Trestily the plane problem as a closed system.

plane finel volume = 100 m<sup>3</sup>

initial state -50°C 0.1 bar

finel state 25°C 0.8 bar

Cp\* = 30 J knol k)

\$0.2 per KWhr initial

state

75

To oz ou 0.6 0.8 1.0

Path A: isothermal compression

isobaric heating.

Closed system energy belonce no kiletic, potential energy
no sheft work
no streams, flows
Theory Balance- DU=Q-SpdV

-> iso ther mel compression

$$W'_{i} = -\int_{\rho}^{V_{2}} \frac{V_{2}}{\rho dV} = -RT \int_{\rho}^{V_{2}} \frac{V_{2}}{V} = -RT \ln \frac{V_{2}}{V_{1}}$$

$$V_{1} = \int_{\rho}^{T_{2}} \frac{V_{2}}{V} = -RT \ln \frac{V_{2}}{V_{1}}$$

$$V_{1} = \int_{\rho}^{T_{2}} -RT \ln \left(\frac{\rho_{2}}{\rho_{1}}\right)$$

$$W'_{1} = \int_{\rho}^{T_{2}} -RT \ln \left(\frac{\rho_{2}}{\rho_{1}}\right)$$

$$V'_{2} = \int_{\rho}^{T_{2}} -RT \ln \frac{V_{2}}{V_{1}}$$

$$V'_{3} = \int_{\rho}^{T_{2}} -RT \ln \left(\frac{\rho_{2}}{\rho_{1}}\right)$$

$$V'_{4} = \int_{\rho}^{T_{2}} -RT \ln \left(\frac{\rho_{2}}{\rho_{1}}\right)$$

$$V'_{5} = \int_{\rho}^{T_{2}} -RT \ln \left(\frac{\rho_{2}}{\rho_{1}}\right)$$

$$V'_{5$$

is obariz heating

$$W_{1}^{1} = - \begin{cases} V_{3} & w_{1}^{2} \\ P_{2} dV = - P_{2} \begin{cases} V_{3} - V_{2} \\ V_{2} \end{cases} = - R(T_{3} - T_{2})$$

$$V_{3} & v_{2} \\ V_{3} & v_{3} \end{cases}$$

and 
$$V_{\underline{z}}$$

$$D\underline{V} = \int_{T_{\underline{z}}}^{T_{\underline{z}}} Cv^{\underline{x}} dT = Cv^{\underline{x}} (T_{\underline{z}} - T_{\underline{z}})$$

$$\int_{T_{\underline{z}}}^{T_{\underline{z}}} Cv^{\underline{x}} dT = \int_{T_{\underline{z}}}^{T_{\underline{z}}} Cv^{\underline{x}} dT$$

$$Q_{ii} = 30 \text{ J/mol.k} \cdot (75 \text{ K}) = 2250 \text{ J/mol}$$

$$Wii = -8.314 \text{ J/mol.k} (75 \text{ K}) = -623.55 \text{ J/mol}$$

$$DM = Q + W \qquad \text{total}$$

$$Q = Q_{i} + Q_{ii} = -3,857.92 + 2250 = -1,607$$

$$W = W_{i} + W_{ii} = 3,857.92 - 623.55 = 3,239.37$$

$$DM = -1,607 + 3234.37$$

$$= |1,627.4 \text{ J/mol}|$$

$$PV = nRT$$

$$\frac{0.8 \text{ ber} \cdot 100 \text{ m}^{3}}{9.319 \cdot 10^{-5} \cdot 248.15} = 3,227.34 \text{ mol} \text{ ain}$$

$$9.319 \cdot 10^{-5} \cdot 248.15 = 5,252,178 \text{ Joules total}$$

$$|KW - k = 3,600,000 \text{ J} = 315,130,180 \text{ Joules/hour.}$$

$$|KW - k = 3,600,000 \text{ J} = 315,130,180 \text{ Joules/hour.}$$

-\$43.76 / Lour

or \$1.37 per minute.