DATE

barryations

PART A

PART	A				
TIME	OBSERVATION				
INITIAL)	Hel and acetone are clear and colouriess. Brz 15 clear and crange.				
(1-0) 1 HIN	immediately turns light yellow (still clear).				
(1-Z) 2 MIN	shil bright yellow and colcurless, slightly paler.				
(2-3) 3MIN	Still bright yellow and colouriess. Continued to become paler-still clear.				
(3-4) 4 MIN	increased transporency of solution very slightly. still golden yellow, butpaler				
(4-5) 5 MIN	final colour is very light, translucent, golden gellow				

6 PART B Absorbance: 395 nm 6 Units (A)

				_
Time	Trial 1	Trigi L	Trigi3	Trialy
455	6.542	1.275	1.198	1.160
60 5	0.538	1.259 2		1.145
90 5	0.525	1.2402	1	1.116
1205	0.507	1 225)	1.118	1.084
150.5	0.489	1.209	1.082	1.053
1805	0.470	4	1.040	1.021

DATE

(alculations and Discussion

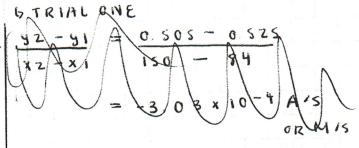
1 initial concentrations of solutions

Trigit	[Brz] H	[Ace tone]H	EHCI] M	Rate m/s
	4×10-3	0 Y	6.2	FEAGLE AVELT
2	8 × 16-3	0.8	0.2	17841
3	8 x 10 -3	0.8	0.4	-1.11×10-3
4	8 × 10 -3	1.6	6.2	-1.11×10-3
		Rate 1	-5.24	3 x 10 -4

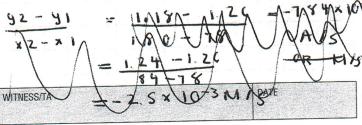
2. Refer to graphs on following page.

3. The rate of reaction can be obtained by a graph of change in absorbance over time. This is because concentration change over time is the definition of the rate of a reaction, and in this case, absorbance is proportional to concentration.

4. STOPE CATCULATIONS
GFOR SAMPLE SEE TRIAL 3+46



STRIAL TWO



SIGNATURE