



## 8.1

$$\frac{d[E]}{dt} = V_E = (k_2 + k_3)[ES] - k_1([E] - [ES])[S]$$

$$\frac{d[S]}{dt} = V_S = k_2[ES] - k_1([E] - [ES])[S]$$

$$\frac{d[ES]}{dt} = V_{ES} = k_1([E] - [ES])[S] - (k_2 + k_3)[ES]$$

$$\frac{d[P]}{dt} = V_P = k_3[ES]$$

## 8.2

```
clc;
%% 解微分方程
x0 = [1, 10, 0, 0];
tspan = [0, 1.5];
[T, X] = ode45(@f, tspan, x0);

%% 绘图
figure;
plot(T, X(:, 1))
ylabel('E');
xlabel('Time');

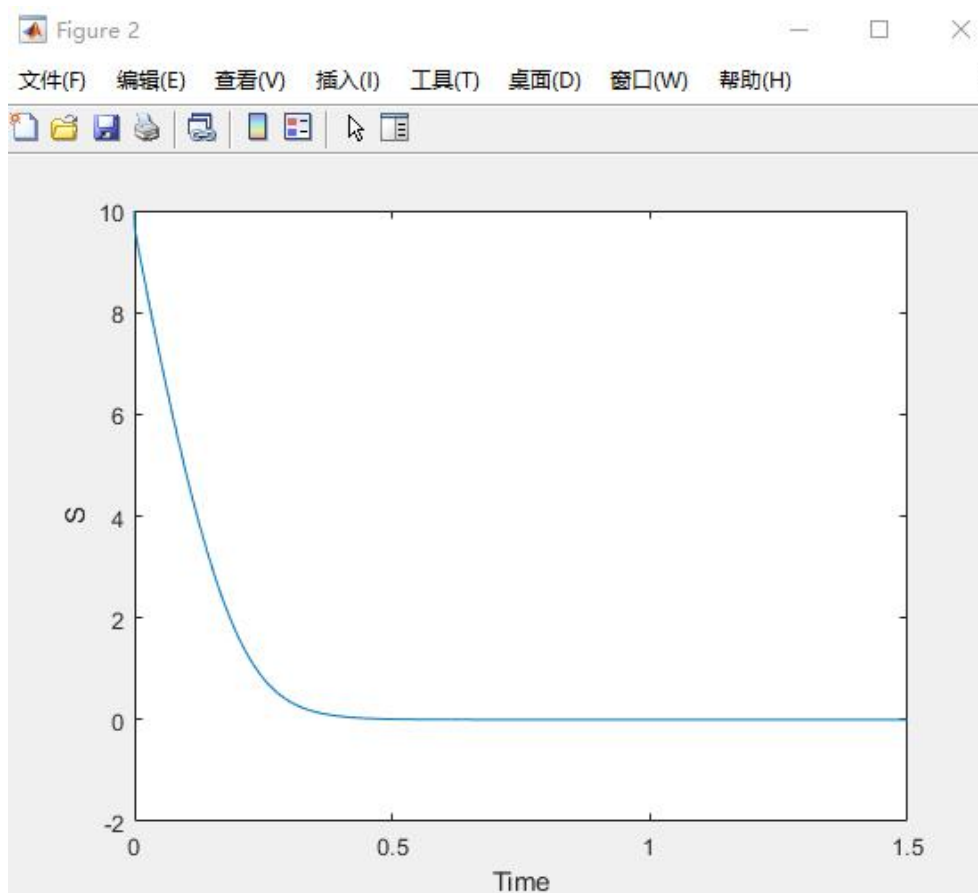
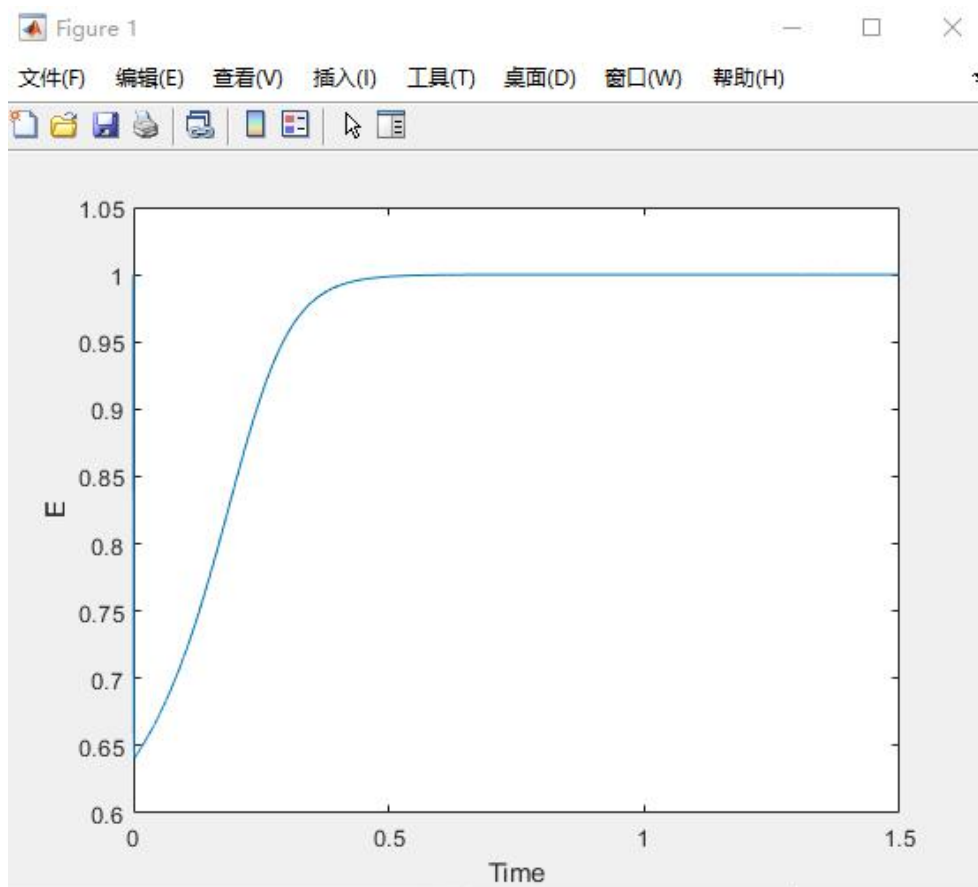
figure;
plot(T, X(:, 2))
ylabel('S');
xlabel('Time');

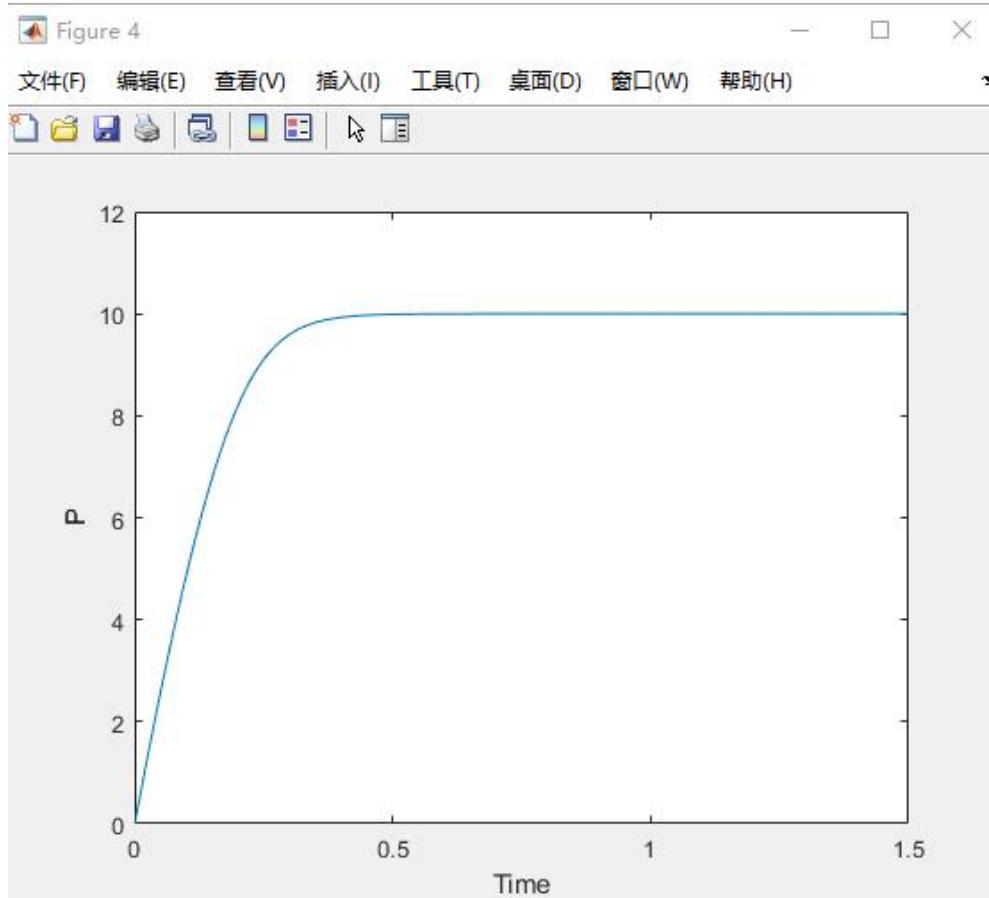
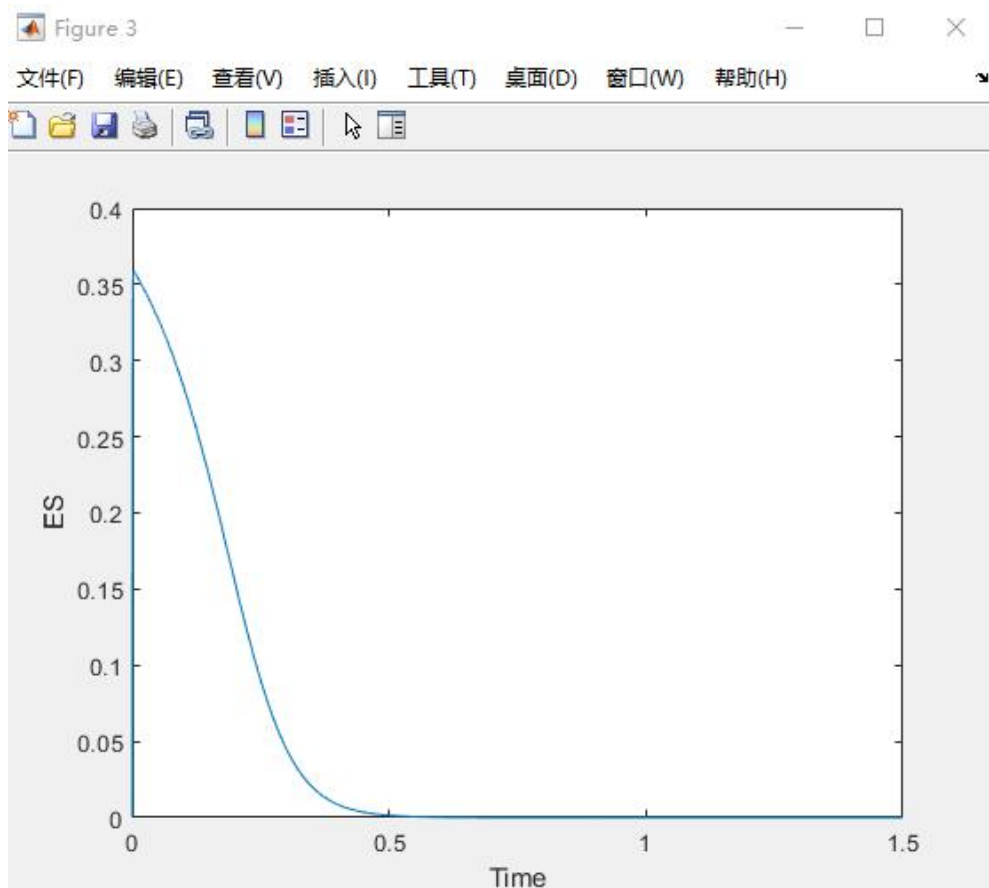
figure;
plot(T, X(:, 3))
ylabel('ES');
xlabel('Time');

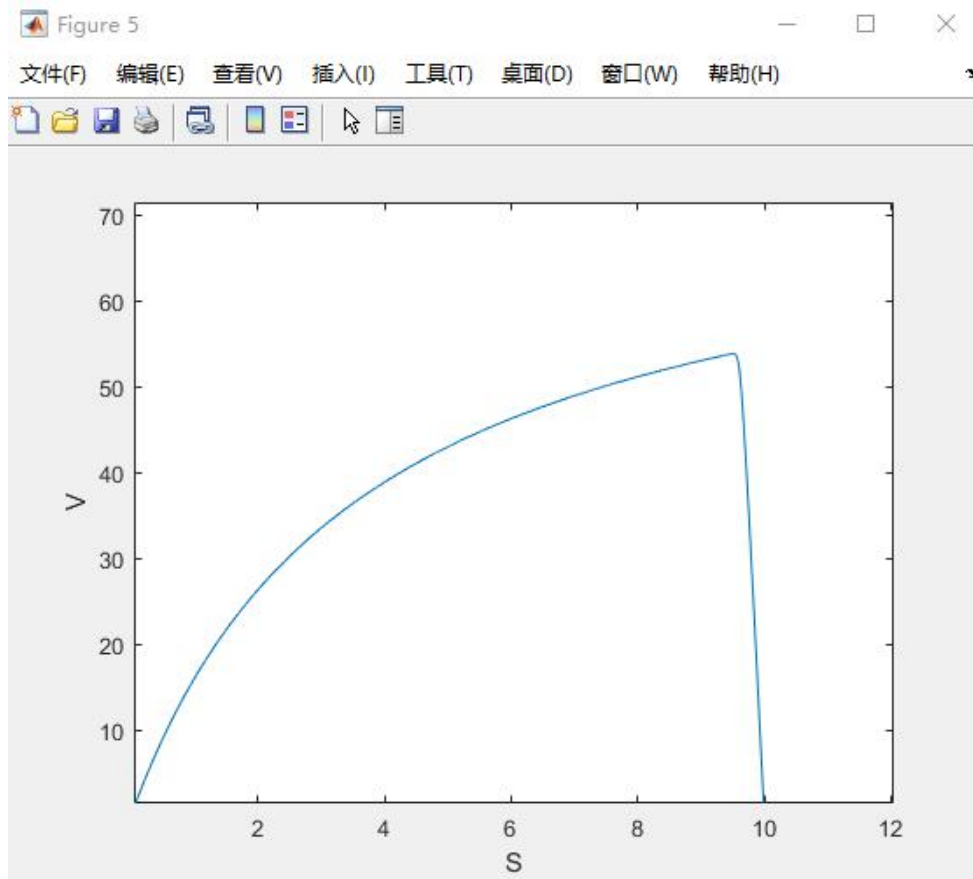
figure;
plot(T, X(:, 4))
ylabel('P');
xlabel('Time');

V=150*X(:, 3);

figure;
plot(X(:, 2), V)
xlabel('S')
ylabel('V')
%% 微分方程函数
function fx = f(t, x)
    % 初始化fx, 需要为列向量
    fx = zeros(4, 1);
    % 四个微分方程组
    fx(1) = 750*x(3) - 100*(x(1)-x(3))*x(2);
    fx(2) = 600*x(3)-100*(x(1)-x(3))*x(2);
    fx(3) = 100*(x(1)-x(3))*x(2)-750*x(3);
    fx(4) = 150*x(3);
end
```







## 8.3

$$V = V_p$$

When the reaction is in dynamic equilibrium, the generation rate of ES is equal to the decomposition rate:

$$k_1([E] - [ES])[S] = (k_2 + k_3)[ES]$$

$$[ES] = \frac{[E][S]}{\frac{k_2 + k_3}{k_1} + [S]}$$

$$\text{Let } k_m = \frac{k_2 + k_3}{k_1}$$

When the S concentration is high, all E is saturated by S and converted into ES ( $[E] = [ES]$ ), at this time, the V reaches  $V_m$ , then  $V_m = k_3[ES] = k_3[E]$ , so:

$$V = k_3[ES] = \frac{k_3[E][S]}{k_m + [S]} = \frac{V_m[S]}{k_m + [S]}$$

Change the above formula into the following form:

$V = V_m - k_m * \frac{V}{[S]}$ ,  $V$  is the vertical axis,  $V/[S]$  is the abscissa axis, so the  $V_m$ :

