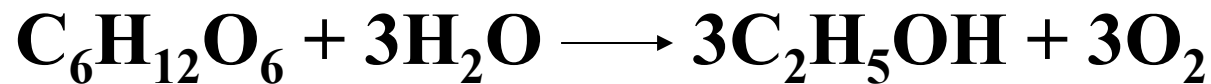


# More examples: batch, PFR, MFR

Your boss wants you to choose a reactor for the enzymatic conversion of glucose to ethanol.



The rxn is run in a large excess of water so the rxn only depends on the concentration of glucose [G].

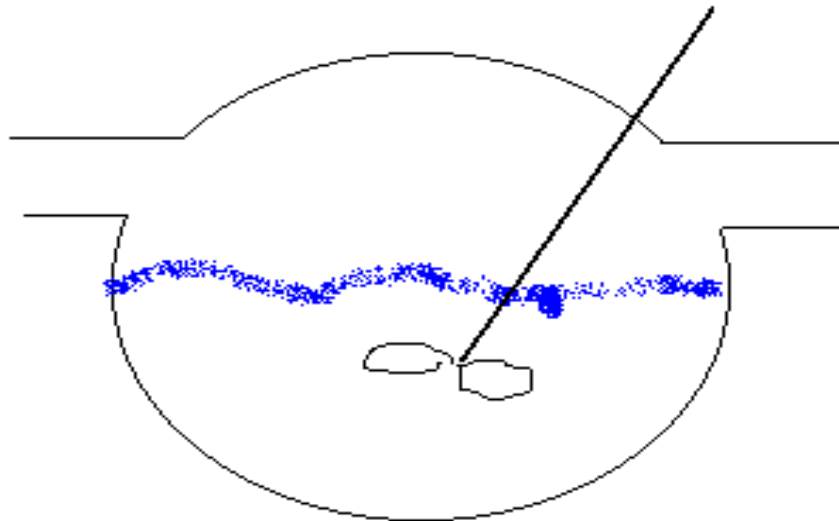
Assume a fixed amount of enzyme.

$$K_m = 5 \text{ mol/m}^3$$

$$V_m = 0.1 \text{ mol/m}^3\text{-min}$$

$$[G_0] = 15 \text{ mol/m}^3$$

- A)** You have a tank of volume 350 L.  
Using it as a batch reactor, find the  
time needed to reach a conversion  
of 75%.

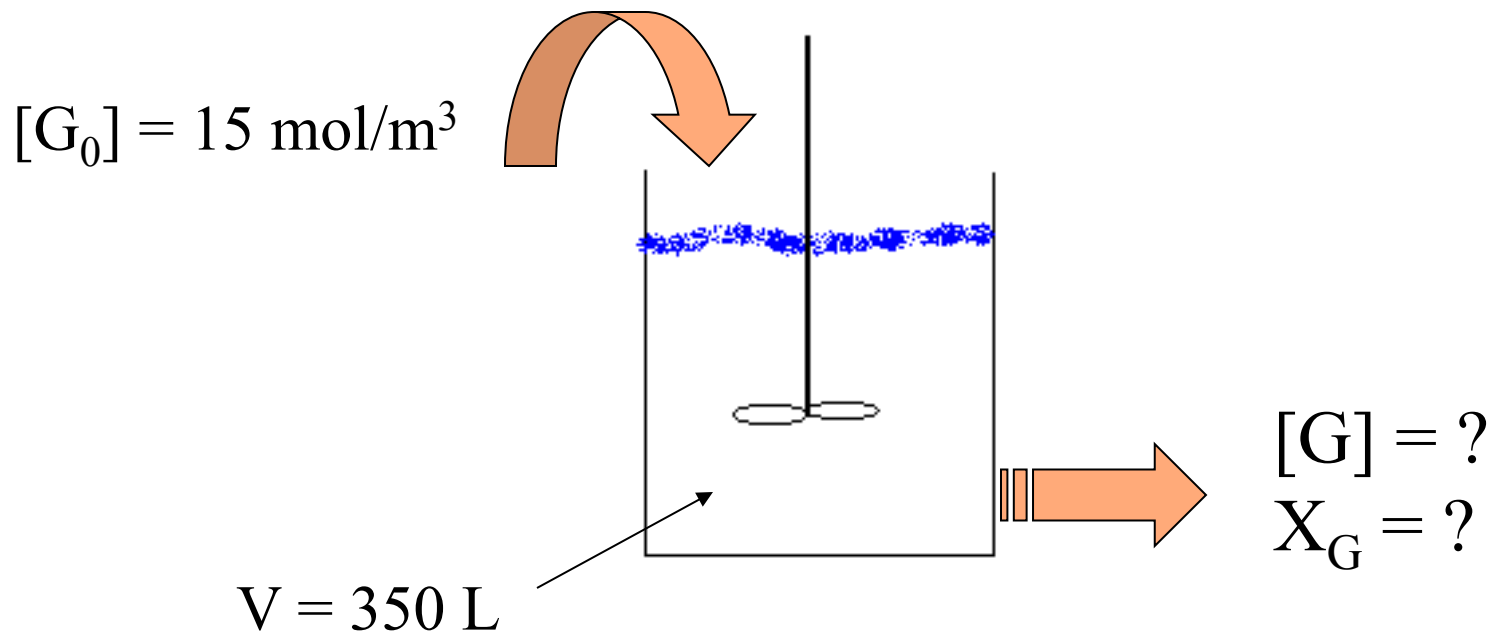


Answer:

$$\begin{aligned} t &= \frac{dC_A}{r_A} = - \int \frac{d[G]}{\frac{V_m[G]}{K_m + [G]}} = - \int_{[G_0]}^{[G]} \frac{K_m d[G]}{V_m[G]} - \int_{[G_0]}^{[G]} \frac{d[G]}{V_m} \\ &= \frac{K_m}{V_m} \ln \frac{[G_0]}{[G]} + \frac{1}{V_m} \{[G_0] - [G]\} \\ &= \frac{K_m}{V_m} \ln \frac{1}{1 - X_G} + \frac{[G_0]}{V_m} X_G = \frac{5}{0.1} \ln \frac{1}{1 - 0.75} + \frac{15}{0.1} (0.75) \end{aligned}$$

$$\mathbf{t = 181.8 \text{ min} = 3.03 \text{ hrs}}$$

B) You can install pumps in the 350 L tank that will provide a flow rate of 25 L/min. If you run the tank as a mixed flow reactor what conversion can you obtain?



Answer:

$$\tau = \frac{C_{A0} - C_A}{-r_A} = \frac{[G_0] - [G]}{\frac{V_m [G]}{K_m + [G]}} = \frac{([G_0] - [G])(K_m + [G])}{V_m [G]}$$

$$= \frac{K_m [G_0]}{V_m [G]} + \frac{[G_0] - K_m}{V_m} - \frac{[G]}{V_m}$$

$$= \frac{K_m}{V_m (1 - X_G)} + \frac{[G_0] - K_m}{V_m} - \frac{[G_0](1 - X_G)}{V_m}$$

$$\tau = \frac{K_m}{V_m} \left( \frac{X_G}{1 - X_G} \right) + \frac{[G_0]X_G}{V_m}$$

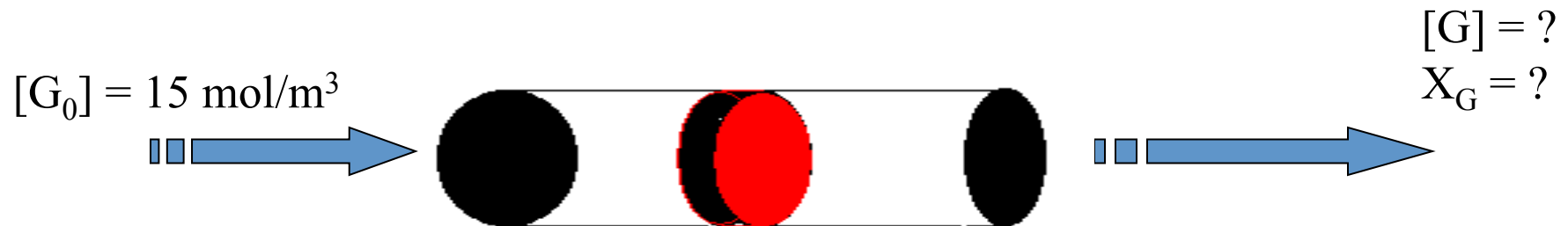
$$\tau = V/q = (350 \text{ L})/(25 \text{ L/min}) = 14 \text{ min}$$

Plug  $\tau$  into above equation, and solve for  $X_G$

$$\mathbf{X_G = 0.0687}$$

C) You may instead use a large pipe as a plug flow reactor. Its length is 20 m and its diameter is 0.6 m.

What conversion can you get?



Answer:

$$\begin{aligned}\text{Volume of reactor} &= \pi r^2 L \\ &= (20 \text{ m}) * (\pi) * (0.3 \text{ m})^2 * (1000 \text{ L}) / (1 \text{ m}^3) = 5655 \text{ L}\end{aligned}$$

$$\tau = V/q = (5655 \text{ L}) / (25 \text{ L/min}) = 226.2 \text{ min}$$

$$\tau = \frac{K_m}{V_m} \ln \left( \frac{1}{1 - X_G} \right) + \frac{[G_0] X_G}{V_m}$$

Plug in  $\tau$  and solve for  $X_G$

$$\Rightarrow \mathbf{X_G = 0.858}$$

PFR gives the best conversion of the three reactors.