



Measuring oxygen, carbon monoxide and hydrogen sulfide diffusion coefficient and solubility in Nafion membranes

Vijay A. Sethuraman^{1,2}, Saahir Khan^{3,4}, Jesse S. Jur⁵, Andrew T. Haug⁶, John W. Weidner*

Center for Electrochemical Engineering, Department of Chemical Engineering, University of South Carolina, Columbia, SC 29208, USA

ARTICLE INFO

Article history:

Received 21 April 2009

Received in revised form 24 June 2009

Accepted 25 June 2009

Available online 3 July 2009

Keywords:

Carbon monoxide

Hydrogen sulfide

Oxygen

Diffusion

Solubility

Nafion

Electrochemical monitoring technique

ABSTRACT

A Devanathan–Stachurski type diffusion cell made from a fuel cell assembly is designed to evaluate the gas transport properties of a proton exchange membrane as a function of cell temperature and gas pressure. Data obtained on this cell using the electrochemical monitoring technique (EMT) is used to estimate solubility and diffusion coefficient of oxygen (O₂), carbon monoxide (CO) and hydrogen sulfide (H₂S) in Nafion membranes. Membrane swelling and reverse-gas diffusion due to water flux are accounted for in the parameter estimation procedure. Permeability of all three gases was found to increase with temperature. The estimated activation energies for O₂, CO and H₂S diffusion in Nafion 112 are 12.58, 20 and 8.85 kJ mol^{−1}, respectively. The estimated enthalpies of mixing for O₂, CO and H₂S in Nafion 112 are 5.88, 3.74 and 7.61 kJ mol^{−1}, respectively. An extensive comparison of transport properties estimated in this study to those reported in the literature suggests good agreement. Oxygen permeability in Nafion 117 was measured as a function of gas pressures between 1 and 3 atm. Oxygen diffusion coefficient in Nafion 117 is invariant with pressure and the solubility increases with pressure and obeys Henry's law. The estimated Henry's constant is 3.5×10^3 atm.

© 2009 Elsevier Ltd. All rights reserved.

1. Introduction

Gas diffusion across a proton exchange membrane (PEM) of a PEM fuel cell has important consequences for its performance and its durability [1,2]. Very thin membranes (~25 μm and less) are presently used in PEM fuel cells in order to decrease the membrane resistance and to increase power density. However, the direct negative consequence of using thin membranes is the increased crossover of reactants from one side to the other. In addition to this loss of fuel due to crossover, recent studies on PEM fuel cell durability suggest that H₂ and O₂ crossover play an important role in the reduction of Pt ions (diffusing from the cathode) inside the mem-

brane [3,4]. The diffusing gases further fuel the hydroxyl (HO•) and hydroxyl-peroxyl (HOO•) formation reactions on metallic catalyst particles present inside the membrane. These radicals are known to attack the tertiary hydrogen at the α-carbon of the perfluorinated membranes commonly used in PEM fuel cells [5,6]. It has been shown that increased gas crossover accelerates membrane degradation [2]. In light of all this, it is vital that one evaluates the gas crossover properties of a proton exchange membrane before they are used in a fuel cell.

Two methods commonly used to measure gas diffusion coefficient and solubility in a polymeric membrane or an ionomer layer are the electrochemical monitoring technique (EMT) [7–12] and the potential step technique (PST) [13–18]. The differences in the two methods include: data acquisition time (~15 min for the EMT vs. ~30 s for the PST), electrode size (micro-electrodes are used for the PST compared to electrodes on the order of 1 cm² for the EMT) and the experimental setup. Further, in PST, the gas is usually dissolved in a liquid electrolyte (i.e., H₃PO₄) and diffuses through the ionomer membrane deposited over the working electrode (i.e., rotating disk electrode), while in the EMT, the membrane is directly in contact with the gas-phase. The two methods have been found to agree well with each other qualitatively as shown in Table 2 for the case of O₂ diffusion in Nafion 117.

In spite of above-mentioned important consequences of gas permeability in fuel cell membranes, there is very little relevant data reported in the literature. There are few theoretical methods aimed

* Corresponding author at: Department of Chemical Engineering, Swearingen Engineering Center, University of South Carolina, 301 Main Street, Columbia, SC 29208, USA. Tel.: +1 803 777 3207; fax: +1 803 777 8265.

E-mail address: weidner@cec.sc.edu (J.W. Weidner).

¹ International Society of Electrochemistry Active Member.

² Present address: Environmental Energy Technology Division, Lawrence Berkeley National Laboratory, Berkeley, CA 94720, USA.

³ NSF-REU participant from Stanford University, Stanford, CA 94305, USA.

⁴ Present address: Feinberg School of Medicine, Northwestern University, Evanston, IL 60208, USA.

⁵ Present address: Department of Chemical Engineering, North Carolina State University, Raleigh, NC 27695, USA.

⁶ Present address: 3M Center, 3M Fuel Center Components, St. Paul, MN 55144, USA.

at understanding gas diffusion through polymeric networks [19] but there are many drawbacks associated with accurately setting up the polymer structure. For example, Nafion membranes are typically operated in the vicinity of their glass transition temperature (T_g) where the rigid frameworks representing their structures do not apply. On the other hand, there is not a robust experimental setup reported anywhere that can be used to estimate gas crossover properties of a membrane in a functional fuel cell for different fuel cell operating conditions (e.g., pressure, temperature, humidity, etc.). In this study, we report the use of the EMT on an apparatus made from fuel cell assembly to measure the diffusion coefficient and solubility values for O_2 , CO and H_2S in a Nafion membrane. The setup designed is similar to that of a Devanathan–Stachurski [7] type cell commonly employed to measure hydrogen permeability in metals. The effect of temperature on the diffusion coefficient and solubility of all three gases in Nafion 112 membrane is evaluated and the associated activation energies and enthalpies of mixing are reported. The effect of pressure on the diffusion coefficient and solubility of O_2 in Nafion is evaluated using a thicker membrane (Nafion 117). The estimated values are compared to those available in the literature. A similar approach was recently adapted by our group towards characterizing SO_2 transport in Nafion 212 and 115 membranes in a functional electrolyzer [20] under open-circuit conditions.

2. Experimental

2.1. Catalyst coated membrane (working electrode) fabrication

Pt catalyst ink with 75% catalyst and 25% Nafion (dry solids content) was prepared with commercially available 40 wt% Pt on Vulcan XC-72R E-TEK catalyst (PEMEAS Fuel Cell Technologies, Somerset, NJ), Nafion in the form of perfluorosulfonic acid-PTFE copolymer (5%, w/w solution, Alfa Aesar, Ward Hill, MA) was used. The catalyst ink was sprayed on to gas diffusion layers (ELAT GDL, 10 cm² active area, PEMEAS Fuel Cell Technologies, Somerset, NJ) with N_2 brush (Paasche Airbrush Company, Chicago, IL), air dried for 30 min and then dried under vacuum (381 mm Hg) at 110 °C for 10 min to evaporate any remaining solvent. This process was repeated until a catalyst loading of 0.5 mg Pt/cm² was achieved. One such GDL with the catalyst was then hot pressed onto a pretreated Nafion 112 or 117 membrane (Alfa Aesar, Ward Hill, MA) at 140 °C and 3450 kPa for 2 min to make a catalyst coated membrane. The catalyst-coated membrane obtained using this procedure is analogous to chemical plating procedure reported by Takenaka et al. [21].

2.2. Devanathan–Stachurski cell

The apparatus used to measure gas diffusion and solubility in a proton exchange membrane was made by modifying an actual fuel cell assembly (Fuel Cell Technologies, Albuquerque, NM) such that one half of the apparatus had gas channels and the other side had a liquid cell. A photograph of the opened up cell (without the electrodes) is shown in Fig. 1 and a cross-sectional schematic is shown in Fig. 2. This is a direct improvement of a similar glassware apparatus reported earlier by Haug and White [11] such that the effect of gas pressure and temperature can now be studied. The effect of gas humidity on gas transport in the membrane can also be studied using this setup with minimal modifications. The catalyst coated membrane was placed between the gas and liquid cells with appropriate gaskets to prevent leakage. A gas diffusion layer was placed between the gas channels and the catalyst coated membrane to ensure uniform gas distribution. The platinized side of the membrane (working electrode) was facing the liquid cell

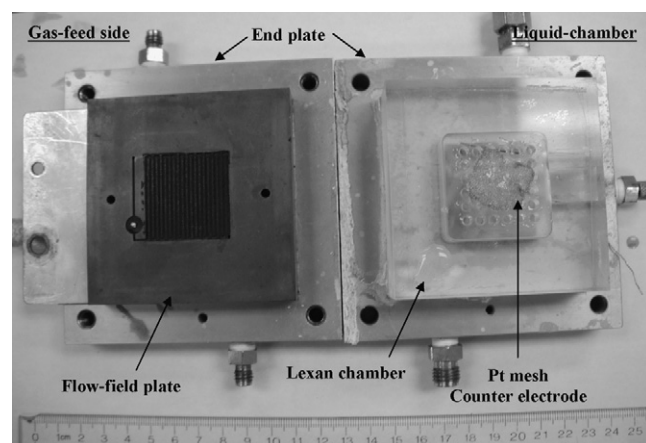


Fig. 1. Photograph of a Devanathan–Stachurski [7] type cell made using a fuel cell assembly for measuring transport and solubility of fuel cell gases in a proton exchange membrane. Shown here is the opened up cell with the gas-feed side on the left and the liquid-chamber on the right. The catalyst coated membrane (working electrode) and the gas diffusion layers are not shown here.

containing 0.1 M H_2SO_4 electrolyte, a Pt mesh counter electrode (1 cm² × 0.2 cm, Alfa Aesar) and saturated KCl–Ag/AgCl reference electrode (~200 mV vs. SHE, Orion Sure-Flow Ag/AgCl Single Junction Half-Cell, Thermo Scientific). The working electrode was held at a constant potential such that the diffusing gas was either oxidized or reduced on its surface. An EG&G Princeton Applied Research Potentiostat/Galvanostat (Model 273A, Ametek Inc., Oak Ridge, TN) was used for this purpose. The respective working electrode potentials for O_2 , CO and H_2S diffusion experiments were 0.1, 0.7 and 0.9 V vs. Ag/AgCl reference electrode. These potentials were chosen such that O_2 was reduced at the working electrode while CO and H_2S were oxidized in each of the diffusion experiments. All gases were of ultra high purity grade (Praxair Inc.). Special safety precautions were put in place for handling pure CO and H_2S . The gas channels had N_2 flowing through them initially while the background current was recorded at the working electrode. Once a steady background current was reached, N_2 was replaced by the gas of interest. This was considered the beginning of the diffusion experiment (i.e., t and i are set to zero). An increase in the current on the working electrode was monitored. Data was recorded at a sampling rate of 10 Hz until a new steady-state (limiting current) was reached. This current–time data was used to estimate the diffusion coefficient and the solubility values. The entire experiment was conducted on a fuel cell test station (Fuel Cell Technologies, Albuquerque, NM)

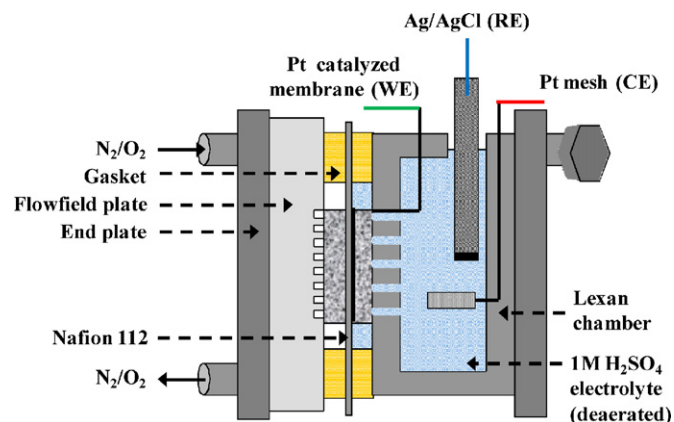


Fig. 2. Schematic (cross-sectional view) of the diffusion cell derived from an actual fuel cell assembly. WE: working electrode, RE: reference electrode and CE: counter electrode.

made for testing the performances of PEM fuel cells. All the gases fed to the diffusion cell were fully humidified (i.e., the gases were in equilibrium with saturated water vapor at the temperature of the diffusion cell). The temperature of the diffusion cell was controlled by heating elements in the end plates in conjunction with cooling fans on either side. Gas pressure was controlled by a combination of back pressure valves. Between 4 and 10 trials were performed at each set of conditions and an average value of the parameters is reported.

It must be noted that one can measure gas crossover properties of a membrane in an actual functioning PEM fuel cell built with a membrane electrode assembly. Hydrogen crossover measurements routinely conducted as part of membrane durability experiments is one such example [2]. However, the following must be considered: (1) the crossover current measured in such experiments is indicative of gas transport through half the membrane electrode assembly (i.e., microporous layer, catalyst layer and the membrane) and not just the membrane and (2) choosing the correct reference electrode is difficult when measuring the crossover of gases other than hydrogen. To mitigate these concerns and to quickly convert a functioning fuel cell into a diffusion cell, the MEA may be replaced by a catalyst-coated membrane (single-sided) with an addition of an appropriate reference electrode. In addition to this, the cell can be used to study chemical plating inside the membrane by choosing an appropriate electrolyte [e.g., $\text{H}_2\text{PtCl}_6 \cdot (\text{H}_2\text{O})_6$] in the liquid cell and a reducing gas (e.g., H_2) on the gas channels.

3. Theory

The electrochemical monitoring technique was used to determine the diffusion coefficients and solubilities for gases in membranes. Fick's law and the appropriate boundary conditions presented in Eqs. (1)–(4) were used to define the system:

$$\frac{\partial c(x, t)}{\partial t} = D_g \frac{\partial^2 c(x, t)}{\partial x^2} \quad (1)$$

$$c(x, t) = 0 \quad \text{for } 0 \leq x \leq L(\lambda) \quad \text{for } t < 0 \quad (2)$$

$$c(x, t) = c_g \quad \text{for } x = 0 \quad \text{for } t \geq 0 \quad (3)$$

$$c(x, t) = 0 \quad \text{for } x = L(\lambda) \quad \text{for } t \geq 0 \quad (4)$$

where D_g is the diffusion coefficient, c_g is the solubility of the diffusing gas, and $L(\lambda)$ is the thickness of the membrane. The steady-state limiting and reaction currents were given by Eqs. (5) and (6), respectively,

$$i_\infty = \frac{n_{e,g} F A D_g c_g}{s_g L(\lambda)} \quad (5)$$

$$i(t) = \frac{n_{e,g} F A D_g}{s_g} \left. \frac{\partial c(x, t)}{\partial x} \right|_{x=L(\lambda)} \quad (6)$$

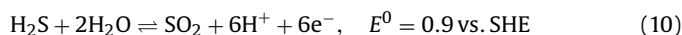
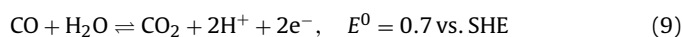
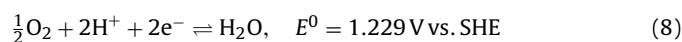
where $n_{e,g}$ is the number of electrons in the electrochemical reaction of interest corresponding to reduction or oxidation of gas g , s_g is the stoichiometric coefficient of the diffusing gas, $L(\lambda)$ is the membrane thickness, A is the cross-sectional area of the working electrode, and F is Faraday's constant. The membrane of interest in this case is Nafion and since Nafion swells upon water uptake, the diffusion length, $L(\lambda)$, is a function of water content [22] and is given as,

$$L(\lambda) = L^0 \left(1 + 0.36 \frac{\hat{\lambda} \bar{V}_0}{\bar{V}_m} \right) \quad (7)$$

where λ is the ratio of moles of water per mole of sulfonic acid sites in the membrane (i.e., water content), L^0 is the dry membrane thickness, $\hat{\lambda}$ is the average water content in the membrane, \bar{V}_0 is the initial volume of the membrane and \bar{V}_m is the partial molar volume

of the membrane. The catalyst layer is assumed to be infinitesimally thin and swelling of ionomer in the catalyst is ignored. The partial molar volume is obtained from the ratio of the molecular weight of the membrane and its density ($\bar{V}_m = M_m / \rho_m$). Myers and Newman have treated swelling in this manner to ensure conservation of membrane mass [23]. The Devanathan–Stachurski type setup used to measure gas diffusion in this work (see Section 2) has the membrane in equilibrium with 0.1 M H_2SO_4 on one side and fully humidified gas on the other side. Therefore, λ is 22 (assuming water activity to be same in 0.1 M H_2SO_4 as in pure liquid water) on the membrane–liquid interface and 14 on the membrane–gas interface [24] and if one assumes a linear profile for λ inside the membrane, the value for $\hat{\lambda}$ results in 18. This assumption is in accord with the experimental observation by Morris and Sun [25]. Temperature effect on membrane water uptake is ignored. The new diffusion lengths, therefore, are 50.83 and 177.86 μm for Nafion 112 and 117 membranes, respectively.

The electrochemical reactions of interest for O_2 [26], CO [27] and H_2S [28], respectively are,



Fan [29] solved the above system of equations using the Laplace transform techniques resulting in Eq. (11),

$$i(t) = \frac{n_{e,g} F A D_g c_g}{s_g L(\lambda)} \left(\frac{2}{\sqrt{\pi \tau}} \sum_{j=0}^{\infty} \exp \left[-\frac{(2j+1)^2}{4\tau} \right] \right) \quad (11)$$

where,

$$\tau = \frac{t D_g}{L(\lambda)^2} \quad (12)$$

Eq. (11) can be used with data for $i(t)$ to obtain D_g and c_g as discussed below.

The method of least squares [30] was used to fit the data from each trial to Eq. (11) and solve for the diffusion coefficient and solubility simultaneously. To determine the accuracy of values obtained for D and c_g , confidence intervals were obtained by using the method described by Kimble et al. [31] shown in Eq. (13),

$$P_k = \hat{P}_k \pm t_\gamma s_{\hat{P}_k} \sqrt{C_{kk}} \quad (13)$$

where \hat{P}_k is the estimate of parameter P_k found through the least squares method, $s_{\hat{P}_k}$ is the standard deviation for the data set, and t_γ is the value of the t -distribution (also known as the student distribution) [10,32,33] with a confidence, γ . Eq. (14) is solved for t_γ to obtain the t -distribution,

$$\int_{t_\gamma}^{\infty} \frac{\Gamma[(f-1)/2]}{\sqrt{\pi f} \Gamma(f/2)} \left(1 + \frac{x^2}{f} \right)^{-(f+1)/2} dx = \alpha \quad (14)$$

$$\alpha = \frac{(1-\gamma)}{2} \quad (15)$$

where f is the degrees of freedom and is equal to $(n-m)$, where n is the number of data points and m is the number of parameters (two in this case, D_g and c_g).

A value for C_{kk} in Eq. (13) can be obtained from the approximate Hessian Matrix [31],

$$\mathbf{N} = \begin{bmatrix} 2 \sum_{j=1}^n \frac{\partial i(j)}{\partial P_{D_g}} \frac{\partial i(j)}{\partial P_{D_g}} & 2 \sum_{j=1}^n \frac{\partial i(j)}{\partial P_{D_g}} \frac{\partial i(j)}{\partial P_{C_g}} \\ 2 \sum_{j=1}^n \frac{\partial i(j)}{\partial P_{C_g}} \frac{\partial i(j)}{\partial P_{D_g}} & 2 \sum_{j=1}^n \frac{\partial i(j)}{\partial P_{C_g}} \frac{\partial i(j)}{\partial P_{C_g}} \end{bmatrix} \quad (16)$$

where $i(j)$ is the current, i , recorded at each data point, j . Eq. (16) is then inverted and the diagonal elements of that matrix, $\mathbf{N}[1,1]$ and $\mathbf{N}[2,2]$, are taken as C_{kk} ($C_{D_g D_g}$ for diffusivity and $C_{C_g C_g}$ for solubility).

Since the membrane is in equilibrium with 0.1 M H_2SO_4 on one side and fully humidified gas on the other, the resulting water flux and the associated transport of dissolved gas from the liquid side to the gas side (counter to the direction of diffusion) need to be quantified. Nguyen and White [34] report the Fickian diffusion coefficient [35] for water in Nafion as,

$$D_{W,F} = (1.76 \times 10^{-5} + 1.94 \times 10^{-4} \lambda) \exp \left[\frac{-2436}{T} \right] \quad (17)$$

The permeability of the gases studied (namely O_2 , CO and H_2S) due to water transport in the direction counter to the direction of gas diffusion can then be given as,

$$P_{g,W} = X_{g,W} c_W D_{W,F} \quad (18)$$

where $X_{g,W}$ is the mole fraction solubility of gas g ($g = \text{O}_2$, CO or H_2S) in pure liquid water, c_W is the solubility of water in Nafion membrane. The mole fraction solubility of the O_2 [36], CO [37] and H_2S [38] are respectively calculated using the following correlations,

$$\ln(X_{\text{O}_2,W}) = A_1 + \frac{B_1}{T^*} + C_1 \ln T^*, \quad T^* = \frac{T}{100} \quad (19)$$

$$\ln(X_{\text{CO},W}) = A_2 + A_3 \frac{100}{T} + A_4 \ln \left(\frac{T}{100} \right) + A_5 \frac{T}{100} + S \left[B_2 + B_3 \frac{T}{100} + B_4 \left(\frac{1}{100} \right)^2 \right] \quad (20)$$

$$\ln(X_{\text{H}_2\text{S},W}) = A_6 + A_7 T + A_8 T^2 + \frac{B_5}{T} + C_2 \ln T \quad (21)$$

A_n , B_n and C_n are coefficients and their values are given in Table 1. The permeability values estimated for the gases from the gas side to the liquid side is corrected for the above. Activation energies for diffusion and mixture enthalpies for the gases were estimated by fitting the estimated diffusion coefficient and solubility values, respectively, to the following Arrhenius relationships,

$$D_g = D_g^0 \exp \left[\frac{E_{a,g}}{RT} \right] \quad (22)$$

$$c_g = c_g^0 \exp \left[\frac{H_g}{RT} \right] \quad (23)$$

where $E_{a,g}$ and H_g are respectively the activation energy for gas diffusion and mixture enthalpy for gas g in a Nafion membrane. Maple 10 (Maplesoft, Waterloo, Ontario) was used to run the parameter estimation routines. Values for parameters used in the data analysis are listed in Table 1.

4. Results and discussion

4.1. Oxygen (O_2)

Fig. 3 shows the background corrected current–time data obtained at the working electrode due to O_2 diffusion through a

Table 1

Parameters used in data analysis for Nafion 112 and 117 membranes.

Parameters	Value	Comments
$A^{\text{Nafion 112}}$	10 cm ²	Measured
$A^{\text{Nafion 117}}$	1.76 cm ²	Measured
A_1	−66.7354	Ref. [36]
A_2	−169.4951	Ref. [37]
A_3	263.5657	Ref. [37]
A_4	159.2552	Ref. [37]
A_5	−25.4967	Ref. [37]
A_6	−3.3747	Ref. [38]
A_7	0.072437	Ref. [38]
A_8	−1.10765 × 10 ^{−4}	Ref. [38]
B_1	87.4755	Ref. [36]
B_2	0.051198	Ref. [37]
B_3	−0.044591	Ref. [37]
B_4	0.0086462	Ref. [37]
B_5	−1549.159	Ref. [38]
C_1	24.4526	Ref. [36]
C_2	0.144237	Ref. [38]
$I_0^{\text{Nafion 112}}$	0.00508 cm	Manufacturer data
$I_0^{\text{Nafion 117}}$	0.01778 cm	Manufacturer data
k_s	3.9 × 10 ^{−8}	Ref. [43]
M_m	1100 g mol ^{−1}	Manufacturer data
n_{e,O_2}	2	Ref. [26]
$n_{e,\text{CO}}$	2	Ref. [27]
$n_{e,\text{H}_2\text{S}}$	6	Ref. [28]
S	0.34	Ref. [37]
S_{O_2}	1/2	Ref. [26]
S_{CO}	1	Ref. [27]
$S_{\text{H}_2\text{S}}$	1	Ref. [28]
$\bar{V}_0^{\text{Nafion 112}}$	0.0508 cm ³	Measured
$\bar{V}_0^{\text{Nafion 117}}$	0.0313 cm ³	Measured
\bar{V}_m	550 cm ³ mol ^{−1}	Estimated
$\hat{\lambda}$	18	Estimated
ρ_m	2 g cm ^{−3}	Manufacturer data
γ	0.95	Confidence

Nafion 112 membrane followed by its reduction according to Eq. (8). Three switches between N_2 and O_2 are shown. The area under the curve marked O_2 or N_2 in the figure represents the corresponding gas in the bulk on the gas-channels side of the diffusion cell. It can be seen that the rise time and the steady-state current (i_∞) are uniform between the switches. Table 2 lists values for diffusion coefficient and solubility values for O_2 in the literature for various temperatures and pressures. The values reported by Haug

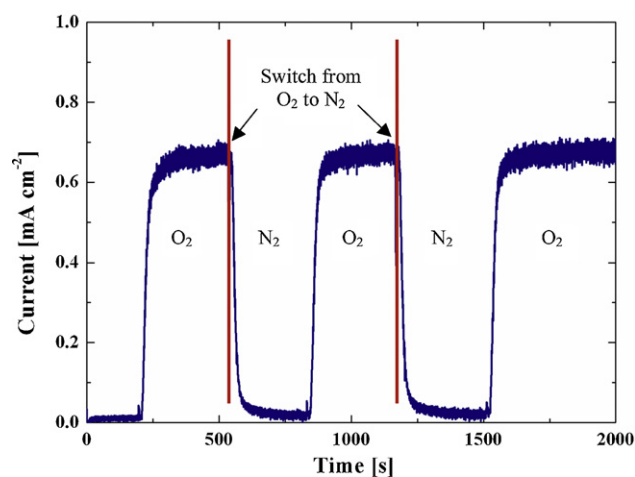


Fig. 3. Background corrected chronoamperometric response at the working electrode for switches between N_2 and O_2 on the gas side at 35 °C. The rise in current followed by a stable value is indicative of O_2 diffusion in the membrane (Nafion 112) and reduction to $\text{H}_2\text{O}/\text{H}_2\text{O}_2$. The working electrode was held at 0.1 V vs. Ag/AgCl reference electrode. This current response from the working electrode was used in conjunction with Eq. (11) to extract O_2 diffusion coefficient and solubility in Nafion 112.

Table 2Values for O₂ diffusion coefficient and solubility in Nafion 117 reported in the literature.

Source	$D (\times 10^6 \text{ cm}^2 \text{ s}^{-1})$	$c_g (\times 10^6 \text{ mol cm}^{-3})$	$T (^\circ\text{C})$	$P (\text{atm})$	Method
Haug and White [11]	0.62	18.7	25	1	EMT ^a
Lehtinen et al. [12]	0.70	13.0	20	1	EMT
	1.9	9.3			PST ^b
Parthasarathy et al. [13]	0.74	26.0	25	1	PST
Parthasarathy et al. [45]	2.88	5.76	40	5	
Ogumi et al. [9] ^c	0.24	7.2	20	1	EMT
	0.29	6.5	30		
	0.44	5.3	40		
	0.52	5.9	50		
Basura et al. [14]	6.0	9.2	30	3	PST
Beattie et al. [15]	5.96	9.16	30	3	PST
	7.87	8.27	40	3	
	9.76	7.53	50	3	
	9.09	8.20	60	3	
	10.31	7.81	70	3	
Beattie et al. [15]	5.24	6.36	60	2	PST
	5.46	7.89		2.5	
	5.48	9.35		3	
	5.47	10.93		3.5	
	5.71	11.83		4	
	6.71	12.10		4.5	
	7.07	12.52		5.0	
Chiou and Paul [40] ^d	0.0457	0.106	35	1	GC ^e
Büchi et al. [46]	2.6	4.8	25	1	PST
Gode et al. [47] ^f	1.1	4.0	25	1	PST
	1.7	3.8	60	1	
Parthasarathy et al. [13]	0.995	9.34	30	5	PST
	2.88	5.76	40		
	3.81	5.30	50		
	5.23	4.96	60		
	6.22	4.92	70		
	8.70	4.43	80		

^a Electrochemical monitoring technique.^b Potential step technique.^c Nafion 120.^d Dry Nafion 117.^e Gas chromatography {See Koros et al. [48] for its design and operation}.^f Reported values at 25 and 60 °C are respectively obtained at 82% and 75% RH.

and White [11] were used as the initial guess for the parameter estimation routine based on O₂ diffusion measurements in this study.

4.2. Effect of temperature

The diffusion measurements for O₂ were conducted as a function of temperature for Nafion 112 and Nafion 117 membranes and the current response at the working electrode is plotted in Fig. 4. Because of the longer diffusion length, the evolution of current for the case of Nafion 117 (Fig. 4b) is slower and distinct with temperature compared to that of Nafion 112 (Fig. 4a). For reasons unknown, the noise to signal ratio is considerably larger for the Nafion 112 membrane. As can be seen, the permeability of O₂ (proportional to the steady-state current) increases with temperature for both membranes. The model fits the data well over the entire temperature range for both membranes. Comparison between data and fit from Eq. (11) is shown for two temperatures for Nafion 112 in Fig. 5. The corresponding error values indicate that the disparity between the model and the data are high during the transience while it is mostly noise during steady-state. The resulting diffusion coefficient and solubility values for various temperatures for Nafion 112 and 117 membranes are plotted in Figs. 6 and 7, respectively. The corresponding confidence intervals and the permeability values for Nafion 112 and Nafion 117 are listed in Table 3a and b, respectively. Oxygen permeability values estimated in this work by the EMT technique is similar to those estimated by Broka and Ekdunge [39] using gas chromatography; by Chiou and Paul [40] as well as by Sakai et al. [41,42] using volumetric methods with high pressure permeation

cells. The exponential fits to data obtained in this work as shown in Fig. 6 correspond to the following,

$$D_{\text{O}_2, \text{Nafion 112}} = 17.45 \times 10^{-6} \exp \left[\frac{-1514}{T} \right] \quad (24)$$

$$D_{\text{O}_2, \text{Nafion 117}} = 24.82 \times 10^{-5} \exp \left[\frac{-1949}{T} \right] \quad (25)$$

The corresponding activation energies for O₂ diffusion in Nafion 112 and 117 are 12.58 and 16.2 kJ mol⁻¹, respectively. The exponential fits to data obtained in this work as shown in Fig. 7 correspond to the following,

$$c_{\text{O}_2, \text{Nafion 112}} = 10.29 \times 10^{-4} \exp \left[\frac{-707.5}{T} \right] \quad (26)$$

$$c_{\text{O}_2, \text{Nafion 117}} = 2.41 \times 10^{-6} \exp \left[\frac{-605}{T} \right] \quad (27)$$

The corresponding enthalpies of mixture for O₂ in Nafion and Nafion 117 are 5.88 and 5.03 kJ mol⁻¹, respectively.

4.3. Effect of pressure

The diffusion cell was used to evaluate the effect of O₂ pressure on its permeability in Nafion 117. A thicker membrane was chosen to withstand the pressure differential since the liquid chamber was not pressurized. The electrochemical response at the working electrode is shown for different O₂ pressures from 1 to 3 atm in Fig. 8 and the corresponding diffusion coefficient, solubility and permeability values along with the confidence intervals are given in Table 3c. The model fit the data very well over the entire pressure

Table 3

Diffusion coefficient, solubility and permeability of O₂ in Nafion 112 and Nafion 117 membranes as a function of temperature and pressure. The confidence intervals for the estimated parameters are given as well.

Temperature (°C)	Diffusion coefficient, <i>D</i> (× 10 ⁶ cm ² s ^{−1})	Confidence interval ^a (× 10 ⁸ cm ² s ^{−1})	Solubility, <i>c_g</i> (× 10 ⁶ mol cm ^{−3})	Confidence interval ^a (× 10 ⁶ mol cm ^{−3})	Permeability, <i>D</i> × <i>c_g</i> (× 10 ¹² mol cm ^{−1} s ^{−1})
(a) Nafion 112					
25	0.1048	0.28	102.63	3.27	10.75
35	0.2049	0.40	84.13	1.87	17.24
45	0.1877	0.32	113.50	2.18	21.32
55	0.1715	0.33	142.72	3.20	24.47
65	0.1895	0.36	139.46	3.10	26.42
75	0.2331	0.77	115.62	4.49	26.95
Temperature (°C)	Diffusion coefficient, <i>D</i> (× 10 ⁶ cm ² s ^{−1})	Confidence interval ^a (× 10 ⁶ cm ² s ^{−1})	Solubility, <i>c_g</i> (× 10 ⁶ mol cm ^{−3})	Confidence interval ^a (× 10 ⁶ mol cm ^{−3})	Permeability, <i>D</i> × <i>c_g</i> (× 10 ¹² mol cm ^{−1} s ^{−1})
(b) Nafion 117					
10	0.256	0.023	23.00	2.24	5.87
25	0.577	0.055	15.50	1.38	8.91
40	1.079	0.170	16.91	2.23	18.19
50	1.349	0.173	15.26	1.98	20.53
60	1.350	0.154	16.54	2.58	22.25
70	1.225	0.291	17.19	5.13	20.78
80	1.316	0.457	14.15	7.49	18.17
Pressure (atm)	Diffusion coefficient, <i>D</i> (× 10 ⁶ cm ² s ^{−1})	Confidence interval ^a (× 10 ⁷ cm ² s ^{−1})	Solubility, <i>c_g</i> (× 10 ⁶ mol cm ^{−3})	Confidence interval ^a (× 10 ⁶ mol cm ^{−3})	Permeability, <i>D</i> × <i>c_g</i> (× 10 ¹² mol cm ^{−1} s ^{−1})
(c) Nafion 117					
1	0.5774	0.54	15.49	1.38	8.93
1.34	0.6105	0.40	22.19	6.12	13.53
1.5	0.6565	0.94	23.36	6.18	15.29
1.68	0.5744	0.33	27.65	6.96	15.86
2	0.6207	1.63	31.44	5.42	19.38
2.5	0.5956	0.83	38.87	2.69	23.11
3	0.6159	1.68	52.03	24.5	31.18

^a 95% confidence.

range. The resulting diffusion coefficient and solubility values for various gas pressures for the Nafion 117 membrane are plotted in Fig. 9. The diffusion coefficient of O₂ in Nafion was invariant with O₂ pressures between 1 and 3 atm. However, the solubility of O₂ in Nafion increased linearly with pressure and followed Henry's law for dilute gases (since the amount of oxygen dissolved in Nafion is very low),

$$K_{O_2} = \frac{P_{O_2}}{x_{O_2}} \quad (28)$$

where P_{O_2} is the partial pressure of O₂ and x is the mole fraction of O₂ in Nafion. The resulting Henry's law constant is 3.504×10^3 atm. Qualitatively, the trend of increasing O₂ solubility and an unchanging O₂ diffusion coefficient with pressure agree well with the results reported by Beattie et al. [15] for Nafion 117.

4.4. Carbon monoxide (CO)

Though there is plenty of data reported in the literature on O₂ diffusion in Nafion, there are not any analogous measurements on CO diffusion. This is not because of lack of need for such parameters. This is because, one can always assume a CO diffusion coefficient value equal to that of O₂ since the size of a CO molecule is only 14% smaller than that of an O₂ molecule. These values are routinely used by the PEM fuel cell modeling community to simulate the behavior of unit cells or fuel cell stacks operating on reformat feed that have as much as 500 ppm CO. We therefore attempt to experimentally obtain these parameters using the EMT.

4.5. Effect of temperature

Fig. 10 shows the current response of the working electrode as a function of temperature between 25 and 65 °C. Unlike O₂, the CO

current transience has two distinct features—the time constants are remarkably different even for Nafion 112 and the current transience shows a step like behavior. Since no such behavior was seen for the case of O₂ diffusion, dual diffusion pathways can be ruled out. This could however be indicative of the way CO electro-oxidation occurs at the working electrode. The model fits the data well except during the transience where the step-like behavior is seen. The fit is shown in Fig. 11 where the large error values in the neighborhood of the step indicative of this can be seen. The estimated diffusion coefficient and solubility values as a function of temperature are plotted in Fig. 12. The corresponding confidence intervals and the permeability values are listed in Table 4. The resulting Arrhenius equations for diffusion coefficient and solubility of CO in Nafion 112, respectively are,

$$D_{CO, Nafion 112} = 4.02 \times 10^{-4} \exp \left[\frac{-2406}{T} \right] \quad (29)$$

$$c_{CO, Nafion 112} = 5.43 \times 10^{-6} \exp \left[\frac{449.8}{T} \right] \quad (30)$$

The corresponding activation energy for CO diffusion through Nafion is ~ 20 kJ mol^{−1} and the mixture of enthalpy is ~ 3.74 kJ mol^{−1}.

4.6. Hydrogen sulfide (H₂S)

H₂S ionization ($H_2S \xrightleftharpoons{K_s} H^+ + HS^-$) in water was taken into account and was assumed to be in equilibrium with the dissociated species. The ionization constant for the first dissociation of H₂S is 3.9×10^{-8} at 0 °C and 3.0×10^{-7} at 100 °C and is assumed to be linear in this range [43]. The second ionization constant was reported in a review by Myers [44] to be $\sim 10^{-19}$. Due to the exceedingly small value for the second ionization constant, we assume that all of H₂S in the aqueous phase and in Nafion are in molecular form. The overall

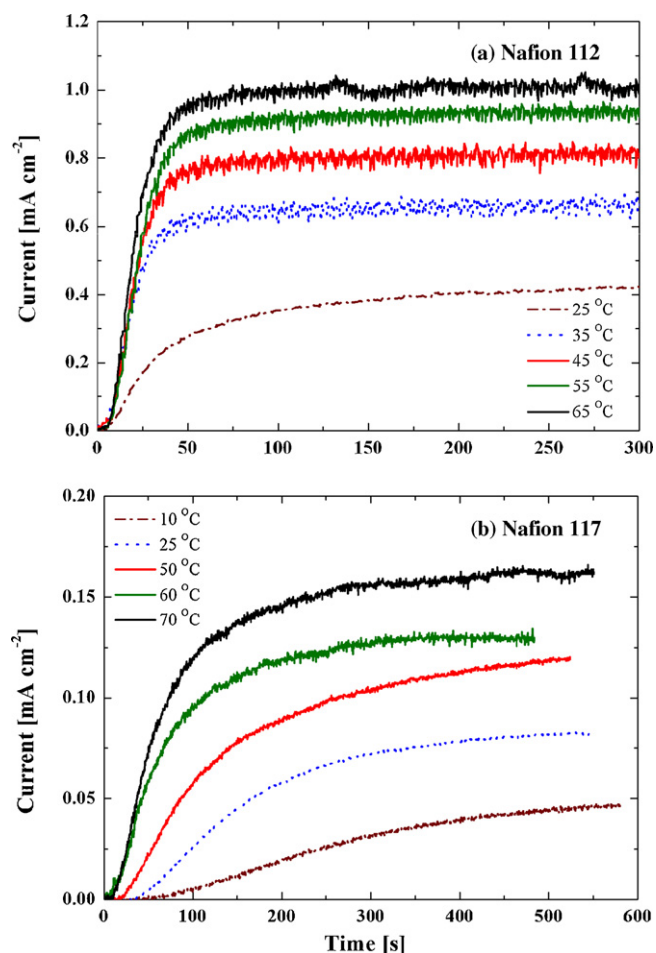


Fig. 4. Current response at the working electrode due to oxygen diffusion through (a) Nafion 112 and (b) Nafion 117 and subsequent reduction at the working electrode at different temperatures. The working electrode was held at 0.1 V vs. Ag/AgCl reference electrode (~ 0.3 V vs. SHE).

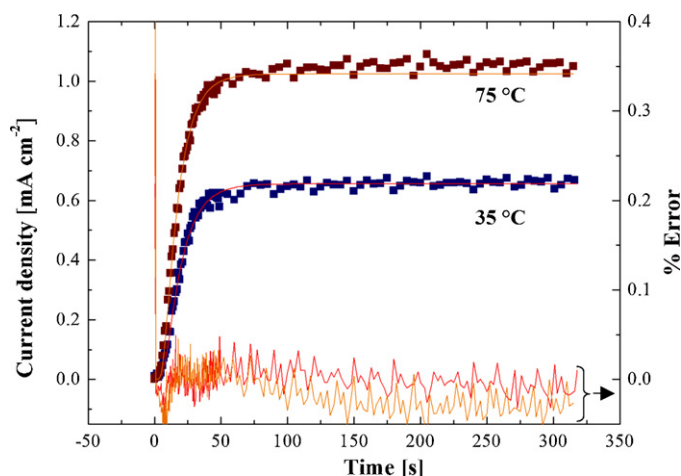


Fig. 5. Comparison of the experimental data (symbols) to Eq. (11) (lines) fitted with the diffusion coefficient and solubility of O₂ in Nafion 112 (for 25 and 65 °C) determined from the method of least squares. The respective error between data and the fit is shown on the secondary axis.

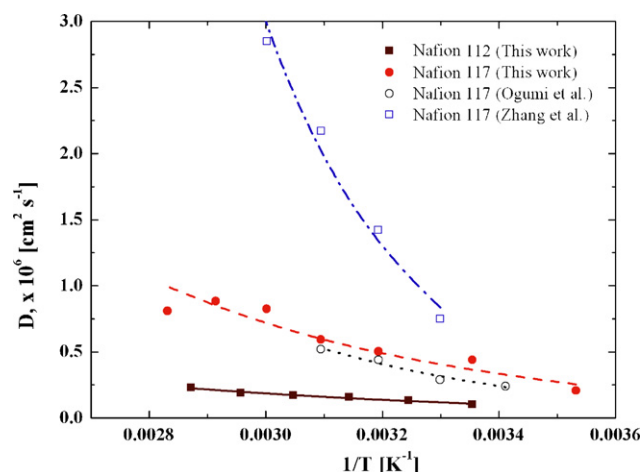


Fig. 6. Oxygen diffusion coefficient in Nafion 112 and 117 estimated using the electrochemical monitoring technique is compared to those reported by Ogumi et al. [9] and Zhang et al. [49]. The symbols correspond to data and the lines correspond to exponential fits. For data obtained in this work the fits correspond to Eqs. (24) and (25). The estimated activation energies for O₂ diffusion through Nafion 112 and Nafion 117 are ~ 12.58 and ~ 16.2 kJ mol⁻¹, respectively.

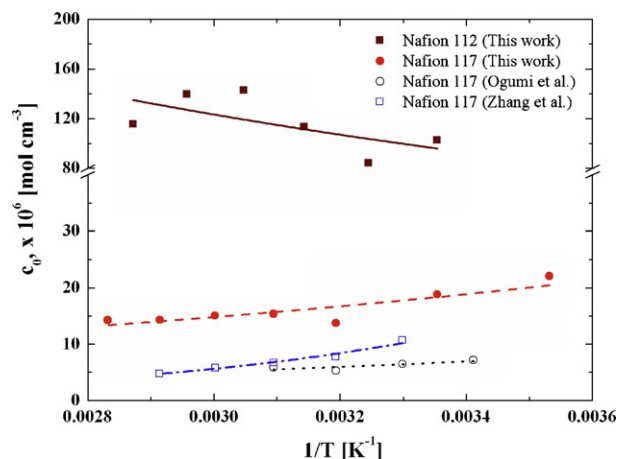


Fig. 7. Oxygen solubility in Nafion 112 and 117 estimated using the electrochemical monitoring technique is compared to those reported by Ogumi et al. [9] and Zhang et al. [49]. The symbols and lines correspond to data and exponential fits, respectively. For data obtained in this work, the fits correspond to Eqs. (26) and (27). The estimated mixture of enthalpies for Nafion 112 and 117 are 5.88 and 5.03 kJ mol⁻¹, respectively.

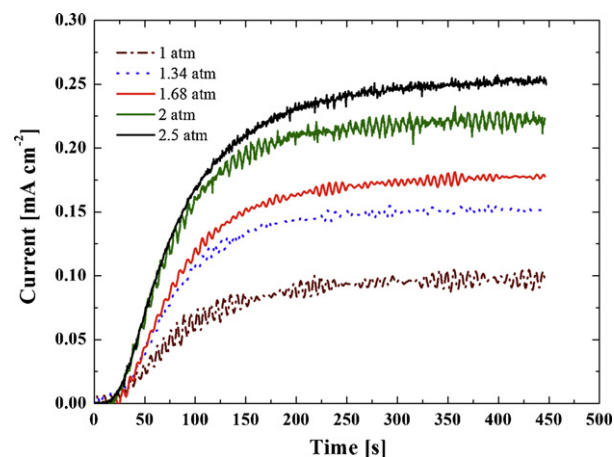


Fig. 8. Current response due to oxygen diffusion through Nafion 117 membrane and reduction at the working electrode for different oxygen pressures from 1 to 3 atm and 25 °C. The working electrode was held at 0.1 V vs. Ag/AgCl reference electrode (~ 0.3 V vs. SHE).

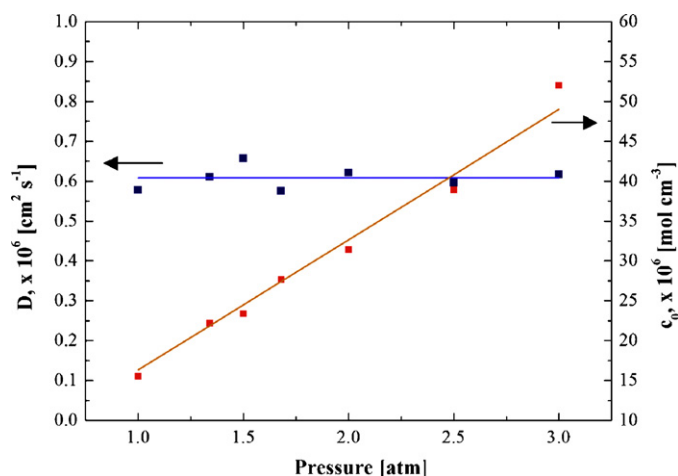


Fig. 9. Diffusion coefficient and solubility of oxygen in Nafion 117 for different oxygen pressures between 1 and 3 atm at 25 °C. The symbols and lines, respectively correspond to data and linear fits.

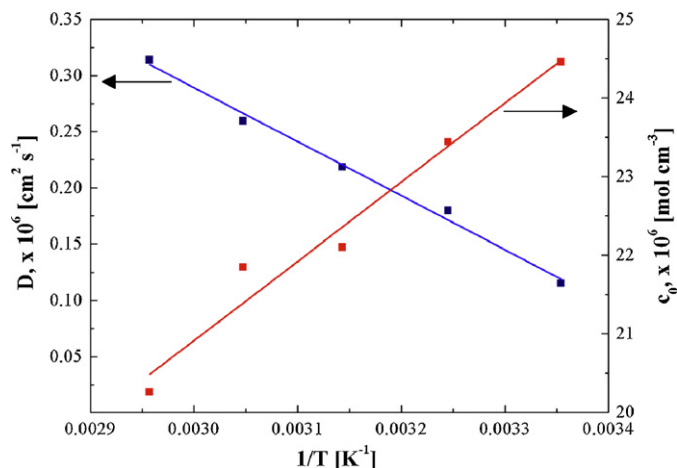


Fig. 12. Carbon monoxide diffusion coefficient and solubility in Nafion 112 estimated from data obtained using the electrochemical monitoring technique. The lines are exponential fits according to Eqs. (29) and (30). The estimated activation energy for CO diffusion through Nafion 112 is $\sim 20 \text{ kJ mol}^{-1}$ and the mixture of enthalpy is $\sim 3.74 \text{ kJ mol}^{-1}$.

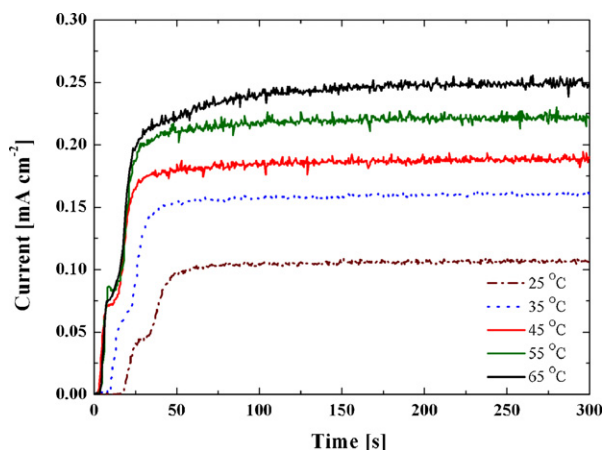


Fig. 10. Current response due to CO diffusion through Nafion 112 and subsequent oxidation to CO_2 at the working electrode at different temperatures. The working electrode was held at 0.9 V vs. Ag/AgCl ($\sim 1.1 \text{ V}$ vs. SHE).

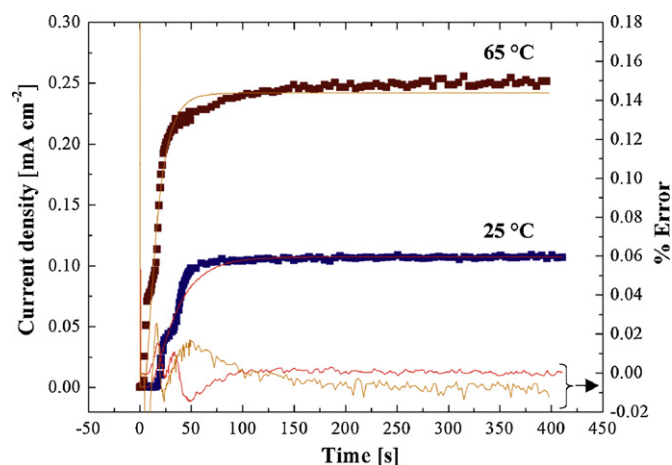


Fig. 11. Comparison of the experimental data (symbols) to Eq. (11) (lines) fitted with the diffusion coefficient and solubility of CO in Nafion 112 (for 25 and 65 °C) determined from the method of least squares. The respective error between data and the fit is shown on the secondary axis.

diffusion of all the species is measured in this study and therefore H_2S diffusion refers to the diffusion of H_2S and its dissociated species.

4.7. Effect of temperature

Fig. 13 shows the current response at the working electrode due to H_2S diffusion through Nafion 112 and subsequent oxidation at the working electrode. The data was fit to Eq. (11) and is shown in Fig. 14. It can be seen from the large error values that the fit is not good in the transient region. Since H_2S dissolves in water, the combined diffusion of the dissociated species and their oxidation is thought to physically take place as one fixed group and therefore treated mathematically with one set of equations. This can lead to significant errors in estimation of the diffusion coefficient and solubility values. It is assumed that the diffusion coefficient of H_2S in Nafion 112 is same as that of HS^- , which may not be true since the size of latter is 3% smaller than the former. The estimated diffusion coefficient and solubility values as a function of temperature are plotted in Fig. 15. The corresponding confidence intervals and the permeability values are listed in Table 5. The resulting Arrhenius

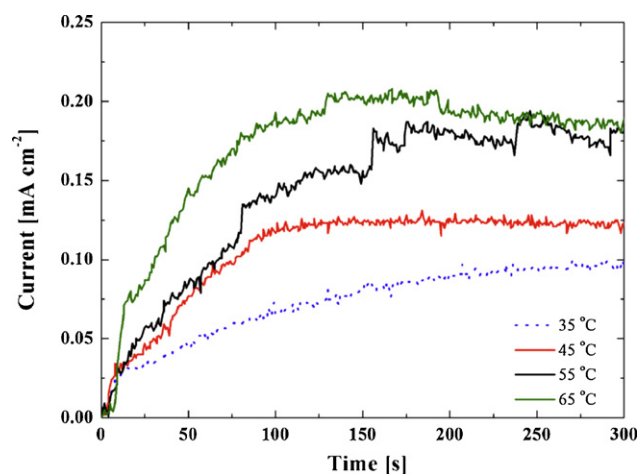
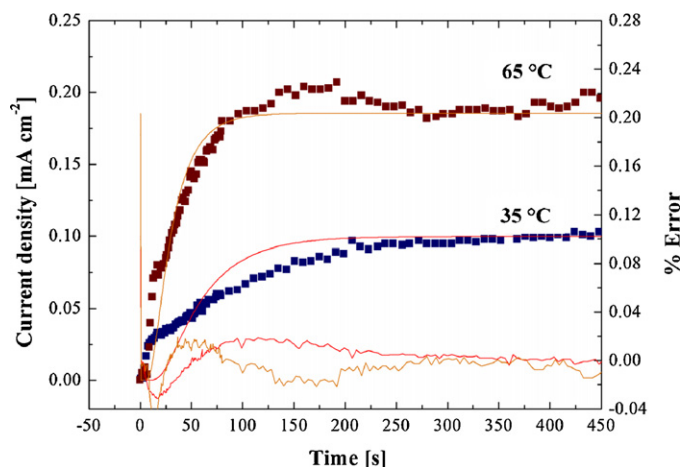
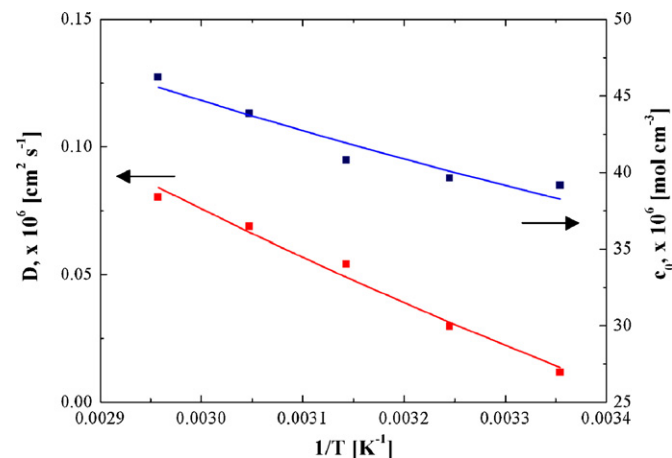


Fig. 13. Current response due to H_2S diffusion through Nafion 112 and subsequent oxidation at the working electrode at different temperatures. The working electrode was held at 0.9 V vs. Ag/AgCl ($\sim 1.1 \text{ V}$ vs. SHE).

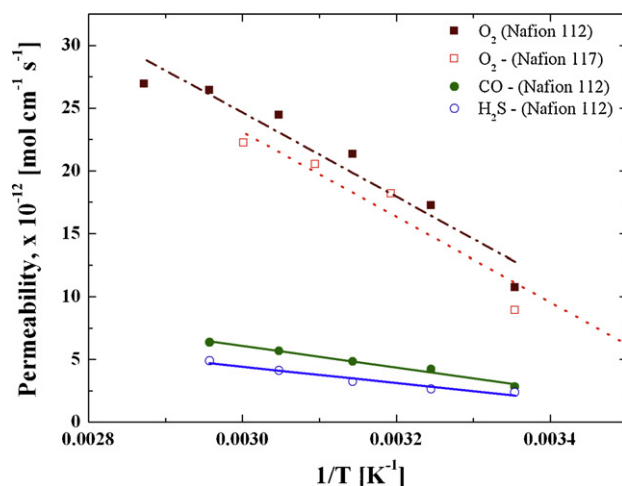
Table 4

Diffusion coefficient, solubility and permeability of CO in a Nafion 112 membrane and the confidence of the estimated parameters.

$T (^{\circ}\text{C})$	Diffusion coefficient, $D (\times 10^6 \text{ cm}^2 \text{ s}^{-1})$	Confidence interval ^a ($\times 10^8 \text{ cm}^2 \text{ s}^{-1}$)	Solubility, c_g ($\times 10^6 \text{ mol cm}^{-3}$)	Confidence interval ^a ($\times 10^6 \text{ mol cm}^{-3}$)	Permeability, $D \times c_g$ ($\times 10^{12} \text{ mol cm}^{-1} \text{ s}^{-1}$)
25	0.1153	0.22	24.46	0.56	2.82
35	0.1796	0.39	23.44	0.60	4.21
45	0.2185	1.08	22.10	0.74	4.83
55	0.2595	1.07	21.85	1.01	5.67
65	0.3139	0.61	20.26	0.81	6.36

^a 95% confidence.**Fig. 14.** Comparison of the experimental data (symbols) to Eq. (11) (lines) fitted with the diffusion coefficient and solubility of H₂S in Nafion 112 (for 35 and 65 °C) determined from the method of least squares. The respective error between data and the fit is shown on the secondary axis.**Fig. 15.** Hydrogen sulfide diffusion coefficient and solubility (symbols) in Nafion 112 estimated from data obtained using the electrochemical monitoring technique. The lines are exponential fits according to Eqs. (31) and (32). The estimated activation energy for H₂S diffusion through Nafion 112 is $\sim 8.85 \text{ kJ mol}^{-1}$ and the mixture of enthalpy is $\sim 7.61 \text{ kJ mol}^{-1}$. Hydrogen sulfide diffusion coefficient and solubility (symbols) in Nafion 112 estimated from data obtained using the electrochemical monitoring technique. The lines are exponential fits according to Eqs. (31) and (32).**Table 5**Diffusion coefficient, solubility and permeability of H₂S in a Nafion 112 membrane and the confidence of the estimated parameters.

$T (^{\circ}\text{C})$	Diffusion coefficient, $D (\times 10^6 \text{ cm}^2 \text{ s}^{-1})$	Confidence interval ^a ($\times 10^8 \text{ cm}^2 \text{ s}^{-1}$)	Solubility, c_g ($\times 10^6 \text{ mol cm}^{-3}$)	Confidence interval ^a ($\times 10^6 \text{ mol cm}^{-3}$)	Permeability, $D \times c_g$ ($\times 10^{12} \text{ mol cm}^{-1} \text{ s}^{-1}$)
25	0.0848	0.34	26.94	2.84	2.37
35	0.0877	0.17	29.96	1.20	2.63
45	0.0948	0.19	34.02	0.79	3.22
55	0.1130	0.27	36.45	3.35	4.12
65	0.1273	0.41	38.36	1.50	4.88

^a 95% confidence.**Fig. 16.** Permeability of O₂, CO and H₂S in Nafion membranes estimated using the electrochemical monitoring technique is plotted as a function of temperature. The lines represent exponential fits (i.e., shows Arrhenius dependence with temperature).

equations for diffusion coefficient and solubility of H₂S in Nafion 112, respectively are,

$$D_{\text{H}_2\text{S, Nafion 112}} = 2.87 \times 10^{-6} \exp \left[\frac{-1065}{T} \right] \quad (31)$$

$$c_{\text{H}_2\text{S, Nafion 112}} = 5.86 \times 10^{-4} \exp \left[\frac{-915}{T} \right] \quad (32)$$

The corresponding activation energy for H₂S diffusion through Nafion is $\sim 8.85 \text{ kJ mol}^{-1}$ and the mixture of enthalpy is $\sim 7.61 \text{ kJ mol}^{-1}$. Unlike O₂ and CO, solubility of H₂S in Nafion increases with temperature.

Fig. 16 shows the permeability of all three gases in Nafion. Since both diffusion coefficient and solubility show Arrhenius dependence with temperature, it is expected for permeability to show the same dependence. Oxygen permeability in Nafion 112 is slightly higher than in Nafion 117 even after correcting for swelling and reverse-osmosis. The increase in O₂ permeability with temperature is the same for both thicknesses (similar activation energies). Overall, the gas permeability in Nafion increases with temperature for all the three gases studied, as expected of gas diffusion in most polymers.

5. Conclusions

The electrochemical monitoring technique has been used on a Devanathan–Stachurski type diffusion cell made from a fuel cell assembly to measure the diffusion coefficient and solubility of O₂, CO and H₂S in Nafion membranes. The diffusion cell demonstrated here has the ability to test gas crossover properties of a membrane as a function of cell temperature and gas pressure. The design can be modified easily to evaluate humidity effects. The current–time data obtained from the EMT has been used in conjunction with a Fickian model to obtain relevant transport parameters for these three gases in Nafion. Membrane swelling and reverse-gas diffusion due to water flux from the liquid cell to the gas side has been accounted for in the parameter estimation routine. The contribution from both the former and the latter were minimal. From diffusion data obtained at various temperatures of interest for PEM fuel cell operation, the activation energy for diffusion and the enthalpy of mixing have been determined for these three gases. The permeability of all three gases was observed to increase with temperature. The estimated parameters agree very well with those reported in the literature. The data reported in this paper can be used by the PEMFC modeling community and the experimental procedure can be adapted as a diagnostic tool. Since gas crossover plays an important role in both the fuel cell power output as well as its long-term durability, we recommend that this procedure be routinely used to evaluate gas transport properties of proton exchange membranes.

Acknowledgements

The authors gratefully acknowledge the support from the National Reconnaissance Office for *Hybrid Advanced Power Sources* under grant # NRO-00-C-1034 and the National Science Foundation for funding Saahir Khan under the *Research Experience for Undergraduates (REU)* Program.

Appendix A. Nomenclature

A	cross-sectional area of the membrane, cm ²
$c(x, t)$	gas concentration at a distance, x , from the membrane and a given time, t
c_g	solubility of gas g in Nafion, mol cm ⁻³
C_{kk}	values of the inverted approximate Hessian matrix for element k where k represents D or c_0
$c_{g,W}$	solubility of gas g in water, mol cm ⁻³
c_W	solubility of water in Nafion, mol cm ⁻³
D_g	diffusion coefficient of gas g , cm ² s ⁻¹
$D_{W,F}$	Fickian diffusion coefficient for water in Nafion, cm ² s ⁻¹
E_a	activation energy, kJ mol ⁻¹
F	Faraday's constant, 96,487 C equiv ⁻¹
$i(t)$	current, A
i_∞	limiting current, A
K_{O_2}	Henry's law constant for O ₂
k_s	first ionization constant for H ₂ S in water
L	thickness of membrane, cm
M_m	molecular weight of the membrane, g mol ⁻¹
m	number of variables
n	number of data points taken in each trial
n_e	number of electrons transferred
$P_{g,W}$	permeability of gas g due to water flux from the liquid side to gas side, mol cm ⁻¹ s ⁻¹
P_k	k th parameter
\hat{P}_k	estimate of the k th parameter
R	universal gas constant, 8.314 J mol ⁻¹ K ⁻¹

s_g	stoichiometric coefficient of gas g
$s_{\hat{P}_k}$	standard deviation
x	distance from the catalyst-layer, cm
t_γ	value of the t distribution
T	temperature, K (or °C)
t	time, s
\bar{V}_0	initial volume of the membrane, cm ³
\bar{V}_m	partial molar volume of the membrane, cm ³ mol ⁻¹
$X_{O_2,W}$	mole fraction solubility of O ₂ in pure liquid water at 101 kPa
$SC_{O,W}$	solubility of CO in pure liquid water, nL L ⁻¹
λ	moles of water per mole of sulfonic acid sites (water content)
$\hat{\lambda}$	average membrane water content
ρ_m	density of the membrane, g cm ⁻³
γ	confidence

References

- [1] V.A. Sethuraman, J.W. Weidner, A.T. Haug, S. Motupally, L.V. Protsailo, J. Electrochem. Soc. 155 (2008) B50.
- [2] V.A. Sethuraman, J.W. Weidner, A.T. Haug, L.V. Protsailo, J. Electrochem. Soc. 155 (2008) B119.
- [3] Y. Shao, G. Yin, Y. Gao, J. Power Sources 171 (2007) 558.
- [4] R. Borup, J. Meyers, B. Pivovar, Y.S. Kim, R. Mukundan, N. Garland, D. Myers, M. Wilson, F. Garzon, D. Wood, P. Zelenay, K. More, K. Stroh, T. Zawodzinski, J. Boncella, J.E. McGrath, M. Inaba, K. Miyatake, M. Hori, K. Ota, Z. Ogumi, S. Miyata, A. Nishikata, Z. Siroma, Y. Uchimoto, K. Yasuda, K. Kimijima, N. Iwashita, Chem. Rev. 107 (2007) 3904.
- [5] J. Yu, B. Yi, D. Xing, F. Liu, Z. Shao, Y. Fu, H. Zhang, Phys. Chem. Chem. Phys. 5 (2003) 611.
- [6] D.A. Schiraldi, J. Macromol. Sci. C: Polym. Rev. 46 (2006) 315.
- [7] M. Devanathan, Z. Stachurski, Proc. R. Soc. Edinb. Ser. A 270 (1962) 90.
- [8] R.S. Yeo, J. McBreen, J. Electrochem. Soc. 126 (1979) 1682.
- [9] Z. Ogumi, Z. Takehara, S. Yoshizawa, J. Electrochem. Soc. 131 (1984) 769.
- [10] Z. Ogumi, T. Kuroe, Z. Takehara, J. Electrochem. Soc. 132 (1985) 2601.
- [11] A.T. Haug, R.E. White, J. Electrochem. Soc. 147 (2000) 980.
- [12] T. Lehtinen, G. Sundholm, S. Holmberg, F. Sundholm, P. Bjornbom, M. Bersell, Electrochim. Acta 43 (1998) 1881.
- [13] A. Parthasarathy, C.R. Martin, S. Srinivasan, J. Electrochem. Soc. 138 (1991) 916.
- [14] V.I. Basura, P.D. Beattie, S. Holdcroft, J. Electroanal. Chem. 458 (1998) 1.
- [15] P.D. Beattie, V.I. Basura, S. Holdcroft, J. Electroanal. Chem. 468 (1999) 180.
- [16] K. Aoki, J. Osteryoung, J. Electroanal. Chem. 122 (1981) 19.
- [17] C.P. Winlove, K.H. Parker, R.K.C. Oxenham, J. Electroanal. Chem. 170 (1984) 293.
- [18] A. Gamez, D. Richard, P. Gallezot, F. Gloaguen, R. Faure, R. Durand, Electrochim. Acta 41 (1996) 307.
- [19] J. Sonnenburg, J. Gao, J.H. Weiner, Macromolecules 23 (1990) 4653.
- [20] J. Staser, J. Weidner, J. Electrochem. Soc. 156 (2009) B836.
- [21] H. Takenaka, E. Torikai, Y. Kawami, N. Wakabayashi, Int. J. Hydrogen Energy 7 (1982) 397.
- [22] A.Z. Weber, J. Newman, J. Electrochem. Soc. 151 (2004) A311.
- [23] J.P. Meyers, J. Newman, J. Electrochem. Soc. 149 (2002) A718.
- [24] M. Bass, V. Freger, Polymer 49 (2008) 497.
- [25] D.R. Morris, X.J. Sun, Appl. Polym. Sci. 50 (1993) 1445.
- [26] P. Vanysek, in: D.R. Lide (Ed.), CRC Handbook of Chemistry and Physics, 70th ed., CRC Press, New York, 1989–1990, p. 153.
- [27] H.A. Gasteiger, N. Markovic, P.N. Ross Jr., E.J. Cairns, J. Phys. Chem. B 98 (1994) 617.
- [28] V.A. Sethuraman, L.A. Wise, S. Balasubramanian, J.W. Weidner, Electrochem. Soc. Trans. 16 (2008) 1.
- [29] D. Fan, R.E. White, N. Gruberger, J. Appl. Electrochem. 22 (1992) 770.
- [30] J.V. Beck, K.J. Arnold, Parameter Estimation in Engineering and Science, John Wiley & Sons, New York, 1977.
- [31] M.C. Kimble, R.E. White, Y.M. Tsou, R.N. Beaver, J. Electrochem. Soc. 137 (1990) 2510.
- [32] R.L. Burden, J.D. Faires, Numerical Analysis, 5th ed., PWS Publishing Co., Boston, MA, 1993.
- [33] I. Guttman, S.S. Wilks, J.S. Hunter, Introductory Engineering Statistics, 3rd ed., John Wiley & Sons, New York, 1982.
- [34] T.V. Nguyen, R.E. White, J. Electrochem. Soc. 140 (1993) 2178.
- [35] S. Motupally, A.J. Becker, J.W. Weidner, J. Electrochem. Soc. 147 (2000) 3171.
- [36] L.H. Gevantman, in: D.R. Lide (Ed.), CRC Handbook of Chemistry and Physics, 79th ed., CRC Press, New York, 1998–1999, p. 8.
- [37] D.A. Wiesenburger, N.L. Guinasso Jr., J. Chem. Eng. Data 24 (1979) 356.
- [38] J.J. Carroll, A.E. Mather, Geochim. Cosmochim. Acta 53 (1989) 1163.
- [39] K. Broka, P. Ekdunge, J. Appl. Electrochem. 27 (1997) 117.
- [40] J.S. Chiou, D.R. Paul, Ind. Eng. Chem. Res. 27 (1988) 2161.
- [41] T. Sakai, H. Takenaka, E. Torikai, J. Electrochem. Soc. 133 (1986) 88.
- [42] T. Sakai, H. Takenaka, N. Wakabayashi, Y. Kawami, E. Torikai, J. Electrochem. Soc. 132 (1985) 1328.

- [43] J.A. Barbero, K.G. McCurdy, P.R. Tremaine, *Can. J. Chem.* 60 (1982) 1872.
- [44] R.J. Myers, *J. Chem. Educ.* 63 (1986) 687.
- [45] A. Parthasarathy, S. Srinivasan, A.J. Appleby, C.R. Martin, *J. Electrochem. Soc.* 139 (1992) 2530.
- [46] F.N. Büchi, M. Wakizoe, S. Srinivasan, *J. Electrochem. Soc.* 143 (1996) 927.
- [47] P. Gode, G. Lindbergh, G. Sundholm, *J. Electroanal. Chem.* 518 (2002) 115.
- [48] W.J. Koros, D.R. Paul, A.A. Rocha, *J. Polym. Sci.* 14 (1976) 687.
- [49] L. Zhang, C. Ma, S. Mukerjee, *Electrochim. Acta* 48 (2003) 1845.