

# Conversion in Non-Ideal Flow Reactors

$$\left( \begin{array}{c} \text{mean concentration} \\ \text{of reactant} \\ \text{in exit stream} \end{array} \right) = \sum_{\substack{\text{all elements} \\ \text{of exit stream}}} \left( \begin{array}{c} \text{concentration of} \\ \text{reactant remaining} \\ \text{in an element of} \\ \text{age between } t \\ \text{and } t + dt \end{array} \right) \left( \begin{array}{c} \text{fraction of exit} \\ \text{stream which is} \\ \text{of age between } t \\ \text{and } t + dt \end{array} \right)$$

$$\left( \frac{\bar{C}_A}{C_{A0}} \right)_{\text{at exit}} = \int_0^{\infty} \left( \frac{C_A}{C_{A0}} \right)_{\substack{\text{for an element or little} \\ \text{batch of fluid of age } t}} \cdot \mathbf{E} \, dt$$

# Conversion in Non-Ideal Flow Reactors

- Conversion at the reactor exit is the sum of the concentration of reactant remaining in an element multiplied by the fraction of exit between  $t$  and  $t+dt$ .
- This is represented mathematically by:

$$\left(\frac{\bar{C}_A}{C_{A0}}\right)_{\text{at exit}} = \int_0^{\infty} \left(\frac{C_A}{C_{A0}}\right)_{\text{for an element or little batch of fluid of age } t} \cdot \mathbf{E} \, dt \quad (1)$$

$$\bar{X}_A = \int_0^{\infty} (X_A)_{\text{element}} \cdot \mathbf{E} \, dt \quad (2)$$

or in a form suitable for numerical integration

$$\frac{\bar{C}_A}{C_{A0}} = \sum_{\text{all age intervals}} \left(\frac{C_A}{C_{A0}}\right)_{\text{element}} \cdot \mathbf{E} \, \Delta t \quad (3)$$

## Conversion in Non-Ideal Flow Reactors (cont.)

- In equation one, the concentration element term is dependent on the order of the reaction.

**1st-Order Reactions:**

$$\left(\frac{C_A}{C_{AO}}\right)_{element} = e^{-kt}$$

**2nd-Order Reactions:**

$$\left(\frac{C_A}{C_{AO}}\right)_{element} = \frac{1}{1 + kC_{AO}t}$$

## Conversion in Non-Ideal Flow Reactors (cont.)

- Finally, the following termed can be applied to equation 1 for any order reaction.

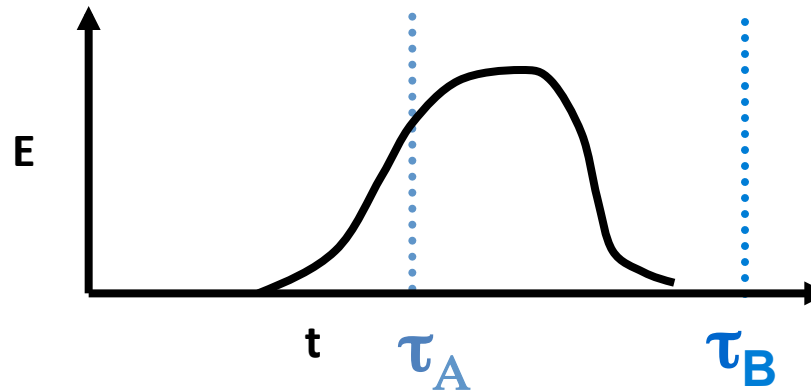
**nth-Order Reactions:**

$$\left(\frac{C_A}{C_{AO}}\right)_{element} = [1 + (n-1)C_{AO}^{n-1}kt]^{\frac{1}{1-n}}$$

# Conversion in Non-Ideal Flow vs. Ideal Flow Reactors

- The space time parameter,  $\tau$ , is equal to the reactor volume divided by the volumetric feed rate.  $\tau$  is an important parameter in calculating conversion in ideal flow reactors.
- Where  $\tau$  falls on the x-axis of an  $E(t)$  curve indicates how much more or less conversion is occurring in a non-ideal reactor as compared to an ideal reactor.

## Conversion in Non-Ideal Flow vs. Ideal Flow Reactors (cont.)



- **Case A:** Particles are spending more time than expected in the reactor. Therefore, a higher conversion is being achieved in a non-ideal situation than would be in an ideal situation.
- **Case B:** Particles are spending less time than expected in the reactor. Therefore, a lower conversion is being achieved in a non-ideal situation than would be in an ideal situation.

## Example 11.4: Determining Conversion in Non-Ideal Reactors

- **Problem Statement:** A liquid is decomposing in a non-ideal reactor in Example 11.1 at a rate which can be expressed by  $-r_A = kC_A$  where  $k = 0.307 \text{ min}^{-1}$ .
- **Determine:** Fraction of reactant unconverted in the real reactor and compare this with the fraction unconverted in a plug flow reactor of the same size.



For the *plug flow reactor* with negligible density change we have

$$\tau = C_{A0} \int_0^{X_A} \frac{dX_A}{-r_A} = -\frac{1}{k} \int_{C_{A0}}^{C_A} \frac{dC_A}{C_A} = \frac{1}{k} \ln \frac{C_{A0}}{C_A}$$

and with  $\tau$  from Example 11.1

$$\frac{C_A}{C_{A0}} = e^{-k\tau} = e^{-(0.307)(15)} = e^{-4.6} = \underline{\underline{0.01}}$$

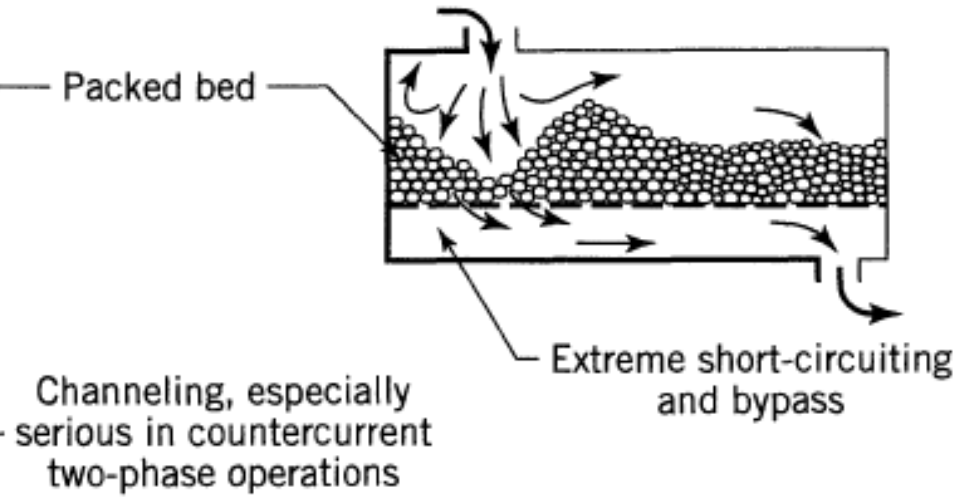
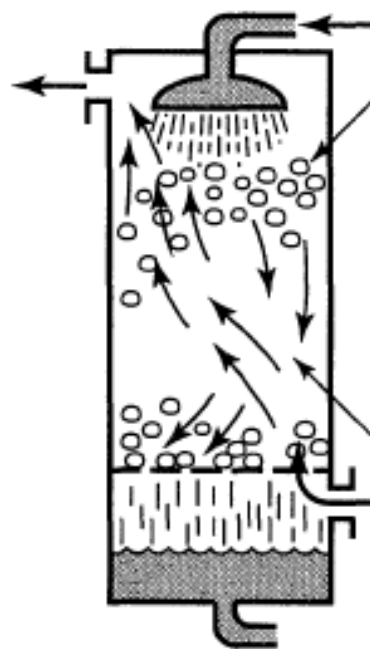
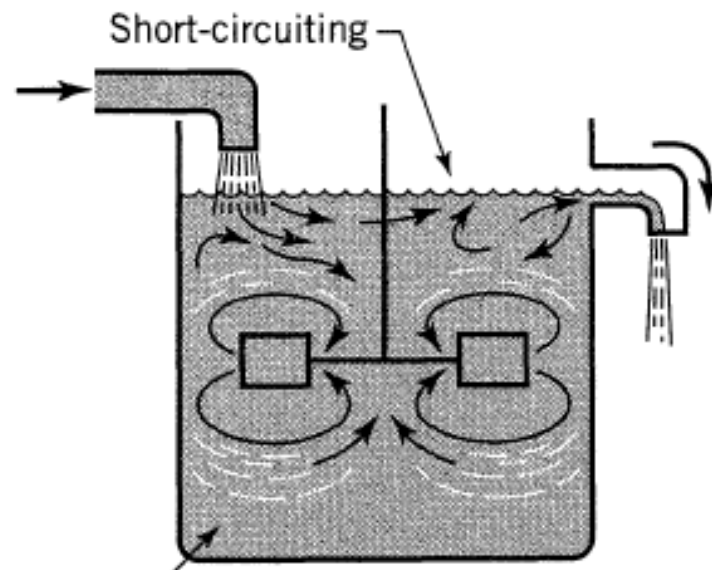
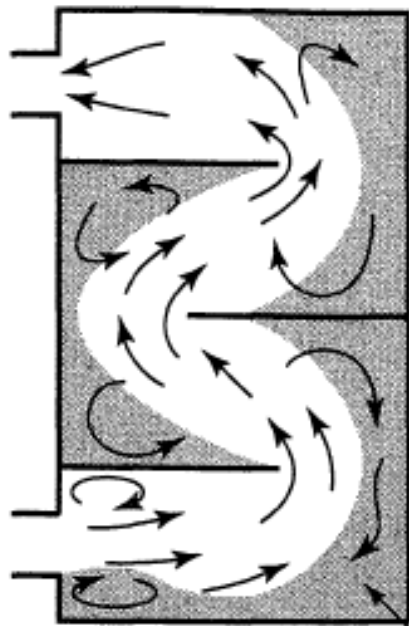
Thus the fraction of reactant unconverted in a plug flow reactor equals 1.0%.

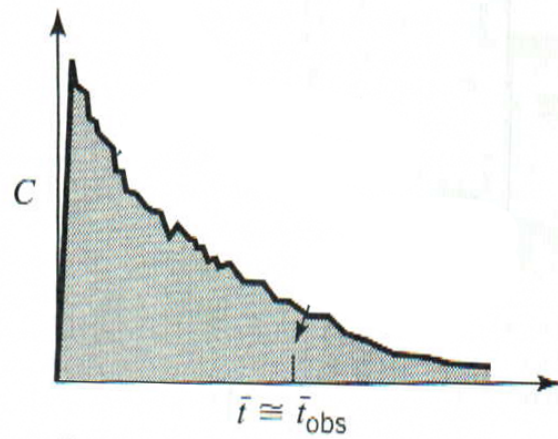
**Table E11.4**

$t$	$\mathbf{E}$	$kt$	$e^{-kt}$	$e^{-kt}\mathbf{E} \Delta t$
5	0.03	1.53	0.2154	$(0.2154)(0.03)(5) = 0.0323$
10	0.05	3.07	0.0464	0.0116
15	0.05	4.60	0.0100	0.0025
20	0.04	6.14	0.0021	0.0004
25	0.02	7.68	0.0005	0.0001
30	0.01	9.21	0.0001	0
<div style="display: flex; align-items: center;"> <div style="text-align: center; margin-right: 10px;"> <math>\underbrace{\hspace{10em}}</math>  given </div> <div> <math display="block">\frac{C_A}{C_{A0}} = \sum e^{-kt}\mathbf{E} \Delta t = \underline{\underline{0.0469}}</math> </div> </div>				

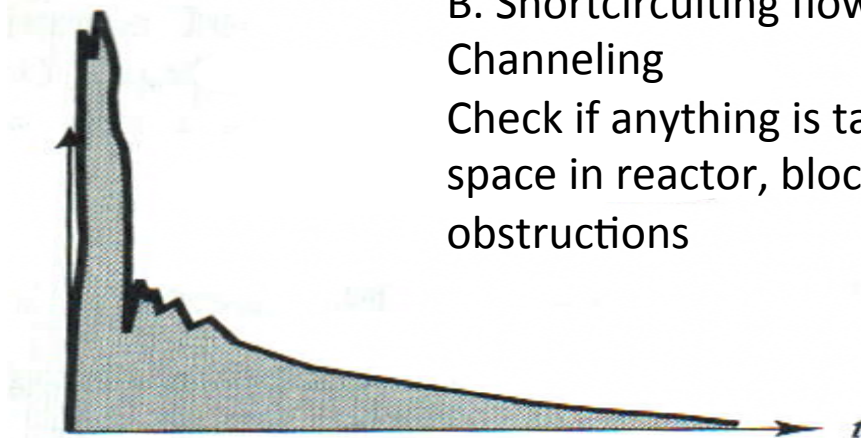
$$\frac{C_A}{C_{A0}} = \underline{\underline{0.047}}$$

From the table we see that the unconverted material comes mostly from the early portion of the E curve.





A. Normal MFR



B. Shortcircuiting flow/  
Channeling

Check if anything is taking up  
space in reactor, blockages/  
obstructions