

Q1:

A: The change in concentration of reactants or products per unit time.

Derivation of integrated rate equation for first-order reaction:

$$\text{Rate} = -d[A]/dt = k[A]$$

$$d[A]/[A] = -k dt$$

$$\text{Integrating both sides: } \ln[A]_o/[A]_t = kt \text{ or } \ln[A]_t = \ln[A]_o - kt$$

Calculation of half-life:

$$t_{1/2} = 0.693/k = 0.693/0.0693 = 10 \text{ minutes}$$

K: rate, integrated rate equation, first-order reaction, half-life

M: 10

Q2:

A: First law: The mass of a substance deposited at an electrode during electrolysis is directly proportional to the quantity of electricity passed.

Second law: The masses of different substances deposited by the same quantity of electricity are proportional to their equivalent weights.

K: electrolysis, quantity of electricity, equivalent weights, Faraday's laws

M: 10

Q3:

A: Lewis structure shows phosphorus with five bonds to chlorine atoms and no lone pairs.

According to VSEPR theory, the electron pair arrangement is trigonal bipyramidal.

The five electron pairs around phosphorus repel each other to minimize repulsion.

Three chlorine atoms lie in the equatorial plane at 120° angles.

Two chlorine atoms lie at the axial positions at 180° to each other.

The molecular geometry is trigonal bipyramidal with bond angles of 120° (equatorial) and 90° (between axial and equatorial).

K: Lewis structure, VSEPR theory, trigonal bipyramidal, bond angles

M: 10