EXP. NUMBER	EXPERIMENT/SUBJECT	DATE
5	Organic Chemistry - 2	[redacted]

NAME	LAB PARTNER	LOCKER/DESK NO.	COURSE & SECTION NO.
[redacted]			Chem 1AA3

Prelab Questions

1. Melting Point of Aspirin: 135°C
Source: Haynes, W., Lide, D. R. Handbook of Chemistry and Physics, 90th ed. Florida: CRC Press, 2009 (3-8)

Purpose:
The purpose of this experiment is to perform several experiments utilizing the aspirin product obtained from experiment # 4.

Procedure
The experiment was carried out as described in the 2011/2012 Chem 1A03/1AA3 Lab Manual, Organic Chemistry - 2.

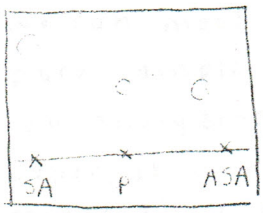
OBSERVATIONS
5. Percent yield calculation
% yield = $\frac{\text{Actual}}{\text{Theoretical}} \times 100$

$= \frac{0.8797\ g}{2.609\ g} \times 100 = 33.7\%$

6. see discussion for explanation.

Melting point: 124 - 126°C.
6. see discussion for explanation.

2. Thin Layer chromatograph (TLC).
TA Demonstration



SA = ~~salycl~~ salicylic acid (starting material)
P = product
ASA = Aspirin.

6. see Discussion for explanation

Reactions of Aspirin
4. b) observations (qualitative)

Test tube 1: Aspirin and NaHCO3
Observation: Bubbles formed on top of solution and almost all aspirin crystals dissolved. Leaves a clear, water-like solution.

Test tube 2: Salicylic Acid and NaHCO3
Observation: Thick layer of bubbles formed immediately followed by the formation of a white precipitate which remained in solution.

Test Tube 3: Benzyl Alcohol and NaHCO3.
Observation: No chemical reaction occured. Remained separate from test tube one solution floated on top of the other.

6. see discussion for analysis of individual reactions.

SIGNATURE	DATE	WITNESS/TA	DATE

NOTE: INSERT DIVIDER UNDER COPY SHEET BEFORE WRITING