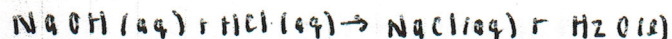


NUMBER	EXPERIMENT/SUBJECT	DATE
	Qualitative Volumetric Analysis	
NAME	LAB PARTNER	LOCKER/DESK NO.
		COURSE & SECTION NO.
		L69 Chem 1A03

Discussion:



During this experiment, an NaOH solution was titrated against an HCl solution. NaOH was released into the HCl solution until it had completely reacted with the HCl, or basically when the equivalence point of the reaction was reached. This point in the reaction was visible as a faint pink coloured solution, as a result of the phenolphthalein indicator which was placed in the HCl solution, prior to the start of the titration. At this point in the reaction, the moles of HCl and NaOH are equal to each other. Since the balanced chemical reaction, the volumes of NaOH, and the standardized concentration of HCl are known, the concentration of NaOH was calculated. To obtain a more scientifically sound results, the experiment was conducted three times, and the mean of NaOH concentration was calculated and recorded.

it was difficult not to exceed the exact equivalence point of the titration, and there was a slight excess of NaOH in the solution.

it is also important to take into consideration the sources of error that would affect results. Specifically, it was difficult to obtain a precise volume of exactly 10.00 mL of HCl solution. Furthermore, due to the nature of phenolphthalein which turns pink when the solution is slightly basic,

SIGNATURE	DATE	WITNESS/TA	DATE