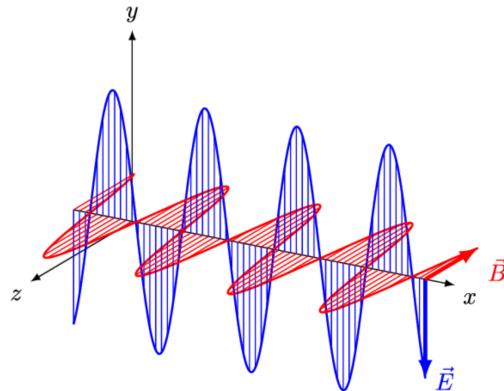


Figure 2.1: The orthogonal electric and magnetic oscillating fields of plan polarised light, propagating along the x-axis. From: [And1mu](<https://upload.wikimedia.org/wikipedia/commons/9/99/EM-Wave.gif>) / [CC BY-SA](<https://creativecommons.org/licenses/by-sa/4.0>)

2.2 Light

As you've already learnt light exhibits properties that have both a wave like and a particle like nature, and light was one of the fundamental pieces that lead towards the quantum theory of the atom. Newton had shown that light from the sun is comprised of a spectrum of colours, we now know that that spectrum goes beyond what we can see with our eyes.

In your studies you have already seen how large amounts of the electromagnetic spectrum is used for different spectroscopic techniques, making use of the quantised transitions within molecules. For example, infra-red light has the same energy as the vibrational transitions within molecules. In photochemistry, we are specifically interested in the visible and near UV part of the EM spectra as

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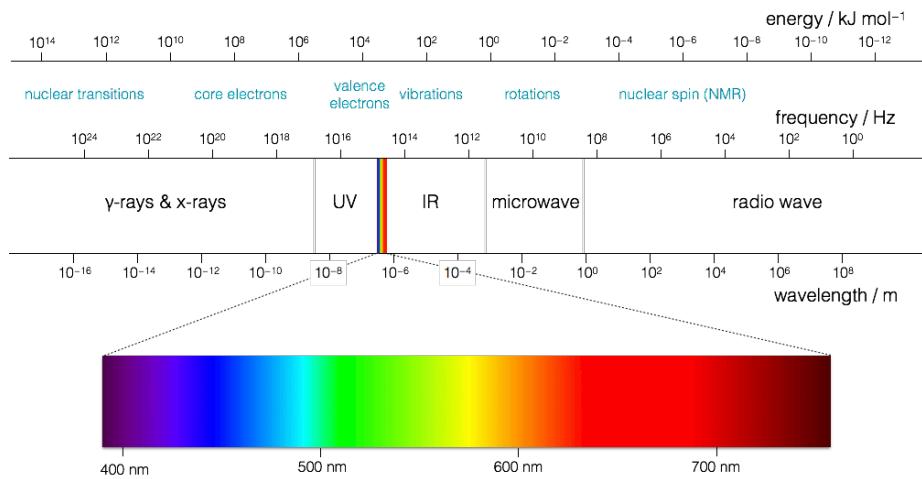


Figure 2.2: TheThe electromagnetic spectrum of light, frequency and energy are equivalent, so high frequency waves have high energy photons

photons with this energy promote transitions for the valance electrons in the molecules we are interested in.

It is important never to forget the wave particle duality of light, however for much of photochemistry and photophysics it is perhaps simpler to consider the light as a stream of incident photons.

Chapter 3

Absorption of light

3.0.1 Learning Objectives

At the end of this section you should be able to:

- Appreciate changes in molecular geometry brought about by light absorption
- Use the Beer Lambert Law to calculate light absorption
- Discuss electronic transitions in terms of Morse energy curves
- Explain the Franck Condon principle and Franck Condon factors
- List selection rules for excitation of a photon and describe factors which affect the molar extinction coefficient

There is a Concept Bite video in Section 3.10 which summarises much of this content.

3.1 Introduction

As you have previously learnt energy levels are quantised, in other words atoms and molecules can only have defined amounts of energy. Different transitions (electronic, vibrational, rotational) in atoms and molecules use different characteristic portions of the electromagnetic spectra, Figure 2.2.

When you saw atomic absorption and emission spectra the lines were very sharp, or in other words the absorption bands are very narrow as shown for the hydrogen absorption and emission spectra (figure 3.1). This isn't the case with molecular absorption spectra, particularly those in solution where absorption bands are typically hundreds of nanometers broad a typical organic fluorophor is shown in figure 3.2. The broad nature of the bands in molecular systems is a product of many individual factors including solvation, vibrational & rotational factors and non minimised structural configurations.

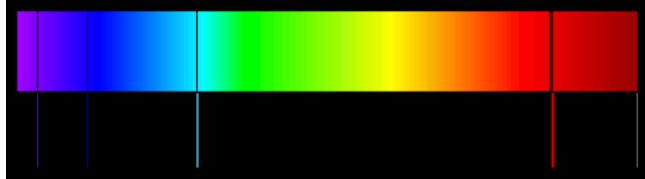


Figure 3.1: The absorption (top) and emission (bottom) spectra of hydrogen showing the very narrow band features, [image](<http://montessorimuddle.org/2012/02/01/emission-spectra-how-atoms-emit-and-absorb-light/>) adapted from [Adrignola](<https://commons.wikimedia.org/wiki/User:Adrignola>) licensed under CC BY 2.0

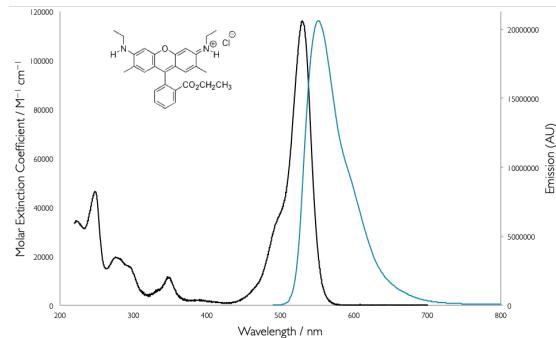


Figure 3.2: The absorption (black) and emission (teal) spectra of Rhodamine 6G in ethanol. The emission spectra was recorded with an excitation wavelength of 480 nm and a bandwidth of 4.25 nm [Adapted from [OMLC](<https://omlc.org/spectra/PhotochemCAD/html/083.html>), [2nd July 2014]]

3.2 Beer-Lambert Law

The empirical equation (equation (4.1), figure 3.3) implies that the probability of a photon being absorbed at any point is the same (much like first order kinetics), and the amount of the total absorption depends upon the concentration of the sample, c , and the ‘path length’, l .

The amount of absorbance, A , is dependent upon the wavelength of the incident light, and the constant of proportionality, ε (here called the molar extinction coefficient), is consequently also wavelength dependent.

$$\log \frac{I_0}{I} = A = \varepsilon cl \quad (3.1)$$

The wavelength of a particular value of the molar extinction coefficient is often represented as a subscript, ε_λ

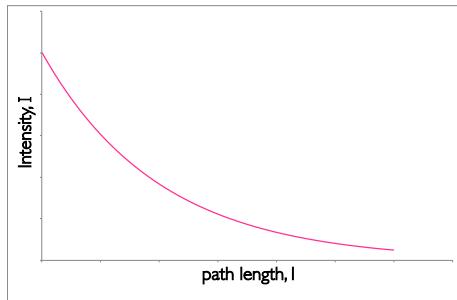


Figure 3.3: The decay of intensity of monochromatic incident light through a uniformly absorbing medium. The decay follows an exponential pattern as elucidated in the Beer-Lambert equation.

The Beer-Lambert law makes a number of assumptions, and this exponential decay of the intensity of light is an important factor. When using the Beer-Lambert law you consider the intensity of the incident radiation, there is an assumption that the intensity of the radiation reaching each part of the sample does not deviate much from this. Hence high absorbing samples tend to show strong deviation from the Beer-Lambert’s linear relationship.

The Beer-Lambert law also has to make a number of other ‘reasonable’ considerations:

- The solutions is well mixed, and absorbers are homogeneously distributed in solution.
- The absorbers do not scatter radiation (all particles will Rayleigh and Raman scatter but this is normally considerably less intense than absorption). Consequently solutions should be optically transparent as optically opaque solutions (such as colloidal solutions) have considerably stronger scattering.

- The absorbers acts independently of each other, this means solutions need to be at a reasonably low concentration (typically less than 0.01 M, or maybe even less depending on the species) so as to avoid electrostatic or π stacking interactions between the chromophores. This is in part important because light is only absorbed when the polarisation of the light is aligned with the transition dipole moment.
- The incident radiation is collimated, and each photon should pass through the same path length.
- The sample holder (cuvette) is optically ‘pure’ such that reflections are avoided (linked to the assumption above).
- The incident radiation is monochromatic, or at the very least has a band width more narrow than the band width of the absorbing transition (this is usually not an issue for molecular systems as bandwidths are usually 10s or more of nm wide, but for atomic or ion spectroscopy where bandwidths are <0.02 nm this is a factor which must be carefully considered).
- The incident radiation does not noticeably affect the concentration of the ground state, in other words the amount of excited states generated must be kept small as when we are talking about the absorption of a chromophore the concentration of that chromophore that appears in the Beer-Lambert equation is the ground state concentration.
- There is no measurable emission from the sample.

However, this empirical relationship can be examined in a deeper way.

3.3 Molecular orbitals, HOMO & LUMO

Ultimately, the absorption (and emission of a photon) which can be modelled empirically, as in the Beer-Lambert law, is actually a property of the quantum mechanics of the system as indicated by the models of absorption and emission of photons by Einstein.

If we think about the size of an atom or molecule (atoms are 100s of pm, molecules frequently less than 1 nm) they are considerably smaller than the wavelength of light required for electronic transitions, so it is often easier to think of an entirely photonic model of light.

If we consider the atom or molecules to be formed of a cloud of electrons bathed in the electromagnetic field, the electron density is attracted to the positive side of the field, and the positively charged nucleus to the negative side. However, electrons are much lighter (about 1/1760 the mass of a proton) and so it is the lighter electrons that ‘feel’ the most effect. As the electric field oscillates it pushes and pulls the electrons within the molecule.

This transition is called an electric-dipole transition since it either creates (absorption) or destroys (emission) an oscillating electric dipole in the atom or molecule. The transition dipole moment, μ_i , arises from the charge displacement during the transition. The magnitude of this transition dipole moment depends

upon the distance the net electronic charge is moved from the ground state average position. This transition dipole is entirely independent of any permanent dipole moment within the molecule.

Figure 3.4 shows the valence molecular orbitals of butadiene, showing the discrete energy levels in the system, absorption of a photon allows an electron to be excited from the highest occupied molecular orbital (HOMO) to any excited state.

Upon excitation of an electron there is a mixing of the molecular orbitals of the fully and partially occupied molecular orbitals, in the case of butadiene excitation of an electron from the HOMO to the LUMO (lowest unoccupied molecular orbital), giving the central C-C bond some double bond character, limiting rotation and leading to a structural change of the excited state.

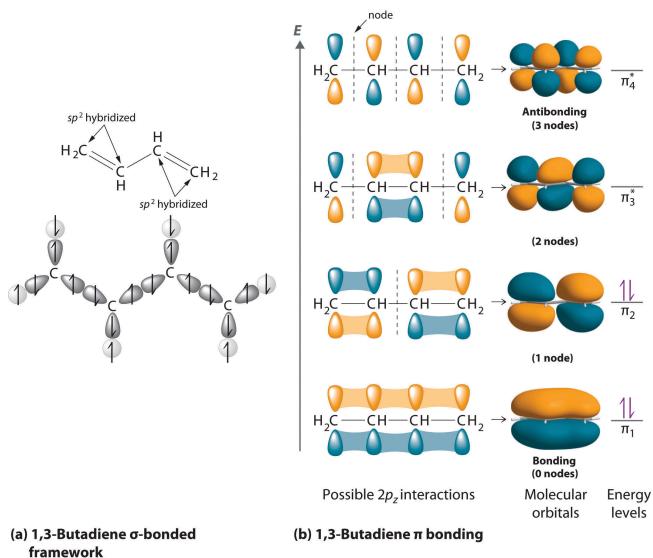


Figure 3.4: MO diagram of butadiene, each carbon atom is assumed to be *sp*² hybridised, with linear combination of the *2p*_z orbitals. As the energy of the levels increases so does the number of nodes. The lowest two energy levels are fully occupied in the ground state, upon absorption of a photon an electron is promoted to one of the anti-bonding orbitals allowing rotation around the double bonds, and giving the centre single bond some double bond character. [[MO diagram of butadiene](<https://2012books.lardbucket.org/books/principles-of-general-chemistry-v1.0m/s13-04-polyatomic-systems-with-multip.html>)]. From by Averill & Eldredge, Principles of General Chemistry licensed under CC BY 2.0. Sept 14.]

In simple organic molecules we only really need to consider π and π^* and σ (non-bonding) molecular orbitals as the energy difference between σ and σ^* orbitals are comparatively very large.

The process of absorption of a photon is very fast (~ 1 fs) (compared with the timescales of vibrations in molecule (~ 10 ps)), this is the basis of the Franck-Condon principle. Consequently when absorption (and emission) processes are sketched, as in figure 3.5, absorption and emission can be indicated by vertical transitions.

The ‘strength’ of the absorption transition, which relates to the extinction coefficient in the Beer-Lambert law, and the Einstein B coefficient is in fact a measure of the ‘overlap integral’ of the wave functions of the ground and excited states.

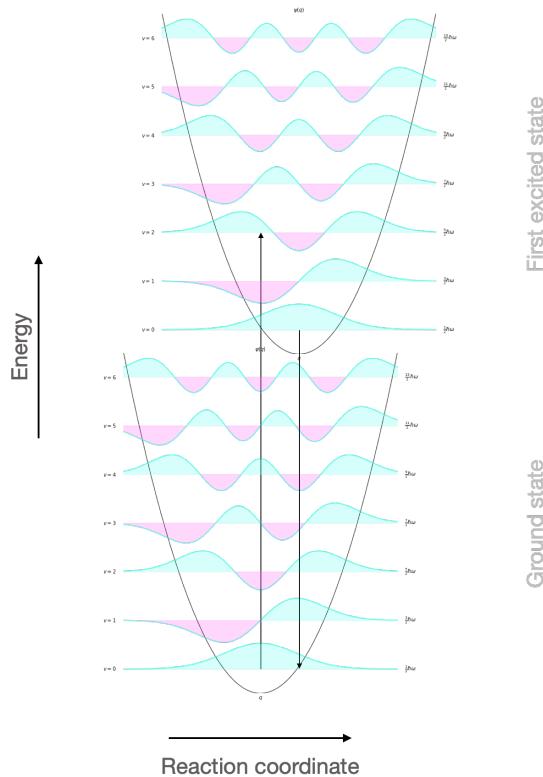


Figure 3.5: The vertical line of the absorption transition as the electron is promoted from the ground state to the excited state. The probability of the electron being excited into each vibrational level is given by the ‘overlap’ of the wave functions of ground and each excited state. [[Franck-Condon Diagram](<https://commons.wikimedia.org/wiki/File:Franck-Condon-diagram.png>)]. From Wikimedia Commons, created by [Mark M. Somoza](<http://www.gnu.org/licenses/fdl-1.3.html>), CC-BY-SA-3.0](<https://creativecommons.org/licenses/by-sa/3.0/>). Sept 14.]

The greater the overlap integral between the ground and excited state the higher the molar extinction coefficient and the more strongly coloured the molecule.

Understanding the variation of possible electronic transitions available when electronic levels are mixed with vibrational and rotational levels, as well as the molecular dynamics allowing multiple ‘non-minimised’ structures explains the relative broadness of molecular electronic absorption and emission spectra.

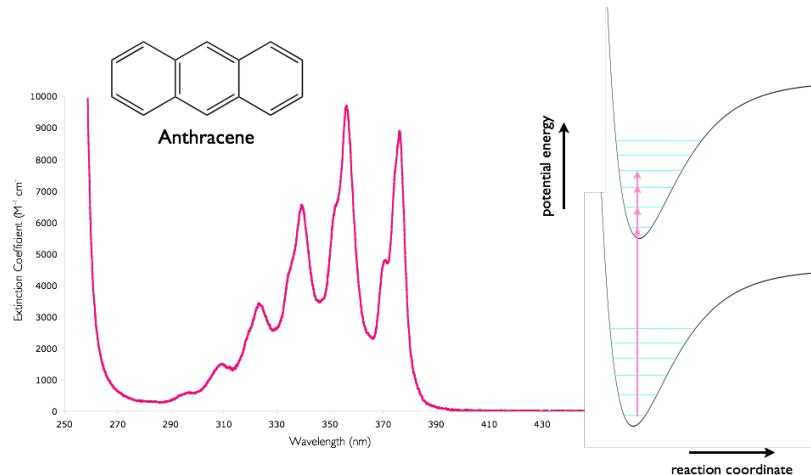


Figure 3.6: The absorption spectra of anthracene in cyclohexane showing the vibrational fine structure, combined with a sketch of the potential energy wells of the ground and excited states. Regions with the highest molar extinction coefficient have the largest overlap integrals of the wavefunctions of the ground and excited states.

3.4 Multiplicity

If you recall Hunds' rule when filling atomic and molecular orbitals you will recall that the lowest energy state occurs when the spin of the electrons is aligned. Usually in an organic molecular orbital the HOMO is a fully occupied π set, or a fully occupied set of non bonding orbitals with all electrons paired.

It is common to refer to the electronic states in a molecule by their *spin multiplicity*, given by $2S+1$, where S is the sum of the electronic spins of the electrons in the orbital. For a species in which all of the electrons are paired $S = 0$, and so $2S + 1 = 1$ and the state is referred to as a *singlet*.

If two electrons are unpaired they will (in the lowest energy configuration) have parallel spins, the total spin $S = 1$ and so the spin multiplicity, $2S + 1 = 3$. This state is referred to as a *triplet* state.

Species such as free radicals have one unpaired electron, consequently $S = \frac{1}{2}$, and $2S + 1 = 2$, and the species is referred to as a *doublet*.

Triplet states with the same electronic configuration (other than spin) have

a lower energy than corresponding singlet states due to an effect called spin correlation. In triplet states the motion of the electrons in the orbitals are organised in such a way as to minimise the coulombic interactions between the charges, by minimising these coulombic interactions the energy of the system is also lowered.

Although formally flipping of electron spin is forbidden in photophysical and non photophysical processes it does occur. Processes that involve a change in electron spin are explained in later sections.

3.5 Conjugation and the HOMO, LUMO energy gap, ΔE



Figure 3.7: The relative energy arrangement of orbitals in a molecule, this simple sketch is usually good enough to gain an understanding of the system under investigation.

The greater the amount of conjugation in a molecule the smaller the energy gap between the HOMO & LUMO and the longer the wavelength of photon absorbed.

This fits with the theory you learnt in your quantum mechanics lectures, where you can model the linear conjugated model quite accurately using the particle in a one dimensional box model.

Table 3.1: The dependence of maximum absorption wavelength on the increasing length of conjugation is clearly shown by moving along the diene, triene, polyene series. Inclusion of a hetero atom with lone pairs introduces non-bonding, n , electrons into the MO, and a second absorption band $n \rightarrow \pi^*$ introduced with a lower energy transition.

$\lambda_{\max} / \text{nm}$ $\pi \rightarrow \pi^*$	$\lambda_{\max} / \text{nm}$ $n \rightarrow \pi^* \pi \rightarrow \pi^*$
217	270 187
227	

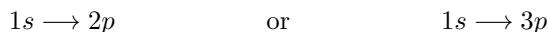
	$\lambda_{\max} / \text{nm}$ $\pi \rightarrow \pi^*$	$\lambda_{\max} / \text{nm}$ $n \rightarrow \pi^* \quad \pi \rightarrow \pi^*$
	263	324 219
	353	

3.6 Selection rules for absorption & emission processes

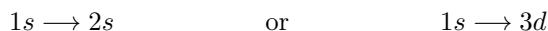
In molecular systems the most important selection rule is $\Delta S = 0$, or there can be no change in the spin multiplicity during an absorption or emission processes, or there can be no ‘flipping’ of electrons spin during a transition. In practice ‘coupling’ due to interactions between the spin of the electron and the orbital mean that this selection rule is not rigidly followed.

In organic molecules containing only C, H, N & O the direct absorption from S_0 to T_1 is so weak that it can’t be seen in the absorption spectra. This ‘spin-orbit coupling’ is more pronounced for heavier elements and molecules containing particularly heavy atoms, such as S, Cl or Br have more relaxed adherence to the $\Delta S = 0$ selection rule.

The Laporte selection rule is more relevant when looking at inorganic transition metal complexes, or for atomic spectra which will not be considered in this work. The Laporte states that there must be a change in angular momentum quantum number, l , upon absorption or emission in a centro symmetric system, or $\Delta l = \pm 1$. So in a centro symmetric molecule, such as $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$, figure 3.6, the allowed transitions would be transitions such as the following:

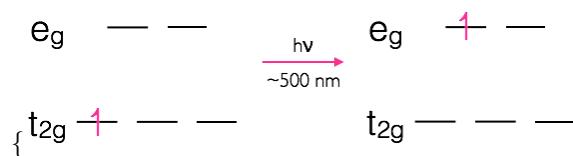


Whilst these transitions would be forbidden according to the Laporte rule:



Consequently the $d-d$ transitions, such as the $t_{2g} \leftrightarrow e_g$ transition in octahedral complexes, which give most transition metal complexes are forbidden according to the Laporte rule.

\begin{figure}



}

\caption{The ligand field splitting of the 3d orbitals in a $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$ complex the solution is pale purple with a molar extinction coefficient, $\sim 6.1 \text{ M}^{-1} \text{ cm}^{-1}$ } \end{figure}

If we compare the magnitude of the molar extinction coefficient for a typical $d - d$ absorption of a transition metal complex with that of a ‘spin allowed’ ‘Laporte allowed’ organic dye molecule we can see the effect of ‘non allowed’ transitions. In table 3.2, the organics, DNA & the synthetic dye TOTO-1 both have extended conjugated molecular orbitals, and the transitions are from singlet to singlet, an allowed transition.

In these cases the Laporte rule does not apply as there is no centre of symmetry. The two examples of transition metal complexes $[\text{Cu}(\text{H}_2\text{O})_6]^{2+}$ & $[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$, have significantly lower molar extinction coefficients.

Molar extinction coefficients, if you recall, were a representation of the orbital overlap between the HOMO & LUMO, or a measure of how ‘allowed’ the transition is. The transitions which are Laporte forbidden are at least two orders of magnitude less intense than the allowed transitions of the organic molecules.

Interestingly, permanganate which is relatively highly coloured when compared to other transition metal complexes cannot have colour due to $d - d$ transitions as the +7 oxidation state of the manganese ion leaves no electrons in either the s or d orbitals. Instead the colour is due to a charge transfer transition, where upon absorption of a photon an electron is briefly transferred from one of the oxygen ligands to the manganese metal centre. Ligand to metal charge transfer complexes (and the associated metal to ligand charge transfer complexes), are not subject to selection rules and so are relatively intense.

Table 3.2: The molar extinction coefficients of a range of molecules, the magnitude is strongly dependent upon the adherence to selection rules, the two organic molecules, DNA bases & TOTO-1 are both ‘spin allowed’, permanganate is a LMCT which has no selection rules, whereas the final three molecules all fall foul of the Laporte selection rule which states $\Delta l = \pm 1$. Since these transitions are ‘forbidden’ the molar extinction coefficient is considerably lower than the allowed transitions.

	$\lambda_{\max} / \text{nm}$	$\varepsilon_\lambda / \text{M}^{-1} \text{ cm}^{-1}$
DNA (per base)	260	6600
TOTO-1	514	117000
KMnO_4	520	1800
CuSO_4	810	20
$[\text{Ti}(\text{H}_2\text{O})_6]^{3+}$	500	6.1

	$\lambda_{\text{max}} / \text{nm}$	$\varepsilon_{\lambda} / \text{M}^{-1} \text{ cm}^{-1}$
PrCl ₃	445	0.07

The weak absorption carries on for the lanthanide complexes where $f - f$ transitions are again forbidden by the Laporte selection rule. However, Ce³⁺ and Tb³⁺ have more intense electronic absorption bands which appear in the UV.

These are due to [Xe]4fⁿ to [Xe]4fⁿ⁻¹5d¹ transitions. Since the electron is moving from an f orbital to an d orbital it is not forbidden by the Laporte selection rule and so is considerably more intense than the forbidden $f - f$ transitions. These transitions occur for these two ions due to the stability provided by either an empty or half filled subshell.

3.7 Structure and Bonding Upon Absorption of a Photon

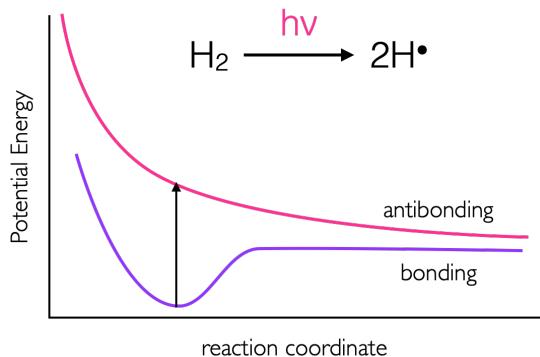


Figure 3.8: The energies of bonding and anti bonding orbitals of molecular hydrogen, as an electron is promoted to the anti bonding * orbital the bond order is reduced to zero and the molecule ‘falls apart’.

Upon absorption of a photon, if an electron is promoted to an anti bonding orbital there is a reduction in the bond order. For a hydrogen molecule, with just two electrons, this leads to radical formation figures 3.8 & 3.9. It should be remembered that there is still an electron in the bonding orbital and the molecular orbital is now a blend of the HOMO and LUMO. The LUMO For molecules with π electrons excitation of an electron to a π^* orbital reduces the bond order to just 1, allowing for rotation around the previously restricted bond. If ethene is excited, the sp² hybridised carbon, takes upon character of sp³ allowing rotation around the C-C bond, and changing the H-C-H bond angles. Rotation around a double bond is the basis of many photochemical

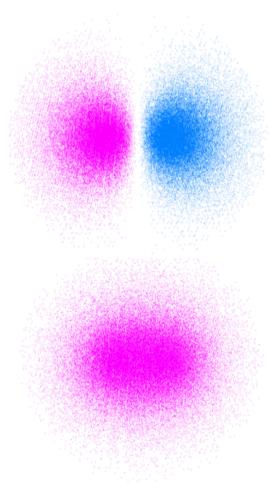


Figure 3.9: The bonding and anti bonding orbitals of molecular hydrogen showing the node along the bond for the anti-bonding orbital.

reactions and is fundamental to vision where retinol, a long conjugated molecule undergoes a cis-trans isomerisation after excitation.

3.8 The Transition Dipole Moment

One thing which hasn't yet been considered is how light physically interacts with a molecule.

Classically it could be considered that the oscillating electric field of a wave could interact with a molecule if there was some resonance of the frequency of light and the frequency of oscillation of the electron in an orbital, it would seem obvious that an electron (carrying an electric charge) and an electromagnetic wave (with its oscillating electric field) should interact; and will interact given the correct wavelength. In simple terms the frequency of light and the energy gap for the transition were the same.

The classical model also explains why it is electrons which are excited (and not protons in the nucleus) because the considerably lighter electrons are more affected (and moved more) in the presence of the electric field.

This approach whilst covering the very basics of a transition does not particularly explain concepts such as the molar extinction coefficient, nor does it particularly consider concepts fundamental to quantum mechanics - such as quantisation of energy - an oscillating electric field is in no way quantised and there is no reason why excitation of the molecule should be quantised either.

However this idea of an oscillating dipole has some relation to quantum

mechanics; because the shape of the molecular orbitals of the HOMO and LUMO are different and there will be a change in the electron distribution and hence a change in dipole of the molecule upon excitation. As such we can define a transition dipole moment for each chromophore - the more similar the ground and excited states the larger the transition dipole moment for the molecule.

A mathematical derivation of the transition dipole moment is beyond the scope of this work¹, but it is worth discussing as transition dipoles are hugely important when it comes to energy transfer in molecules (see Förster Resonance energy transfer, Section @ref({sec:forster})).

If we consider light as a wave polarisation is well understood, however there is a quantum mechanical explanation for polarisation as well - in short we can use wave behaviour to consider polarisation and photon behaviour to explain quantisation - just to stick with concepts which should be familiar to us.

Light is only absorbed by a molecule if the polarisation of the light aligns with the transition dipole moment on the molecule. Figure 3.10 shows CS₂ a simple linear molecule - where the transition dipole moment runs along the long axis of the molecule.

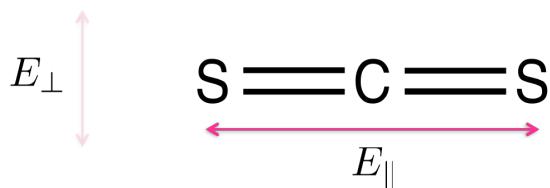


Figure 3.10: Carbon disulfide (CS₂) is a linear molecule - due to the shape of the molecule there is a transition dipole which runs down the length of the long axis. Light aligned such that the electric field runs parallel with the long axis of the molecule E_{||} will be absorbed, light which in which the electric field runs perpendicular to the long axis of the molecule E_⊥ will not be absorbed.

For more complicated molecules each of the transitions from the HOMO, HOMO-1 to the LUMO etc. occur with different transition dipole moments across the chromophore, figure @ref{fig:adenosine}. Each transition is only excited when light is aligned with that transition; in most cases this isn't something we need to consider as most incident light we consider is isotropic, but alignment of transition dipoles (either between light and molecules - or between two different molecules) is an important consideration.

As already discussed the transition dipole moment is derived from the difference in electron density of the ground and excited state. Figure 3.12 shows the HOMO and LUMO of an organic chromophore, and a metal complex MLCT state.

¹see Principles of Photochemistry by Turro *et al* if you wish to know more

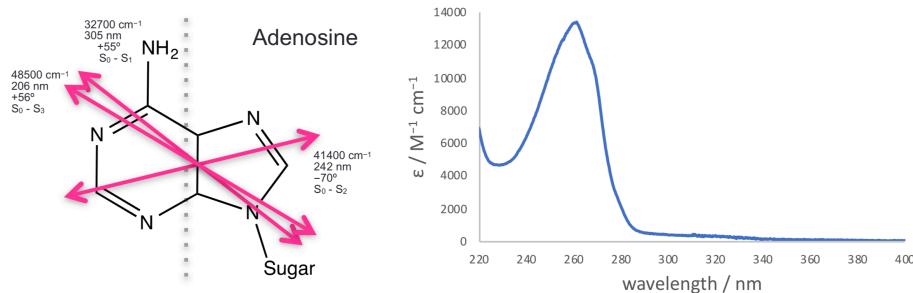


Figure 3.11: The three lowest energy transitions of adenosine each indicated with their transition dipole moment (all in the plane of the molecule, calculated values). These match with the observed spectrum with a weak transition around 310 nm, a much stronger transition around 260 nm and a third transition starting at the edge of the measured spectrum. [Spectrum [Adapted from OMLC](<https://omlc.org/spectra/PhotochemCAD/html/033.html>), 31st October 2018]

3.9 Factors Which Affect the Absorption

In the condensed phase the solvent is an important factor in determining the wavelength of maximum absorption, this effect, called solvatochromism, can be due to many properties of the solvent but is most likely due to solvent polarity. The LUMO energy level is unchanged in the different solvents, but the HOMO can either be more stabilised or less stabilised by the change in solvent (Figure 3.13).

This leads to a change in the HOMO LUMO energy gap and consequently a change in the wavelength of maximum absorption.

If the absorption maxima shifts to increasing wavelength (red shift, decreasing energy of transition) it is referred to as a bathochromic shift, whereas if the absorption maxima shifts to decreasing wavelength (blue shift, increasing energy of transition) it is referred to as a hypsochromic shift.

Another major factor that can affect the absorption of a sample are interactions between the chromophores. This is most commonly encountered with DNA, which shows a significantly higher absorption at high temperatures (Figure 3.14). This is because the bases of DNA ‘stack’ to form the rungs of a ladder, this stacking has favourable $\pi - \pi$ interactions which help to stabilise the macromolecule. As the DNA is heated the double helix begins to melt and the rigid ladder stack is broken. The favourable $\pi - \pi$ interactions are lost and so the absorption increases. There is little or no change in the wavelength of the maximum absorbance in this process only an increase in the absorbance due to a higher value of molar extinction coefficient. This effect as the absorbance of DNA increases upon melting is called hyperchromicity, the

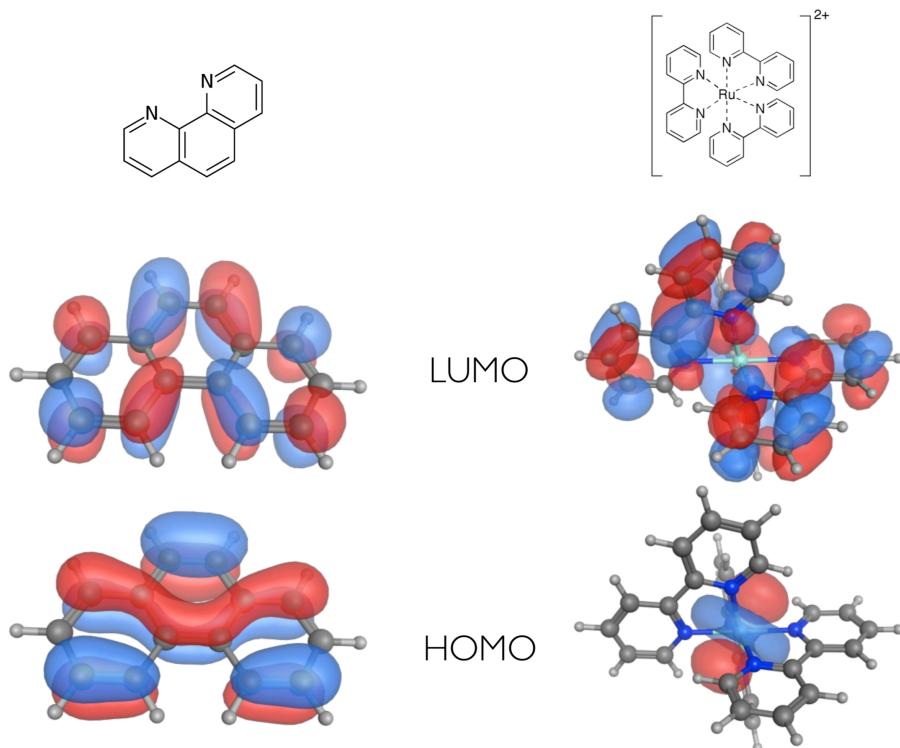


Figure 3.12: The calculated HOMO (bottom) and LUMO (top) of 1,10-phenanthroline (left) a model planar organic chromophore and $[\text{Ru}(\text{bpy})_3]^{2+}$ (bpy - bipyridine) which has an MLCT state. For the ruthenium complex the HOMO is clearly centred on $4d_{z^2}$ orbital, whereas after absorption of a photon (and transfer of the electron) the electron density is then centred over the ligands in the LUMO.

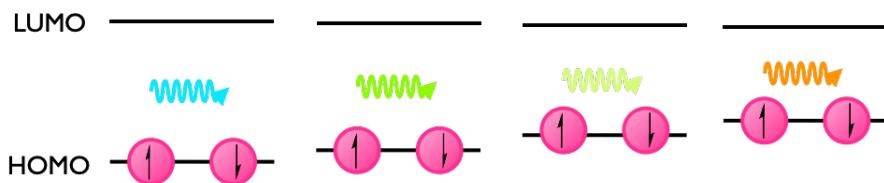


Figure 3.13: The change in absorption wavelength as the HOMO is destabilised in a solvent series, here is worth noting that the colour observed and the wavelength of maximum absorption are not the same thing. As the HOMO is destabilised there is a red shift in the wavelength required to excite the electron.

converse, a decrease in absorption, is called hypochromicity.

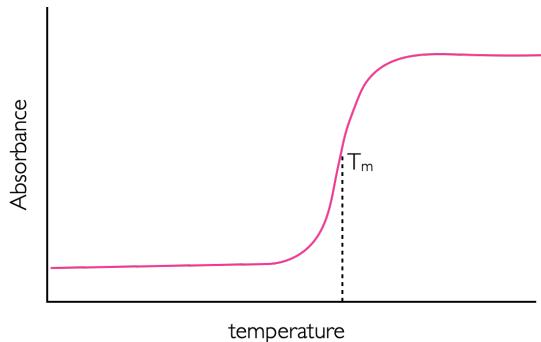


Figure 3.14: A sketch of the melting of DNA leading to a large increase in absorbance. The melting temperature T_m , is defined much as the end point of a titration curve is, as halfway between the two turning points. The increase in absorbance is due to the breaking of the stack which exists in double stranded helical DNA.

3.10 Before Completing this Section

To support the material in this section it is suggested you read chapter 2 of Wardle ‘Principles and Application of Photochemistry’ and the section in Atkins’ ‘Physical Chemistry’ on Einstein coefficients. If you need a refresher of some MO theory it is recommended you look at the early chapters of Chemistry³.

Chapter 4

Workshop Questions for Week 2 Absorption

4.1 Questions

4.1.1 Short mathematical question - Beer Lambert law

- How far can monochromatic 489 nm light travel through a 0.100 M solution of fluorescein with an extinction coefficient at 489 nm of $92000 \text{ M}^{-1} \text{ cm}^{-1}$ before 90 % of it is absorbed?

(I will use MCQs and UniDoodle to ask this in class)

4.1.2 Short conceptual question - molar extinction coefficient

- Modify the molecule in figure 4.1 to increase the molar extinction coefficient (do not worry about what may happen to wavelength).

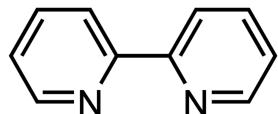


Figure 4.1: The structure of bipyridine (also known as bpy).

(I will use UniDoodle's drawing feature to ask this in class)

4.1.3 Short conceptual question - intensity of colour

- What factors influence the ‘intensity of colour’ of the following solutions?

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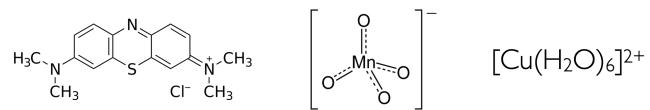


Figure 4.2: The structures of the organic dye methylene blue (left), potassium permanganate (centre) and copper hexa-aqua (right).



Figure 4.3: 1 mM solutions of the organic dye methylene blue (right), potassium permanganate (centre) and copper hexa-aqua (left).

(This will be a discussion question)

4.1.4 Short conceptual question - line width

- Why are some spectra very broad (figure 4.4), whereas others have sharp peaks (figure 4.5)?

You will need to look at the x-scale to truly note the difference in the width of these emission spectra.

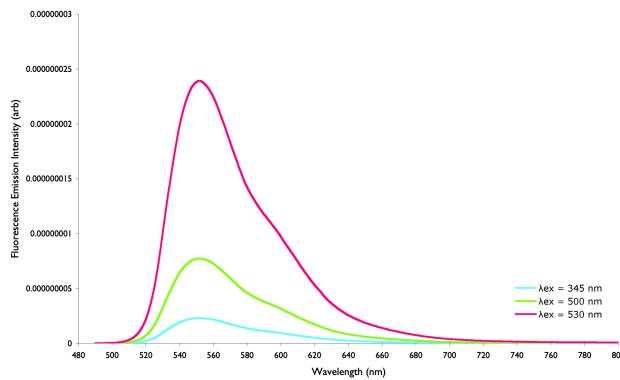


Figure 4.4: The emission spectrum of rhodamine 6G

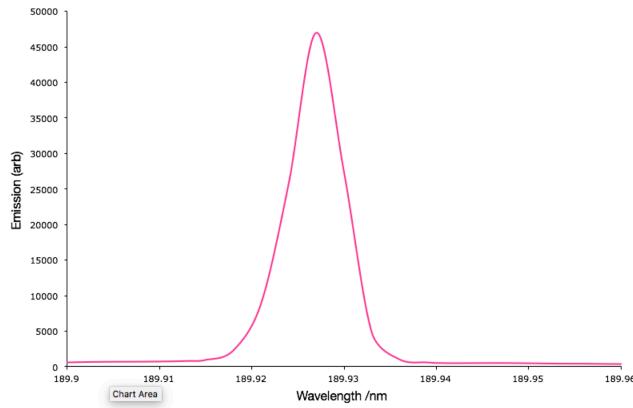


Figure 4.5: The emission spectrum of Sn(II)

(This will be a discussion question)

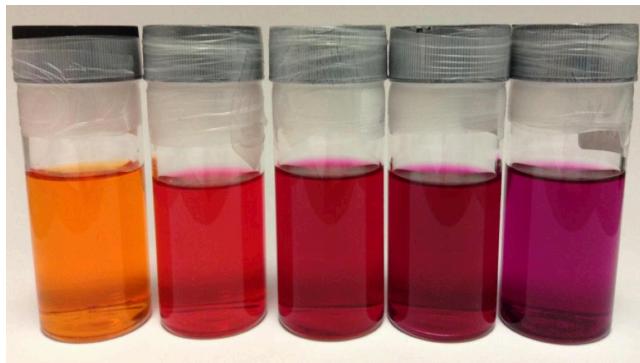


Figure 4.6: Ethidium bromide dissolved in from right; water(orange), methanol, ethanol, propanol and butanol(purple)

4.1.5 Short conceptual question - the effect of solvation on absorbance

- Why does the observed colour of ethidium bromide depend upon the solvent (figure 4.6))?

Think about the effect of solvation on the energy levels and why those energy levels matter! Remember that if light is transmitted through a solution that is the colour we observe...

(This will be a discussion question)

4.1.6 Extended question - Azobenzene

Azobenzene undergoes the following cis-trans isomerisation, the isomerisation occurs in the ps timescale.

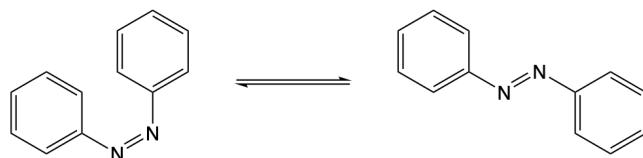


Figure 4.7: The cis-trans isomerisation of azobenzene

- Why would you expect the absorption spectrum of each isomer to be different?
- Suggest why the trans conformation is more stable than the cis isomer.
- Use the following data to predict the proportion of each isomer under 360 nm excitation.

Table 4.1: The molar extinction coefficient of the two isomers of azobenzene.

	$360 / M^{-1} cm^{-1}$	$460 / M^{-1} cm^{-1}$
trans-azobenzene	22000	4500
cis-azobenzene	2100	5500

- Would there be more or less trans azobenzene at 460 nm? Justify your answer.
- It has been suggested the 360 nm absorption is an $S_0 \rightarrow S_2$ absorption, and the 460 nm band is an $S_0 \rightarrow S_1$ absorption. Suggest which energy levels are involved for each of the two transitions and compare it to stilbene which has a similar structure, but the cis and trans absorptions are 280 & 295 nm respectively.

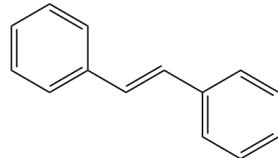


Figure 4.8: The structure of stilbene

(This will be a discussion question)

4.2 Answers

4.2.1 Short mathematical question - Beer Lambert law

- How far can monochromatic 489 nm light travel through a 0.100 M solution of fluorescein with an extinction coefficient at 489 nm of $92000 M^{-1} cm^{-1}$ before 90 % of it is absorbed?

If 90% of the incident light is being absorbed then 10% of the light is transmitted...

The Beer-Lambert law is (Equation (4.1)):

$$\log \frac{I}{I_0} = -\varepsilon cl \quad (4.1)$$

rearranging this:

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$$l = \frac{\log \frac{I_0}{I}}{\varepsilon c}$$

$$l = \frac{\log \frac{100\%}{10\%}}{92000 \text{ M}^{-1} \text{ cm}^{-1} \times 0.100 \text{ M}}$$

$$l = \frac{1}{9200000 \text{ M}^{-1} \text{ m}^{-1} \times 0.100 \text{ M}}$$

$$l = 1.0 \text{ m}$$

4.2.2 Short conceptual question - molar extinction coefficient

- Modify the molecule in figure 4.1 to increase the molar extinction coefficient (do not worry about what may happen to wavelength).

Image may be found in Section 4.1.2

If a molecule is rigid the difference between the ground and excited state structure is smaller and this usually leads to a higher molar extinction coefficient, so any structural changes which increase this rigidity will likely lead to a higher molar extinction coefficient.

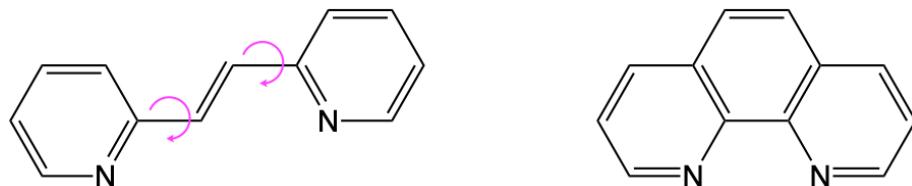


Figure 4.9: Two possible modifications of bpy, the stilbene derivative on the left has a very different ground and excited state structure due to the possible rotations around the two conjugated single bonds, the phenanthroline modification on the right does not allow much deviation in structure and so the excited state is only bent as the bond order is reduced leading to a larger overlap integral.

Excitation of an electron from a bonding or non-bonding orbital into an anti-bonding orbital reduces the bond order, and will ‘kink’ aromatic systems.

4.2.3 Short conceptual question - intensity of colour

- What factors influence the ‘intensity of colour’ of the following solutions?

Images may be found in Section 4.1.3

CuSO_4 in solution forms the hexa aqua complex listed in figure ???. This is a centrosymmetric molecule which therefore runs foul of the Laporte selection rule which says that “allowed transitions in centrosymmetric molecules must involve a change in parity”. So either $g \rightarrow u$ or $u \rightarrow g$ are allowed, and $g \rightarrow g$ and $u \rightarrow u$ are forbidden.

You will not be expected to work out symmetry types for orbitals but can ‘back justify’ to explain with this.

These symmetry labels appear on the energy level diagram for CU^{2+} (figure 4.10).

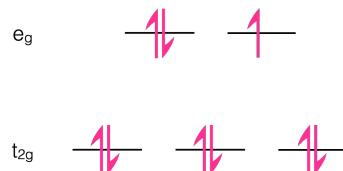


Figure 4.10: The splitting of the d-orbitals in octahedral Cu^{2+} , the transition is between the HOMO t_{2g} level and the PUMO e_g level. The subscript g on each of these orbitals is a symmetry label indicating there is no change or parity and so the transition is Laporte forbidden.

Both of the split energy levels are gerade (g) and so there is no change of parity for the transition from t_{2g} to e_g , therefore the transition is selection rule forbidden and has a low molar extinction coefficient. The ‘breaking of symmetry’ by possible rotamers of the water ligands mean this value is higher than for the gas phase atomic molar absorption coefficient. The maximum molar extinction coefficient of copper sulfate solutions is around $20 \text{ M}^{-1} \text{ cm}^{-1}$.

The permanganate (MnO_4^-) is a ligand to metal charge transfer transition, since excitation of an electron effectively moves an electron from one of the oxygen ligands to the manganese ‘empty’ metal centre there is no limitation from the spin selection rule and the complex is not subject to the Laporte selection rule. This leads to moderate to ‘high’ molar extinction coefficients.

The maximum molar extinction coefficient of potassium permanganate solutions is around $1800 \text{ M}^{-1} \text{ cm}^{-1}$.

The final molecule is the synthetic dye methylene blue, this is not centrosymmetric and so is not subject to the Laporte selectron rule, further the structure with fused aromatic rings and a large amount of conjugation is fairly rigid and so there is a large overlap of the wavefunctions of the ground

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and excited state. The maximum molar extinction coefficient of methylene blue solutions is around $75000 \text{ M}^{-1} \text{ cm}^{-1}$.

4.2.4 Short conceptual question - line width

- *Why are some spectra very broad (figure ??), whereas others have sharp peaks (figure ??)?*

You will need to look at the x-scale to truly note the difference in the width of these emission spectra.

Images may be found in Section 4.1.4

We are looking at two very different spectra here, first off we have an atomic transition and a molecular transition. Secondly one is solvated, the other in the gas (or more accurately plasma) phase. Both of these factors are important...

When we have a transition for either absorption or emission then on an atomic level those energy levels are incredibly well defined, this leads to the energy required to promote an electron between these levels to be equally well defined.

As soon as we solvate this the energy of the occupied (HOMO) is softened as intermolecular interactions will stabilise this level to different amounts, solvation may also allow previously absolutely parity forbidden processes to occur.

The molecular spectrum is much broader because each of the possible rotamers, vibratamers and isomers has a slightly different energy, the myriad possible energies are expanded still by solvation and temperature effects. The more possible structural isomers and rotamers the 'softer' any defining features in the spectrum. For example if we compare the rhodamine spectrum (figure ??) to that of anthracene (figure 4.11) we can see that there is more 'fine structure' in the anthracene spectrum, this is because of the smaller possible number of structural variations, there are no possible rotamers and so transitions are more defined.

4.2.5 Short conceptual question - the effect of solvation on absorbance

Image may be found in Section 4.1.5

- *Why does the observed colour of ethidium bromide depend upon the solvent (figure ??))?*

Think about the effect of solvation on the energy levels and why those energy levels matter! Remember that if light is transmitted through a solution that is the colour we observe...

This is in essence answering the same question as in question ??, solvation affects the energy levels of occupied orbitals. Without going into any more

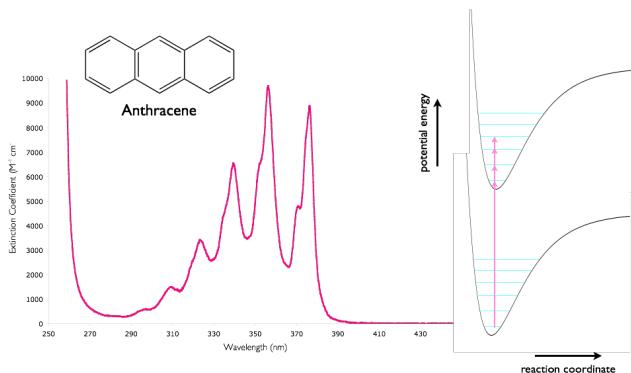


Figure 4.11: The absorption and emission spectra of anthracene indicating fine structure in the spectrum

detail than essential the different solvents stabilise the HOMO of ethidium bromide different amounts - this is likely due to different hydrogen bonding of the alcohols. The more stabilised the ground state the bigger the energy gap leading to a blue shift in the absorption spectra. The orange solution is absorbing blue light ($\lambda_{\text{max}} = 480 \text{ nm}$) and the purple solution absorbing orange ($\lambda_{\text{max}} = 550 \text{ nm}$) indicating the energy gap of the aqueous solution is larger and therefore the solvent stabilisation is biggest.

It is worth noting that the energy of ‘empty’ orbitals is unaffected and so the energy of the LUMO is unchanged (figure 3.13).

Figure above in Section 4.1.5

4.2.6 Extended question - Azobenzene

Azobenzene undergoes the following cis-trans isomerisation, the isomerisation occurs in the ps timescale.

Image may be found in Section 4.1.6

- Why would you expect the absorption spectrum of each isomer to be different?

Trans azobenzene is centrosymmetric and so is subject to the Laporte selection rule, cis azobenzene is non-centrosymmetric and so it is not parity forbidden. Consequently the HOMO-LUMO transition which is allowed for the cis is forbidden for the trans and so the absorbance spectra is different.

- Suggest why the trans conformation is more stable than the cis isomer.

Steric hinderance will mean that the cis-isomer is not planar, if it isn’t planar it isn’t fully conjugated, and if it isn’t fully conjugated it isn’t as stable. There is no steric hinderance in the trans isomer.

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- Use the following data to predict the proportion of each isomer under 360 nm excitation.

Table: The molar extinction coefficient of the two isomers of azobenzene.

	$360 / \text{M}^{-1} \text{cm}^{-1}$	$460 / \text{M}^{-1} \text{cm}^{-1}$
trans-azobenzene	22000	4500
cis-azobenzene	2100	5500

For the photostationary state the absorbance of the two isomers will be the same.

The excited state for both molecules will be the same, and the probability of the excited state decaying into any given isomer is the same.

So if $A = \varepsilon cl$ and $A_{trans} = A_{cis}$ then at 360 nm:

$$\varepsilon_{trans} c_{trans} = \varepsilon_{cis} c_{cis}$$

$$c_{trans} = 0.095 c_{cis}$$

- Would there be more or less trans azobenzene at 460 nm? Justify your answer.

There would be more trans at the photostationary state when excited at 460 nm, this is because the relative ratio of molar extinction coefficients favors formation of more of the trans isomer.

- It has been suggested the 360 nm absorption is an $S_0 \rightarrow S_2$ absorption, and the 460 nm band is an $S_0 \rightarrow S_1$ absorption. Suggest which energy levels are involved for each of the two transitions and compare it to stilbene which has a similar structure, but the cis and trans absorptions are 280 & 295 nm respectively.

Image may be found in Section 4.1.6

Azobenzene has n , π and π^* orbitals whereas stilbene only has π and π^* orbitals. Given the different orbital types available in azobenzene the expectation that the transitions will have a wider energy spacing than in stilbene where we just have the bonding and antibonding π set, and so we would expect these energy levels to be more closely spaced.

Chapter 5

Emission of Light

5.0.1 Learning Objectives

At the end of this block you should be able to:

- Define singlet and triplet states in terms of molecular orbitals
- Explain the mirror symmetry observed in absorption/fluorescence spectra.
- Construct and explain the Jablonski diagram.
- Define the rates of the processes in the Jablonski scheme and give typical values for the rate constants.
- Define singlet and triplet lifetimes and describe how they are measured.
- Analyse decay curves to obtain lifetimes of excited states.
- Define and derive the quantum yield of fluorescence and phosphorescence.

Only the light absorbed by a molecule can produce photochemical change in a molecule. If a species absorbs radiation then one particle is excited for each quantum of radiation absorbed. This concept of light having to be absorbed

before anything else can happen is an obvious one, but is an important concept in itself, it goes on to define the concept of quantum yield of a process, effectively the proportion of excited states which ‘decay’ to any given pathway.

The reaction coordinate, which appears on our x axis of any plot with potential wells, is essentially a measure of the net difference between the wavefunction of ground and excited state, the bigger the difference the bigger the offset on the reaction coordinate axis.

Wavefunctions are hard to visualise though, and so we may also consider it the sum of the net difference in position of every atom between the ground and excited state, the bigger the structural change, the bigger the difference in wavefunction, and so the bigger the offset on the reaction coordinate axis.

5.1 Fluorescence & Phosphorescence

Just as there are ‘spin forbidden’ and ‘spin allowed’ processes for absorption, when the molecule is in an excited state there are allowed and forbidden path ways of deactivation of the excited state.

If we look at a Jablonski diagram, figure 19, which you should already be familiar with, it shows the range of photophysical deactivation pathways of an excited state.

A Jablonski diagram (figure 5.1) usually uses a short hand to refer to the excited states, where a ground state, if it is a singlet, is referred to as S_0 , and each subsequent excited state is referred to as S_1 , S_2 etc.

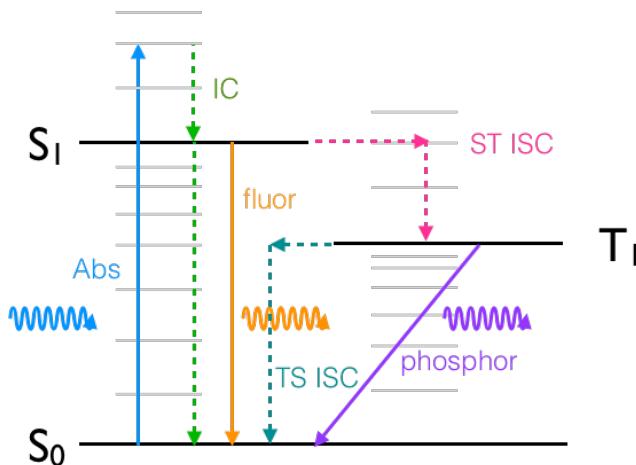


Figure 5.1: Jablonski diagram showing: A, the absorption of a photon, and the photophysical deactivation pathways, IC, internal conversion, ISC, intersystem crossing, F, fluorescence & P, phosphorescence.

Triplet states are referred to as T_1 , T_2 ... if the ground state is not a triplet then there is no state labelled T_0 . All excited states are referred to by integer values of 1 or more.

A note on Jablonski diagrams: transitions which are either allowed or non-radiative and therefore not governed by selection rules appear as either horizontal or vertical transitions, forbidden transitions (such as phosphorescence) are indicated by transitions which occur at an angle. Radiative transitions (those that occur with either the absorption or emission of a photon) are shown as solid lines with the incident or excident photon shown as a wave like arrow. Non radiative transitions are shown as dashed (or alternatively wiggly) lines to clearly differentiate them from radiative transitions. Transitions with a change of spin are isoenergetic (shown as horizontal lines) before decaying via internal conversion (vertical lines). A

Jablonski diagram is a sketch of the energy levels and should not include details of the potential wells of either the ground or excited state.

The spin selection rule which applies for absorption transitions $\Delta S = 0$, so formally emission transitions from the S_1 to S_0 are spin allowed, allowed emission transitions are referred to a fluorescence.

Emission transitions that involve a change in spin of the electron, such as from T_1 to S_0 , are spin forbidden, and referred to as phosphorescence. The excited state can also decay via non-radiative transitions and other routes.

Table 5.1 the deactivation pathways from excited states:

Table 5.1: The excitation and decay pathways in molecules.

<i>'Allowed transitions'</i>		
Singlet-singlet absorption	fluorescence	$S_0 + h\nu \rightarrow S_1$
Singlet-singlet emission		$S_1 \rightarrow S_0 + h\nu'$
<i>'Forbidden transitions'</i>		
Singlet-triplet absorption	fluorescence	$S_0 + h\nu \rightarrow T_1$
Triplet-singlet emission		$T_1 \rightarrow S_0 + h\nu''$
<i>'Other transitions'</i>		
Internal conversion	(vibrational relaxation)	$S_1 \rightarrow S_0 + \text{heat}$
	(vibrational relaxation)	$S_1 \rightarrow T_1 + \text{heat}$
		$T_1 \rightarrow S_0 + \text{heat}$
<i>Other pathways</i>		
Quenching of excited state		$S_1 + Q \rightarrow$
Chemistry from excited state		$S_0 + Q + \text{heat}$
		$S_1 + Q \rightarrow$
		$S_0 + Q^* + \text{heat}$
		$T_1 + Q \rightarrow$
		$S_0 + Q + \text{heat}$
		$T_1 + Q \rightarrow$
		$S_0 + Q^* + \text{heat}$
		$S_1 \rightarrow \text{new/changed molecule}$

The spin selection rule has important consequences on the lifetime of singlet and triplet states. The spin allowed process of fluorescence usually has a short lifetime (less than 10 ns), whereas the spin forbidden process of phosphorescence has a significantly longer lifetime, which can be as long as several seconds.

Vibrational relaxation occurs in the ps timescale, consequently the excited

state decays vibrationally to the S_1 , $v' = 0$ state and emission essentially always occurs from the ground vibrational electronically excited state. This is an important concept in photophysics and is called Kasha's rule.

5.1.1 Kasha's Rule

An excited state always emits from the lowest energy level of that spin multiplicity state.

The Frank-Condon principle applies equally to emission processes and the excited state will decay to a number of available vibrationally excited states of the ground electronic state. After relaxation to the $v' = 0$ level emission occurs to each of the available vibrational states of the ground state. The relative intensity of each band depends upon the coupling or overlap integral between the ground and excited states, as shown in figure 5.2.

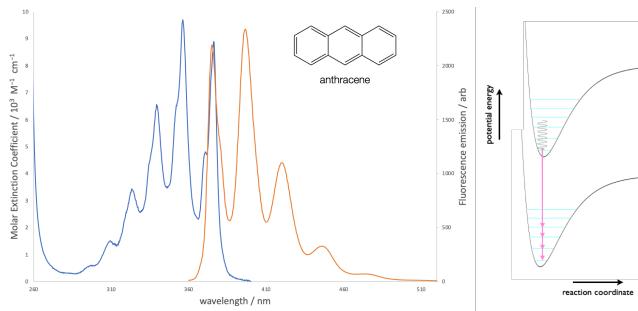


Figure 5.2: The absorption and emission spectra of anthracene in ethanol solution. Both the absorption and emission spectra show vibrational fine structure, due to transitions to discrete vibrational energy levels in the excited and ground states respectively.

5.2 Stoke's shift

The Franck-Condon factors in absorption and emission is clearly evident by the symmetry shown in the emission of most organic chromophores. The Stoke's shift, the difference between the $v = 0$ to $v' = 0$ absorption peak and the $v' = 0$ to $v = 0$ emission peak, can simply be explained by the offset of the ground and excited state potential wells (figure 5.3). Absorption happens from the centre of the $v = 0$ ground state and emission from the centre of the $v' = 0$ excited state, however this does not explain why the two curves are offset from one another.

The axis title of 'reaction coordinate' is used ubiquitously, but is poorly understood. The simple idea that a bond becomes longer (as would be the

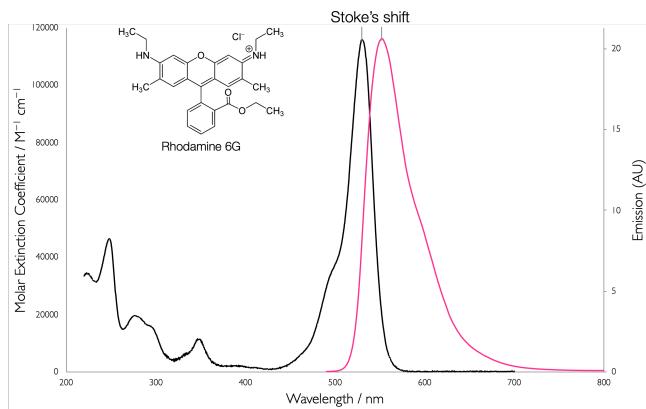


Figure 5.3: The absorption and emission spectra of rhodamine 6G showing the Stokes shift of the two spectra.[Data from [OMLC](<https://omlc.org/spectra/PhotochemCAD/html/083.html>), accessed July 2020]

case for a diatomic molecule) does not apply as we are considering the whole molecule.

The Stokes shift can be far more easily accounted for if we consider what is going on in the molecule. Initially the molecule exists in a state optimised for that state, the shape and structure of the dye molecule and the arrangement of solvent around it are all minimised to give the lowest energy configuration. We shall consider the most simple case of the transitions between the ground vibrational states of the HOMO & LUMO.

Absorption of a photon is a very fast process, in the order of femtoseconds, and is significantly faster than molecular vibrations or structural reorganisation. Consequently the molecule becomes excited into a state which has exactly the same structural configuration and salvation shell as the ground state, however this is now no longer the most optimal state, figure 5.4.

Since the excited state is relatively stable (ns) when compared with the time taken for vibrations and molecular restructuring (ps) the excited state dye molecule and the solvent will reorganise to minimise the energy of the excited state, however in doing so this destabilises the ground state of the molecule.

The actual process of emission is again very fast (ps) and so there is no opportunity for the structure of the chromophore or solvation sphere to reorganise. Upon deactivation this state, optimised for the excited state reorganises, lowering the energy of the ground state and destabilising the excited state.

As a consequence of this stabilising action of the ground and excited states the energy gap of absorption is larger than the energy gap of emission, or the

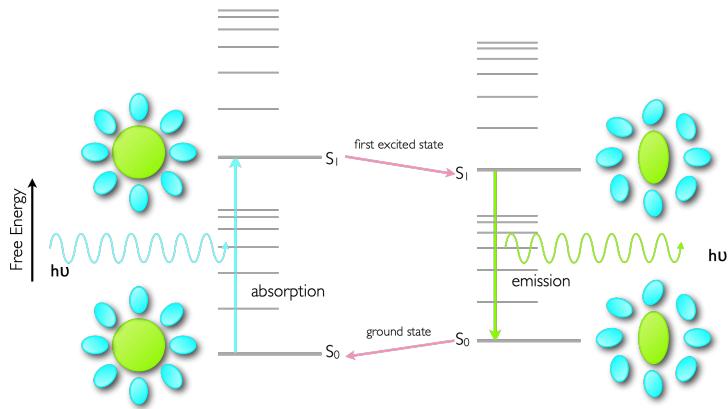


Figure 5.4: Sketch of the photophysical processes involved in the Stokes shift of emission. The ground state structure of the dye and solvent sphere (lower left) is maintained upon almost instantaneous absorption of a photon, however this then relaxes to an optimised configuration of dye and solvent for the excited state (top right), however this destabilises the ground state.

emission of a photon is redshifted from the absorption. This explains the origin of Stokes shift, and explains the offset the Morse curves of the ground and excited states which have previously been used to explain the origin of the Stokes shift.

5.3 Emission from S_2

Kasha's rule is almost always true, but there are known exceptions to this rule, one of these examples is azulene. Azulene has a broad absorption band between 500-600 nm, however emission from this molecule occurs at below 400 nm. If we look at figure 5.5 which shows the absorption and emission spectra of azulene it can be seen that the emission spectra is a mirror image of S_0-S_2 (excitation into the second excited electronic state) absorption band, indicating the emission is also from the S_2 to S_0 states.

The molar extinction coefficient is also significantly higher for this S_0-S_2 absorption indicating a much greater overlap integral of these two states.

What causes the unusual photophysical example? In short there is a very large energy gap between the S_2 and S_1 excited states this has the effect of greatly reducing the rate of internal conversion. There is a very weak emission from the S_1 to S_0 , in this case the energy gap is relatively small and there is an efficient deactivation of the excited state by internal conversion.

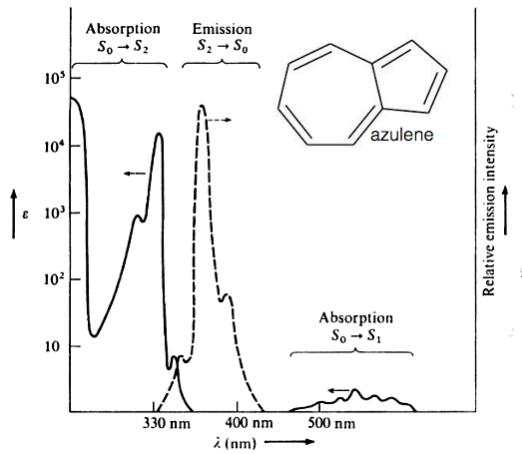


Figure 5.5: The absorption (solid line) and emission (dashed line) spectrum of azulene in ethanol. Note the emission appears in the near UV part of the spectrum.

5.4 Rates of Photophysical Processes & Quantum Yield

Fluorescence (and phosphorescence) are first order processes, just like radioactive decay. The probability of a molecule decaying from an excited state depends only the concentration of that excited state, $[A^*]$.

Rate of decay of $[A^*]$:

$$\frac{d[A^*]}{dt} = -k_f [A^*] \quad (5.1)$$

It is usual to refer to the ‘lifetime’ of photophysical processes, and it is a term already used at the start of this chapter, the lifetime of a process, usually given the symbol τ , is the reciprocal of the rate constant for that process, so $\tau_f = 1/k_f$. Often the natural lifetime of a dye is used, this assumes that there are no other deactivation pathways from the excited state and all of the energy is lost as fluorescence (or the fluorescence quantum yield is 1). The natural lifetime of fluorescence is given the symbol τ_0^0 , it is the maximum a lifetime can be and is solvent dependent.

If there is more than one possible deactivation pathway from the excited state is often useful to be able to quantify these values. The quantum yield of a process describes the efficiency of each process, when a molecule absorbs a photon of light it generates an excited state; there are a number of different

ways the molecule can decay from this excited state (see section 5.1, figure 5.1). The quantum yield for each process is essentially the ratio of reactions proceeding down each pathway and the number of photons absorbed, quantum yield has no units.

$$\phi = \frac{\text{number of reactions proceeding down a given pathway}}{\text{number of photons absorbed}} \quad (5.2)$$

Table 5.2: The deactivation pathways of an excited state, with the associated rate constants and quantum yields.

		Rate of Process	Quantum Yield
Absorption	$S_0 + h\nu \rightarrow S_1$	I	
Fluorescence	$S_1 \rightarrow S_0 + h\nu'$	$k_f[S_1]$	ϕ_f
Internal conversion	$S_1 \rightarrow S_0 + \text{heat}$	$k_{ic}[S_1]$	ϕ_{ic}
Intersystem crossing <i>singlet to triplet</i>	$S_1 \rightarrow T_1 + \text{heat}$	$k_{ST}[S_1]$	ϕ_{ST}
Phosphorescence	$T_1 \rightarrow S_0 + h\nu''$	$k_p[T_1]$	ϕ_p
Intersystem crossing <i>internal conversion triplet to singlet</i>	$T_1 \rightarrow S_0 + \text{heat}$	$k_{TS}[T_1]$	ϕ_{TS}

According to the Stark-Einstein law the total quantum yield of all processes is unity (one), however if secondary reactions occur (such as radical propagation) it may be more than 1. The quantum yield and rate of a process are intrinsically linked. It should be implicit that the faster the rate of a process (the bigger the rate constant) the bigger the quantum yield of that process; in fact the quantum yield of a process, such as fluorescence, is the rate of fluorescence divided by the sum of the rates of all of the deactivation processes (equation (6.1)).

$$\phi_f = \frac{k_f^0}{k_f^0 + k_{ic} + k_{ST} + k_{\text{other}}} \quad (5.3)$$

(Examples of expressions for other processes may be found in Section ?? equations (??), (??) & (??).)

When there are more processes occurring then the singlet lifetime (sometimes called the fluorescence lifetime) will be smaller than the natural fluorescence lifetime, because now the denominator term is bigger (equation (6.2)).

$$\tau_s = \frac{1}{k_f^0 + k_{ic} + k_{ST} + k_{\text{other}}} \quad (5.4)$$

If we assume that there is no quenching of the system (this is discussed later) and no chemistry the only decay pathways from the excited state are emission (fluorescence), internal conversion and intersystem crossing to the triplet state, and so consequently only these terms would appear in the equation for the quantum yield. Processes that occur from the triplet state are independent, and therefore none of these terms occur in quantum yield or lifetime equations from the singlet, fluorescent state.

For processes that occur from the triplet state, then obviously, the quantum yield will depend upon the total number of triplet states produced, which is given by the quantum yield of intersystem crossing. So the quantum yield of phosphorescence is given by:

$$\phi_p = \frac{k_p^0}{k_p^0 + k_{TS} + k_{\text{other}}} \quad (5.5)$$

Quantum yields of photophysical processes can be difficult to measure, in part because it is difficult to easily quantify the number of photons going into a sample, and because emission occurs at all angles from the sample. A technique called ‘actinometry’ is usually used where the emission from an unknown sample is compared with that of a well characterised sample.

A good actinometer should have a good overlap in both the absorption and emission spectra, a high molar absorption coefficient, it should be photochemically stable and have a high emission quantum yield, a selection of well characterised actinometers are included in figure 7.1. Measurements on the absorption at the excitation wavelength, and the full emission spectra are required for both the unknown and the actinometer, as well as details on the reactive index of the solvents used. For best results a monochromatic light source should be used (a laser) and the same excitation wavelength must be used for both samples; the absorbance of the solutions should also be low (<0.1).

Usually emission spectra are measured as intensity against wavelength, but when ‘integrating’ the emission of the unknown and actinometer it is important that they are integrated over a linear space (energy or frequency) which can make calculations more difficult (but perfectly doable in something like Excel). Equation (5.6) shows the relationship between the fluorescence quantum yield of an unknown, Φ , with that of a reference actinometer, Φ_R . A is the absorbance at the excitation wavelength, and n the refractive index of

the solvent. Here, because all terms appear as ratios there is no need to worry about units with anything.

$$\phi = \phi_R \frac{\int I d\bar{\nu}}{\int I_R d\bar{\nu}} \frac{A_R}{A} \frac{n^2}{n_R^2} \quad (5.6)$$

The integrated emission intensity, $\int I d\bar{\nu}$, of the unknown and reference is here shown as being integrated over wavenumber, but so long as it is integrated over linear space¹ it does not matter if frequency or energy is used instead as any units from the terms cancel.

Table 5.3: The photophysical details of well characterised actinometers.

Reference	Solvent	λ_{ex}/nm	Φ_{em}	$\lambda_{em,max} / \text{nm}$
quinine sulfate	0.1 M H ₂ SO ₄	350	0.577	444
fluorescein	0.1 M H ₂ SO ₄	496	0.95	521
rhodamine 6G	ethanol	488	0.94	530
cresyl violet	methanol	540	0.54	622

5.5 Fluorescence lifetime

Spontaneous emission (either fluorescence or phosphorescence) from an excited state is a first order process, and so shows similar kinetics to that which you have already studied, figure 5.6 & equation (5.7).

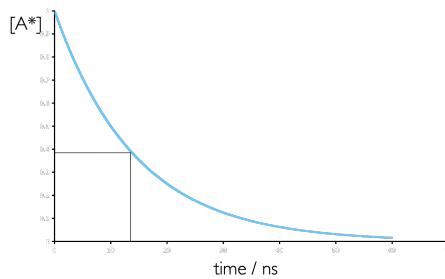


Figure 5.6: Emission from an excited state is a first order process, whereby an excited state decays spontaneously. The emission lifetime is the time required for the concentration of an excited state to fall to $1/e$ of its initial value.

¹The term linear space implies the steps are the same size, energy is the more fundamental unit, since wavelength is a reciprocal function of energy wavelength space is non linear.

However, instead of using rate constants when talking about this decay (as seen in Section 5.4 photochemists tend to talk about emission lifetimes, τ .

The fluorescence lifetime is defined as the time required for the concentration of excited state to fall to $1/e$ (~ 0.3679) of the initial concentration.

So for an excited singlet state the fluorescence lifetime is related to the concentrations as follows:

$$[S_1]_t = [S_1]_0 e^{-\frac{t}{\tau}} \quad (5.7)$$

where τ is the value given in equation (6.2).

5.6 Before Completing this Section

To support the material in this section it is suggested you read chapters 3 & 4 of Wardle ‘Principles and Application of Photochemistry’.

Chapter 6

Workshop Questions for Week 3

6.1 Questions

6.1.1 Short mathematical question - Quantum Yield and lifetime

The quantum yield and lifetime of a dye were measured to be 0.43 & 2.6 ns respectively. What is the natural lifetime?

(I will use MCQs and UniDoodle to ask this in class)

You can find the answer to this question in 6.2.1

6.1.2 Short conceptual question - Effect of other processes

For a given value of Φ_0 what happens to the lifetime and quantum yield as k_{IC} and k_{ISC} increases?

(This will be a discussion question)

You can find the answer to this question in 6.2.2

6.1.3 Conceptual question - Effect of structure

Fluorescein (figure 6.1) in basic aqueous solution has a quantum yield of fluorescence, Φ_f , of 0.95, and fluorescence lifetime, τ_f , of 4.1 ns.

(This will be a discussion question)

You can find the answer to this question in 6.2.3

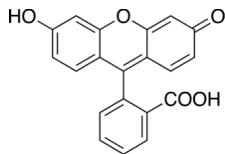


Figure 6.1: The structure of the fluorescent molecule fluorescein

6.1.4 Conceptual question - lack of symmetry in spectra

The absorption and emission spectrum of fluorescein is shown in figure (6.2)

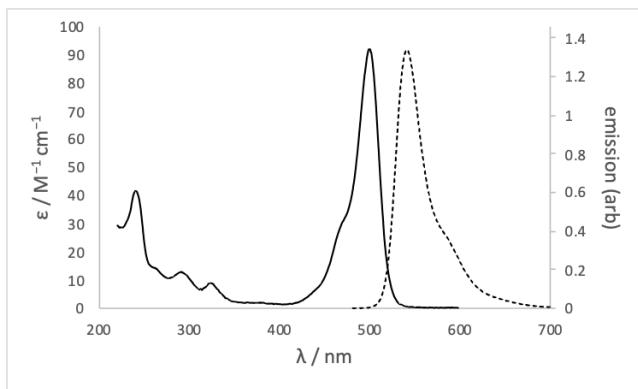


Figure 6.2: The absorption (solid) and emission (dashed) spectrum of fluorescein in basic ethanol.

Why are the absorption bands between 200 – 350 nm not reflected in the emission spectrum?

You can find the answer to this question in 6.2.4

6.1.5 Conceptual question - Stokes' shift

The inorganic dye $[\text{Ru}(\text{bpy})_3]^{2+}$ has a measured lifetime in water of 580 ns and a natural lifetime of 13.8 μs . The spectrum is shown in figure 6.3. What is the origin of the large Stokes' shift in this system?

Data from Shi *et al.*, Synthesis and characterization of phosphorescent two coordinate copper(I) complexes bearing diamidocarbene ligands. Dalton Trans., 2017, 46, 745-752.

You can find the answer to this question in 6.2.5

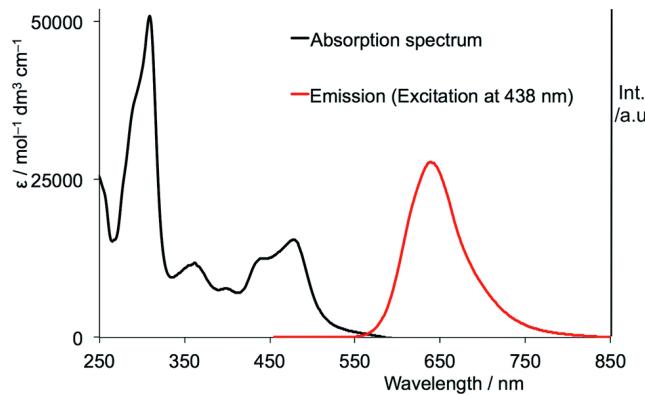


Figure 6.3: The absorption (black) and emission (red) spectrum of ruthenium tris bipyridine in water.

6.1.6 Conceptual question - the effect of binding on emission

The asymmetric cyanine dye YO-Pro-1 is a DNA stain because it has a large increase on fluorescence emission when binding to DNA. The lifetime in free solution is around 2 ps and when bound to DNA is 2.4 ns. What is the structural origin of the large increase of emission upon binding?

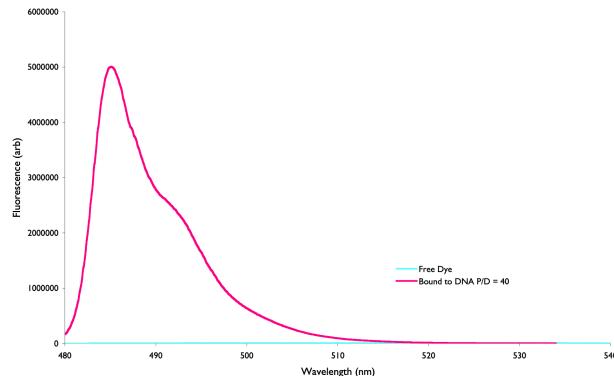


Figure 6.4: The emission spectrum of the chromophore YO-Pro-1 when free in aqueous solution (blue) and when bound to DNA (pink)

(This will be a discussion question)

You can find the answer to this question in 6.2.6

6.1.7 Extended question - Properties of Ethidium Bromide (Example Exam Question)

Ethidium bromide (EB, figure 6.5) is used as a DNA stain, which is essentially non-fluorescent in aqueous solution, but shows a strong enhancement of emission upon binding to double stranded DNA (which has a negatively charged backbone).

Emission is almost exclusively from the singlet excited state, but a triplet state has been shown to exist, which emits with a low quantum yield ($\Phi_P = 0.00006$).

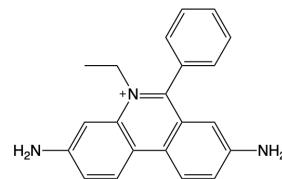


Figure 6.5: The structure of the cationic ethidium bromide chromophore.

- Sketch a Jablonski diagram for the processes you know to occur.
- The molar extinction coefficient, ϵ , of EB has been measured to be $78500 \text{ M}^{-1} \text{ cm}^{-1}$. What factors contribute to EB having such a high extinction coefficient?

The following spectra, lifetimes and quantum yield have been measured for EB in different free solution and DNA systems:

Table 6.1: The lifetimes and quantum yields of ethidium bromide in aqueous solution and when bound to DNA in protiated and deuterated systems.

	/ ns	Φ_f
H ₂ O (no DNA)	1.6	0.012
D ₂ O (no DNA)	6.3	
DNA	28.3	0.220
DNA (deuterated)	38.4	

- What factors likely lead to an enhancement of fluorescence quantum yield upon binding to DNA?
- Show that the natural lifetime of EB is 129 ns.
- What is the origin of the large Stoke's shift ($\sim_{\text{max ex}} = 520 \text{ nm}$, $\sim_{\text{max em}} = 608 \text{ nm}$)

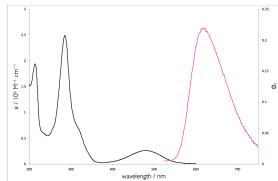


Figure 6.6: The absorption (black) and emission (pink) spectrum of ethidium bromide when bound to DNA.

- What transitions are responsible for the absorption features in the: * 400-600 nm range * 200-350 nm range
- Why are the features in the 200-350 nm range not replicated in the emission spectrum?
- Why does deuterating the solvent (or DNA) effect the lifetime of the excited state?
- What effect would freezing the samples have on the lifetime, fluorescence quantum yield & phosphorescence quantum yield.

A study of the thermodynamics of the dye DNA system measured the binding constant of EB with DNA to be 1.05×10^6 .

- Why is the measured quantum yield for a system containing 2 μM EB and 20 mM DNA only 0.18?
- Why would increasing the ionic strength of the solution, increase the fluorescence intensity of EB in solution with DNA?

You can find the answer to this question in 6.2.8

6.2 Answers

6.2.1 Short mathematical question - Quantum Yield and lifetime

The fluorescence quantum yield (equation (6.2)) and fluorescence lifetime (equation (6.2)) are related to the rate constants for the different decay processes.

$$\phi_f = \frac{k_f^0}{k_f^0 + k_{ic} + k_{ST} + k_{\text{other}}} \quad (6.1)$$

$$\tau_f = \frac{1}{k_f^0 + k_{ic} + k_{ST} + k_{\text{other}}} \quad (6.2)$$

The natural lifetime is the lifetime when the only decay pathway is fluorescence emission (equation (6.3)).

$$\tau_f^0 = \frac{1}{k_f^0} \quad (6.3)$$

Combining equations (6.1) and (6.2):

$$\frac{\tau_f}{\phi_f} = \frac{\frac{1}{k_f^0 + k_{ic} + k_{ST} + k_{other}}}{\frac{k_f^0}{k_f^0 + k_{ic} + k_{ST} + k_{other}}} = \frac{1}{k_f^0} = \tau_f^0$$

$$\text{Consequently } \tau_f^0 = \frac{\tau_f}{\phi_f} = \frac{2.6}{0.43} = 6.0 \text{ ns}$$

6.2.2 Short conceptual question - Effect of other processes

If k_f^0 is fixed (which it is for a given chromophore) then as we adjust the other rate constants in equation (6.2) the emission lifetime (the time it takes for the intensity to reduce to $1/e$ of the initial intensity) decreases, as shown in figure 6.7.

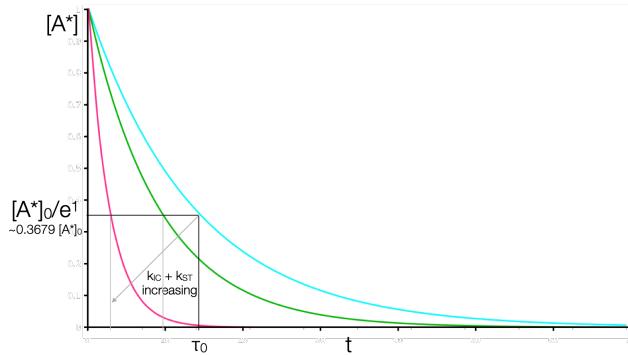


Figure 6.7: The variation of the emission lifetime (the time it takes to decay to $1/e$ of the initial concentration, for a constant rate of fluorescence emission but increasing rates of non-radiative decay).

6.2.3 Conceptual question - Effect of structure

The only things that affect the quantum yield and lifetime (equations (6.2) and (6.2)) are the different rate constants involved. The quantum yield is a measure of the proportion of excited states which decay by each pathway.

Therefore if the quantum yield of fluorescence is near unity then neither the rate of intersystem crossing or internal conversion is large. Intersystem crossing may occur due to the change in orbital type offered, but given the lack of heavy atoms this rate is unlikely to be high and will probably be a negligible proportion given the high rates of other processes.

The rate of intersystem crossing is more efficient for rotations than vibrations (due to the closer spacing of rotational energy levels), therefore if rotations can occur the rate of internal conversion tends to be quite high. In this case there is only one likely bond that could rotate to loose the excited state of the system. This is the single bond linking the two ring systems. Rotation around this bond is severely inhibited though due to the presence of the acid group which restricts rotation.

Please note that it is rotations around bonds that cross the conjugation within the molecule that allow for the most efficient energy loss.

6.2.4 Conceptual question - lack of symmetry in spectra

Kasha's rule states "*An excited state always emits from the lowest energy level of that spin multiplicity state.*" Consequently we can excite into any available orbital, but the excited state will rapidly decay into the lowest energy level of that multiplicity. So although we are exciting from S_0 into S_1 , S_2 , S_3 ... any excited state will rapidly decay by internal conversion down to the S_1 excited state and all emission will occur from this state.

Kasha's rule is a consequence of the energy gap law which I alluded to when talking about the special case of azulene, but this will be covered in a video specifically talking about the energy gap law.

6.2.5 Conceptual question - Stokes' shift

Firstly we need to separate two things - the first is just the origin of the Stokes' shift due to the transition of an electron occurring very quickly leaving an optimised state to occupy an unoptimised state, which later relaxes. A typical Stokes' shift is ~ 20 nm, the Stokes' shift in this example is more than 150 nm... why?

If we consider the system that we have at hand, in this case $[\text{Ru}(\text{bpy})_3]^{2+}$ we can note the HOMO-LUMO transition is an MLCT. This has a large difference in structure of the HOMO & LUMO with a redistribution of charge which will also have a large effect on the arrangement of the solvent shell. Given this large difference in optimised structure there would be a large shift.

However this isn't the only factor, the natural lifetime of 13.8 μs would indicate that the emission transition is phosphorescence, this spin forbidden transition has electrons aligned in parallel which is lower than the corresponding singlet state. So after excitation there is a change in multiplicity

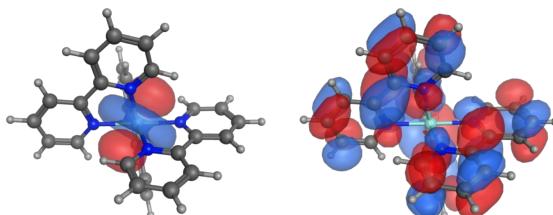


Figure 6.8: The HOMO (left) and LUMO (right) of ruthenium (II) tris bipyridine clearly showing the MLCT nature of the transition.

of the excited state, this lowers the energy of the system further moving the λ_{max} of emission to lower energy (longer wavelength).

Data from Shi *et al.*, Synthesis and characterization of phosphorescent two coordinate copper(I) complexes bearing diamidocarbene ligands. Dalton Trans., 2017, 46, 745-752.

6.2.6 Conceptual question - the effect of binding on emission

The chromophore YO-Pro-1 has a methine bridge between the two halves of the chromophore, this offers efficient loss of the excited state *via* internal conversion (see question 6.1.3). As noted earlier free rotation around a bond which cuts across the conjugation is particularly efficient in deactivating an excited state.

If this pathway is inhibited by binding to the DNA (the dye intercalates between the base pairs but any dye-DNA interaction would inhibit this mechanism of excited state decay) then the fluorescence lifetime will increase (equation (6.2)).

6.2.7 Short conceptual question - effect of polar solvents on emission

6.2.8 Extended question - Properties of Ethidium Bromide (Example Exam Question)

If phosphorescence occurs there must be some intersystem crossing in both directions.

There is a rigid structure indicating a similar ground and excited state structure, which would lead to a good overlap integrals between the HOMO and LUMO. The molar extinction coefficient is a measure of this overlap integral.

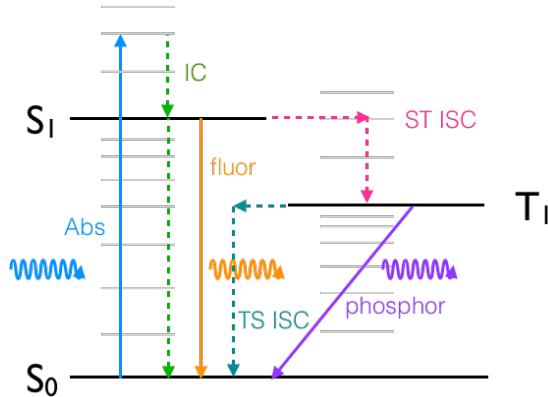


Figure 6.9: A Jablonski diagram indicating the allowed transitions within the system, absorbance , internal conversion, intersystem crossing, fluorescence and phosphorescence.

The chromophore single bond between the two halves of the chromophore, this offers efficient loss of the excited state *via* internal conversion (see question 6.1.3. As noted earlier free rotation around a bond which cuts across the conjugation is particularly efficient in deactivating an excited state.

If this pathway is inhibited by binding to the DNA (the dye intercalates between the base pairs but any dye DNA interaction would inhibit this mechanism of excited state decay) then the fluorescence lifetime will increase (equation (6.2)).

$$\phi_f = \frac{k_f^0}{k_f^0 + k_{ic} + k_{ST} + k_{\text{other}}}$$

$$\tau_f = \frac{1}{k_f^0 + k_{ic} + k_{ST} + k_{\text{other}}}$$

The natural lifetime is the lifetime when the only decay pathway is fluorescence emission (equation (6.3)).

$$\tau_f^0 = \frac{1}{k_f^0}$$

Combining equations (6.2) and (6.2):

$$\frac{\tau_f}{\phi_f} = \frac{\frac{1}{k_f^0 + k_{ic} + k_{ST} + k_{other}}}{\frac{k_f^0}{k_f^0 + k_{ic} + k_{ST} + k_{other}}} = \frac{1}{k_f^0} = \tau_f^0$$

Consequently $\tau_f^0 = \frac{\tau_f}{\phi_f} = \frac{1.6}{0.012} = 129$ ns

We are told the transition is fluorescence, but it has a very long natural lifetime. This long lifetime would indicate that despite the transition being selection rule allowed there is a structural explanation for the inefficiency of this process.

This would indicate that there is now a poor overlap integral indicating a poor overlap of the excited state and ground state structures, or there is a large difference in the structures of the excited and ground states.

This large difference in structure would be accounted for in the Stokes' shift - a large deviation in structure would lead to a larger energy loss upon the optimisation of the excited state.

The molecule has hetero atoms and so there will be both $n - \pi^*$ and $\pi - \pi^*$ transitions, the $n - \pi^*$ transitions will occur at lower energy (higher wavelength) and be the HOMO-LUMO transition.

Kasha's rule states *An excited state always emits from the lowest energy level of that spin multiplicity state*. Consequently we can excite into any available orbital, but the excited state will rapidly decay into the lowest energy level of that multiplicity. So although we are exciting from S_0 into S_1 , S_2 , S_3 ... any excited state will rapidly decay by internal conversion down to the S_1 excited state and all emission will occur from this state.

Ultimately because it affects the rates of internal conversion, and so the lifetime is changed as described in equation @??eq:lifetimefluor). We will learn more as to why when we cover the energy gap law.

You would expect the emission lifetime and quantum yield to increase on freezing (as described in equation @??eq:lifetimefluor)) as the efficiency (and therefore rate) of internal conversion mechanisms (particularly rotational relaxation) would be dramatically reduced in solid systems.

This will be due to there being two states - one where the chromophore is bound, one where it is not. It is an equilibrium so not all dye will be bound to DNA. When not bound there is no fluorescence enhancement. The quantum yield we see is in fact a weighted average between the two states.

This is a problem solving question and any reasonable attempt would have been given credit. It is demonstrating an ability to think about data and the problem in a scientific manner.

This would push the equilibrium to form the dye:DNA complex increasing binding, as the binding increases the proportion of molecules displaying the higher quantum yield increases and the overall emission intensity of the sample increases.

This is a problem solving question and any reasonable attempt would have been given credit. It is demonstrating an ability to think about data and the problem in a scientific manner.

Chapter 7

Non-emmisiive Deactivation of the Excited State

7.0.1 Learning Objectives

At the end of this section you should be able to:

- Discuss the heavy atom effect.
- Understand the difference between static and dynamic quenching.
- Graphically determine quenching constants for static, dynamic and mixed method quenching.
- Describe mathematically the effect of quenching on emission quantum yields and lifetime.
- Describe the mechanisms of non-collisional energy transfer.
- Demonstrate understanding of the factors which affect the rates of non-collisional energy transfer.

7.1 The Energy Gap Law

The natural lifetime of fluorescence assumes that there is no other available pathway for deactivation of the excited state. However, the effects of internal conversion and intersystem crossing have already been seen by the presence of rate constants for internal conversion (IC) and intersystem crossing (ST) in the equations for quantum yield of fluorescence. The rate of non-radiative decay is given by the energy gap law (a formalised equation is not given here) which depends upon a number of factors but principally the energy gaps (ΔE) between electronic levels.

Non radiative decay become increasingly unfavourable as the energy gaps become bigger, and the rate of non-radiative decay decreases exponentially with increasing energy gaps. The energy gaps between higher electronic states

66CHAPTER 7. NON-EMMISIVE DEACTIVATION OF THE EXCITED STATE

(S_2 , S_3 , S_4 etc.) are usually significantly smaller than the S_1 - S_0 energy gap leading to rapid internal conversion down to the lowest (singlet) excited state, which is the origin of Kasha's rule.

Table 7.1: The lifetime of the excited state increases dramatically as the energy gap between the ground and excited state increases due to a decrease in internal conversion. The quantum yield of emission increases as the rate constant for internal conversion decreases..

	abs / nm	em / nm	ΔE / eV	/ nw	Φ_{em}
$[Os(phen)_3]^{2+}$	650	720	0.186	260	0.016
$[Os(phen)_2(dppene)]^{2+}$	455	609	0.69	1830	0.138
$[Os(phen)(dppene)_2]^{2+}$	400	530	0.761	3600	0.518

The rate of non-radiative decay also depends upon the vibrations within a molecule (this should be logical since the energy is lost as heat), as well as the size of any structural changes required upon deactivation of the excited state. Since the frequency of vibrations is a factor in the rate of non-radiative decay there are clear isotope effects; particularly when hydrogens are replaced with deuteriums. C-H stretches occur with a frequency around 3000 cm^{-1} , whereas this is much lower at around 2200 cm^{-1} for C-D stretches; because of this the efficiency of vibrational relaxation in the hydrogen sample is considerably greater and the lifetime of the excited state is consequently longer in the deuterated sample.

Now a quantitative value may be determined for the rate of internal conversion it is easy to quantify the rate of intersystem crossing.

For intersystem crossing the efficiency is again determined by the size of the energy gap between the singlet and triplet states.

El Sayed's rule states:

the rate of intersystem crossing is relatively large if the radiationless transition involves a change in the orbital type.*

Consequently intersystem crossing is considerably more efficient in molecules where there are hetero atoms as can be seen when comparing the rates of singlet triplet intersystem crossing between anthracene and 9-acetoanthracene, figure 7.2. However, the presence of a hetero atom does not guarantee a high rate of intersystem crossing, for systems where there is no change in orbital type there rate constants are in the order of 100 to 1000 times more slow than those with a change in orbital type.

Table 7.2: The rates of intersystem crossing for a range of organic chromophores illustrating El Sayed's rule. When orbitals are listed as (for example) , * then it is looking at the pair of electrons in the highest orbitals (the valence HOMO electrons in the ground state), upon excitation an electron is excited from the to the * orbital.

	transition	transition type	k_{ST} / s^{-1}
anthracene	$S_1 (, *)$ $T_1 (, *)$	forbidden	1.4×10^8
acetone	$S_1 (n, *)$ $T_1 (n, *)$	forbidden	5×10^8
benzil	$S_1 (n, *)$ $T_1 (n, *)$	forbidden	5×10^8
biacetyl	$S_1 (n, *)$ $T_1 (n, *)$	forbidden	7×10^7
9-acetoanthracene	$S_1 (, *)$ $T_1 (n, *)$	allowed	$\sim 10^{10}$
benzophenone	$S_1 (, *)$ $T_1 (n, *)$	allowed	$\sim 10^{11}$

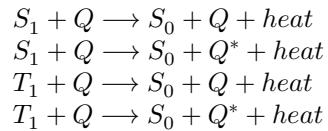
7.2 Collisional Deactivation of the Excited State

The photophysical deactivation pathways described above (emission, internal conversion and intersystem crossing) are not the only ways to deactivate an excited state of a molecule, and you perhaps have already noted the use of *kother* in equations (6.1) - (5.5). Another common pathway of deactivation, particularly in the solution phase, is collisional deactivation commonly known as quenching.

Table 5.1, part of which is repeated below, details all of the possible decay pathways from an excited state. Quenching competes with other processes, and consequently reduces the lifetime and yield of emission processes.

Other pathways

Quenching of excited state



Quenching of the excited state competes with emission from the excited state, consequently reducing the quantum yield of fluorescence (or phosphorescence).

This bi-molecular, second order quenching process follows Stern-Volmer kinetics, equation (7.1), a derivation of which can be found in Wardle, p 88-90).

$$\frac{I_0}{I} = 1 + k_q \tau_0 [Q] \quad (7.1)$$

Where I_0/I is the ratio of the unquenched and quenched steady state intensity of the emission. For dynamic quenching this value will also be the same as ϕ_0/ϕ and τ_0/τ , the ratios of the quantum yield and lifetime respectively, equation (7.2).

$$\frac{I_0}{I} = \frac{\phi_0}{\phi} = \frac{\tau_0}{\tau} = 1 + k_q \tau_0 [Q] \quad (7.2)$$

Since this is a second order term when it appears in the equations (6.1) - (5.5), it appears as $k_q[Q]$, such as indicated in equation (7.3).

$$\phi_f = \frac{k_f^0}{k_f^0 + k_{ic} + k_{ST} + k_q[Q]} \quad (7.3)$$

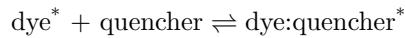
It is often useful to know if the quenching is diffusion or activation controlled, reactions in the solution phase cannot occur faster than the diffusion controlled rate, with second order rate constant k_d , as described in equation 19.

$$k_d = \frac{8RT}{3\eta} \quad (7.4)$$

where T is the absolute temperature, and η the viscosity of the solvent. (It should be noted that there is an issue with units here where a conversion from m^3 to dm^3 is required). Diffusion controlled rates in aqueous, methanol and ethanol solutions are normally in the order of $10^{10} \text{ mol}^{-1} \text{ dm}^3$. Values of rate constant lower than this are termed as activation controlled as only a fraction of the collisions result in a quenching of the excited state of the molecule.

7.3 Static Quenching

Dynamic quenching effects the lifetime of the excited state, as described in equation (7.2). However this is not the case for static quenching; in static quenching an equilibrium exists between dye bound to the quencher and unbound dye and quencher as described below, with equilibrium constant K_q .



$$K_q = \frac{[\text{dye:quencher}]}{[\text{dye}][\text{quencher}]}$$

In this case free dye is emissive with an unquenched lifetime, and dye that exists in complex with the quencher is completely non-emissive. Static quenching follows a very similar form to Stern-Volmer as described above, and again a full derivation may be found in Wardle.

$$\frac{I_0}{I} = 1 + K_s [Q] \quad (7.5)$$

As with dynamic quenching the change in steady state emission is reflected in the change in quantum yield of fluorescence, but this time because the dye exists in either a quenched (completely non fluorescent) or unquenched (fluorescent) state then the fluorescent lifetime is unchanged in static quenching.

In some cases a quencher can act as both a static and dynamic quencher; in this case the lifetime is only affected from the dynamic quenching and so the rate constant may easily be determined, from knowing this it is simple to determine the equilibrium constant. Systems where there is a combination of static and dynamic quenching are easy to spot as plots of I_0/I against the $[Q]$ are curved.

7.4 Förster Resonance Energy Transfer

In addition to static and dynamic quenching, as described above, where the excited state of the dye molecule is lost as heat, it is possible for a dye to be quenched by transferring the energy between two chromophores. There are two distinct mechanisms for this happening the first of which is described here; Förster Resonance Energy Transfer (FRET).

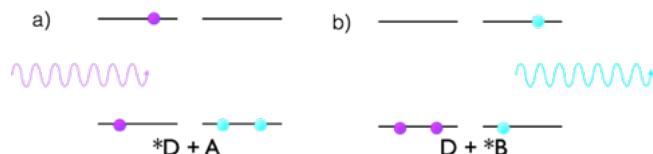


Figure 7.1: The mechanism of Förster resonance energy transfer. a) A donor molecule (purple) is excited by absorption of a photon into an excited state, there is a dipole-dipole interaction with the acceptor molecule (blue) which is initially in the ground state. b) the product of energy transfer with the acceptor now in an excited state which is now capable of emission of a photon of lower energy.

The first thing to note is that FRET is not a collisional process, energy is transferred between a donor and acceptor moiety at a distance, nor is this an

emission of a photon by the donor and reabsorption of this photon by the acceptor.

FRET is a dipole-dipole interaction where the acceptor quenches the excited state of the donor by energy transfer between the two moieties, figure 7.1, leaving the donor in an excited state which then decays as described in the a chapter on emission. This dipole-dipole interaction is a co-aligning of the transition dipole moments (see page Section @ref(#sec:transdipole) on the donor and acceptor molecules - this gives a large dependence on the orientation factors of these two dipoles, the term κ in equation (7.7). In order for there to be an energy transfer there has to be an overlap between the emission spectra of the donor and the absorption spectra of the acceptor (equation (7.6)), the greater, as illustrated in figure 7.2; this is in effect the overlap between the wave function of the excited state donor and the wave function of the ground state donor. The greater this overlap the more efficient the energy transfer.

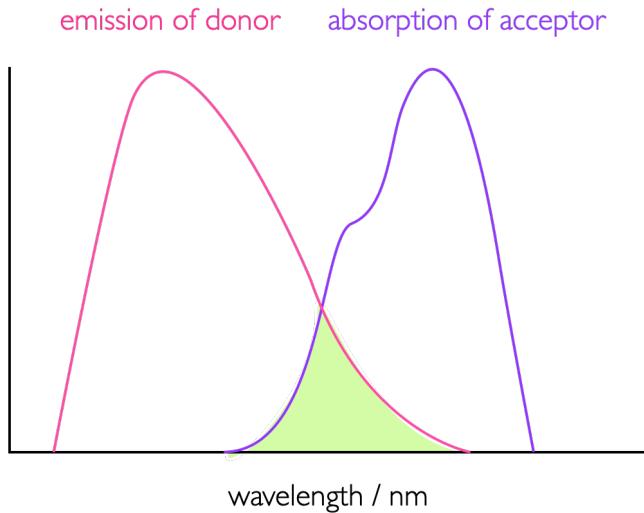


Figure 7.2: The overlap integral, J , (green shaded area) between the emission spectrum of a donor and absorption spectrum of an acceptor as used in FRET, the greater this overlap integral more more efficient the energy transfer.

$$J(\bar{\nu}) = \int_0^{\infty} I_D(\bar{\nu})\varepsilon_A(\bar{\nu})d\bar{\nu} \quad (7.6)$$

The rate of Förster resonance energy transfer has been found empirically to depend upon a number of factors in addition to the overlap integral, J , such as refractive index of the solvent, n , and the dipole orientation factor κ (equation (7.7)). Perhaps more obviously it also depends upon the separation of donor

and acceptor, r , quantum yield of emission of the donor, ϕ_D , and the unquenched lifetime of the donor molecule, τ_D .

$$k_{ET} = \frac{\phi_D \kappa^2}{\tau_D r^6} \frac{9 \ln 10 e^4}{128 \pi^5 N_A n^4} J \quad (7.7)$$

However, equation (7.7), is not simple to use, nor particularly descriptive and a simplified version of the equation tends to be used which has defined a 'Förster radius', R_0 , equations 23 & 24. The Förster radius is the radius at which half of the emission is quenched by the acceptor.

$$R_0^6 = \frac{9 \ln 10 e^4 \phi_D \kappa^2}{128 \pi^5 N_A n^4} J \quad (7.8)$$

$$k_{ET} = \frac{1}{\tau_D} \frac{R_0^6}{r^6} \quad (7.9)$$

Förster distances tend to be between 20 - 90 Å, and due to the parity of this with the size of biological macromolecules FRET is an excellent technique for studying biomolecules. The quantum yield of Förster energy transfer, Φ_{ET} (equation (7.10)), is usually just given the symbol E , for efficiency. As described in equations (7.7) - (7.9) this can be defined by:

$$\phi_E T = \frac{k_E T}{k_f^0 + k_{ic} + k_{ST} + k_{ET}} \quad (7.10)$$

By combining equations (7.9) & (7.10) it becomes easy to see how the efficiency of Förster energy transfer (E or Φ_{ET}) depends upon the distance, and we can see that the efficiency of energy transfer is 0.5 at the Förster distance. Figure 7.3 sketches the distance dependence of efficiency of energy transfer, where it can be seen that at twice the Förster distance the efficiency has dropped almost to 0.

$$E = 1 - \frac{\tau_{DA}}{\tau_D} \quad (7.11)$$

$$E = \frac{R_0^6}{R_0^6 + r^6} \quad (7.12)$$

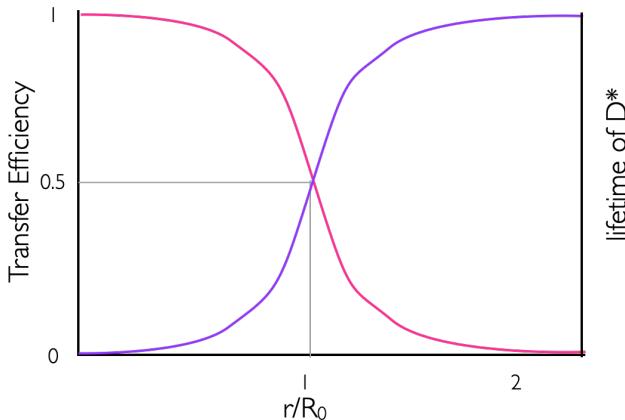


Figure 7.3: The distance dependence on the efficiency of energy transfer between a donor and acceptor in Förster resonance energy transfer (pink line), the purple line represents the distance dependence variation in the measured lifetime of the donor molecule with a maximum value of the unquenched donor of D^* .

The efficiency of energy transfer is calculated by measuring the quenched and unquenched fluorescence lifetime of the donor molecule. More detailed derivations of equations (7.11) & (7.12) may be found in Turro, Principles of Molecular Photochemistry.

Experimentally the process is confirmed to energy transfer at a distance, and not collisional energy transfer, as the rate constant of energy transfer k_{ET} is independent of solvent viscosity (as seen in equation (7.4) & Stern-Volmer kinetics) and rates can be significantly higher than the rate limiting diffusion controlled rates. The energy is transferred by a dipole-dipole interaction.

7.5 Dexter Energy Transfer

Forster resonance energy transfer transfers energy between a donor and acceptor moiety by a dipole-dipole interaction, however there is a second method of energy transfer which takes place by a slightly different mechanism. Dexter energy transfer transfers energy from a donor to acceptor moiety by a concerted exchange of electrons between the chromophores, figure 7.4.

The rate of Dexter energy transfer is again distance dependent and again is dependent upon the overlap integral, J , between the emission of the donor and the absorption of the acceptor (figure 7.2, equation (7.6)). The rate of Dexter energy transfer is given by:

$$k_{ET} = K J e^{-\frac{2r}{L}} \quad (7.13)$$

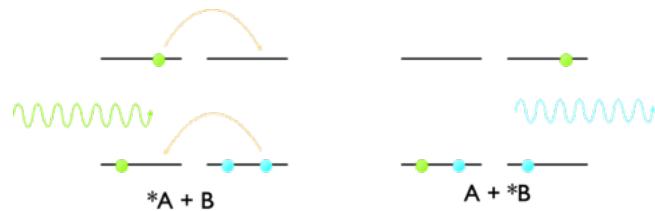


Figure 7.4: The mechanism of Dexter energy transfer, where there is a concerted exchange of electrons from the LUMO of the donor to the empty LUMO of the acceptor and at the same time from the HOMO of the acceptor to the HOMO of donor. The two electrons move at the same time leading to a net exchange of energy from the donor to acceptor.

where J is the overlap integral, K an empirical constant, r the separation between the donor and acceptor, and L another constant which represents the closest possible separation of the donor and acceptor (the sum of the van der Waals radii of the two chromophores).

From examining equation (7.13) it can be seen that the rate of Dexter energy transfer decreases rapidly with increasing separation of donor and acceptor and occurs daily at distances less than 20 Å. Dexter energy transfer is the most dominant mechanism of triplet-triplet quenching.

7.6 Triplet-Triplet Annihilation

Excited state triplets are long lived, due to the spin-forbidden relaxation pathways back to the S_0 state. However, there is a deactivation pathway that occurs by annihilation of two triplet states leading to formation of an excited state singlet and a ground state singlet.

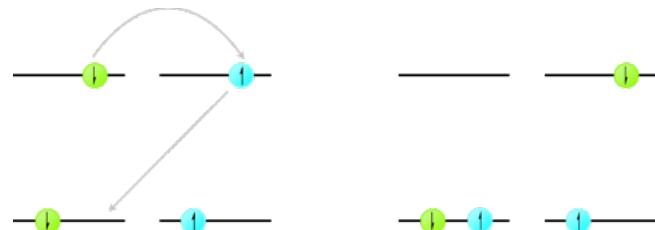
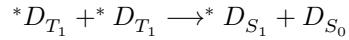


Figure 7.5: The mechanism of triplet triplet annihilation by Dexter energy transfer, concerted electron exchange between a donor and acceptor molecule leading to formation of a singlet excited state on one molecule and a ground state singlet on the other molecule.



Triplet-triplet annihilation is the reason that quantum yields of phosphorescence are almost never unity (1).

7.7 Quenching by Molecular Oxygen

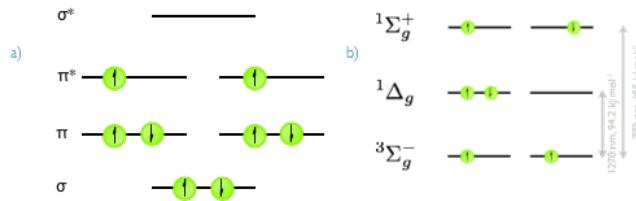
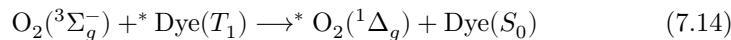


Figure 7.6: a) The valence set of molecular orbitals in molecular oxygen, showing the full electronic arrangement for the ${}^3\Sigma_g^-$ triplet ground state. b) The ${}^*\text{HOMO}$ orbital of molecular oxygen, with the ${}^3\Sigma_g^-$ triplet ground state and the singlet excited states ${}^1\Delta_g$ and ${}^1\Sigma_g^+$. The relative energies of the two excited states are indicated.

Molecular oxygen is unusual because it is a ground state triplet. It is an efficient quencher of excited triplet states because it has two accessible excited singlet states, figure 7.6. It is also capable of quenching excited singlet states, however since the quenching occurs by a Dexter mechanism the overlap integral between excited singlet states and ground state molecular oxygen tends to be very small leading to very small quantum yields of singlet oxygen formation.



Usually the molecular oxygen is excited into the ${}^3\Sigma_g^-$ state, but this rapidly decays to the ${}^1\Delta_g$ state. The ${}^1\Delta_g$ is usually simply referred to as singlet oxygen and is relatively stable because of the spin forbidden deactivation, having a lifetime of a few μs in aqueous solvent and hours in the gas phase.

Chemically singlet oxygen formation is very interesting as it is extremely reactive, and has been shown to be an important factor in a number of mechanistic pathways leading to damage of DNA and proteins.

7.8 Before Completing this Section

To support the material in this section it is suggested you read chapter 6 of Wardle ‘Principles and Application of Photochemistry’.

Chapter 8

Workshop Questions for Week 4

8.1 Short conceptual question - Electronic-vibrational overlap integrals

On the figure sketch the vibrational energy levels in the ground and the excited state.

How would the following affect this overlap integral? 1. Energy gap. 1. The vibrational energy gaps 1. Reaction coordinate (the difference in structure between ground and excited state)

(I will use drawing in UniDoodle to ask this in class, with the second part being a discussion question)

8.2 Short conceptual question - The effect of structural changes on quantum yield

5,10-dihydroindeno[2,1-a]indene and trans-stilbene (figure 8.2) are similar in structure but have very different fluorescent quantum yields of 1.00 and 0.05 respectively, however for trans-stilbene this increases to 0.75 at 77 K. Suggest a reason for the difference in quantum yield of: - the two molecules - the two temperatures

(This will be a discussion question)

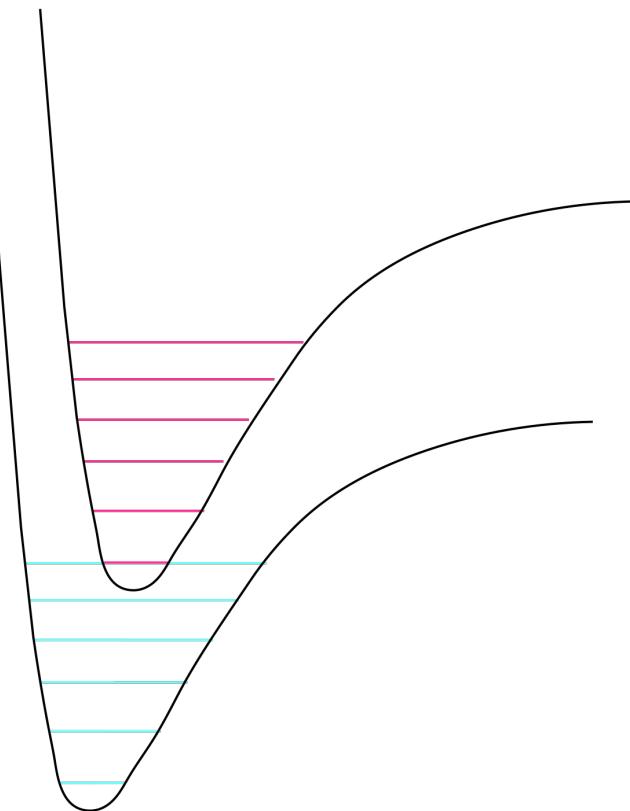


Figure 8.1: The ground and excited state potential wells and the vibrational levels within them.

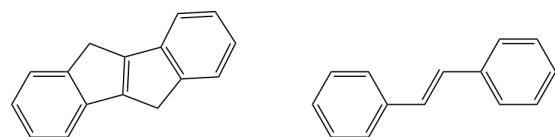


Figure 8.2: 5,10-dihydroindeno[2,1-a]indene (left) and trans-stilbene (right)

8.5. SHORT CONCEPTUAL QUESTION - DEACTIVATION OF EXCITED STATE AROMATIC HYDROCARBS

8.3 Short mathematical question - Effect of the rate of singlet triplet intersystem crossing on th quantum yield of phosphorescence

Why is it likely that the quantum yield of phosphorescence of a sample would increase after the sample is frozen.

(This will be a discussion question)

8.4 Short mathematical question - Determining the quantum yield of phosphorescence

A molecule decays by a combination of internal conversion, intersystem crossing and phosphorescence. What is the quantum yield of phosphorescence?

- $k_{IC} = 2.1 \times 10^{11} \text{ s}^{-1}$
- $k_{ST} = 2.9 \times 10^9 \text{ s}^{-1}$
- $k_{TS} = 7.4 \times 10^6 \text{ s}^{-1}$
- $k_p^\circ = 6.2 \times 10^8 \text{ s}^{-1}$

(This will be a discussion question)

8.5 Short conceptual question - Deactivation of excited state aromatic hydrocarbons

Table 8.1: The quantum yields of various deactivation processes in small organic molecules measured at 77 K in a glass matrix.

	Φ_f	Φ_{ST}	$\Delta E / \text{kJ mol}^{-1}$
Naphthalene	0.20	0.80	385
Anthracene	0.70	0.30	318
Pyrene	0.6	low	322
Tetracene	0.1	0.65	251
Pentacene	0.10	0.15	209

When examining the data above suggest why it is likely why the quantum yields of both fluorescence and singlet to triplet intersystem crossing decrease with increasing molecule size.

(This will be a discussion question)

8.6 Short conceptual question - Affect of deuteration of solvents.

Singlet oxygen has a phosphorescence wavelength of around 1070 nm and a lifetime of 2 μ s in water, how would you expect this lifetime to change for singlet oxygen in D₂O?

(This will be a discussion question)

8.7 Short conceptual question - Isotope effects on deactivation of an excited state

The fluorescence quantum yield and singlet state lifetime of both protonated and deuterated pyrene are 0.90 and 450 ns respectively. Why does deuteration of the sample have no measureable affect on these values?

Conversely for naphthalene phosphorescence (in glass at 77 K) the quantum yield of phosphorescence increases from 0.05 to ~0.80 on deuteration of the sample. ($\Delta E = 251 \text{ kJ mol}^{-1}$). Explain this observation with respect to the energy gap law.

(This will be a discussion question)

8.8 Short conceptual question - The effect of heavy atoms on the rate of intersystem crossing

Table 8.2: The affect of substitution of different halogens on the rates of phosphorescence and singlet to triplet intersystem crossing.

	k _p	k _{ST}	Φ_p / Φ_f
Naphthalene	0.05	0.39	0.09
1-fluoronaphthalene	0.23	0.42	0.07
1-chloronaphthalene	1.1	2.35	5.2
1-bromonaphthalene	13.5	36.5	169
1-iodonaphthalene	190	310	>760

Briefly explain why the rates of these processes increase as we move down the group.

(This will be a discussion question)

8.9. SHORT CONCEPTUAL QUESTION - THE EFFECT OF ABSORBANCE AND EMISSION WAVELENGTHS

8.9 Short conceptual question - The effect of absorbance and emission wavelengths on the quantum yield of emission

Table 8.3: The spectroscopic details of a family of osmium complexes.

	abs / nm	em / nm	ΔE / eV	/ ns	Φ_{em}
$[\text{Os}(\text{phen})_3]^{2+}$	650	720	0.186	260	0.016
$[\text{Os}(\text{phen})_2(\text{dppene})]^{2+}$	455	609	0.69	1830	0.138
$[\text{Os}(\text{phen})(\text{dppene})_2]^{2+}$	400	530	0.761	3600	0.518

Why does the fluorescence lifetime increase as the phenanthroline ligands are replaced with dppene ligands?

(This will be a discussion question)

Chapter 9

Other Excited State Systems

9.0.1 Learning Objectives

At the end of this section you should be able to:

- Explain the origin of excimer and exciplex emission spectra.
- Describe why excimer and exciplex formation are concentration dependent.
- Interpret changes in emission spectra with changing concentration to justify formation of either excimers or exciplexes.
- Show an awareness of excited state reactions and rearrangements.

9.1 Excimers

An excimer is an excited state dimer formed between two identical molecules. A photon is delocalised between two molecules which in the excited state are weakly bound together. Excimers are relatively long lived species, and were first noticed when increasing the concentration of certain fluorophores.

It was found that upon increasing the concentration of the chromophore the fluorescence did not increase linearly, and upon reaching higher concentrations the fluorescence decreased; however, this decrease in fluorescence occurred with an increase in a new emission band of lower energy. Excimer bands tend to be broad featureless and Gaussian in profile, figure 33. It is important to note that there is no bonding between the chromophores in the ground state; excimers only occur at high concentrations as an excited state molecule has to interact with a second ground state molecule in the lifetime of the excited state. There is no change in the absorption of the sample (as there is only excitation of monomer species), only formation of a new band in the emission spectrum.

Excimers exist because of formation of a weak bond between the two chromophores as shown in figure 9.1. By formation of this bond any emission from the excimer has to be lower in energy than the fluorescence of the monomer chromophore. The ground state consists of two molecules in an dissociated state, upon emission from the excimer the structure of the excimer remains (as with the Stokes shift) and the two molecules diffuse apart.

It is important to note that there is no charge separation in the excimer and that the excited state is shared across both molecules.

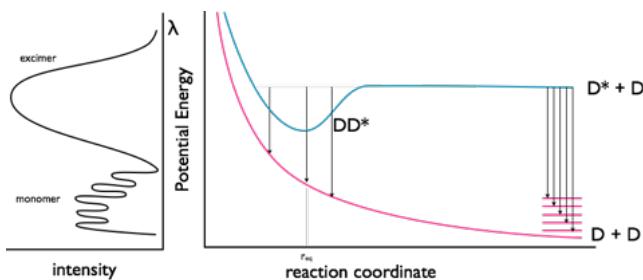


Figure 9.1: The monomer and excimer emission of pyrene, and an energy level profile showing the fluorescence from the non-bonded D^* state as well as the excimer emission from the DD^* state.

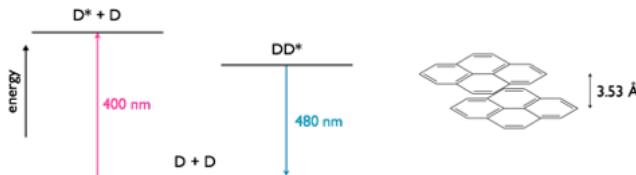


Figure 9.2: A pyrene excimer with a separation of 3.53 Å between the two planes of nuclei, in the excited state dimer the two ring systems of pyrene have a weak interaction between them. The emission maxima of fluorescence is 400 nm whereas the emission maxima for the excimer is 480 nm.

A concept bite video briefly covering the concepts of eximers and exciplexes
(video length 7m05s)

9.2 Exciplexes

Excimers consist of an excited state shared between two identical monomer units, whereas exciplexes are excited state complexes of two different chromophores. Due to the differences in molecular orbitals of the chromophores exciplexes have a dipole across the excited state complex; this dipole can lead to formation of a charge separated species, figure 9.3.

Exciplexes can undergo emission from this state, leading to formation of dissociated monomer units or there can be a solvent reorganisation leading to formation of a solvent separated radical ion pair, at this point there can still be charge recombination and relaxation of the system (this time with the excess energy lost as heat) to the original monomer chromophore units.

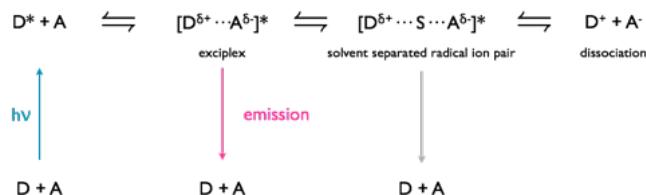


Figure 9.3: Formation of an exciplex from an excited state donor and ground state acceptor chromophore. The exciplex has a small charge separation leading to a dipole over the excited state complex, under some circumstances with the introduction of solvent between the chromophores in the exciplex this can then go on to form a solvent separated radical ion pair which may then go on to form a charge separated redox pair.

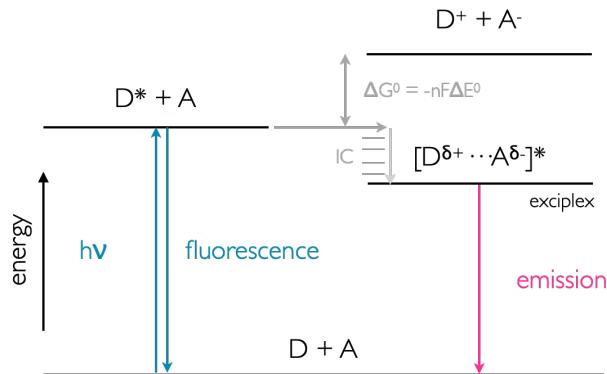


Figure 9.4: Energy profile of exciplex formation showing the lower energy state of the exciplex over the $D^* + A$ system. Emission from the exciplex is lower energy than fluorescence from the monomer donor chromophore.

Exciplexes are common in organic solvent as non polar solvents are not good at stabilising ions in solution, as the polarity of the solvent increases then the charge separated state becomes more favoured. Exciplexes and other charge separated species are of interest in chemistry for use in molecular wires and energy storage devices.

Please see video above in Section 9.1 for a review of this topic (timecode 4m52s).

9.3 Photoinduced Isomerisation

When discussing the absorption of light it was noted that by promotion of an electron into an anti-bonding orbital the bond order was reduced. In the case of double () bonds in molecules such as retinal (figure), the bond order of one of the double bonds is reduced and therefore allows rotation around this now single bond, as the energy is lost (either by internal conversion or emission) the -bond is restored. A simple Lewis model of bonding is convenient, however we have to consider the molecular orbital, and this version of bonding highlights a particular bond which is weakened in the excited state, consequently rotation occurs around a specific bond in a conjugated chain.

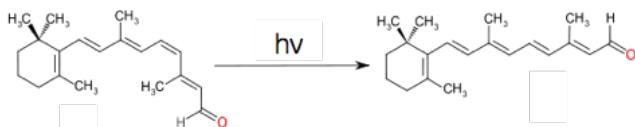


Figure 9.5: The cis-trans isomerisation of retinal, which when combined with the protein rhodopsin is vital in vision. Absorption of a photon, reduces the bond order of a specific bond due to characteristics of the excited state molecular orbital.

The steady state photochemical yield of any such cis-trans isomerisation will depend upon the absorption of each of the chromophore isomers at the excitation wavelength, the composition of this steady state is referred to as the photostationary state. The composition of this photostationary state may be easily calculated making only minor assumptions.

The most important of these assumptions is that when in the excited state there is an equal probability that the molecule will relax to the cis & trans states. The second assumption is that (near) monochromatic light is used, because we have already seen that the molar extinction coefficient is very wavelength dependent. Finally the excitation has to occur long enough for the (steady state) photostationary state to be established.

Figure 9.6 shows the absorption spectra of the two isomers of stilbene, there are a number of interesting photophysical details about this molecule, but it should be noted that the molar extinction coefficient is different for each of the isomers. The point where the two spectra cross is called the isobestic point , for stilbene this is 288 nm, if the system is excited at this wavelength then the yield of each isomer will be the same. However at all other wavelengths the yield of one isomer will be higher than the other. If an excitation wavelength is used whereby the molar extinction coefficient of the trans isomer is highest then this isomer is excited preferentially and there is less of this isomer in the photostationary state.

In summary the higher the molar absorption coefficient at the excitation

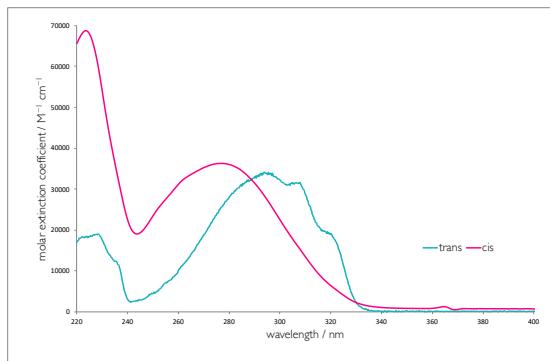


Figure 9.6: The absorption spectra of trans & cis stilbene in hexane.

wavelength the lower the yield of that product in the photostationary state.

The dynamics of cis-trans isomerisation may be studied by using some of the photophysics already discussed. Figure 9.7 shows a cis / trans system, in the cis position the emission from a fluorophore is quenched so there is no visible emission from the dye. In the trans position position of the quencher now means that there is either no quenching or the quenching is greatly reduced, consequently the emission from the fluorophore is greatly increased.

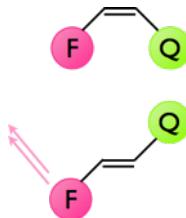


Figure 9.7: The quenching of emission of a fluorophore by a nearby quencher, in the cis configuration the emission from the fluorophore is completely quenched, whereas in the trans configuration emission from the fluorophore is observed.

9.4 Before Completing this Section

To support the material in this section it is suggested you read pages 90-96 of Wardle ‘Principles and Application of Photochemistry’.

Chapter 10

Workshop Questions for Week 5

10.1 Short mathematical question - Determining rates of diffusion controlled quenching

The rate of diffusion (k_d) in solution is given by equation (10.1), where η is the viscosity of the solution.

$$k_d = \frac{8RT}{3\eta} \quad (10.1)$$

Determine the maximum possible rate of diffusion controlled quenching at 20 °C in:

- a. water ($\eta = 1.0016 \text{ mPa s}$)
- b. methanol ($\eta = 0.594 \text{ mPa s}$)

(I will use UniDoodle to ask this in class)

10.2 Short mathematical question - Determining the effect of diffusion controlled quenching on emission intensity

What emission intensity would you expect if 50 mM of a quencher quenches the emission of a chromophore dissolved in basic ethanol at 10 °C, with

natural lifetime of 13.2 ns and emission quantum yield, Φ_f , of 0.32, if the unquenched intensity is 35240.

$$\eta_{\text{EtOH}, 10 \text{ } ^\circ\text{C}} = 1.394 \text{ mPa s}$$

(This will be a discussion question)

10.3 Short conceptual question - effect of temperature

YO-Pro-1 is quenched in the presence of molecular oxygen, as the temperature increases this quenching increases, what does this indicate about the mechanism of quenching and why?

(This will be a discussion question)

10.4 Short conceptual question - change in behaviour on freezing

As the concentration of a species increases the wavelength of emission increases in the solution phase but this same shift in wavelength is not when the solution is frozen. What photochemical process may be occurring to explain this effect?

(This will be a discussion question)

10.5 Short conceptual question - static quenching

When in the presence of a quencher the intensity of emission of a chromophore ($\tau_0 = 5.2 \text{ ns}$) was reduced from 4200 cps to 2100 cps at $25 \text{ } ^\circ\text{C}$ and 1800 cps at $10 \text{ } ^\circ\text{C}$. What would the lifetime of the quenched chromophore be at $25 \text{ } ^\circ\text{C}$?

(This will be a discussion question)

10.6 Short conceptual question - isoemissive point

As the concentration of pyrene dissolved in toluene increases a new broad band with a broad featureless spectrum is observed, which is from excimer emission. What evidence is there that emission is only from single molecule and excimer emission and no other states?

How would the absorption spectrum change as the concentration increases?

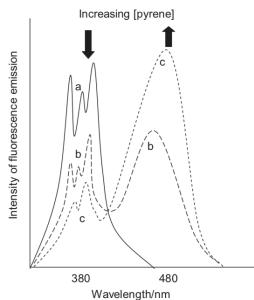


Figure 10.1: The emission spectrum of pyrene in toluene as low (solid line) and higher (dotted lines) concentrations.

10.7 Short conceptual question - static quenching

10-methylacridinium chloride (MAC) is quenched in the presence of adenosine monophosphate (AMP, figure (10.3)). The effect of quenching is enhanced by addition of sodium sulfate.

In a separate experiment the emission of MAC is unaffected by addition of sodium sulfate (when not in the presence of AMP).

Suggest the mechanism of quenching, justifying this with reference to the experimental data.

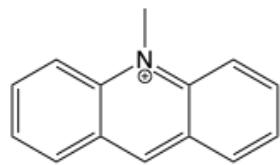


Figure 10.2: The structure of methyl acridinium.

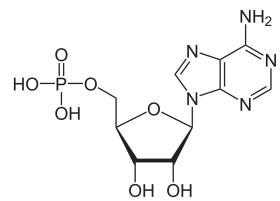


Figure 10.3: The structure of adenosine monophosphate.

(This will be a discussion question)

10.8 Long mathematical question - Quenching of emission of acridone

Acridone (figure 10.4) is found to be quenched in the presence of potassium iodide in aqueous solution at 26 oC. Solutions were maintained at constant ionic strength by use of KNO₂, the KNO₂ does not affect the emission intensity of the solution.

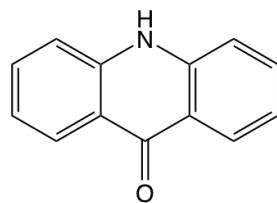


Figure 10.4: The structure of acridone.

The following data were collected for the emission of acridone.

Table 10.1: The effect of potassium iodide concentration on emission intensity and fluorescence lifetime of acridone in aqueous solution.

[KI] / M	[KNO ₂] / M	Emission intensity / arb	/ ns
0	1.100	16580	17.60
0.040	1.060	3753	3.90
0.100	1.000	1566	1.80
0.200	0.900	721	0.95
0.300	0.800	446	0.64
0.500	0.600	242	0.39
0.800	0.300	121	0.25

1. Using an appropriate plot (or plots) determine if the quenching is static, dynamic or a combination of both mechanisms.
2. Determine any relevant quenching constants (k_d and/or K_S)

(This will be a discussion question)

Chapter 11

Workshop Questions for Week 6

11.1 Short conceptual question - O₂ quenching

1. Given molecular oxygen has an emission with λ_{max} around 1280 nm why is it such a good quencher of excited states?
2. Why is the efficiency of quenching lower for excited singlet states?

Are you all happy with the energy level diagram of molecular oxygen?

(This will be a discussion question)

11.2 Short mathematical question - Förster resonance energy transfer

A donor, D, with an unquenched lifetime of 5.0 ns, was found to have a steady state emission intensity of 20.5 in free solution and 4.1 when in the presence of a quencher, Q. Assuming the Förster distance is 50 Å determine :

1. the transfer efficiency, E.
2. the lifetime of D in the presence of the quencher Q
3. the equilibrium separation of D & Q
4. the rate constant for energy transfer

(This will be a discussion question)

11.3 Short conceptual question - Förster distance

Determine the quantum yield of emission of the donor in a FRET pair system where donor and acceptor are separated by the Förster distance, the lifetime of the unquenched system is 8.4 ns. The unquenched quantum yield is 0.58*

(I will poll this question on UniDoodle and then maybe discuss some of the responses)

11.4 Short conceptual question - redox chemistry

Why does formation of an excited state decrease both the oxidation and reduction potential of a species?

11.5 Short conceptual question - Donor-acceptor system

In a single molecule study of a donor bound to an acceptor via a flexible, short aliphatic chain it was found that lifetime of the excited state varied (time between single photon detection of emission after an excitation pulse). Suggest why this is the case.

(This will be a discussion question)

11.6 Short conceptual question - Quenching of ruthenium

Ruthenium tris bipyridine $[\text{Ru}(\text{bpy})_3]^{2+}$ has a \sim_{max} , $\text{abs} \sim = 470 \text{ nm}$ and \sim_{max} , $\text{em} \sim = 465 \text{ nm}$. The natural lifetime is $13.6 \mu\text{s}$.

Suggest why molecular oxygen is an efficient quencher of the emmission of $[\text{Ru}(\text{bpy})_3]^{2+}$.

(This will be a discussion question)

11.7 Short conceptual question - Energy transfer and spectra

How would you expect the absorption and emission spectra observed in an experiment to change when an emissive acceptor is added?

11.8. SHORT CONCEPTUAL QUESTION - EFFECT OF SEPARATION 95

If you keep the emission wavelength the same and scan the intensity of this band with changing excitation wavelength how will this (excitation) spectrum vary from the absorbance and how can it be used to confirm energy transfer?

(This will be a discussion question)

11.8 Short conceptual question - Effect of separation

A ‘cyanine-3’ and ‘cyanine-5’ are cyanine dyes which are frequently used as markers in experiments with DNA.

Cyanine-3 has a \sim max, $\text{abs} \sim = 554 \text{ nm}$ and \sim max, $\text{em} \sim = 568 \text{ nm}$

Cyanine-5 has a \sim max, $\text{abs} \sim = 649 \text{ nm}$ and \sim max, $\text{em} \sim = 666 \text{ nm}$

In an experiment they are tethered to either end of a DNA oligomer of 4, 5, 6, 7, 8, 9, 10 11 & 12 base pairs. However the pattern expected of increasing lifetime as the separation increased was not observed. Suggest why this may be occurring.

(This will be a discussion question)

11.9 Short conceptual question - lifetimes

The emission spectra of 2-phenylindole shows a marked difference with changes in concentration.

At a concentration of $1 \times 10^{-5} \text{ M}$ the measured lifetime is 0.86 ns, whereas at $5 \times 10^{-3} \text{ M}$ the measured lifetime is 3.42 ns.

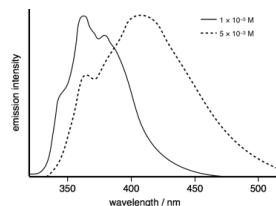


Figure 11.1: The emission spectrum of pyrene in toluene as low (solid line) and higher (dotted lines) concentrations.

Why does the measured lifetime depend upon concentration?

(This will be a discussion question)

11.10 Short conceptual question - effect of time delay

If a solution of pyrene in cyclohexane is excited with a very short (ps) pulse of light after 1 ns the emission spectrum observed is mainly that of the monomer, whereas after 100 ns emission is principally from the excimer. Why is this the case?

(This will be a discussion question)

11.10.1 Short conceptual question - effect of polar solvents on emission

Solutions containing anthracene and diethylaniline are shown to have broad emission at around 450 nm in toluene, but in dichloromethane no emission is observed.

The emission from anthracene has λ_{max} of 375 nm.

Suggest the processes going on which account for these observations.

You can find the answer to this question in 6.2.7