Boron Nitride Modulates Polymer Electrolyte Conductivity

Colby A. Snyder, Dr. Shreyas S. Pathreeker, Dr. George Papamokos, & Professor Russell J. Composto

November 29, 2023

Penn Arts & Sciences
University of Pennsylvania
Penn Engineering

The case for safer sodium-ion electrolytes

- Sodium is 100x less expensive than lithium
- Current electrolytes are flammable and risk thermal runaway
- Safer electrolytes are important for grid-scale energy storage

Thermal runaway is a major battery safety issue

Penn Engineering

Usiskin et al. (2021), Gebert et al. (2021), Scutum London (2023)

Composite polymer electrolytes are promising

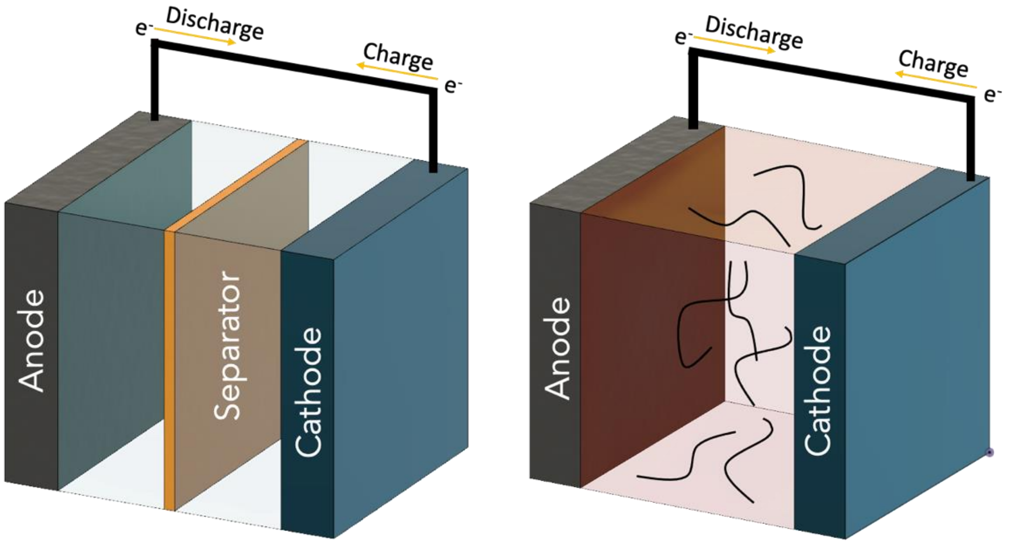


Fig. 1. Liquid (left) vs. polymer (right) electrolyte

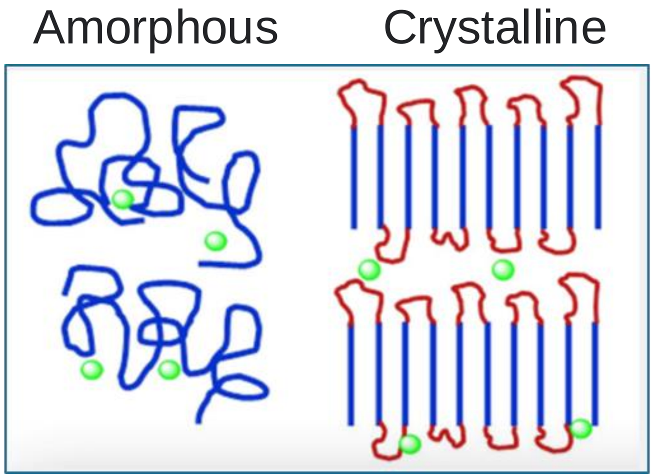
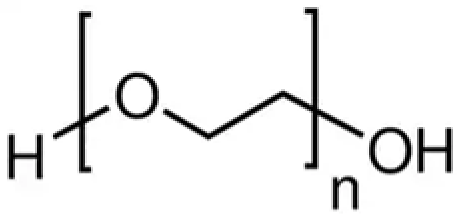


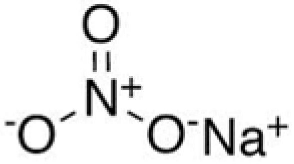
Fig. 2. (adapted from Cheng et al., 2014). Ions can only move through amorphous domains

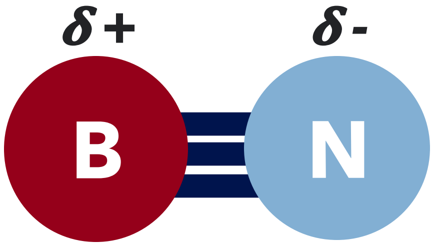
Model designed in Autodesk Fusion 360

Boron nitride is an interesting CPE filler

- Hexagonal boron nitride (BN) may interact with Na+ and NO3- ions
- BN increases the transference number of PEO–LiTFSI electrolytes
  - Li^+–BN binding energy: -157.6 kJ/mol
  - $TFSI^-$–BN binding energy: -166.51 kJ/mol





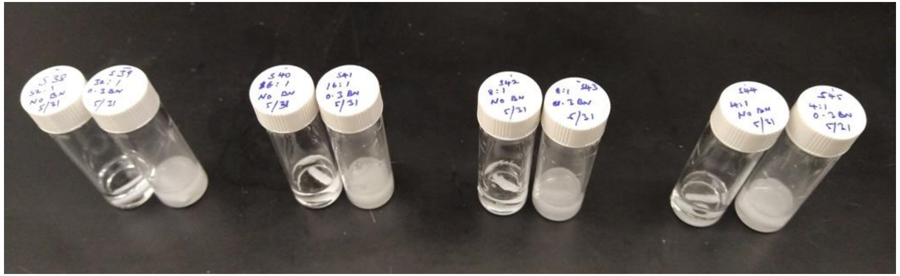


Sigma Aldrich (2023)

Methods

Phase 1: Prepare and cast solutions

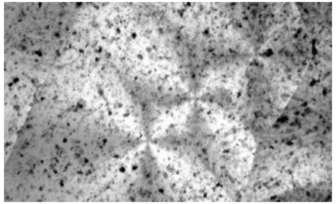
- PEO, $NaNO_3$, and h-BN mixed in water
- 0.3 wt.% and 3 wt% h-BN, 4:1 and 24:1 ether $O:Na^+$ ratios, and controls (9 samples)
- Solutions cast on 1 in^2 glass slides (381 µm-thick)



Phase 2: Dry in vacuum oven

- Hot plate at 120ºC for 1h
- Vacuum oven at 120ºC for 16h
- Vacuum chamber at 25ºC for 1h

PEO Film, 25 mm^2



PEO crystal with 3% BN, 1.5 mm^2

Penn Engineering

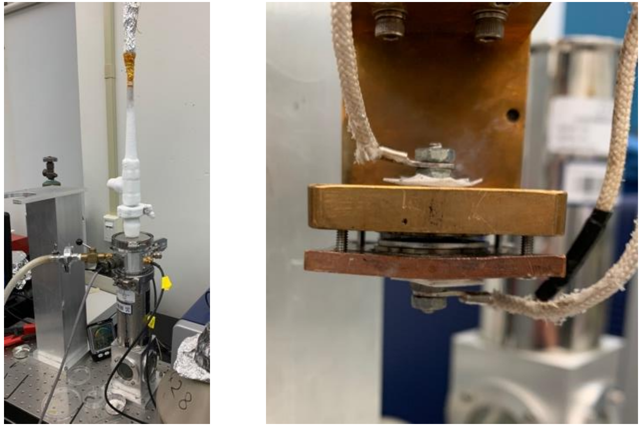
Phase 3: Characterize Samples

Structural Characterization

- Differential Scanning Calorimetry (DSC)
- X-Ray Diffraction (XRD)
- Fourier Transform Infrared Spectroscopy (FTIR)

Electrochemical Characterization

- Electrochemical Impedance Spectroscopy (EIS)



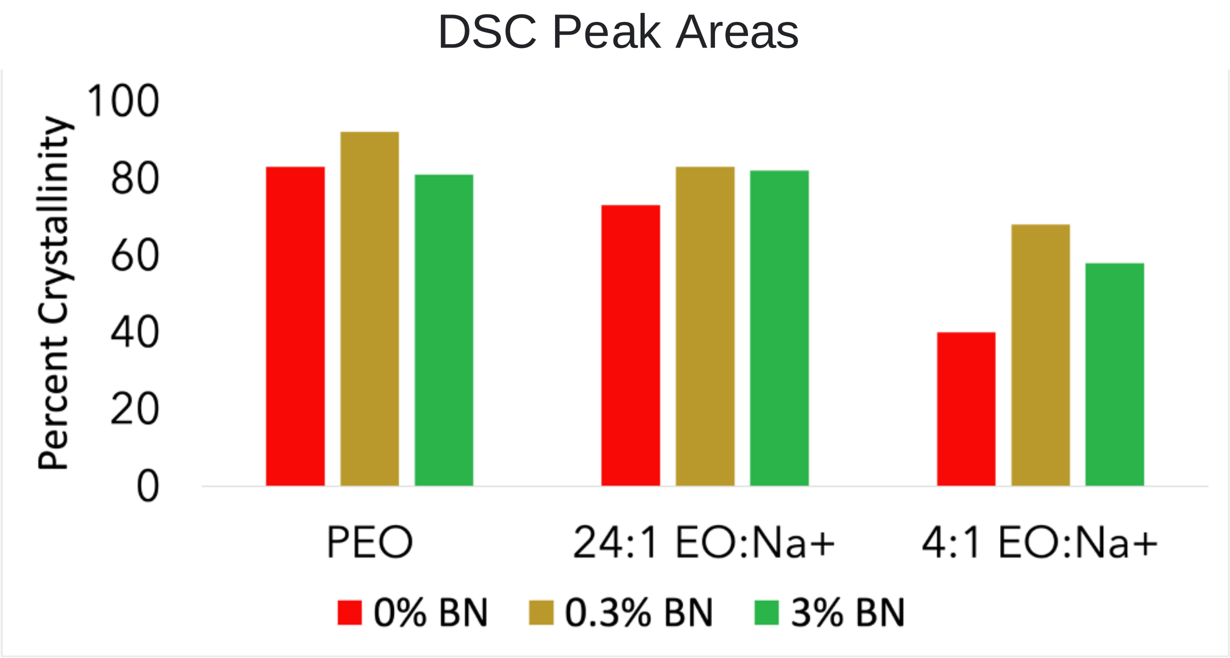
Cryostat (EIS)

Penn Engineering

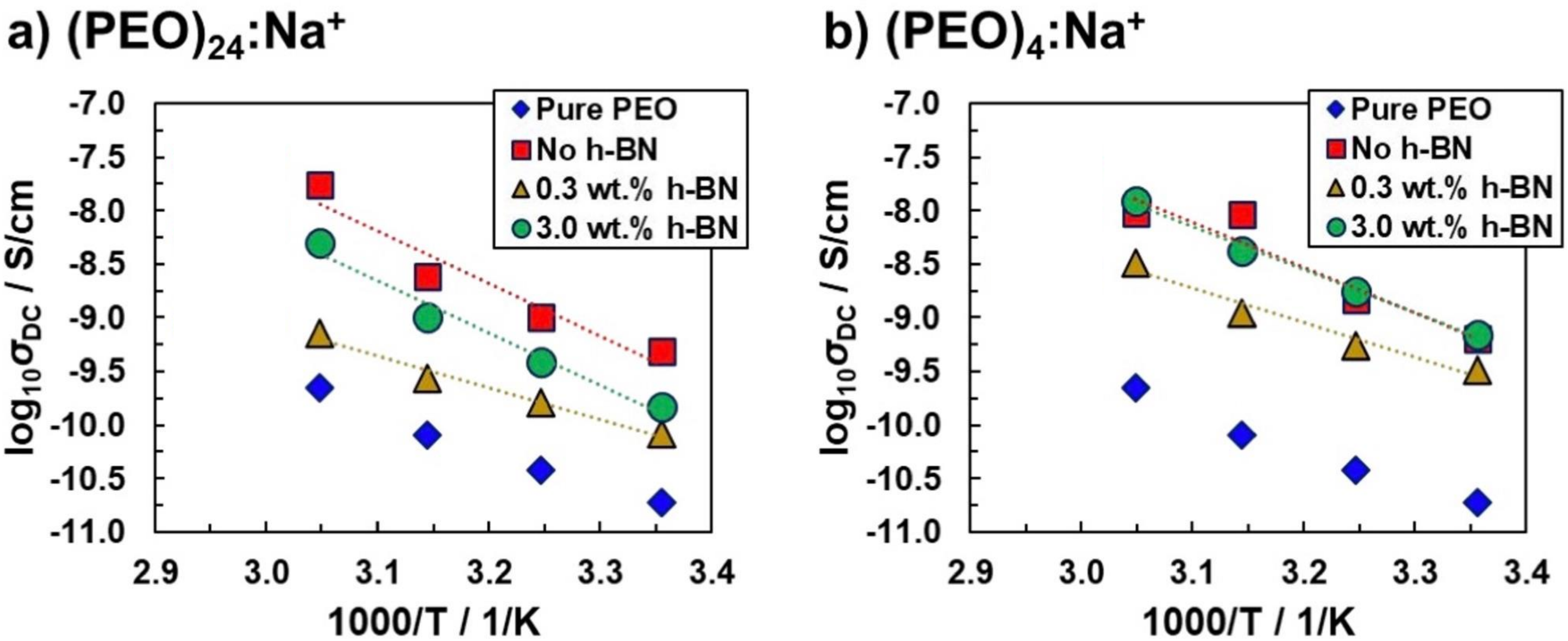
Results

Adding BN affects PEO crystallinity

- **Differential Scanning Calorimetry** shows heat of melting
- Increasing BN to 0.3% increases crystallinity
- Crystallinity decreases at 3%
- Overall increased crystallinity could hinder ion mobility



Crystallinity affects ionic conductivity

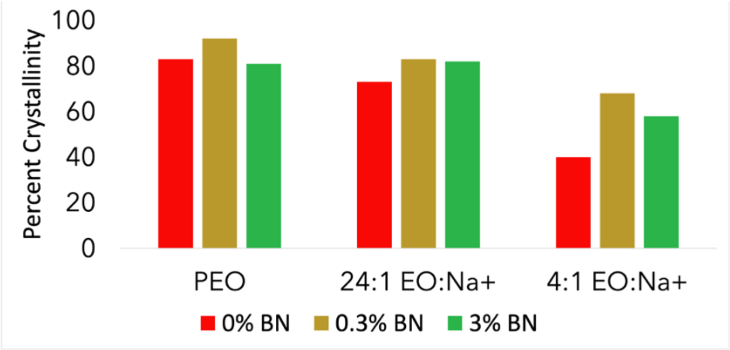


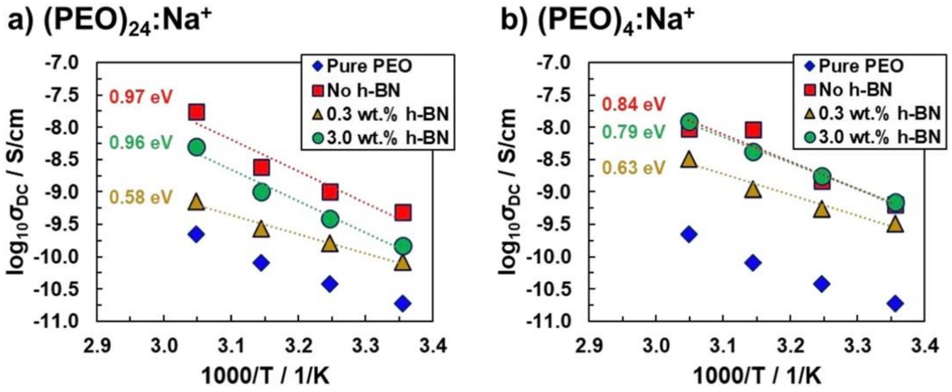
Ionic conductivity is highest for BN-free samples, showing the effect of increased crystallinity

Conclusions

- Adding BN can increase PEO crystallinity via enhanced nucleation
- The effect of increased crystallinity dominates, decreasing ionic conductivity
- Sodium CPE properties may be tailored by changing the geometry of filler materials

DSC Peak Areas





Acknowledgements

I'd like to thank the following people and organizations for making this project possible: Vagelos Integrated Program in Energy Research (VIPER), NSF grant FMRG-2134715, Ben Ferko, Steve Szewczyk, Mohamed Hassan, Professor Eric Detsi, Katie Sun, and my mentor, Shreyas Pathreeker.

VIPER @ Penn
Arts & Sciences · Engineering
UNIVERSITY of PENNSYLVANIA

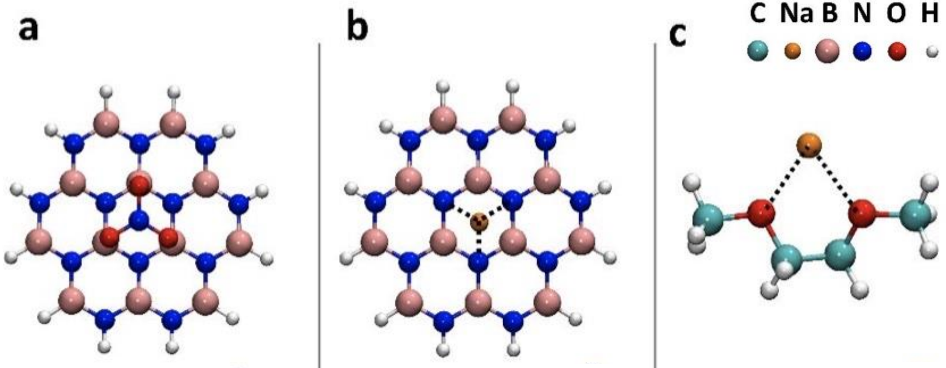
NSF

Supplementary Information

Phase 4: Compare to DFT Calculations

- Significant BN-salt binding energies
- Trends differ from lithium-ion literature

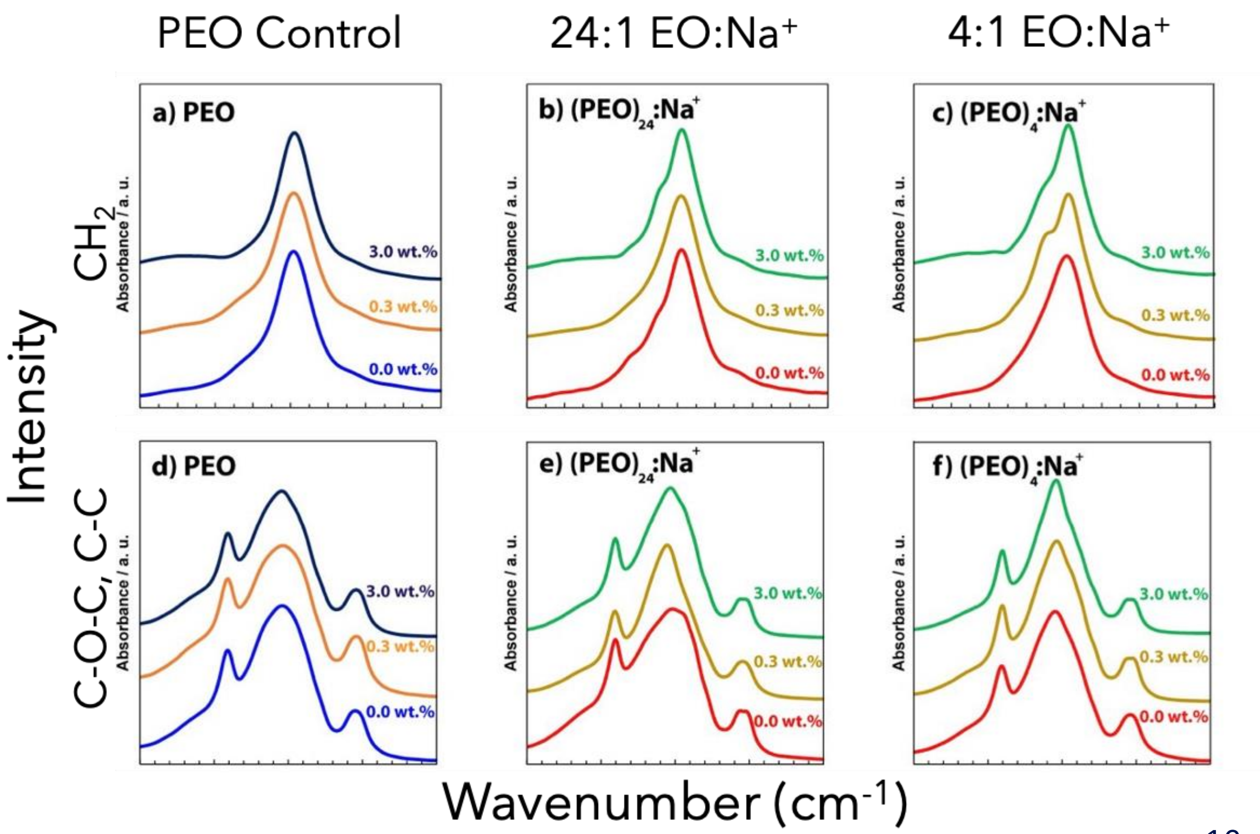
Complexation Energy (kcal/mol)



BN-NO_3^-	BN-Na^+	PEO-Na^+
-13.4	-29.3	-44.1

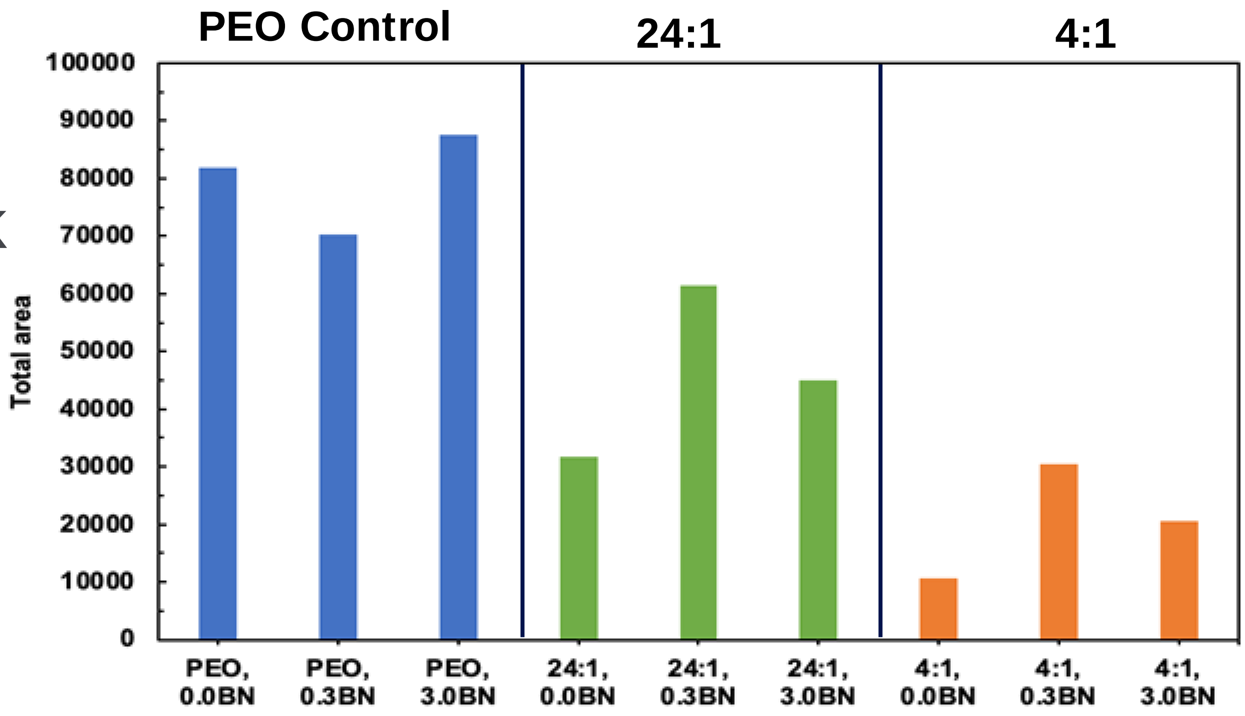
BN decreases PEO-$NaNO_3$ complexation

- PEO complexes with Na^+, decreasing crystallinity
- FTIR peaks (840 and 1100 cm^{-1}) broaden with salt
- Peaks narrow when BN is added



XRD peak area varies with h-BN loading

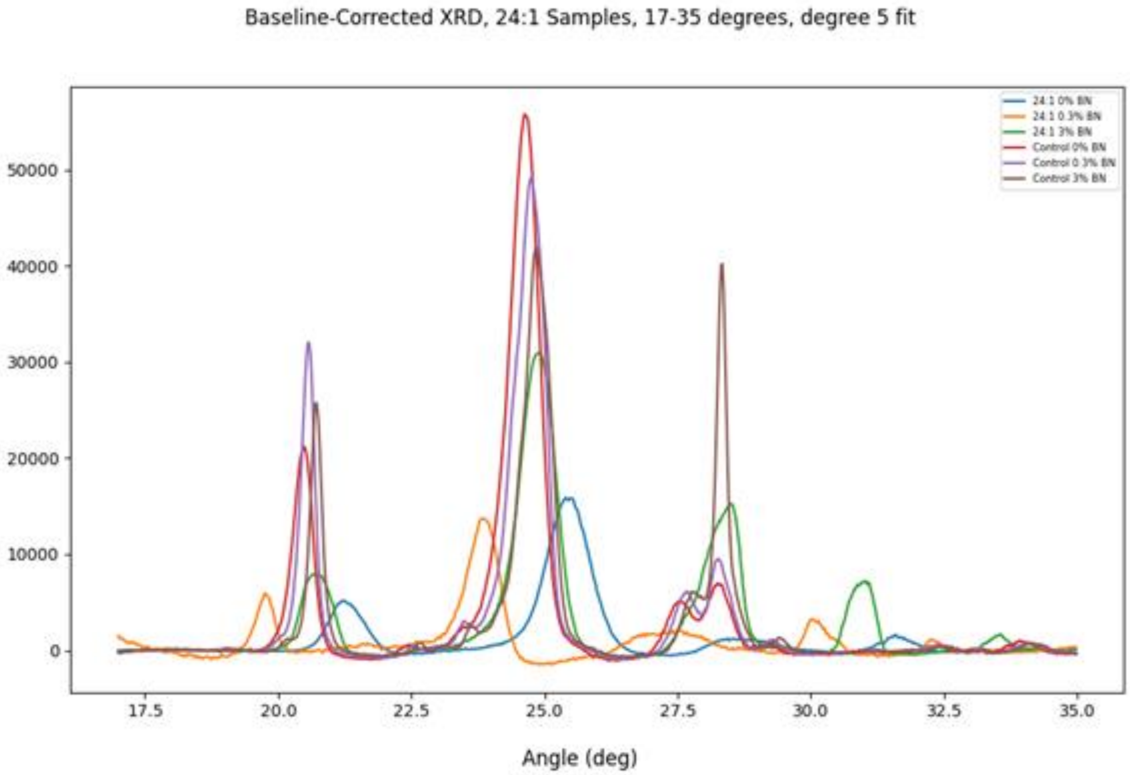
For salt-doped samples, total peak area (crystallinity) increases with 0.3% h-BN then decreases with 3% h-BN



Penn Engineering

Additional XRD from Trial 2

No significant NaNO3 peaks appear in the 24:1 XRD traces

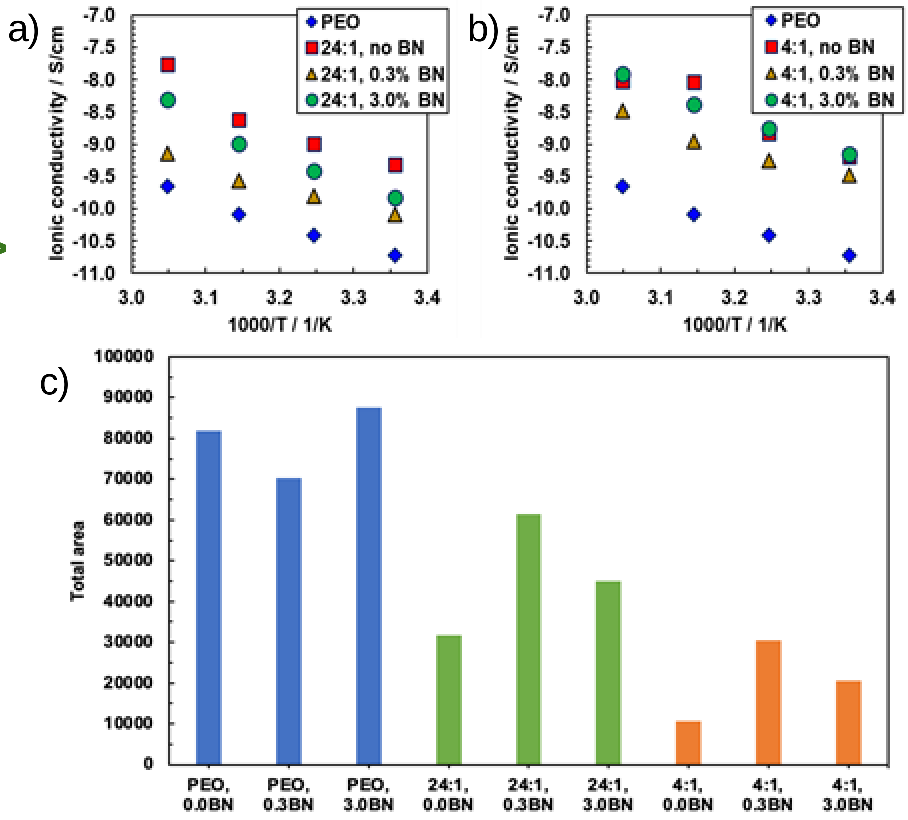


Penn Engineering

Ionic Conductivity and XRD Peak Area

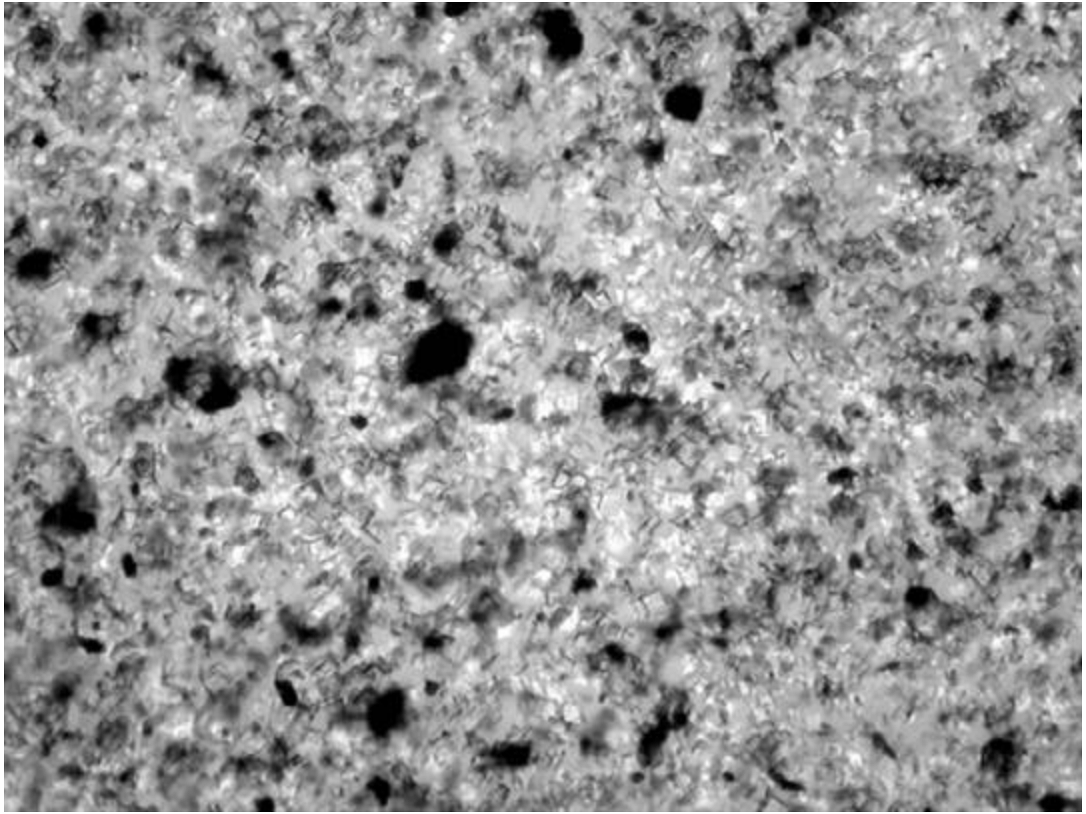
- Ionic conductivity follows the trend $IC_{\text{No h-BN}} > IC_{3.0\%} > IC_{0.3\%}$

- Interestingly, crystallinity follows the inverse trend $X_{\text{No h-BN}} < X_{3.0\%} < X_{0.3\%}$

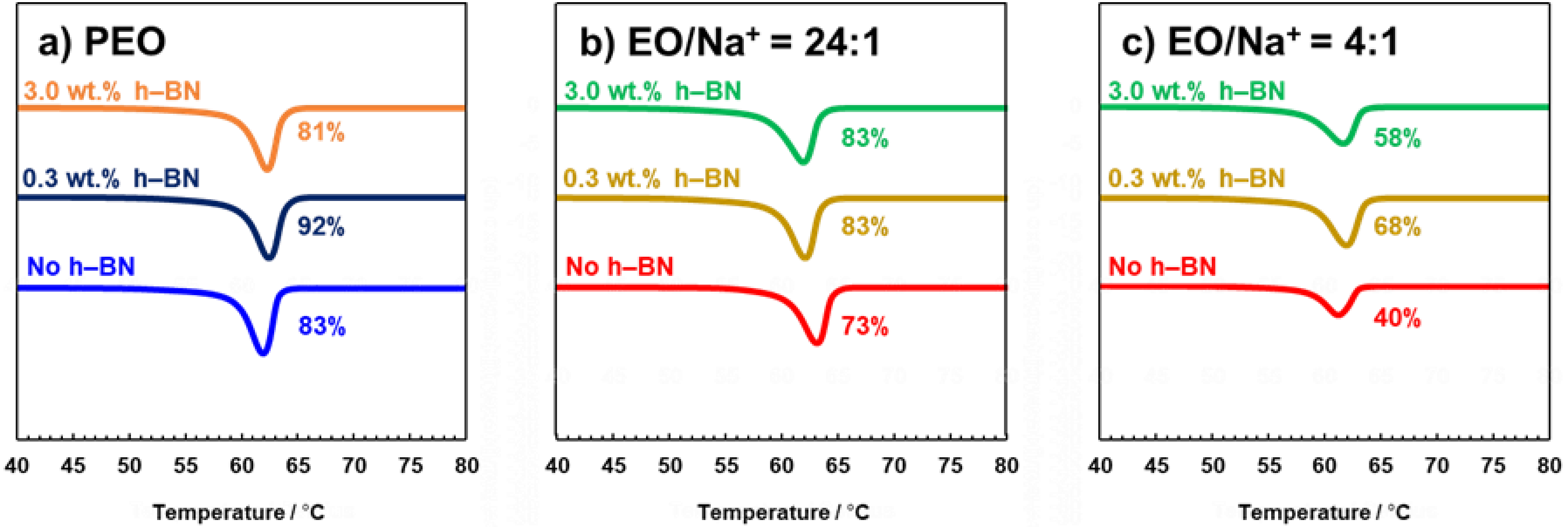


Penn Engineering

Salt Crystal image (4:1 NaNO3, 3% BN)

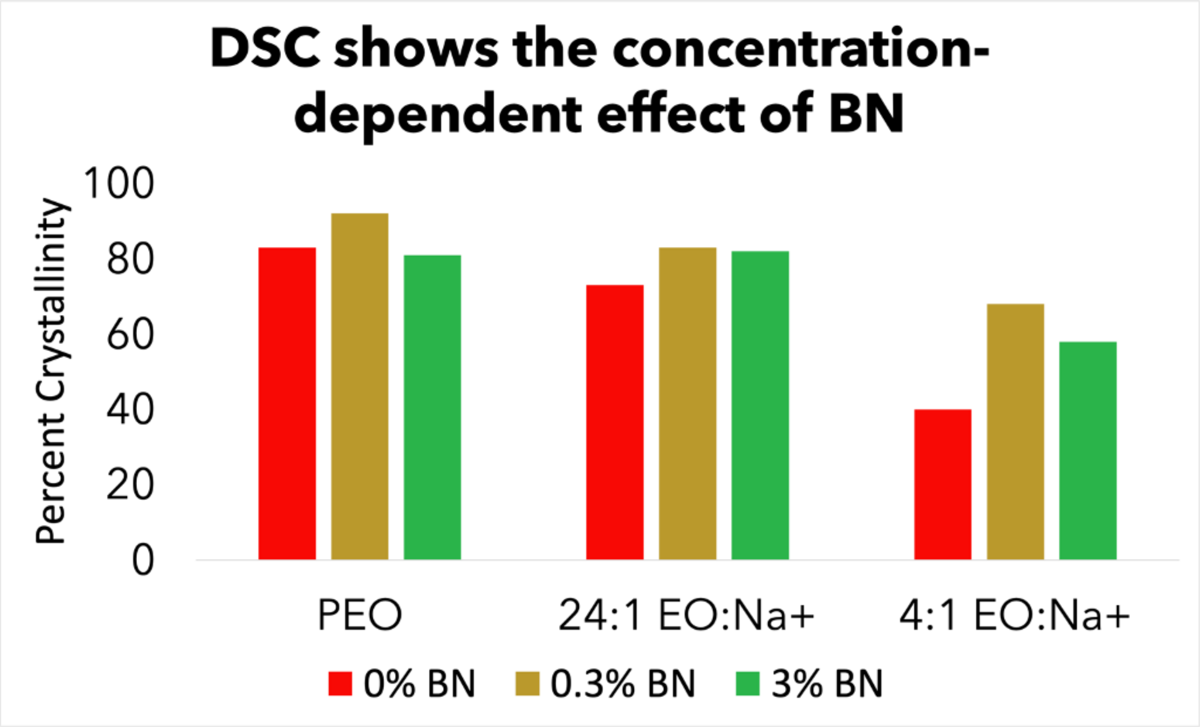


DSC on polymer electrolytes



Penn Engineering

DSC shows the concentration-dependent effect of BN



PEO-$NaNO_3$ system is soluble in H_2O