

(A) Heat is given out

Q 1.

hysicsaholics

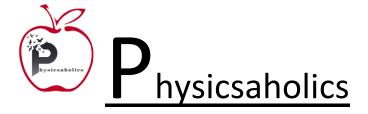
(B) Heat is taken in place whether heat is taken in or given out

If specific heat of a substance is infinite, it means--



DPP – 1

	(C) No change in tem (D) All of the above	perature takes					
Q 2.	Two spheres made of capacities are in the r (A) 1:2 (C) 1:4	f same substance have diameters in the ratio $1:2$. Their thermal atio of - (B) $1:8$ (D) $2:1$ -					
Q 3.	-	at 30°C and 20°C. When mixed in equal masses, the temperature ad to be 26°C. Their specific heats are in the ratio of - (B) 1:1 (D) 4:3					
Q 4.	and 28°C respectively	qual masses of three different liquids A,B and C are 12°C, 19°C y. The temperature when A and B are mixed is 16°C, when B 8°C; what is the temperature when A and C are mixed? (B) 20.26°C (D) 28°C					
Q 5.	Heat required to convert one gram of ice at 0°C into steam at 100°C is (given $L_{steam} = 536 \text{ cal/gm}$)- (A) 100 calorie (B) 0.01 kilocalorie (C) 716 calorie (D) 1 kilocalorie						
Q 6.	300 gm of water at 25 mixture is -:- (A) 0 °C	5°C is added to 100 gm of ice at 0°C. The final temperature of the (B) 2°C (C) 1°C (D) 3°C					
Q 7.	A 1 g of ice is mixe temperature of the mi A) 100°C (C) 75°C	d with 1 g of steam. After thermal equilibrium is achieved, the xture is: - (B) 55°C (D) 0°C					
Q 8.	If x grams of steam at 100°C becomes water at 100°C which converts y grams of ice at 0°C into water at 100°C, then the ratio x/y will be – (A) 1/3 (B) 1/2 (C) 1/4 (D) none						
Q 9.		is passed into 6 g of ice at 0°C. If the latent heats of steam and 80 cal/g, then the final temperature is— (B) 50°C (D) 100°C					





Q 10. The amount of heat required to raise the temperature of 1 kg of water through 1° C is called

(A) kilocalorie

(B) calorie

(C) B.T.U.

(D) calorie/°C

Answer Key

Q.1	C	0.2	h	Q.3 a	Q.4 b Q.5 c
V.1		V. -	D	Q.o u	
Q.6	9	Q.7	9	Q.8 a	Q.9 d Q.10 a
Q.0	а	Q. 7	a	V.0 \\ a	Q.10 a)